

## 2-(2-Pyridylamino)pyridinium tetrachloridozincate(II)

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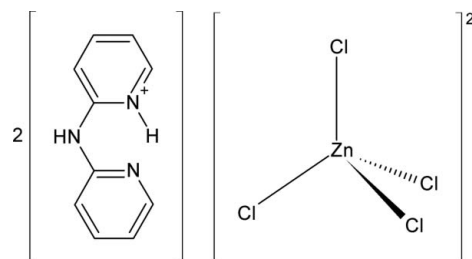
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Key indicators: single-crystal X-ray study;  $T = 150$  K; mean  $\sigma(\text{C}-\text{C}) = 0.004$  Å;  $R$  factor = 0.032;  $wR$  factor = 0.080; data-to-parameter ratio = 18.7.

The structure of the title compound,  $(\text{C}_{10}\text{H}_{10}\text{N}_3)_2[\text{ZnCl}_4]$ , is composed of  $\text{C}_{10}\text{H}_9\text{N}_3\text{H}^+$  (DPAH<sup>+</sup>) cations and  $[\text{ZnCl}_4]^{2-}$  anions. The two pyridyl rings of DPAH<sup>+</sup> are approximately coplanar, with a dihedral angle of  $7.2(2)^\circ$  between their corresponding least-squares planes. The proton is disordered in a one-to-one ratio over the two chemically equivalent pyridyl N atoms. An intramolecular hydrogen bond is formed between the pyridinium H atom and the pyridyl N atom of the other pyridyl ring. The Zn atom lies on a twofold rotation axis. There are also some weak  $\text{N}-\text{H}\cdots\text{Cl}$  hydrogen bonds. These interactions lead to the formation of an alternating zigzag chain in the solid state. The results clearly show that reducing agents normally used in hydrothermal syntheses, such as metallic zinc employed here, are also active in terms of coordination chemistry.

### Related literature

For related literature, see: Bock *et al.* (1998); Bose *et al.* (2004); Camus *et al.* (2000); Chowdhury *et al.* (2005); Du & Zhao (2004); Gillon *et al.* (2000); Marinescu *et al.* (2005); Rahaman *et al.* (2005); Rice *et al.* (2002); Visser *et al.* (1997); Willett (1995); Youngme *et al.* (2005).



### Experimental

#### Crystal data

$(\text{C}_{10}\text{H}_{10}\text{N}_3)_2[\text{ZnCl}_4]$   
 $M_r = 551.61$   
 Monoclinic,  $C2/c$   
 $a = 14.620(3)$  Å  
 $b = 11.260(2)$  Å  
 $c = 14.765(3)$  Å  
 $\beta = 101.13(3)^\circ$

$V = 2384.9(8)$  Å<sup>3</sup>  
 $Z = 4$   
 Mo  $K\alpha$  radiation  
 $\mu = 1.50$  mm<sup>-1</sup>  
 $T = 150(2)$  K  
 $0.1 \times 0.1 \times 0.1$  mm

#### Data collection

Siemens SMART CCD area-detector diffractometer  
 Absorption correction: multi-scan (SADABS in SAINT-NT; Bruker, 1999)  
 $T_{\min} = 0.861$ ,  $T_{\max} = 0.861$

8907 measured reflections  
 2718 independent reflections  
 2063 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.029$

#### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.031$   
 $wR(F^2) = 0.079$   
 $S = 1.17$   
 2718 reflections  
 145 parameters

H atoms treated by a mixture of independent and constrained refinement  
 $\Delta\rho_{\max} = 0.53$  e Å<sup>-3</sup>  
 $\Delta\rho_{\min} = -0.31$  e Å<sup>-3</sup>

**Table 1**

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{N1}-\text{H1N}\cdots\text{N3}$	1.06	1.82	2.612 (3)	129
$\text{N3}-\text{H3N}\cdots\text{N1}$	1.03	1.83	2.612 (3)	130
$\text{N1}-\text{H1N}\cdots\text{Cl2}$	1.06	2.73	3.548 (2)	134
$\text{N3}-\text{H3N}\cdots\text{Cl2}$	1.03	2.74	3.477 (3)	129
$\text{N2}-\text{H2N}\cdots\text{Cl1}^\dagger$	0.93	2.39	3.313 (3)	170

Symmetry code: (i)  $x, -y + 1, z - \frac{1}{2}$ .

Data collection: SMART-NT (Bruker, 2001); cell refinement: SAINT-NT (Bruker, 1999); data reduction: SAINT-NT; program(s) used to solve structure: SHELXTL-NT (Sheldrick, 2008); program(s) used to refine structure: SHELXTL-NT; molecular graphics: SHELXTL-NT; software used to prepare material for publication: SHELXTL-NT.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: ZL2105).

## References

- Bock, H., Schodel, H., Van, T. T. H., Dienelt, R. & M. Gluth (1998). *J. Prakt. Chem. Chem. Z.* **340**, 722–732.
- Bose, D., Rahaman, S. H., Mostafa, G., Walsh, R. D. B., Zaworotko, M. J. & Ghosh, B. K. (2004). *Polyhedron*, **23**, 545–552.
- Bruker (1999). *SAINT-NT* and *SADABS*. Bruker AXS Inc., Madison, Wisconsin, USA.
- Bruker (2001). *SMART-NT*. Bruker AXS Inc., Madison, Wisconsin, USA.
- Camus, A., Marsich, N., Lanfredi, A. M. M., Ugozzoli, F. & Massera, C. (2000). *Inorg. Chim. Acta*, **309**, 1–2.
- Chowdhury, H., Ghosh, R., Rahaman, S. H., Rosair, G. M. & Ghosh, B. K. (2005). *Indian J. Chem. Sect. A*, **44**, 687–692.
- Du, M. & Zhao, X.-J. (2004). *Acta Cryst.* **E60**, o439–o441.
- Gillon, A. L., Lewis, G. R., Orpen, A. G., Rotter, S., Starbuck, J., Wang, X. M., Rodríguez-Martín, Y. & Ruiz-Pérez, C. (2000). *J. Chem. Soc. Dalton Trans.* pp. 3897–3905.
- Marinescu, G., Andruh, M., Julve, M., Lloret, F., Llusar, R., Uriel, S. & Vaissermann, J. (2005). *Cryst. Growth Des.* **5**, 261–267.
- Rahaman, S. H., Bose, D., Chowdhury, H., Mostafa, G., Fun, H.-K. & Ghosh, B. K. (2005). *Polyhedron*, **24**, 1837–1844.
- Rice, C. R., Onions, S., Vidal, N., Wallis, J. D., Senna, M. C., Pilkington, M. & Stoeckli-Evans, H. (2002). *Eur. J. Inorg. Chem.* pp. 1985–1997.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Visser, H. G., Purcell, W., Basson, S. S. & Claassen, Q. (1997). *Polyhedron*, **16**, 2851–2855.
- Willett, R. D. (1995). *Acta Cryst.* **C51**, 1517–1519.
- Youngme, S., Chailuecha, C., van Albada, G. A., Pakawatchai, C., Chaichit, N. & Reedijk, J. (2005). *Inorg. Chim. Acta*, **358**, 1068–1078.

## supporting information

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### S1. Comment

Core to the crystal engineering of supramolecular compounds are the different strategies that are used to control the molecular aggregation processes. For example, the selection of metal ions that promote the formation of extended arrays, and organic molecules capable of forming hydrogen bonding or close packed structures are fundamental to control the molecular aggregation (Guillon *et al.*, 2000, Rice *et al.*, 2002). Therefore, the strategies adopted in the preparation of crystal structures involve two main concepts. One of them is related with the use of the shape-controlled close packing of molecules, and the second one with the use of the specific interactions to control aggregation of molecular species (Guillon *et al.*, 2000).

2,2-Dipyridylamine (DPA) is used as an organic linker presenting different coordination modes ranging from monodentate to chelating bidentate or bridging tridentate. Furthermore, this organic ligand contains the amino group and aromatic rings that may induce intermolecular hydrogen bonding and  $\pi$ - $\pi$  stacking interactions that may lead to different structures. (Camus, *et al.* 2000, Bose *et al.*, 2004, Rahaman *et al.* 2005, Chowdhury *et al.*, 2005, Youmgme *et al.*, 2005, Marinescu *et al.*, 2005).

In this work we inform on the synthesis of a novel crystal structure with a tetrachlorozincate anion obtained by hydrothermal synthesis. The reaction conditions described in the experimental section lead to the oxidation of the metallic zinc used as the reducing agent and, to the formation of a crystalline tetrahalide species, which is present together with the cationic counterion, DPAH<sup>+</sup>.

The organic cation (DPAH<sup>+</sup>) exhibits two pyridinium rings in a highly planar arrangement with a dihedral angle between them of only 7.2 (2)°. This is consistent with the value observed for the isostructural cobalt derivative (DPAH<sup>+</sup>)<sub>2</sub>[CoCl<sub>4</sub>]<sup>2-</sup> of the title compound (Visser *et al.*, 1997); both compounds being isostructural. The pyridinium nitrogen atoms on the cation are in a "face to face" (or *U*) arrangement, allowing the existence of an intramolecular hydrogen bond (see hydrogen bonding table), also observed for other 2-(pyridin-2-ylamino)pyridinium salts such as the Cl<sup>-</sup>, squarate (C<sub>4</sub>O<sub>4</sub>H<sup>-</sup>), tetraphenylborate (B(C<sub>6</sub>H<sub>5</sub>)<sub>4</sub>)<sup>-</sup> (Bock *et al.*, 1998) or NO<sub>3</sub><sup>-</sup> (Du & Zhao, 2004) compounds. The parameters for this interaction within these salts are basically identical. A different situation is observed for (DPAH<sup>2+</sup>)[CuCl<sub>4</sub>]<sup>2-</sup> (Willett, 1995) and (DPAH<sup>2+</sup>)(CF<sub>3</sub>SO<sub>3</sub>)<sup>-</sup> (Bock *et al.*, 1998) where the molecule is diprotonated, leading to an *S*-like conformation which precludes intramolecular hydrogen bonding.

As expected, the tetrachlorozincate anion, with the zinc atom lying on the twofold axis, displays an almost perfect tetrahedral coordination environment. The [ZnCl<sub>4</sub>]<sup>2-</sup> anion interacts with the cations through weak hydrogen bonds between Cl1 and Cl2 with the DPAH<sup>+</sup> cations, as summarized in the hydrogen bonding table and depicted in Figure 1. Each tetrachlorozincate anion is bonded in this way to four cations. This differs from what is observed in (DPAH<sup>2+</sup>)[CuCl<sub>4</sub>]<sup>2-</sup>, where the values N(amine)⋯Cl = 2.133 Å, N(pyridinium)⋯Cl = 2.307, 2.470 Å) suggest stronger interactions.

The rather strong distortion from square planar geometry in this latter molecule has been addressed to the hydrogen bonding interactions (Willett, 1995).

As can be seen in Figure 1, there are two pairs of strictly parallel cations, which are separated by approximately 3.3 Å, a value that is in the range for  $\pi$ - $\pi$  interactions (Marinescu *et al.*, 2005), but somewhat shorter than the 3.534 (5) Å reported for (DPAH<sup>+</sup>)NO<sub>3</sub> (Du & Zhao, 2004). Each pair of DPAH<sup>+</sup> cations is connected with a neighboring pair in a "head to tail" fashion (see Figure 1), thus leading to a packing arrangement with a zigzag chain with alternating organic cations and tetrachlorozincate anions as seen in Figure 2. These chains interact in the solid by means of  $\pi$ - $\pi$  contacts.

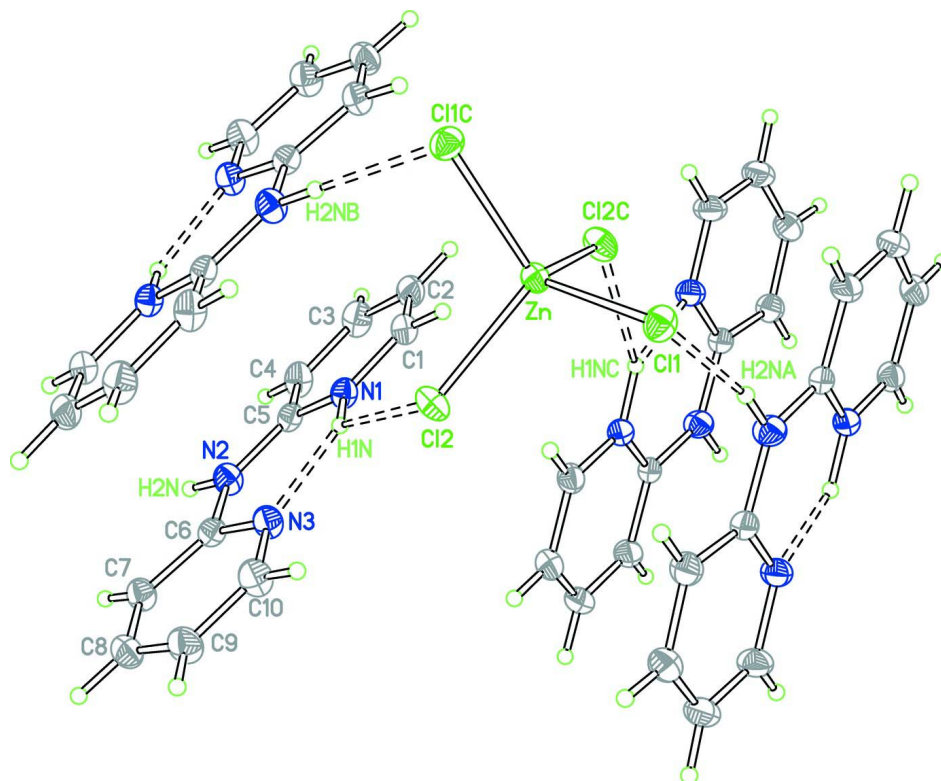
## S2. Experimental

The compound was obtained by hydrothermal synthesis employing the following reagents: CrCl<sub>3</sub>.6H<sub>2</sub>O, DPA, V<sub>2</sub>O<sub>5</sub>, Zn, H<sub>3</sub>PO<sub>4</sub> (85%) and H<sub>2</sub>O (all Aldrich, used without further purification) in the following quantities: 0.1523 g, 0.28 g, 0.1438 g, 0.1054 g, 0.58 ml and 5 ml respectively. This mixture was sealed in a 23 ml PTFE-lined stainless steel autoclave, and heated to 393 K (120°C) for 72 h. The yield of the studied compound was approximately 10%, and X-ray quality crystals were directly selected from the bulk mixture of products.

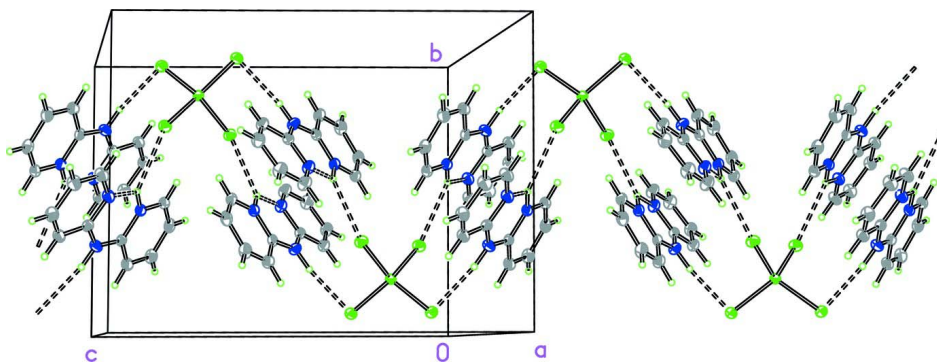
## S3. Refinement

The hydrogen atoms positions were calculated after each cycle of refinement with *SHELXL* (Bruker, 1999) using a riding model with C—H distance equal to 0.96 Å.  $U_{\text{iso}}(\text{H})$  values were set equal to 1.2 $U_{\text{eq}}$  of the parent carbon atom. At the final stages of refinement, the hydrogen bonded to the nitrogen atoms becomes evident in the Fourier Difference Map, with some disorder. This was modeled considering two half-occupied hydrogen sites, one on each pyridyl nitrogen atom.

These N—H hydrogen atoms were then refined using a riding model with C—N—H angles idealized for amide H atoms, but the N—H distances were allowed to refine freely.

**Figure 1**

Molecular drawing for I showing hydrogen bonding between  $[\text{ZnCl}_4]^{2-}$  and the organic cation. Atom numbering scheme is included. Displacement ellipsoids at 20%. The second hydrogen position H3N is not included in the diagram for clarity. Symmetry codes: A:  $x, 1 - y, 1/2 + z$ ; B:  $1 - x, 1 - y, 1 - z$ ; C  $1 - x, y, 3/2 - z$ .

**Figure 2**

Packing diagram showing the zigzag chains defined by weak intermolecular interactions in the solid state.

## 2-(2-Pyridylamino)pyridinium tetrachloridozincate(II)

### Crystal data

$(\text{C}_{10}\text{H}_{10}\text{N}_3)_2[\text{ZnCl}_4]$

$M_r = 551.61$

Monoclinic,  $C2/c$

Hall symbol:  $-C 2yc$

$a = 14.620 (3) \text{ \AA}$

$b = 11.260 (2) \text{ \AA}$

$c = 14.765 (3) \text{ \AA}$

$\beta = 101.13 (3)^\circ$

$V = 2384.9 (8) \text{ \AA}^3$

$Z = 4$

$F(000) = 1120$   
 $D_x = 1.536 \text{ Mg m}^{-3}$   
 Mo  $K\alpha$  radiation,  $\lambda = 0.71073 \text{ \AA}$   
 Cell parameters from 12557 reflections  
 $\theta = 2.9\text{--}27.5^\circ$

$\mu = 1.50 \text{ mm}^{-1}$   
 $T = 150 \text{ K}$   
 Prism, yellow  
 $0.1 \times 0.1 \times 0.1 \text{ mm}$

#### Data collection

Siemens SMART CCD area-detector  
 diffractometer  
 Radiation source: fine-focus sealed tube  
 Graphite monochromator  
 phi and  $\omega$  scans  
 Absorption correction: multi-scan  
 (SADABS in SAINT-NT; Bruker, 1999)  
 $T_{\min} = 0.861$ ,  $T_{\max} = 0.861$

8907 measured reflections  
 2718 independent reflections  
 2063 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.029$   
 $\theta_{\max} = 27.5^\circ$ ,  $\theta_{\min} = 3.6^\circ$   
 $h = -18 \rightarrow 18$   
 $k = -14 \rightarrow 14$   
 $l = -18 \rightarrow 19$

#### Refinement

Refinement on  $F^2$   
 Least-squares matrix: full  
 $R[F^2 > 2\sigma(F^2)] = 0.031$   
 $wR(F^2) = 0.079$   
 $S = 1.17$   
 2718 reflections  
 145 parameters  
 0 restraints  
 Primary atom site location: structure-invariant  
 direct methods

Secondary atom site location: difference Fourier  
 map  
 Hydrogen site location: inferred from  
 neighbouring sites  
 H atoms treated by a mixture of independent  
 and constrained refinement  
 $w = 1/[\sigma^2(F_o^2) + (0.0297P)^2 + 1.3722P]$   
 where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\max} = 0.001$   
 $\Delta\rho_{\max} = 0.53 \text{ e \AA}^{-3}$   
 $\Delta\rho_{\min} = -0.31 \text{ e \AA}^{-3}$

#### Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry.

An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Least-squares planes ( $x, y, z$  in crystal coordinates) and deviations from them (\* indicates atom used to define plane)

$-7.5257 (0.0143) x - 6.4929 (0.0111) y + 10.6583 (0.0124) z = 0.0740 (0.0091)$

\*  $-0.0004 (0.0017) N3$  \*  $0.0057 (0.0017) C6$  \*  $-0.0049 (0.0019) C7$  \*  $-0.0009 (0.0021) C8$  \*  $0.0062 (0.0022) C9$  \*  
 $-0.0057 (0.0021) C10$

Rms deviation of fitted atoms = 0.0046

$-8.0734 (0.0131) x - 5.2989 (0.0109) y + 11.5436 (0.0110) z = 0.7176 (0.0115)$

Angle to previous plane (with approximate e.s.d.) =  $7.15 (0.16)$

\*  $0.0093 (0.0015) N1$  \*  $0.0017 (0.0017) C1$  \*  $-0.0085 (0.0019) C2$  \*  $0.0045 (0.0020) C3$  \*  $0.0061 (0.0019) C4$  \*  $-0.0132 (0.0016) C5$

Rms deviation of fitted atoms = 0.0081

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) etc. and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

#### Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$	Occ. (<1)
Zn	0.5000	0.81308 (3)	0.7500	0.03912 (11)	
Cl1	0.42368 (4)	0.92732 (5)	0.84063 (4)	0.05165 (16)	

Cl2	0.39576 (4)	0.69311 (5)	0.65703 (4)	0.05397 (16)	
N1	0.49641 (13)	0.42148 (17)	0.60363 (12)	0.0447 (4)	
H1N	0.437 (3)	0.478 (3)	0.5901 (7)	0.054*	0.50
C1	0.57114 (16)	0.4524 (2)	0.66942 (15)	0.0479 (5)	
H1	0.5699	0.5234	0.7013	0.057*	
C2	0.64808 (18)	0.3812 (3)	0.68967 (17)	0.0575 (6)	
H2	0.6992	0.4034	0.7343	0.069*	
C3	0.64815 (19)	0.2743 (3)	0.64175 (18)	0.0645 (7)	
H3	0.6995	0.2241	0.6552	0.077*	
C4	0.57324 (19)	0.2427 (2)	0.57501 (17)	0.0568 (6)	
H4	0.5732	0.1712	0.5434	0.068*	
C5	0.49701 (15)	0.3193 (2)	0.55522 (15)	0.0428 (5)	
N2	0.42142 (14)	0.29509 (19)	0.48707 (14)	0.0515 (5)	
H2N	0.4271 (2)	0.228 (2)	0.4519 (12)	0.062*	
C6	0.33748 (15)	0.3559 (2)	0.46259 (15)	0.0435 (5)	
C7	0.26855 (17)	0.3177 (2)	0.38963 (17)	0.0533 (6)	
H7	0.2780	0.2514	0.3550	0.064*	
C8	0.18643 (18)	0.3801 (3)	0.37007 (18)	0.0602 (7)	
H8	0.1393	0.3559	0.3219	0.072*	
C9	0.17357 (19)	0.4798 (3)	0.42241 (19)	0.0640 (7)	
H9	0.1180	0.5224	0.4103	0.077*	
C10	0.24470 (17)	0.5133 (3)	0.49185 (19)	0.0612 (7)	
H10	0.2370	0.5805	0.5262	0.073*	
N3	0.32622 (14)	0.45227 (19)	0.51276 (14)	0.0508 (5)	
H3N	0.378 (3)	0.4795 (17)	0.566 (3)	0.061*	0.50

*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Zn	0.0415 (2)	0.03647 (19)	0.03747 (19)	0.000	0.00295 (13)	0.000
Cl1	0.0604 (4)	0.0447 (3)	0.0527 (3)	0.0036 (3)	0.0181 (3)	-0.0025 (3)
Cl2	0.0527 (3)	0.0485 (3)	0.0531 (3)	-0.0037 (3)	-0.0089 (3)	-0.0057 (3)
N1	0.0472 (10)	0.0473 (11)	0.0400 (10)	0.0084 (9)	0.0093 (8)	0.0012 (8)
C1	0.0519 (13)	0.0508 (13)	0.0397 (12)	0.0013 (11)	0.0059 (10)	-0.0005 (10)
C2	0.0503 (14)	0.0758 (18)	0.0438 (13)	0.0054 (13)	0.0025 (10)	0.0067 (13)
C3	0.0570 (15)	0.0811 (19)	0.0537 (15)	0.0292 (14)	0.0068 (12)	0.0073 (14)
C4	0.0635 (16)	0.0573 (15)	0.0494 (14)	0.0217 (13)	0.0103 (12)	-0.0014 (12)
C5	0.0450 (11)	0.0492 (12)	0.0357 (10)	0.0077 (10)	0.0114 (9)	0.0059 (10)
N2	0.0539 (12)	0.0482 (12)	0.0517 (12)	0.0065 (9)	0.0089 (9)	-0.0097 (9)
C6	0.0437 (12)	0.0453 (12)	0.0427 (12)	0.0040 (10)	0.0111 (9)	0.0040 (10)
C7	0.0550 (14)	0.0547 (14)	0.0486 (13)	0.0022 (12)	0.0062 (11)	-0.0063 (12)
C8	0.0507 (14)	0.0729 (18)	0.0527 (14)	0.0006 (13)	-0.0006 (11)	-0.0021 (13)
C9	0.0477 (14)	0.0779 (19)	0.0627 (16)	0.0164 (13)	0.0016 (12)	-0.0026 (14)
C10	0.0524 (14)	0.0685 (17)	0.0607 (16)	0.0195 (13)	0.0061 (12)	-0.0118 (14)
N3	0.0457 (11)	0.0574 (12)	0.0478 (11)	0.0096 (9)	0.0050 (9)	-0.0066 (10)

*Geometric parameters (Å, °)*

Zn—C12	2.2856 (8)	C5—N2	1.371 (3)
Zn—C12 <sup>i</sup>	2.2856 (8)	N2—C6	1.391 (3)
Zn—C11 <sup>i</sup>	2.2949 (7)	N2—H2N	0.9333
Zn—C11	2.2949 (7)	C6—N3	1.342 (3)
N1—C5	1.355 (3)	C6—C7	1.394 (3)
N1—C1	1.360 (3)	C7—C8	1.373 (4)
N1—H1N	1.0584	C7—H7	0.9300
C1—C2	1.366 (3)	C8—C9	1.396 (4)
C1—H1	0.9300	C8—H8	0.9300
C2—C3	1.397 (4)	C9—C10	1.365 (4)
C2—H2	0.9300	C9—H9	0.9300
C3—C4	1.371 (4)	C10—N3	1.359 (3)
C3—H3	0.9300	C10—H10	0.9300
C4—C5	1.395 (3)	N3—H3N	1.0328
C4—H4	0.9300		
C12—Zn—C12 <sup>i</sup>	107.54 (4)	N2—C5—C4	121.9 (2)
C12—Zn—C11 <sup>i</sup>	108.90 (3)	C5—N2—C6	129.7 (2)
C12 <sup>i</sup> —Zn—C11 <sup>i</sup>	109.79 (3)	C5—N2—H2N	115.2
C12—Zn—C11	109.79 (3)	C6—N2—H2N	115.2
C12 <sup>i</sup> —Zn—C11	108.90 (3)	N3—C6—N2	116.8 (2)
C11 <sup>i</sup> —Zn—C11	111.82 (4)	N3—C6—C7	121.9 (2)
C5—N1—C1	120.5 (2)	N2—C6—C7	121.3 (2)
C5—N1—H1N	119.7	C8—C7—C6	118.5 (2)
C1—N1—H1N	119.7	C8—C7—H7	120.8
N1—C1—C2	121.4 (2)	C6—C7—H7	120.8
N1—C1—H1	119.3	C7—C8—C9	120.1 (2)
C2—C1—H1	119.3	C7—C8—H8	120.0
C1—C2—C3	118.4 (2)	C9—C8—H8	120.0
C1—C2—H2	120.8	C10—C9—C8	118.2 (2)
C3—C2—H2	120.8	C10—C9—H9	120.9
C4—C3—C2	120.6 (2)	C8—C9—H9	120.9
C4—C3—H3	119.7	N3—C10—C9	122.6 (2)
C2—C3—H3	119.7	N3—C10—H10	118.7
C3—C4—C5	119.0 (2)	C9—C10—H10	118.7
C3—C4—H4	120.5	C6—N3—C10	118.6 (2)
C5—C4—H4	120.5	C6—N3—H3N	120.7
N1—C5—N2	118.03 (19)	C10—N3—H3N	120.7
N1—C5—C4	120.1 (2)		
C5—N1—C1—C2	-1.0 (3)	C5—N2—C6—N3	-1.0 (4)
N1—C1—C2—C3	-0.7 (4)	C5—N2—C6—C7	179.3 (2)
C1—C2—C3—C4	1.0 (4)	N3—C6—C7—C8	-1.0 (4)
C2—C3—C4—C5	0.4 (4)	N2—C6—C7—C8	178.7 (2)
C1—N1—C5—N2	-177.3 (2)	C6—C7—C8—C9	0.4 (4)
C1—N1—C5—C4	2.4 (3)	C7—C8—C9—C10	0.7 (4)



C3—C4—C5—N1	-2.1 (4)	C8—C9—C10—N3	-1.2 (5)
C3—C4—C5—N2	177.6 (2)	N2—C6—N3—C10	-179.1 (2)
N1—C5—N2—C6	-5.5 (4)	C7—C6—N3—C10	0.6 (4)
C4—C5—N2—C6	174.7 (2)	C9—C10—N3—C6	0.6 (4)

Symmetry code: (i)  $-x+1, y, -z+3/2$ .

*Hydrogen-bond geometry (Å, °)*

<i>D—H...A</i>	<i>D—H</i>	<i>H...A</i>	<i>D...A</i>	<i>D—H...A</i>
N1—H1 <i>N</i> ...N3	1.06	1.82	2.612 (3)	129
N3—H3 <i>N</i> ...N1	1.03	1.83	2.612 (3)	130
N1—H1 <i>N</i> ...C12	1.06	2.73	3.548 (2)	134
N3—H3 <i>N</i> ...C12	1.03	2.74	3.477 (3)	129
N2—H2 <i>N</i> ...C11 <sup>ii</sup>	0.93	2.39	3.313 (3)	170

Symmetry code: (ii)  $x, -y+1, z-1/2$ .