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Safflorite, (Co,Ni,Fe)As₂, isomorphous with marcasite

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Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{As}-\text{As}) = 0.001$ Å; some non-H atoms missing; R factor = 0.026; wR factor = 0.061; data-to-parameter ratio = 19.5.

Safflorite, a naturally occurring cobalt-nickel-iron diarsenide (Co,Ni,Fe)As₂, possesses the marcasite-type structure, with cations ($M = \text{Co} + \text{Ni} + \text{Fe}$) at site symmetry $2/m$ and As anions at m . The $M\text{As}_6$ octahedra share two edges, forming chains parallel to c . The chemical formula for safflorite should be expressed as (Co,Ni,Fe)As₂, rather than the end-member format CoAs₂, as its structure stabilization requires the simultaneous interaction of the electronic states of Co, Ni, and Fe with As₂²⁻ dianions.

Related literature

For related literature, see: Anawar *et al.* (2003); Carlon & Bleeker (1988); Darmon & Wintenberger (1966); Ennaciri *et al.* (1995); Goodenough (1967); Grorud (1997); Hem *et al.* (2001); Holmes (1947); King (2002); Kjekshus (1971); Kjekshus *et al.* (1974, 1979); Lutz *et al.* (1987); Makovicky (2006); O'Day (2006); Ondrus *et al.* (2001); Palenik *et al.* (2004); Petruk *et al.* (1971); Radcliffe & Berry (1968, 1971); Reich *et al.* (2005); Robinson *et al.* (1971); Swanson *et al.* (1966); Tossell (1984); Tossell *et al.* (1981); Vaughan & Rosso (2006); Wagner & Lorenz (2002).

Experimental

Crystal data

As_{1.99}Co_{0.61}Fe_{0.17}Ni_{0.22}S_{0.01}
 $M_r = 207.77$ Orthorhombic, $Pnmm$ $a = 5.0669$ (6) Å $b = 5.8739$ (7) Å $c = 3.1346$ (4) Å $V = 93.29$ (2) Å³ $Z = 2$ Mo $K\alpha$ radiation $\mu = 43.75$ mm⁻¹ $T = 293$ (2) K $0.06 \times 0.05 \times 0.04$ mm

Data collection

Bruker APEX2 CCD area-detector diffractometer

Absorption correction: multi-scan (TWINABS; Sheldrick, 2007)

 $T_{\min} = 0.179$, $T_{\max} = 0.274$

(expected range = 0.114–0.174)

1232 measured reflections

254 independent reflections

227 reflections with $I > 2\sigma(I)$ $R_{\text{int}} = 0.041$

Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.025$ $wR(F^2) = 0.061$ $S = 0.91$

254 reflections

13 parameters

 $\Delta\rho_{\text{max}} = 1.44$ e Å⁻³ $\Delta\rho_{\text{min}} = -1.82$ e Å⁻³

Data collection: APEX2 (Bruker, 2003); cell refinement: SAINT (Bruker, 2005); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: Xtal-Draw (Downs & Hall-Wallace, 2003); software used to prepare material for publication: SHELXTL (Sheldrick, 2008).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: MG2054).

References

- Anawar, H. M., Akai, J., Komaki, K., Terao, H., Yoshioka, T., Ishizuka, T., Safiullah, S. & Kato, K. (2003). *J. Geochem. Expl.* **77**, 109–131.
- Bruker (2003). SMART. Bruker AXS Inc., Madison, Wisconsin, USA.
- Bruker (2005). SAINT. Bruker AXS Inc., Madison, Wisconsin, USA.
- Carlon, C. J. & Bleeker, W. (1988). *Geol. Mijnbouw* **67**, 279–292.
- Darmon, R. & Wintenberger, M. (1966). *Bull. Soc. Fr. Minér. Cristal.* **89**, 213–215.
- Downs, R. T. & Hall-Wallace, M. (2003). *Am. Mineral.* **88**, 247–250.
- Ennaciri, A., Barbanson, L. & Touray, J. C. (1995). *Mineralium Deposita* **30**, 75–77.
- Goodenough, J. B. (1967). *Solid State Commun.* **5**, 577–580.
- Grorud, H. F. (1997). *Nor. Geol. Tidsskr.* **77**, 31–38.
- Hem, S. R., Makovicky, E. & Gervilla, F. (2001). *Can. Mineral.* **39**, 831–853.
- Holmes, R. J. (1947). *Geol. Soc. Amer. Bull.* **58**, 299–392.
- King, R. J. (2002). *Geol. Today*, **18**, 72–75.
- Kjekshus, A. (1971). *Acta Chem. Scand.* **25**, 411–422.
- Kjekshus, A., Peterzens, P. G., Rakke, T. & Andresen, A. F. (1979). *Acta Chem. Scand.* **A33**, 469–480.
- Kjekshus, A., Rakke, T. & Andresen, A. F. (1974). *Acta Chem. Scand.* **28**, 996–1000.
- Lutz, H. D., Jung, M. & Waschenbach, G. (1987). *Z. Anorg. Allg. Chem.* **554**, 87–91.
- Makovicky, E. (2006). *Rev. Miner. Geochem.* **61**, 7–125.
- O'Day, P. A. (2006). *Elements*, **2**, 77–83.
- Ondrus, P., Vavrin, I., Skala, R. & Veselovsky, F. (2001). *N. Jahrb. Miner. Monatsh.* **2001**, 169–185.
- Palenik, C. S., Utsunomiya, S., Reich, M., Kesler, S. E. & Ewing, R. C. (2004). *Am. Mineral.* **89**, 1359–1366.
- Petruk, W., Harris, D. C. & Stewart, J. M. (1971). *Can. Mineral.* **11**, 150–186.
- Radcliffe, D. & Berry, L. G. (1968). *Am. Mineral.* **53**, 1856–1881.
- Radcliffe, D. & Berry, L. G. (1971). *Can. Mineral.* **10**, 877–881.
- Reich, M., Kesler, S. E., Utsunomiya, S., Palenik, C. S., Chryssoulis, S. & Ewing, R. C. (2005). *Geochim. Cosmochim. Acta*, **69**, 2781–2796.
- Robinson, K., Gibbs, G. V. & Ribbe, P. H. (1971). *Science*, **172**, 567–570.
- Sheldrick, G. M. (2007). TWINABS. University of Göttingen, Germany.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Swanson, H. E., Morris, M. C. & Evans, E. H. (1966). *Natl Bur. Stand. Monogr.* **25**, 10–11.
- Tossell, J. A. (1984). *Phys. Chem. Miner.* **11**, 75–80.
- Tossell, J. A., Vaughan, D. J. & Burdett, J. K. (1981). *Phys. Chem. Miner.* **7**, 177–184.
- Vaughan, D. J. & Rosso, K. M. (2006). *Rev. Miner. Geochem.* **61**, 231–264.
- Wagner, T. & Lorenz, J. (2002). *Mineral. Mag.* **66**, 385–403.

supporting information

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Safflorite, (Co,Ni,Fe)As₂, isomorphous with marcasite**Hexiong Yang, Robert T. Downs and Carla Eichler****S1. Comment**

Minerals in the FeAs₂—NiAs₂—CoAs₂ system include löllingite FeAs₂, rammelsbergite NiAs₂, pararammelsbergite NiAs₂, clinosafflorite CoAs₂, and safflorite CoAs₂. These diarsenide minerals, together with Fe—Ni—Co disulfides and sulfarsenides, are commonly found in complex Co—Ni—As ore deposits, such as Håkansboda, Sweden (Carlson & Bleeker, 1988), the Cobalt District, Ontario (Petruk *et al.*, 1971), Bou-Azzer, Morocco (Ennaciri *et al.*, 1995), Modum, Norway (Grorud, 1997), and Spessart, Germany (Wagner & Lorenz, 2002). When precipitated from hydrothermal solutions, these minerals can incorporate considerable amounts of trace metals, especially so-called "invisible" gold (*e.g.*, Palenik *et al.*, 2004; Reich *et al.*, 2005). Under oxidizing conditions, however, they can release significant amounts of arsenic into natural water and soils, in some cases producing serious arsenic poisoning and contamination (King, 2002; Anawar *et al.*, 2003; O'Day, 2006). Therefore, the crystal structures and bonding models of Fe—Ni—Co disulfides, diarsenides, and sulfarsenides have been a subject of extensive experimental and theoretical studies (Vaughan & Rosso, 2006; Makovicky, 2006, and references therein)

The crystal structures of all minerals, except safflorite, in the FeAs₂—NiAs₂—CoAs₂ system have been determined. Topologically, löllingite FeAs₂ (Kjekshus *et al.*, 1979; Lutz *et al.*, 1987; Ondrus *et al.*, 2001) and rammelsbergite NiAs₂ (Kjekshus *et al.*, 1974, 1979) possess the marcasite (FeS₂)-type structure with space group *Pnmm*, whereas clinosafflorite CoAs₂ (Darmon & Wintenberger, 1966; Kjekshus, 1971) is isostructural with the modified marcasite-type structure of arsenopyrite (FeAsS) with space group *P2₁/c* (Hem *et al.*, 2001; Makovicky, 2006). From the unit-cell dimensions measured from X-ray diffraction, safflorite was assumed to be isotypic with marcasite (Holmes, 1947; Radcliffe & Berry, 1968, 1971). Chemical analyses of various natural and synthetic samples reveal that *Pnmm* safflorite always contains some amounts of Fe and Ni, whereas materials with 80–100% (mole) CoAs₂ crystallize in monoclinic *P2₁/c* symmetry (Holmes, 1947; Swanson *et al.*, 1966; Radcliffe & Berry, 1971). This study presents the first structure determination of safflorite based on single-crystal X-ray diffraction data.

Safflorite is isomorphous with marcasite. Each cation ($M = \text{Co, Ni, and Fe}$) at site symmetry $2/m$ is octahedrally coordinated by six anions (As) at site symmetry m and each anion is tetrahedrally bonded to another anion (forming As—As dianion units) plus three M cations. The $M\text{As}_6$ octahedra share two edges, forming chains parallel to c , and two vertices with adjacent chains (Fig. 1). The average M —As bond distance (2.360 Å) is identical to that in clinosafflorite (Kjekshus, 1971), but slightly shorter than that in löllingite (2.379 Å) (Kjekshus *et al.*, 1979; Lutz *et al.*, 1987) or rammelsbergite (2.378 Å) (Kjekshus *et al.*, 1979). Notably, as the d -orbital electrons in M cations increase from Fe ($d = 6$) in löllingite to Co ($d = 7$) in safflorite, and Ni ($d = 8$) in rammelsbergite, the M — M separation along the chain direction increases significantly from 2.882 to 3.134, and 3.545 Å, respectively, while the As—As edge length shared by the two M octahedra concomitantly decreases from 3.808 to 3.547, and 3.219 Å. The octahedral distortion, measured by the octahedral angle variance (OAV) and quadratic elongation (OQE) (Robinson *et al.*, 1971), decreases. The OAV and OQE values are 92.87 and 1.0265 for FeAs₆, 21.04 and 1.0058 for CoAs₆, and 16.22 and 1.0049 for NiAs₆.

The variation of the M—M separation with the number of *d*-orbital electrons in marcasite-type disulfides, diarsenides, and sulfarsenides has been a matter of discussion (see Vaughan & Rosso, 2006 for a thorough review). Theoretical calculations based on molecular orbital and band models predict that, due to the interaction between the $3d\sigma(e_g)$ orbitals of M^{2+} and the π_b orbitals of As_2^{2-} , the M—As—M angle subtending the M—M separation across the shared octahedral edge should be substantially smaller for $FeAs_2$ than for $CoAs_2$ and $NiAs_2$, resulting in the so-called "compressed marcasite-type" structure (Tossell *et al.*, 1981; Tossell, 1984). Indeed, this angle is 74° in $FeAs_2$ löllingite (Lutz *et al.*, 1987), but 83° in $(Co,Ni,Fe)As_2$ safflorite and 96° in rammelsbergite (Kjekshus *et al.*, 1979). It is intriguing to note that the end-member $CoAs_2$ has been found to only crystallize in the arsenopyrite-type structure ($P2_1/c$) (Holmes, 1947; Swanson *et al.*, 1966; Radcliffe & Berry, 1971), rather than the marcasite-type structure ($Pnmm$). This observation may be explained by the existence of an unpaired electron occupying one of the π_b orbitals, which splits into a lower-energy filled band and a higher-energy empty band (Goodenough, 1967), thus resulting in the symmetry reduction from $Pnmm$ to $P2_1/c$. In other words, the presence of some Ni/Fe in place of Co appears to be an essential requirement for the $CoAs_2$ system to crystallize in the $Pnmm$ symmetry. The pure system will otherwise be stabilized energetically in the clinosafflorite structure.

Another outstanding feature of the safflorite structure is the prominent anisotropic displacement ellipsoid of the *M* cation, the $U_{11}:U_{22}:U_{33}$ ratio being approximately 3:1:9, with the ellipsoid axial directions roughly parallel to the unit cell axes. This ratio can be compared to the differences of three unit-cell dimensions between $FeAs_2$ löllingite and $NiAs_2$ rammelsbergite [$(a_{Lol} - a_{Ram})/a_{Lol} : (b_{Lol} - b_{Ram})/b_{Lol} : (c_{Lol} - c_{Ram})/c_{Lol}$] (Kjekshus *et al.*, 1974, 1979; Lutz *et al.*, 1987), which is about 3:1:8. Accordingly, the marked anisotropy of the displacement parameters of the *M* cation in safflorite is interpreted as a consequence of positional disorder with Fe and Ni occupying apparent different positions, which in turn results from the different interactions of their *d*-electrons with the As_2^{2-} dianions.

S2. Experimental

The safflorite specimen used in this study is from Timiskaming County, Ontario, Canada, and is in the collection of the RRUFF project (deposition No. R070611; <http://rruff.info>), donated by James Shigley. The average chemical composition (15 point analyses), $(Co_{0.61}Ni_{0.22}Fe_{0.17})_{\Sigma=1}(As_{1.99}S_{0.01})_{\Sigma=2}$, was determined with a CAMECA SX50 electron microprobe (<http://rruff.info>).

S3. Refinement

Due to the similar X-ray scattering powers for Co, Ni, and Fe, all cations were assumed to be Co and their site occupancies were not determined during the refinement. All crystals examined were twinned, with {011} as twin plane. The structure refinements were performed based on X-ray diffraction data collected from a twinned crystal, which were processed with *TWINABS* (Sheldrick, 2007). The ratio of two twin components is 0.73:0.27. The highest residual peak in the difference Fourier maps was located at (0.133, 0.370, 0.256), 0.85 Å from atom As, and the deepest hole at (0.133, 0.473, 0), 0.69 Å from As.

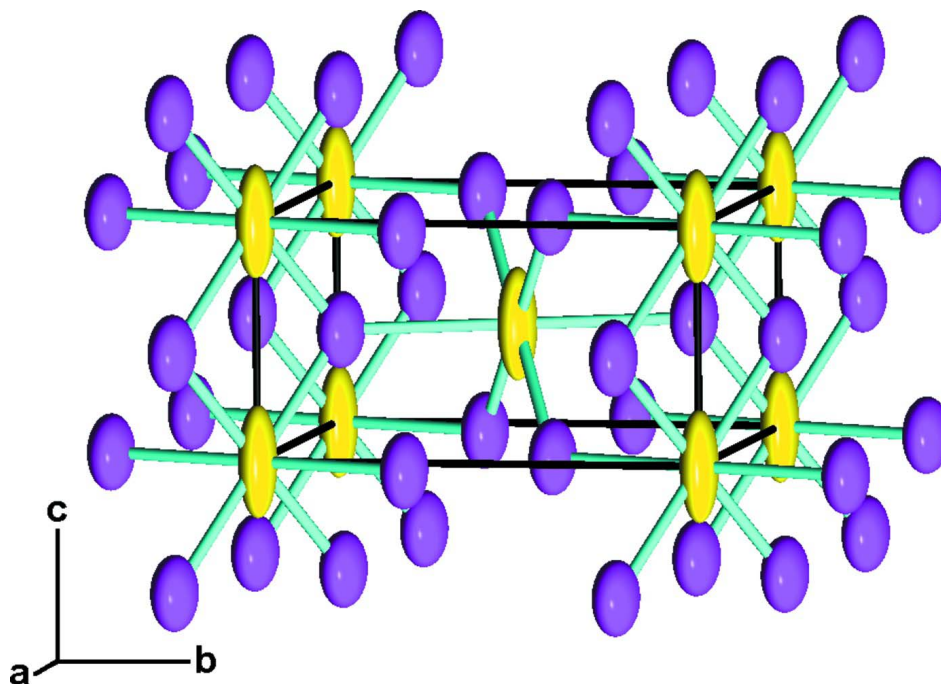


Figure 1

Crystal structure of safflorite, with displacement ellipsoids drawn at the 99.9% probability level. The M ($=\text{Co}+\text{Ni}+\text{Fe}$) cations (yellow spheres) are situated in octahedra coordinated by six As atoms (pink spheres).

Cobalt-iron-nickel diarsenide

Crystal data

$\text{As}_{1.99}\text{Co}_{0.61}\text{Fe}_{0.17}\text{Ni}_{0.22}\text{S}_{0.01}$

$M_r = 207.77$

Orthorhombic, $Pnmm$

Hall symbol: $-P\ 2\ 2n$

$a = 5.0669$ (6) Å

$b = 5.8739$ (7) Å

$c = 3.1346$ (4) Å

$V = 93.29$ (2) Å³

$Z = 2$

$F(000) = 168$

$D_x = 7.396$ Mg m⁻³

Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 153 reflections

$\theta = 5.0\text{--}31.4^\circ$

$\mu = 43.75$ mm⁻¹

$T = 293$ K

Granular, black

$0.06 \times 0.05 \times 0.04$ mm

Data collection

Bruker APEX2 CCD area-detector
diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

φ and ω scan

Absorption correction: multi-scan
(TWINABS; Sheldrick, 2007)

$T_{\min} = 0.179$, $T_{\max} = 0.274$

1232 measured reflections

254 independent reflections

227 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.041$

$\theta_{\max} = 36.4^\circ$, $\theta_{\min} = 5.3^\circ$

$h = -8 \rightarrow 7$

$k = -8 \rightarrow 9$

$l = -5 \rightarrow 5$

Refinement

Refinement on F^2
 Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.025$
 $wR(F^2) = 0.061$
 $S = 0.91$
 254 reflections
 13 parameters
 0 restraints
 Primary atom site location: structure-invariant
 direct methods

Secondary atom site location: difference Fourier
 map
 $w = 1/[\sigma^2(F_o^2) + (0.0428P)^2]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} < 0.001$
 $\Delta\rho_{\max} = 1.44 \text{ e } \text{Å}^{-3}$
 $\Delta\rho_{\min} = -1.82 \text{ e } \text{Å}^{-3}$
 Extinction correction: *SHELXL*,
 $F_c^* = kF_c[1 + 0.001x F_c^2 \lambda^3 / \sin(2\theta)]^{-1/4}$
 Extinction coefficient: 0.016 (6)

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) etc. and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
M	0.0000	0.0000	0.0000	0.0130 (2)
As	0.18637 (9)	0.36589 (7)	0.0000	0.01033 (17)

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
M	0.0090 (4)	0.0033 (4)	0.0267 (4)	0.0001 (3)	0.000	0.000
As	0.0140 (3)	0.0046 (2)	0.0124 (2)	-0.00002 (14)	0.000	0.000

Geometric parameters (Å , $^\circ$)

M—As	2.3475 (5)	M—As ⁱⁱⁱ	2.3669 (4)
M—As ⁱ	2.3475 (5)	M—As ^{iv}	2.3669 (4)
M—As ⁱⁱ	2.3669 (4)	M—As ^v	2.3669 (4)
As—M—As ⁱⁱ	88.016 (9)	M—As—M ^{vi}	125.085 (13)
As ⁱ —M—As ⁱⁱ	91.984 (9)	M ^{vi} —As—M ^{vii}	82.931 (17)
As ⁱⁱ —M—As ^{iv}	82.931 (17)	M—As—As ^{viii}	106.12 (3)
As ⁱⁱⁱ —M—As ^{iv}	97.069 (17)	M ^{vi} —As—As ^{viii}	107.599 (2)

Symmetry codes: (i) $-x, -y, -z$; (ii) $x-1/2, -y+1/2, z-1/2$; (iii) $-x+1/2, y-1/2, -z+1/2$; (iv) $x-1/2, -y+1/2, z+1/2$; (v) $-x+1/2, y-1/2, -z-1/2$; (vi) $-x+1/2, y+1/2, -z+1/2$; (vii) $-x+1/2, y+1/2, -z-1/2$; (viii) $-x, -y+1, -z$.