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A bi-TTF with a bipyridine spacer: 4,4'-bis[(3,6,7-trimethylsulfanyl)tetrathia- fulvalen-2-yl)sulfanylmethyl]-2,2'- bipyridine

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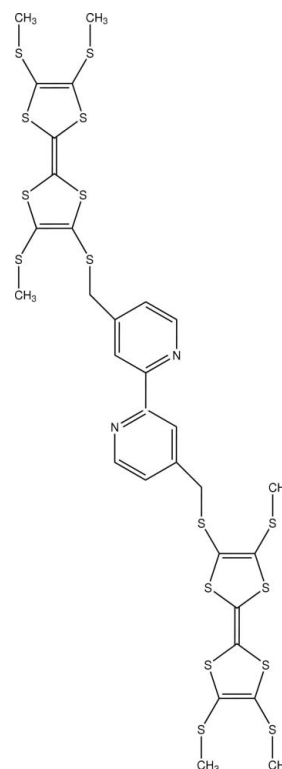
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Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{C}-\text{C}) = 0.004$ Å; R factor = 0.035; wR factor = 0.071; data-to-parameter ratio = 15.4.

The title compound, $\text{C}_{30}\text{H}_{28}\text{N}_2\text{S}_{16}$, is a precursor to hybrid magnetic materials. The complete molecule is generated by a crystallographic inversion centre. In the crystal structure, the TTF core is not planar and adopts a chair conformation; the two C_3S_2 rings are folded around the $\text{S}\cdots\text{S}$ hinges, the dihedral angles being 17.14 (8) and 13.46 (7)°. There is a short $\text{S}\cdots\text{S}$ contact [3.4863 (14) Å] in the crystal structure.

Related literature

For general background, see: Yagubskii (1993); Williams *et al.* (1992); Sakata *et al.* (1998); Fabre (2002). For coordination complexes of TTF with nitrogen aromatic substituents, see: Setifi *et al.* (2003); Liu *et al.* (2003); Boudiba *et al.* (2005). For the double Wittig coupling reaction used in the synthesis of the bi-TTF(bipyridine), see: Ikeda *et al.* (1993); Gonzales *et al.* (2000). For the synthesis of the precursors, see: Doria *et al.* (1986); Hudhomme *et al.* (2006); Blanchard *et al.* (1993).



Experimental

Crystal data

$\text{C}_{30}\text{H}_{28}\text{N}_2\text{S}_{16}$	$\gamma = 72.245$ (13)°
$M_r = 929.50$	$V = 965.2$ (3) Å ³
Triclinic, $P\bar{1}$	$Z = 1$
$a = 7.4840$ (12) Å	Mo $K\alpha$ radiation
$b = 7.7691$ (11) Å	$\mu = 0.92$ mm ⁻¹
$c = 17.707$ (3) Å	$T = 293$ (2) K
$\alpha = 88.973$ (12)°	$0.19 \times 0.11 \times 0.06$ mm
$\beta = 80.071$ (13)°	

Data collection

Oxford Diffraction XCalibur diffractometer with CCD detector	6690 measured reflections
Absorption correction: multi-scan (Blessing, 1995)	3391 independent reflections
$T_{\min} = 0.884$, $T_{\max} = 0.937$	1942 reflections with $I > 2\sigma(I)$
	$R_{\text{int}} = 0.038$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.035$	220 parameters
$wR(F^2) = 0.070$	H-atom parameters constrained
$S = 0.83$	$\Delta\rho_{\text{max}} = 0.23$ e Å ⁻³
3391 reflections	$\Delta\rho_{\text{min}} = -0.21$ e Å ⁻³

Data collection: *CrysAlis CCD* (Oxford Diffraction, 2006); cell refinement: *CrysAlis RED* (Oxford Diffraction, 2006); data reduction: *CrysAlis RED*; program(s) used to solve structure: *SIR92* (Altomare *et al.*, 1993); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEPIII* (Burnett & Johnson, 1996), *CAMERON* (Watkin *et al.*, 1993) and *ORTEP-3* (Farrugia, 1997); software used to prepare material for publication: *WinGX* (Farrugia, 1999).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: NC2124).

References

- Altomare, A., Cascarano, G., Giacovazzo, C. & Guagliardi, A. (1993). *J. Appl. Cryst.* **26**, 343–350.
- Blanchard, P., Dugauy, G., Cousseau, J., Sallé, M., Jubault, M., Gorgues, A., Boubekeur, K. & Batail, P. (1993). *Synth. Met.* **55–57**, 2113–2117.
- Blessing, R. H. (1995). *Acta Cryst.* **A51**, 33–38.
- Boudiba, L., Gouasmia, A. K., Kaboub, L., Cador, O., Ouahab, L. & Fabre, J.-M. (2005). *Synth. Met.* **150**, 317–320.
- Burnett, M. N. & Johnson, C. K. (1996). *ORTEP III*. Report ORNL-6895. Oak Ridge National Laboratory, Tennessee, USA.
- Doria, G., Passarotti, C., Sala, R., Magrini, R., Sberze, P., Tibolla, M., Ceserani, R. & Castello, R. (1986). *Farm. Ed. Sci.* **41**, 417–428.
- Fabre, J.-M. (2002). *J. Solid State Chem.* **168**, 367–383.
- Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.
- Farrugia, L. J. (1999). *J. Appl. Cryst.* **32**, 837–838.
- Gonzalez, A., Segura, J. L. & Martin, N. (2000). *Tetrahedron Lett.* **41**, 3083–3086.
- Hudhomme, P., Sallé, M., Gautier, N., Belyasmine, A. & Gorgues, A. (2006). *Arkoivoc*, **iv**, 49–72.
- Ikeda, K., Kawabata, K., Tanaka, K. & Mizutani, M. (1993). *Synth. Met.* **55–57**, 2007–2012.
- Liu, H.-X., Dolder, S., Franz, P., Neels, A., Stoeckli-Evans, H. & Decurtins, S. (2003). *Inorg. Chem.* **42**, 4801–4803.
- Oxford Diffraction (2006). *CrysAlis CCD* and *CrysAlis RED*. Oxford Diffraction Ltd, Abingdon, Oxfordshire, England.
- Sakata, J. I., Sato, H., Misayaki, A., Enoki, T., Okano, Y. & Kato, R. (1998). *Solid State Commun.* **108**, 377–381.
- Setifi, F., Ouahab, L., Gohlen, S., Yoshida, Y. & Saito, G. (2003). *Inorg. Chem.* **42**, 1791–1793.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Watkin, D. M., Pearce, L. & Prout, C. K. A. (1993). *CAMERON*. University of Oxford, England.
- Williams, J. M., Ferraro, J. R., Thorn, R. I., Carlson, K. D., Geiser, U., Wang, H. H. A. M., Kini, A. M. & Wangbo, M. H. (1992). *Organic Superconductors (including Fullerenes), Synthesis, Structure, Properties and Theory*. Englewood Cliffs, NJ: Prentice Hall.
- Yagubskii, E. B. (1993). *Mol. Cryst. Liq. Cryst.* **230**, 139–156.

supporting information

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A bi-TTF with a bipyridine spacer: 4,4'-bis[(3,6,7-trimethylsulfanyltetrathiafulvalen-2-yl)sulfanylmethyl]-2,2'-bipyridine

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S1. Comment

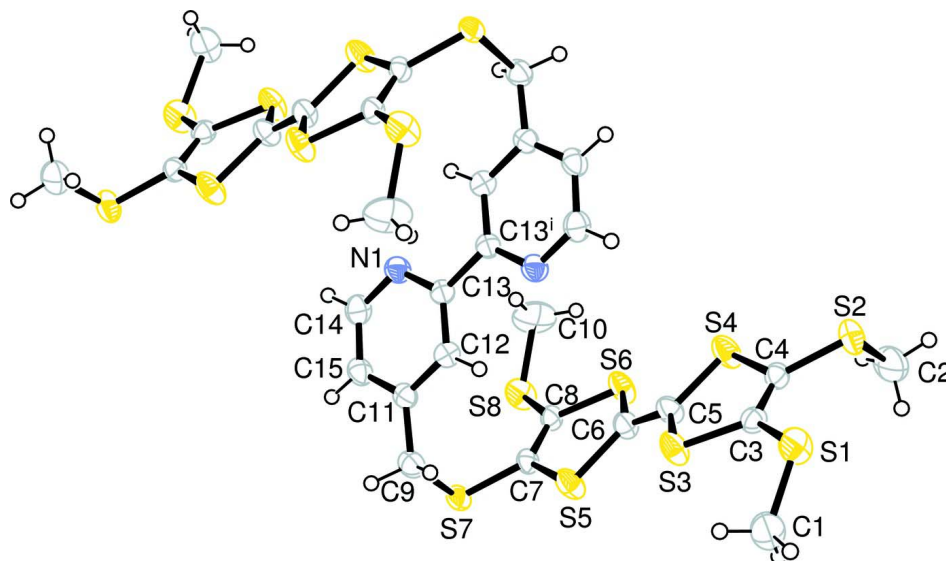
To date the search for solids presenting two physical properties such as magnetism and electrical conductivity inside a same material has expanded greatly, particularly with materials involving tetrathiafulvalene (TTF) derivatives well known to provide conducting and even superconducting salts (Williams *et al.*, 1992; Yagubskii, 1993; Sakata *et al.*, 1998; Fabre, 2002). To introduce a magnetic network, involving localized spins, inside such conducting salts, a particularly promising way is to build a coordination complex between a transition metal (Cu, Co ...) and a pyridine or bipyridine moiety bonded to a TTF core (Setifi *et al.*, 2003; Liu *et al.*, 2003; Boudiba *et al.*, 2005). Following this strategy we synthesized the title precursor: bi-TTF(bipyridine) **1** and studied its crystal structure to verify if the molecular geometry could allow a subsequent easy formation of the target coordinating complex. The molecular structure is shown in Fig. 1. As expected two TTF cores bearing methylsulfanyl substituents are connected by a bipyridine spacer. The molecule lies on a crystallographic centre of symmetry located at the centre of the bipyridine moiety, the asymmetric unit is thus composed of half a molecule. As a result the bipyridine spacer is in the *trans* conformation. The TTF cores deviate strongly from planarity and take a chair conformation. The two C₃S₂ rings are folded around the S...S hinges: the central group S3/S4/C5/C6/S5/S6 is planar and the external planes S3/S4/C3/C4 and S5/S6/C7/C8 make dihedral angles of 17.14 (8)° and 13.46 (7)° respectively. There is a short S...S contact [3.4863 (14) Å] in the crystal structure.

S2. Experimental

The bi-TTF(bipyridine) **1** was synthesized (37% yield) by using a double Wittig coupling reaction (Ikeda, 1993; Gonzalez, 2000) between two appropriate formyl-TTF units and 4,4'-bis(methyltripenylphosphonium)-2,2'-bipyridine-dibromide previously obtained as described in the literature (Doria, 1986). Red crystals (m.p.: 158°C) of **1** were obtained as thin platelets by slow evaporation of a solution of **1** in a mixture of dichloromethane-acetonitrile.

S3. Refinement

H atoms were located in a difference map then positioned geometrically and refined using a riding model with C—H distances set to 0.96 Å (*sp*³) and 0.93 Å (*sp*²), and *U*_{iso}(H) equal to 1.2 times the equivalent *U*_{iso} of the atom of attachment.

**Figure 1**

The molecular structure of the title compound, with atom labels and 50% probability displacement ellipsoids for non-H atoms. Symmetry code: $i = -x+1, -y+1, -z+1$.

4,4'-bis[(3,6,7-trimethylsulfonyltetrathiafulvalen-2-yl)sulfanylmethyl]-2,2'-bipyridine

Crystal data

$C_{30}H_{28}N_2S_{16}$

$M_r = 929.50$

Triclinic, $P\bar{1}$

Hall symbol: $-P\ 1$

$a = 7.4840$ (12) Å

$b = 7.7691$ (11) Å

$c = 17.707$ (3) Å

$\alpha = 88.973$ (12)°

$\beta = 80.071$ (13)°

$\gamma = 72.245$ (13)°

$V = 965.2$ (3) Å³

$Z = 1$

$F(000) = 478$

$D_x = 1.599$ Mg m⁻³

Melting point: 431 K

Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 1223 reflections

$\theta = 2.9$ – 25.0 °

$\mu = 0.92$ mm⁻¹

$T = 293$ K

Block, red

$0.19 \times 0.11 \times 0.06$ mm

Data collection

Oxford Diffraction XCalibur
diffractometer with CCD detector
Radiation source: fine-focus sealed tube
Graphite monochromator

φ and ω scans

Absorption correction: multi-scan
(Blessing, 1995)

$T_{\min} = 0.884$, $T_{\max} = 0.937$

6690 measured reflections

3391 independent reflections

1942 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.039$

$\theta_{\max} = 25.0$ °, $\theta_{\min} = 2.9$ °

$h = -8 \rightarrow 8$

$k = -9 \rightarrow 9$

$l = -18 \rightarrow 21$

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.035$

$wR(F^2) = 0.071$

$S = 0.83$

3391 reflections

220 parameters

0 restraints

Primary atom site location: structure-invariant
direct methods

Secondary atom site location: difference Fourier map
 Hydrogen site location: inferred from neighbouring sites
 H-atom parameters constrained

$$w = 1/[\sigma^2(F_o^2) + (0.0277P)^2]$$

where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} = 0.013$
 $\Delta\rho_{\max} = 0.23 \text{ e } \text{Å}^{-3}$
 $\Delta\rho_{\min} = -0.22 \text{ e } \text{Å}^{-3}$

Special details

Experimental. Excalibur (Oxford Diffraction) four-circle Kappa geometry diffractometer equipped with an area CCD detector. Crystal-detector distance (mm): 70.0

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å²)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
C1	−0.3976 (5)	0.7011 (5)	0.9533 (2)	0.0573 (11)
H1A	−0.3692	0.6081	0.9142	0.069*
H1B	−0.5251	0.7212	0.9806	0.069*
H1C	−0.3095	0.6642	0.9884	0.069*
C2	0.0146 (6)	1.2151 (5)	0.9481 (2)	0.0681 (13)
H2A	−0.0135	1.1531	0.9938	0.082*
H2B	−0.0070	1.3402	0.9609	0.082*
H2C	0.1455	1.1612	0.9247	0.082*
C3	−0.1411 (4)	0.8506 (4)	0.86485 (17)	0.0307 (8)
C4	−0.0438 (4)	0.9702 (4)	0.85327 (17)	0.0300 (8)
C5	0.1986 (4)	0.6686 (4)	0.79718 (18)	0.0339 (8)
C6	0.3591 (4)	0.5317 (4)	0.77605 (17)	0.0325 (8)
C7	0.6075 (4)	0.2257 (4)	0.72586 (16)	0.0271 (8)
C8	0.7017 (4)	0.3475 (4)	0.71193 (17)	0.0277 (8)
C9	0.5564 (5)	−0.0019 (4)	0.62192 (17)	0.0366 (8)
H9A	0.5827	−0.1235	0.6028	0.044*
H9B	0.4232	0.0432	0.6433	0.044*
C10	0.9345 (5)	0.4866 (5)	0.6095 (2)	0.0711 (13)
H10A	0.8469	0.4962	0.5748	0.085*
H10B	1.0597	0.4718	0.5809	0.085*
H10C	0.8944	0.5944	0.6417	0.085*
C11	0.6020 (4)	0.1130 (4)	0.55723 (17)	0.0296 (8)
C12	0.5025 (4)	0.2947 (4)	0.55804 (17)	0.0308 (8)
H12	0.4020	0.3454	0.5979	0.037*
C13	0.5516 (4)	0.4015 (4)	0.49995 (17)	0.0276 (8)
C14	0.7883 (5)	0.1602 (4)	0.44029 (19)	0.0381 (9)
H14	0.8869	0.1118	0.3995	0.046*
C15	0.7491 (4)	0.0458 (4)	0.49646 (17)	0.0351 (8)

H15	0.8209	-0.0757	0.4935	0.042*
N1	0.6944 (4)	0.3353 (3)	0.44078 (14)	0.0343 (7)
S1	-0.37725 (13)	0.90564 (13)	0.91017 (6)	0.0500 (3)
S2	-0.13591 (13)	1.19803 (12)	0.88259 (5)	0.0442 (3)
S3	-0.02471 (12)	0.63191 (11)	0.82438 (5)	0.0408 (2)
S4	0.18635 (13)	0.89647 (12)	0.79792 (5)	0.0451 (3)
S5	0.37241 (12)	0.30319 (11)	0.77697 (5)	0.0419 (3)
S6	0.58066 (12)	0.56994 (11)	0.74683 (5)	0.0409 (3)
S7	0.69659 (12)	-0.00222 (11)	0.69692 (5)	0.0333 (2)
S8	0.93880 (12)	0.29726 (12)	0.66720 (5)	0.0453 (3)

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.051 (2)	0.054 (3)	0.064 (3)	-0.026 (2)	0.017 (2)	-0.008 (2)
C2	0.108 (4)	0.050 (3)	0.059 (3)	-0.032 (3)	-0.034 (3)	0.001 (2)
C3	0.0287 (19)	0.033 (2)	0.0275 (19)	-0.0055 (15)	-0.0035 (15)	0.0019 (15)
C4	0.0294 (19)	0.0278 (19)	0.0287 (19)	-0.0028 (15)	-0.0052 (15)	-0.0004 (15)
C5	0.0280 (19)	0.0282 (19)	0.045 (2)	-0.0107 (16)	-0.0004 (16)	0.0009 (16)
C6	0.0274 (19)	0.0291 (19)	0.038 (2)	-0.0093 (16)	0.0030 (16)	-0.0033 (16)
C7	0.0237 (18)	0.0235 (19)	0.0299 (19)	-0.0010 (14)	-0.0044 (15)	-0.0019 (14)
C8	0.0224 (18)	0.0268 (19)	0.0293 (18)	-0.0038 (15)	0.0014 (15)	-0.0034 (15)
C9	0.042 (2)	0.031 (2)	0.041 (2)	-0.0131 (16)	-0.0135 (18)	0.0027 (16)
C10	0.046 (3)	0.083 (3)	0.079 (3)	-0.027 (2)	0.011 (2)	0.025 (3)
C11	0.0317 (19)	0.0278 (19)	0.033 (2)	-0.0109 (15)	-0.0132 (16)	0.0020 (15)
C12	0.0281 (18)	0.033 (2)	0.0307 (19)	-0.0090 (15)	-0.0052 (15)	0.0016 (15)
C13	0.0264 (18)	0.0279 (18)	0.0290 (18)	-0.0063 (14)	-0.0098 (15)	0.0016 (15)
C14	0.031 (2)	0.038 (2)	0.041 (2)	-0.0062 (17)	-0.0012 (17)	-0.0046 (17)
C15	0.036 (2)	0.0276 (19)	0.040 (2)	-0.0027 (16)	-0.0142 (18)	-0.0004 (17)
N1	0.0322 (16)	0.0298 (17)	0.0366 (17)	-0.0053 (13)	-0.0019 (14)	0.0005 (13)
S1	0.0310 (5)	0.0449 (6)	0.0615 (7)	-0.0045 (4)	0.0131 (5)	-0.0036 (5)
S2	0.0453 (6)	0.0301 (5)	0.0517 (6)	-0.0035 (4)	-0.0073 (5)	-0.0098 (4)
S3	0.0268 (5)	0.0278 (5)	0.0627 (6)	-0.0086 (4)	0.0076 (5)	-0.0077 (4)
S4	0.0335 (5)	0.0282 (5)	0.0667 (7)	-0.0098 (4)	0.0105 (5)	-0.0042 (5)
S5	0.0310 (5)	0.0280 (5)	0.0620 (7)	-0.0109 (4)	0.0086 (5)	-0.0017 (4)
S6	0.0281 (5)	0.0266 (5)	0.0636 (7)	-0.0094 (4)	0.0061 (5)	-0.0077 (4)
S7	0.0382 (5)	0.0230 (5)	0.0360 (5)	-0.0030 (4)	-0.0107 (4)	0.0023 (4)
S8	0.0246 (5)	0.0390 (5)	0.0627 (7)	-0.0041 (4)	0.0075 (5)	-0.0003 (5)

Geometric parameters (Å, °)

C1—S1	1.788 (3)	C8—S8	1.741 (3)
C1—H1A	0.9599	C8—S6	1.755 (3)
C1—H1B	0.9599	C9—C11	1.497 (4)
C1—H1C	0.9599	C9—S7	1.828 (3)
C2—S2	1.783 (4)	C9—H9A	0.9600
C2—H2A	0.9599	C9—H9B	0.9600
C2—H2B	0.9599	C10—S8	1.771 (3)

C2—H2C	0.9599	C10—H10A	0.9599
C3—C4	1.338 (4)	C10—H10B	0.9599
C3—S1	1.736 (3)	C10—H10C	0.9599
C3—S3	1.756 (3)	C11—C12	1.380 (4)
C4—S2	1.744 (3)	C11—C15	1.380 (4)
C4—S4	1.758 (3)	C12—C13	1.381 (4)
C5—C6	1.340 (4)	C12—H12	0.9300
C5—S4	1.744 (3)	C13—N1	1.342 (4)
C5—S3	1.765 (3)	C13—C13 ⁱ	1.489 (6)
C6—S5	1.747 (3)	C14—N1	1.327 (4)
C6—S6	1.762 (3)	C14—C15	1.376 (4)
C7—C8	1.339 (4)	C14—H14	0.9300
C7—S7	1.743 (3)	C15—H15	0.9300
C7—S5	1.761 (3)		
S1—C1—H1A	109.5	C11—C9—H9B	109.4
S1—C1—H1B	109.5	S7—C9—H9B	109.4
H1A—C1—H1B	109.5	H9A—C9—H9B	108.0
S1—C1—H1C	109.5	S8—C10—H10A	109.5
H1A—C1—H1C	109.5	S8—C10—H10B	109.5
H1B—C1—H1C	109.5	H10A—C10—H10B	109.5
S2—C2—H2A	109.5	S8—C10—H10C	109.5
S2—C2—H2B	109.5	H10A—C10—H10C	109.5
H2A—C2—H2B	109.5	H10B—C10—H10C	109.5
S2—C2—H2C	109.5	C12—C11—C15	117.1 (3)
H2A—C2—H2C	109.5	C12—C11—C9	120.7 (3)
H2B—C2—H2C	109.5	C15—C11—C9	122.2 (3)
C4—C3—S1	123.6 (2)	C11—C12—C13	120.3 (3)
C4—C3—S3	116.7 (2)	C11—C12—H12	119.9
S1—C3—S3	119.55 (18)	C13—C12—H12	119.9
C3—C4—S2	124.6 (2)	N1—C13—C12	122.5 (3)
C3—C4—S4	117.5 (2)	N1—C13—C13 ⁱ	116.2 (3)
S2—C4—S4	117.70 (18)	C12—C13—C13 ⁱ	121.3 (4)
C6—C5—S4	124.5 (2)	N1—C14—C15	124.1 (3)
C6—C5—S3	121.8 (2)	N1—C14—H14	118.0
S4—C5—S3	113.67 (18)	C15—C14—H14	118.0
C5—C6—S5	124.6 (2)	C14—C15—C11	119.3 (3)
C5—C6—S6	121.5 (2)	C14—C15—H15	120.3
S5—C6—S6	113.88 (18)	C11—C15—H15	120.3
C8—C7—S7	125.6 (2)	C14—N1—C13	116.8 (3)
C8—C7—S5	117.1 (2)	C3—S1—C1	104.34 (16)
S7—C7—S5	117.26 (17)	C4—S2—C2	101.90 (17)
C7—C8—S8	124.6 (2)	C3—S3—C5	94.78 (15)
C7—C8—S6	117.1 (2)	C5—S4—C4	94.69 (15)
S8—C8—S6	118.20 (17)	C6—S5—C7	95.17 (14)
C11—C9—S7	111.1 (2)	C8—S6—C6	95.04 (15)

C11—C9—H9A	109.4	C7—S7—C9	99.50 (14)
S7—C9—H9A	109.4	C8—S8—C10	102.28 (16)

Symmetry code: (i) $-x+1, -y+1, -z+1$.