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2-(4-Ethoxybenzylidene)butanoic acid

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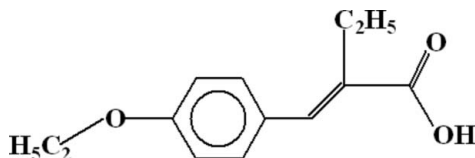
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Key indicators: single-crystal X-ray study; $T = 296$ K; mean $\sigma(\text{C}-\text{C}) = 0.002$ Å; R factor = 0.050; wR factor = 0.123; data-to-parameter ratio = 20.1.

In the crystal structure of the title compound, $\text{C}_{13}\text{H}_{16}\text{O}_3$, dimers are formed due to intermolecular $\text{O}-\text{H}\cdots\text{O}$ hydrogen bonding. There exists an intramolecular $\text{C}-\text{H}\cdots\text{O}$ hydrogen bond which forms a five-membered ring. There is also present a $\text{C}-\text{H}\cdots\pi$ interaction between a methyl CH group of an ethyl group and the centroid of the aromatic ring.

Related literature

For related literature, see: Bernstein *et al.* (1995); Burt (2004); Muhammad *et al.* (2007, 2008a,b,c); Niaz *et al.* (2008).



Experimental

Crystal data

$\text{C}_{13}\text{H}_{16}\text{O}_3$
 $M_r = 220.26$
Monoclinic, $P2_1/c$
 $a = 10.3192$ (7) Å
 $b = 22.0761$ (15) Å
 $c = 5.2362$ (3) Å
 $\beta = 100.751$ (4)°

$V = 1171.91$ (13) Å³
 $Z = 4$
Mo $K\alpha$ radiation
 $\mu = 0.09$ mm⁻¹
 $T = 296$ (2) K
 $0.26 \times 0.20 \times 0.16$ mm

Data collection

Bruker Kappa APEXII CCD diffractometer
Absorption correction: multi-scan (SADABS; Bruker, 2005)
 $T_{\min} = 0.975$, $T_{\max} = 0.985$

13990 measured reflections
3009 independent reflections
1599 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.049$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.050$
 $wR(F^2) = 0.122$
 $S = 1.00$
3009 reflections
150 parameters

H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{\max} = 0.14$ e Å⁻³
 $\Delta\rho_{\min} = -0.21$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{O1}-\text{H1}\cdots\text{O2}^i$	1.01 (2)	1.61 (2)	2.6184 (15)	176.8 (15)
$\text{C7}-\text{H7}\cdots\text{O1}$	0.93	2.30	2.7218 (19)	107
$\text{C12}-\text{H12B}\cdots\text{CgA}^{ii}$	0.96	2.82	3.6534 (19)	145

Symmetry codes: (i) $-x + 1, -y, -z - 1$; (ii) $x, y, z + 1$. CgA is the centroid of the C1-C6 ring.

Data collection: APEX2 (Bruker, 2007); cell refinement: APEX2; data reduction: SAINT (Bruker, 2007); program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: ORTEP-3 for Windows (Farrugia, 1997) and PLATON (Spek, 2003); software used to prepare material for publication: WinGX (Farrugia, 1999) and PLATON.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: AT2588).

References

- Bernstein, J., Davis, R. E., Shimoni, L. & Chang, N.-L. (1995). *Angew. Chem. Int. Ed. Engl.* **34**, 1555–1573.
Bruker (2005). SADABS. Bruker AXS Inc. Madison, Wisconsin, USA.
Bruker (2007). APEX2 and SAINT. Bruker AXS Inc. Madison, Wisconsin, USA.
Burt, S. (2004). *Int. J. Food Microbiol.* **94**, 223–253.
Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.
Farrugia, L. J. (1999). *J. Appl. Cryst.* **32**, 837–838.
Muhammad, N., Ali, S., Tahir, M. N. & Zia-ur-Rehman (2008a). *Acta Cryst.* **E64**, o1373.
Muhammad, N., Tahir, M. N., Ali, S. & Zia-ur-Rehman (2008b). *Acta Cryst.* **E64**, m946–m947.
Muhammad, N., Tahir, M. N., Ali, S. & Zia-ur-Rehman (2008c). *Acta Cryst.* **E64**, m978.
Muhammad, N., Zia-ur-Rehman, Ali, S. & Meetsma, A. (2007). *Acta Cryst.* **E63**, o2174–o2175.
Niaz, M., Tahir, M. N., Zia-ur-Rehman, Ali, S. & Khan, I. U. (2008). *Acta Cryst.* **E64**, o733.
Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
Spek, A. L. (2003). *J. Appl. Cryst.* **36**, 7–13.

supporting information

Acta Cryst. (2008). E64, o1458 [doi:10.1107/S160053680802103X]

2-(4-Ethoxybenzylidene)butanoic acid

Niaz Muhammad, M. Nawaz Tahir, Zia-ur-Rehman and Saqib Ali

S1. Comment

Cinnamic acid derivatives play an important role in the production of lignins for higher plants. They are well documented for their antibacterial, antifungal and antiparasitic activities (Burt, 2004). We have recently reported the various derivatives of cinamic acids (Niaz *et al.*, 2008, Muhammad *et al.*, 2008*a*) and tin complexes (Muhammad *et al.*, 2008*b*, 2008*c*). In continuation of our project, we herein report the structure of the title compound (I).

The crystal structure of 3-(4-bromophenyl)-2-ethylacrylic acid (II) (Muhammad *et al.*, 2007) differs from (I) due to the attachment of ethoxy group at *para* position instead of Br-atom. Similarly, (E)-2-(2-fluorobenzylidene)butanoic acid (III) (Niaz *et al.*, 2008) differs as the attachment of F-atom is on *meta* position. Thus, (I) is very different from the reported compounds (II) and (III).

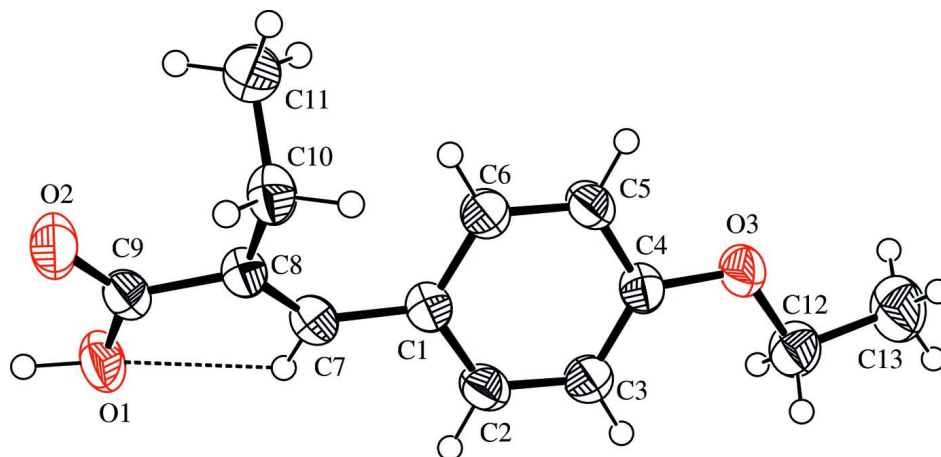
In the crystal structure of the title compound (I), the C—C bonds are in the range 1.464 (2) to 1.511 (2) Å, and C=C have a value of 1.337 (2) Å. The value of C—O bond in the carboxylate group is 1.287 (2) Å, whereas C=O is of 1.248 (2) Å. The attached ethoxy group have 1.3595 (19) and 1.431 (2) Å as the C—O bond distances. In the asymmetric unit, there is an intermolecular H-bond of C—H \cdots O type (Table 1, Fig 1) which completes a five membered ring (O1/C9/C8/C7/H7). Centrosymmetric dimers, $R_2^2(8)$ (Bernstein *et al.* 1995) are formed due to the intermolecular hydrogen bonding, O1—H1 \cdots O2ⁱ [symmetry code: (i) = -x + 1, -y, -z - 1]. These dimers have a π -interaction C12—H12B \cdots CgAⁱⁱ [symmetry code: (ii) = x, y, z + 1] where CgA is the centroid of the aromatic ring A(C1—C6). The dihedral angle between (O1/C9/O2) and (C8/C10/C11) is 75.55 (15)°, whereas it is 30.19 (17)° between (C7/C8/C9) and (O3/C12/C13). The ethoxy group makes a dihedral angle of 2.77 (11)° with the aromatic ring. The formation of centrosymmetric dimers and packing of (I) is shown in Fig 2.

S2. Experimental

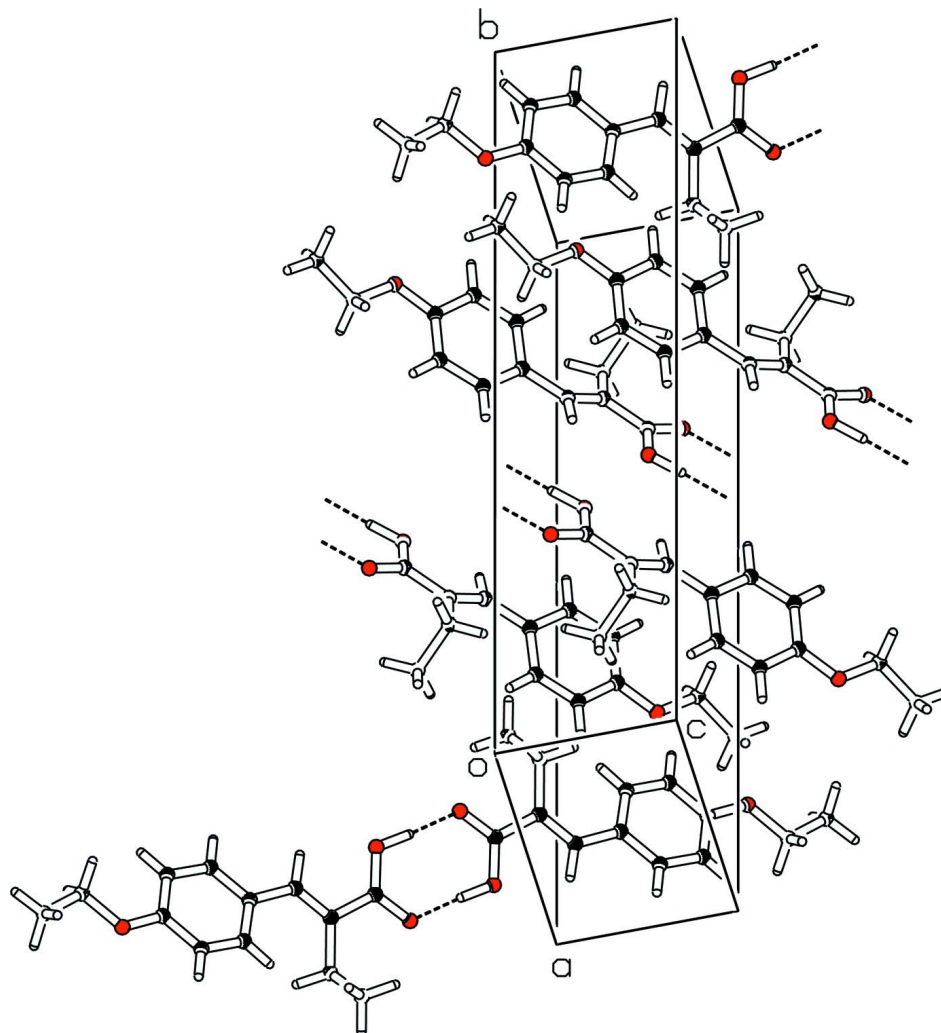
Compound (I) was prepared according to the reported procedure (Muhammad *et al.*, 2007). A mixture of 4-ethoxybenzaldehyde (1.39 ml, 10 mmol), ethylmalonic acid (2.64 g, 20 mmol) and piperidine (1.98 ml, 20 mmol) in a pyridine (12.5 ml) solution was heated on a steam-bath for 24 h. The reaction mixture was cooled and added to a mixture of 25 ml of concentrated HCl and 50 g of ice. The precipitate formed in the acidified mixture was filtered off and washed with ice-cold water. The product was recrystallized from ethanol [yield: 66%].

S3. Refinement

The H-atom of carboxylate group is taken from fourier difference map and coordinates were refined. The other H atoms were positioned geometrically, with C—H = 0.93, 0.97, and 0.96 Å for aromatic, methylene and methyl H, respectively, and constrained to ride on their parent atoms, with $U_{iso}(H) = xU_{eq}(C, O)$, where $x = 1.5$ for methyl H, and $x = 1.2$ for all other H atoms.

**Figure 1**

ORTEP drawing of the title compound, with the atom numbering scheme. The thermal ellipsoids are drawn at the 50% probability level. H-atoms are shown by small circles of arbitrary radii. The intramolecular H-bonds are shown by dotted lines.

**Figure 2**

The packing figure (*PLATON*: Spek, 2003) which shows the dimeric nature of the compound and the interlinkages of the dimers.

2-(4-Ethoxybenzylidene)butanoic acid

Crystal data

$C_{13}H_{16}O_3$

$M_r = 220.26$

Monoclinic, $P2_1/c$

Hall symbol: $-P 2_1bc$

$a = 10.3192 (7) \text{ \AA}$

$b = 22.0761 (15) \text{ \AA}$

$c = 5.2362 (3) \text{ \AA}$

$\beta = 100.751 (4)^\circ$

$V = 1171.91 (13) \text{ \AA}^3$

$Z = 4$

$F(000) = 472$

$D_x = 1.248 \text{ Mg m}^{-3}$

Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$

Cell parameters from 3009 reflections

$\theta = 2.0\text{--}28.7^\circ$

$\mu = 0.09 \text{ mm}^{-1}$

$T = 296 \text{ K}$

Prismatic, colourless

$0.26 \times 0.20 \times 0.16 \text{ mm}$

Data collection

Bruker KappaAPEXII CCD
 diffractometer
 Radiation source: fine-focus sealed tube
 Graphite monochromator
 Detector resolution: 7.4 pixels mm⁻¹
 ω scans
 Absorption correction: multi-scan
 (SADABS; Bruker, 2005)
 $T_{\min} = 0.975$, $T_{\max} = 0.985$

13990 measured reflections
 3009 independent reflections
 1599 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.049$
 $\theta_{\max} = 28.7^\circ$, $\theta_{\min} = 2.0^\circ$
 $h = -13 \rightarrow 13$
 $k = -29 \rightarrow 29$
 $l = -7 \rightarrow 6$

Refinement

Refinement on F^2
 Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.050$
 $wR(F^2) = 0.122$
 $S = 1.01$
 3009 reflections
 150 parameters
 0 restraints
 Primary atom site location: structure-invariant
 direct methods

Secondary atom site location: difference Fourier
 map
 Hydrogen site location: inferred from
 neighbouring sites
 H atoms treated by a mixture of independent
 and constrained refinement
 $w = 1/[\sigma^2(F_o^2) + (0.0465P)^2 + 0.113P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} < 0.001$
 $\Delta\rho_{\max} = 0.15 \text{ e } \text{\AA}^{-3}$
 $\Delta\rho_{\min} = -0.21 \text{ e } \text{\AA}^{-3}$

Special details

Geometry. Bond distances, angles *etc.* have been calculated using the rounded fractional coordinates. All su's are estimated from the variances of the (full) variance-covariance matrix. The cell e.s.d.'s are taken into account in the estimation of distances, angles and torsion angles

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
O1	0.62469 (12)	-0.00644 (5)	-0.2194 (2)	0.0545 (4)
O2	0.47639 (13)	0.06481 (5)	-0.3546 (2)	0.0577 (5)
O3	1.08170 (11)	0.17222 (5)	1.0247 (2)	0.0506 (4)
C1	0.82240 (15)	0.09272 (7)	0.4115 (3)	0.0388 (5)
C2	0.89602 (16)	0.05704 (7)	0.6035 (3)	0.0437 (6)
C3	0.98214 (17)	0.08106 (7)	0.8122 (3)	0.0447 (6)
C4	0.99847 (16)	0.14314 (7)	0.8316 (3)	0.0396 (5)
C5	0.92792 (17)	0.17994 (7)	0.6416 (3)	0.0457 (6)
C6	0.84164 (17)	0.15537 (7)	0.4371 (3)	0.0461 (6)
C7	0.73819 (16)	0.06338 (7)	0.1901 (3)	0.0420 (6)
C8	0.63076 (16)	0.08379 (7)	0.0301 (3)	0.0385 (5)
C9	0.57279 (17)	0.04572 (7)	-0.1946 (3)	0.0411 (5)
C10	0.55968 (18)	0.14224 (7)	0.0567 (3)	0.0456 (6)
C11	0.5866 (2)	0.19159 (8)	-0.1265 (3)	0.0613 (7)
C12	1.16254 (18)	0.13667 (8)	1.2209 (3)	0.0523 (6)
C13	1.2471 (2)	0.17960 (10)	1.4013 (4)	0.0696 (8)

H1	0.5832 (17)	-0.0279 (8)	-0.385 (4)	0.0654*
H2	0.88696	0.01518	0.59100	0.0525*
H3	1.02879	0.05571	0.93877	0.0537*
H5	0.93908	0.22172	0.65263	0.0548*
H6	0.79466	0.18098	0.31205	0.0553*
H7	0.76348	0.02428	0.15510	0.0503*
H10A	0.58522	0.15670	0.23393	0.0547*
H10B	0.46558	0.13434	0.02599	0.0547*
H11A	0.56467	0.17725	-0.30240	0.0919*
H11B	0.67825	0.20238	-0.08728	0.0919*
H11C	0.53385	0.22647	-0.10632	0.0919*
H12A	1.21722	0.10923	1.14239	0.0627*
H12B	1.10791	0.11292	1.31526	0.0627*
H13A	1.30278	0.20185	1.30692	0.1044*
H13B	1.30068	0.15718	1.53886	0.1044*
H13C	1.19191	0.20730	1.47359	0.1044*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
O1	0.0617 (9)	0.0404 (7)	0.0544 (7)	0.0025 (6)	-0.0076 (6)	-0.0162 (5)
O2	0.0623 (9)	0.0493 (8)	0.0515 (7)	0.0067 (6)	-0.0149 (7)	-0.0126 (6)
O3	0.0561 (8)	0.0421 (7)	0.0456 (7)	-0.0019 (6)	-0.0109 (6)	-0.0018 (5)
C1	0.0377 (10)	0.0378 (9)	0.0402 (9)	-0.0003 (7)	0.0053 (8)	-0.0036 (7)
C2	0.0484 (11)	0.0316 (9)	0.0501 (10)	-0.0018 (8)	0.0066 (9)	-0.0015 (7)
C3	0.0477 (11)	0.0397 (10)	0.0434 (10)	0.0012 (8)	-0.0002 (8)	0.0037 (7)
C4	0.0391 (10)	0.0395 (9)	0.0387 (9)	-0.0009 (8)	0.0034 (8)	-0.0049 (7)
C5	0.0504 (11)	0.0312 (9)	0.0507 (10)	-0.0010 (8)	-0.0027 (9)	-0.0027 (8)
C6	0.0481 (11)	0.0380 (9)	0.0470 (9)	0.0008 (8)	-0.0044 (8)	0.0016 (7)
C7	0.0478 (11)	0.0347 (9)	0.0430 (9)	-0.0017 (8)	0.0075 (8)	-0.0048 (7)
C8	0.0414 (10)	0.0367 (9)	0.0365 (9)	-0.0037 (8)	0.0053 (8)	-0.0042 (7)
C9	0.0444 (10)	0.0344 (9)	0.0433 (9)	-0.0038 (8)	0.0048 (8)	-0.0047 (7)
C10	0.0460 (11)	0.0468 (10)	0.0425 (9)	0.0009 (8)	0.0041 (8)	-0.0099 (8)
C11	0.0821 (15)	0.0456 (11)	0.0513 (11)	0.0044 (10)	-0.0003 (10)	-0.0010 (8)
C12	0.0490 (12)	0.0516 (11)	0.0507 (10)	0.0063 (9)	-0.0054 (9)	0.0003 (8)
C13	0.0603 (14)	0.0705 (14)	0.0663 (13)	0.0025 (11)	-0.0183 (11)	-0.0055 (11)

Geometric parameters (Å, °)

O1—C9	1.287 (2)	C12—C13	1.498 (3)
O2—C9	1.248 (2)	C2—H2	0.9300
O3—C4	1.3595 (19)	C3—H3	0.9300
O3—C12	1.431 (2)	C5—H5	0.9300
O1—H1	1.01 (2)	C6—H6	0.9300
C1—C2	1.387 (2)	C7—H7	0.9300
C1—C6	1.400 (2)	C10—H10A	0.9700
C1—C7	1.464 (2)	C10—H10B	0.9700
C2—C3	1.379 (2)	C11—H11A	0.9600

C3—C4	1.382 (2)	C11—H11B	0.9600
C4—C5	1.382 (2)	C11—H11C	0.9600
C5—C6	1.370 (2)	C12—H12A	0.9700
C7—C8	1.337 (2)	C12—H12B	0.9700
C8—C9	1.479 (2)	C13—H13A	0.9600
C8—C10	1.504 (2)	C13—H13B	0.9600
C10—C11	1.511 (2)	C13—H13C	0.9600
C4—O3—C12	118.54 (12)	C1—C6—H6	119.00
C9—O1—H1	112.9 (10)	C5—C6—H6	119.00
C2—C1—C6	116.29 (14)	C1—C7—H7	115.00
C6—C1—C7	124.44 (14)	C8—C7—H7	115.00
C2—C1—C7	119.13 (14)	C8—C10—H10A	109.00
C1—C2—C3	122.74 (14)	C8—C10—H10B	109.00
C2—C3—C4	119.41 (15)	C11—C10—H10A	109.00
O3—C4—C5	115.64 (14)	C11—C10—H10B	109.00
C3—C4—C5	119.30 (15)	H10A—C10—H10B	108.00
O3—C4—C3	125.05 (14)	C10—C11—H11A	109.00
C4—C5—C6	120.54 (14)	C10—C11—H11B	109.00
C1—C6—C5	121.71 (15)	C10—C11—H11C	109.00
C1—C7—C8	130.35 (15)	H11A—C11—H11B	109.00
C7—C8—C9	118.18 (14)	H11A—C11—H11C	109.00
C9—C8—C10	115.44 (14)	H11B—C11—H11C	109.00
C7—C8—C10	126.37 (14)	O3—C12—H12A	110.00
O1—C9—O2	121.94 (14)	O3—C12—H12B	110.00
O1—C9—C8	118.35 (14)	C13—C12—H12A	110.00
O2—C9—C8	119.72 (14)	C13—C12—H12B	110.00
C8—C10—C11	114.16 (15)	H12A—C12—H12B	109.00
O3—C12—C13	107.38 (15)	C12—C13—H13A	109.00
C1—C2—H2	119.00	C12—C13—H13B	109.00
C3—C2—H2	119.00	C12—C13—H13C	109.00
C2—C3—H3	120.00	H13A—C13—H13B	110.00
C4—C3—H3	120.00	H13A—C13—H13C	109.00
C4—C5—H5	120.00	H13B—C13—H13C	109.00
C6—C5—H5	120.00		
C12—O3—C4—C3	2.0 (2)	O3—C4—C5—C6	179.67 (15)
C12—O3—C4—C5	-176.99 (14)	C3—C4—C5—C6	0.6 (3)
C4—O3—C12—C13	177.96 (14)	C4—C5—C6—C1	-0.5 (3)
C6—C1—C2—C3	1.2 (2)	C1—C7—C8—C9	174.19 (16)
C7—C1—C2—C3	177.04 (16)	C1—C7—C8—C10	-6.8 (3)
C2—C1—C6—C5	-0.4 (2)	C7—C8—C9—O1	3.1 (2)
C7—C1—C6—C5	-176.02 (16)	C7—C8—C9—O2	-176.51 (16)
C2—C1—C7—C8	156.09 (18)	C10—C8—C9—O1	-176.06 (14)
C6—C1—C7—C8	-28.4 (3)	C10—C8—C9—O2	4.4 (2)
C1—C2—C3—C4	-1.1 (3)	C7—C8—C10—C11	103.29 (19)
C2—C3—C4—O3	-178.82 (15)	C9—C8—C10—C11	-77.65 (19)
C2—C3—C4—C5	0.1 (2)		

Hydrogen-bond geometry (Å, °)

<i>D</i> —H \cdots <i>A</i>	<i>D</i> —H	H \cdots <i>A</i>	<i>D</i> \cdots <i>A</i>	<i>D</i> —H \cdots <i>A</i>
O1—H1 \cdots O2 ⁱ	1.01 (2)	1.61 (2)	2.6184 (15)	176.8 (15)
C7—H7 \cdots O1	0.93	2.30	2.7218 (19)	107
C12—H12 <i>B</i> \cdots CgA ⁱⁱ	0.96	2.82	3.6534 (19)	145

Symmetry codes: (i) $-x+1, -y, -z-1$; (ii) $x, y, z+1$.