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## (E)-3-(3,5-Dimethoxyphenyl)acrylohydrazone

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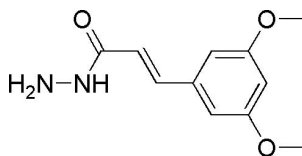
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 Key indicators: single-crystal X-ray study;  $T = 150$  K; mean  $\sigma(\text{C}-\text{C}) = 0.007$  Å;  $R$  factor = 0.111;  $wR$  factor = 0.274; data-to-parameter ratio = 17.4.

In the title compound,  $\text{C}_{11}\text{H}_{14}\text{N}_2\text{O}_3$ , the planar hydrazone group is oriented with respect to the benzene ring at a dihedral angle of  $48.00(3)^\circ$ . In the crystal structure, intermolecular  $\text{N}-\text{H}\cdots\text{O}$  hydrogen bonds link the molecules.

### Related literature

For related literature, see: Zheng *et al.* (2003); Al-Talib *et al.* (1990); Yousif *et al.* (1986); Ahmad *et al.* (2001); Al-Soud *et al.* (2004); El-Emam *et al.* (2004); Furniss *et al.* (1978). For bond-length data, see: Allen *et al.* (1987).



### Experimental

#### Crystal data

$\text{C}_{11}\text{H}_{14}\text{N}_2\text{O}_3$   
 $M_r = 222.24$   
 Monoclinic,  $P2_1/c$   
 $a = 4.8910(19)$  Å  
 $b = 30.358(11)$  Å  
 $c = 8.3440(14)$  Å  
 $\beta = 113.02(3)^\circ$

$V = 1140.4(7)$  Å<sup>3</sup>  
 $Z = 4$   
 Mo  $K\alpha$  radiation  
 $\mu = 0.10$  mm<sup>-1</sup>  
 $T = 150(1)$  K  
 $0.90 \times 0.17 \times 0.12$  mm

#### Data collection

Bruker-Nonius KappaCCD area-detector diffractometer  
 Absorption correction: gaussian (Coppens, 1970)  
 $T_{\min} = 0.961$ ,  $T_{\max} = 0.993$   
 7864 measured reflections  
 2522 independent reflections  
 1547 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.139$

#### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.111$   
 $wR(F^2) = 0.274$   
 $S = 1.13$   
 2522 reflections  
 145 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{\text{max}} = 0.33$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.41$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{N1}-\text{H1}\cdots\text{O1}^i$	0.86	2.02	2.870 (3)	168

 Symmetry code: (i)  $x - 1, y, z$ .

Data collection: COLLECT (Hooft, 1998) and DENZO (Otwinowski & Minor, 1997); cell refinement: DIRAX/LSQ (Duisenberg, 1992); data reduction: EvalCCD (Duisenberg, 1992); program(s) used to solve structure: SIR92 (Altomare *et al.*, 1994); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: PLATON (Spek, 2003); software used to prepare material for publication: SHELXL97.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HK2530).

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## supporting information

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**(E)-3-(3,5-Dimethoxyphenyl)acrylohydrazide**

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**S1. Comment**

Aromatic hydrazides are important intermediates in heterocyclic chemistry and have been used for the synthesis of various biologically active five-membered heterocycles such as 2,5-disubstituted-1,3,4-oxadiazoles (Zheng *et al.*, 2003; Al-Talib *et al.*, 1990) and 5-substituted-2-mercapto-1,3,4-oxadiazoles (Yousif *et al.*, 1986; Ahmad *et al.*, 2001; Al-Soud *et al.*, 2004; El-Emam *et al.*, 2004). In view of the versatility of these compounds, we have synthesized the title compound, and report herein its crystal structure.

In the molecule of the title compound, (Fig. 1), the bond lengths (Allen *et al.*, 1987) and angles are generally within normal ranges. The benzene ring (C4-C9) is oriented with respect to the planar hydrazide group (O1/N1/N2/C1) at a dihedral angle of 48.00 (3)°.

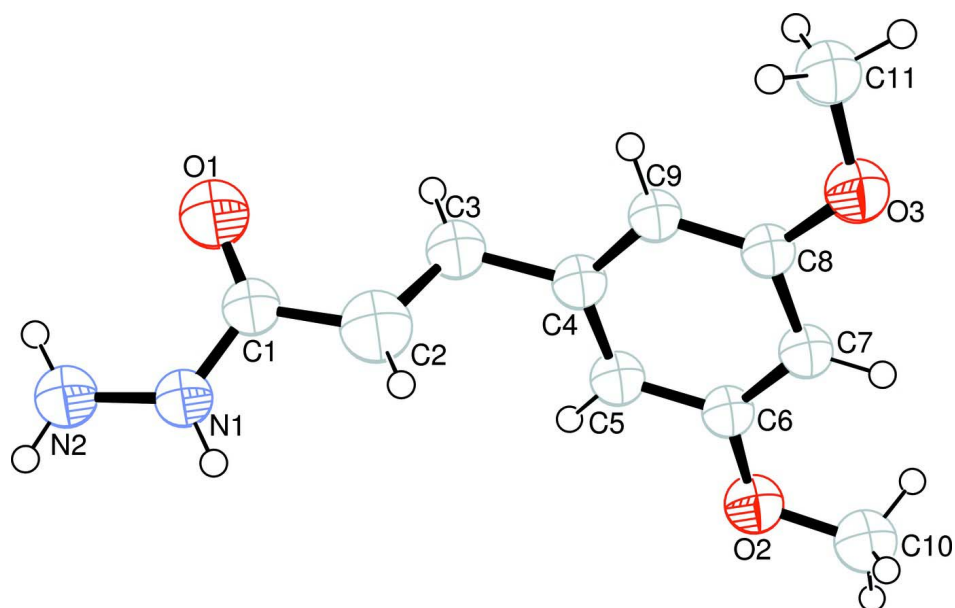
In the crystal structure, intermolecular N-H...O hydrogen bonds (Table 1) link the molecules (Fig. 2), in which they may be effective in the stabilization of the structure.

**S2. Experimental**

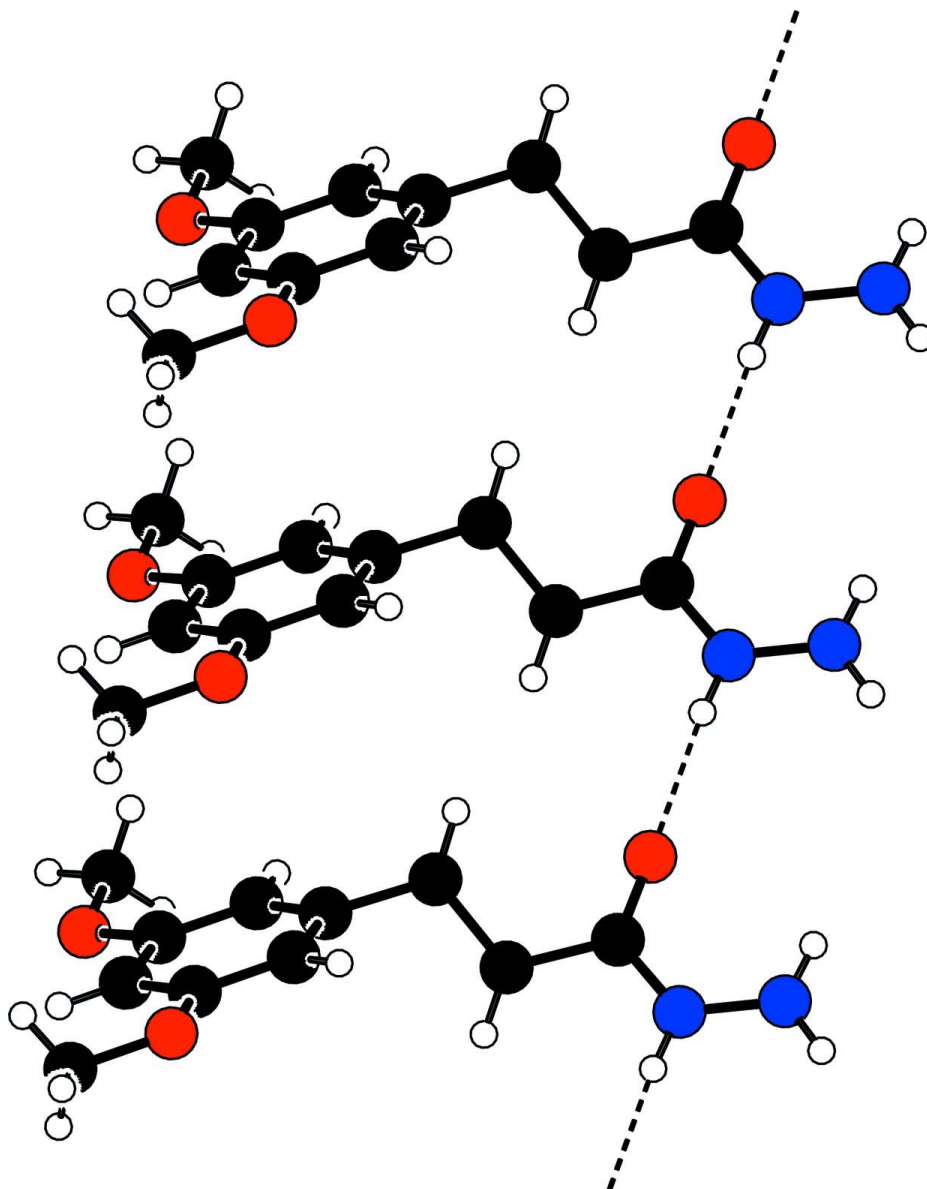
The title compound was synthesized by the reaction of methyl ester of (E)-3-(3,5-dimethoxyphenyl)acrylic acid with hydrazine hydrate according to the literature method (Furniss *et al.*, 1978). For the preparation of the title compound, a mixture of (E)-methyl 3-(3,5-dimethoxyphenyl)acrylate (2.22 g, 10 mmol) and hydrazine hydrate (15 ml, 80%) in absolute ethanol (50 ml) was refluxed for 5 h at 413-423 K. The excess solvent was removed by distillation. The solid residue was filtered off, washed with water and recrystallized from ethanol (30%) to give the title compound (yield; 1.55 g, 70%, m.p. 401-402 K). Colorless single crystals were obtained by slow evaporation of an ethanol solution at room temperature.

**S3. Refinement**

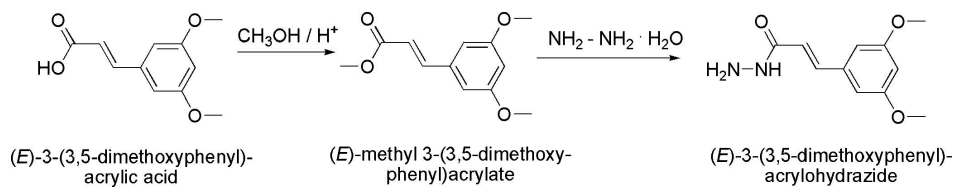
H atoms were positioned geometrically, with N-H = 0.86 Å (for NH and NH<sub>2</sub>) and C-H = 0.93 and 0.96 Å for aromatic and methyl H, respectively, and constrained to ride on their parent atoms with  $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C},\text{N})$ .

**Figure 1**

The molecular structure of the title molecule with the atom-numbering scheme.

**Figure 2**

A packing diagram of the title compound. Hydrogen bonds are shown as dashed lines.

**Figure 3**

Reaction scheme.

**(E)-3-(3,5-dimethoxyphenyl)acrylohydrazide***Crystal data*C<sub>11</sub>H<sub>14</sub>N<sub>2</sub>O<sub>3</sub> $M_r = 222.24$ Monoclinic,  $P2_1/c$ Hall symbol:  $-P\ 2ybc$  $a = 4.8910\ (19)\ \text{\AA}$  $b = 30.358\ (11)\ \text{\AA}$  $c = 8.3440\ (14)\ \text{\AA}$  $\beta = 113.02\ (3)^\circ$  $V = 1140.4\ (7)\ \text{\AA}^3$  $Z = 4$  $F(000) = 472$  $D_x = 1.294\ \text{Mg m}^{-3}$ 

Melting point: 401(1) K

Mo  $K\alpha$  radiation,  $\lambda = 0.71073\ \text{\AA}$ 

Cell parameters from 7914 reflections

 $\theta = 1\text{--}27.5^\circ$  $\mu = 0.10\ \text{mm}^{-1}$  $T = 150\ \text{K}$ 

Needle, colorless

 $0.90 \times 0.17 \times 0.12\ \text{mm}$ *Data collection*Bruker–Nonius KappaCCD area-detector  
diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

Detector resolution: 9.091 pixels  $\text{mm}^{-1}$  $\varphi$  and  $\omega$  scansAbsorption correction: gaussian  
(Coppens, 1970) $T_{\min} = 0.961$ ,  $T_{\max} = 0.993$ 

7864 measured reflections

2522 independent reflections

1547 reflections with  $I > 2\sigma(I)$  $R_{\text{int}} = 0.139$  $\theta_{\max} = 27.5^\circ$ ,  $\theta_{\min} = 3.0^\circ$  $h = -5 \rightarrow 6$  $k = -39 \rightarrow 35$  $l = -10 \rightarrow 9$ *Refinement*Refinement on  $F^2$ 

Least-squares matrix: full

 $R[F^2 > 2\sigma(F^2)] = 0.111$  $wR(F^2) = 0.274$  $S = 1.13$ 

2522 reflections

145 parameters

0 restraints

Primary atom site location: structure-invariant  
direct methodsSecondary atom site location: difference Fourier  
mapHydrogen site location: inferred from  
neighbouring sites

H-atom parameters constrained

 $w = 1/[\sigma^2(F_o^2) + (0.0574P)^2 + 2.6221P]$ where  $P = (F_o^2 + 2F_c^2)/3$  $(\Delta/\sigma)_{\max} < 0.001$  $\Delta\rho_{\max} = 0.33\ \text{e \AA}^{-3}$  $\Delta\rho_{\min} = -0.41\ \text{e \AA}^{-3}$ *Special details*

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) etc. and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
O1	0.5287 (7)	0.28398 (13)	0.5987 (6)	0.0682 (12)
O2	-0.4491 (7)	0.08832 (12)	0.3236 (4)	0.0538 (9)
O3	0.3626 (7)	0.04751 (12)	0.8379 (5)	0.0544 (10)

N1	0.0720 (8)	0.30545 (13)	0.5611 (6)	0.0468 (10)
H1	-0.0993	0.2975	0.5571	0.056*
N2	0.1165 (9)	0.34981 (14)	0.5371 (6)	0.0548 (11)
H2A	0.2856	0.3587	0.5404	0.066*
H2B	-0.0256	0.3683	0.5186	0.066*
C1	0.2785 (10)	0.27498 (16)	0.5905 (7)	0.0469 (12)
C2	0.1853 (14)	0.22916 (18)	0.6146 (10)	0.074 (2)
H2	0.0374	0.2236	0.6559	0.089*
C3	0.3475 (11)	0.19327 (16)	0.5674 (7)	0.0516 (13)
H3	0.5095	0.1979	0.5373	0.062*
C4	0.2192 (10)	0.14822 (15)	0.5752 (6)	0.0437 (11)
C5	-0.0544 (10)	0.13674 (16)	0.4480 (6)	0.0451 (11)
H5	-0.1530	0.1561	0.3573	0.054*
C6	-0.1797 (9)	0.09645 (15)	0.4560 (6)	0.0408 (10)
C7	-0.0380 (9)	0.06710 (16)	0.5864 (6)	0.0417 (10)
H7	-0.1233	0.0400	0.5912	0.050*
C8	0.2389 (10)	0.07918 (15)	0.7127 (6)	0.0404 (10)
C9	0.3672 (9)	0.11891 (15)	0.7090 (6)	0.0391 (10)
H9	0.5512	0.1261	0.7945	0.047*
C10	-0.5868 (11)	0.0472 (2)	0.3233 (7)	0.0605 (15)
H10A	-0.4515	0.0237	0.3308	0.073*
H10B	-0.7611	0.0443	0.2179	0.073*
H10C	-0.6420	0.0461	0.4217	0.073*
C11	0.6609 (10)	0.05504 (18)	0.9613 (7)	0.0547 (13)
H11A	0.7879	0.0602	0.8999	0.066*
H11B	0.7292	0.0296	1.0344	0.066*
H11C	0.6652	0.0802	1.0318	0.066*

*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
O1	0.0387 (18)	0.061 (2)	0.117 (4)	-0.0048 (17)	0.043 (2)	0.001 (2)
O2	0.0457 (18)	0.065 (2)	0.0394 (19)	-0.0032 (17)	0.0038 (15)	-0.0008 (16)
O3	0.0441 (18)	0.056 (2)	0.049 (2)	-0.0058 (16)	0.0027 (15)	0.0103 (16)
N1	0.0296 (18)	0.050 (2)	0.062 (3)	-0.0042 (17)	0.0198 (18)	-0.005 (2)
N2	0.043 (2)	0.049 (2)	0.072 (3)	0.0035 (19)	0.022 (2)	0.009 (2)
C1	0.040 (2)	0.050 (3)	0.056 (3)	-0.005 (2)	0.024 (2)	-0.003 (2)
C2	0.076 (4)	0.048 (3)	0.134 (6)	-0.010 (3)	0.080 (4)	-0.008 (3)
C3	0.048 (3)	0.046 (3)	0.069 (4)	0.000 (2)	0.032 (3)	0.006 (2)
C4	0.043 (2)	0.047 (3)	0.049 (3)	-0.003 (2)	0.026 (2)	-0.004 (2)
C5	0.048 (3)	0.051 (3)	0.035 (2)	0.008 (2)	0.015 (2)	0.008 (2)
C6	0.038 (2)	0.049 (3)	0.037 (2)	0.002 (2)	0.0157 (19)	-0.006 (2)
C7	0.040 (2)	0.043 (2)	0.043 (3)	-0.005 (2)	0.017 (2)	-0.004 (2)
C8	0.042 (2)	0.045 (3)	0.035 (2)	0.003 (2)	0.0155 (19)	0.0017 (19)
C9	0.033 (2)	0.045 (3)	0.038 (2)	-0.0019 (19)	0.0119 (18)	-0.0043 (19)
C10	0.043 (3)	0.078 (4)	0.049 (3)	-0.008 (3)	0.006 (2)	-0.009 (3)
C11	0.046 (3)	0.064 (3)	0.045 (3)	-0.003 (2)	0.007 (2)	0.006 (2)

*Geometric parameters (Å, °)*

O1—C1	1.230 (5)	C4—C5	1.389 (7)
O2—C6	1.371 (5)	C5—H5	0.9301
O2—C10	1.416 (6)	C6—C5	1.381 (7)
O3—C8	1.374 (5)	C6—C7	1.367 (7)
O3—C11	1.438 (6)	C7—H7	0.9298
N1—N2	1.391 (6)	C8—C7	1.402 (6)
N1—C1	1.320 (6)	C8—C9	1.365 (6)
N1—H1	0.8600	C9—C4	1.389 (7)
N2—H2A	0.8601	C9—H9	0.9300
N2—H2B	0.8600	C10—H10A	0.9598
C1—C2	1.502 (7)	C10—H10B	0.9600
C2—H2	0.9300	C10—H10C	0.9600
C3—C2	1.489 (7)	C11—H11A	0.9601
C3—C4	1.517 (7)	C11—H11B	0.9600
C3—H3	0.9300	C11—H11C	0.9600
N2—N1—H1	118.2	C7—C6—C5	121.3 (4)
C1—N1—N2	123.6 (4)	O2—C6—C5	115.2 (4)
C1—N1—H1	118.2	C6—C7—C8	117.9 (4)
N1—N2—H2A	120.1	C6—C7—H7	120.9
N1—N2—H2B	119.9	C8—C7—H7	121.1
H2A—N2—H2B	120.0	C9—C8—O3	124.4 (4)
C6—O2—C10	117.8 (4)	C9—C8—C7	122.1 (4)
C8—O3—C11	117.1 (4)	O3—C8—C7	113.4 (4)
O1—C1—N1	121.8 (5)	C8—C9—C4	119.0 (4)
O1—C1—C2	123.1 (5)	C8—C9—H9	120.4
N1—C1—C2	115.1 (4)	C4—C9—H9	120.6
C1—C2—H2	122.6	O2—C10—H10A	109.9
C3—C2—C1	114.9 (4)	O2—C10—H10B	109.6
C3—C2—H2	122.5	O2—C10—H10C	108.9
C2—C3—C4	112.1 (4)	H10A—C10—H10B	109.5
C2—C3—H3	124.0	H10A—C10—H10C	109.5
C4—C3—H3	123.9	H10B—C10—H10C	109.5
C5—C4—C3	119.0 (4)	O3—C11—H11A	109.3
C5—C4—C9	119.8 (4)	O3—C11—H11B	109.3
C9—C4—C3	121.2 (4)	O3—C11—H11C	109.8
C4—C5—H5	120.1	H11A—C11—H11B	109.5
C6—C5—C4	119.9 (4)	H11A—C11—H11C	109.5
C6—C5—H5	120.0	H11B—C11—H11C	109.5
C7—C6—O2	123.6 (4)		

*Hydrogen-bond geometry (Å, °)*

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
N1—H1 $\cdots$ O1 <sup>i</sup>	0.86	2.02	2.870 (3)	168

Symmetry code: (i)  $x-1, y, z$ .