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# 1*H*-Pyrrole-2-carboxylic acid

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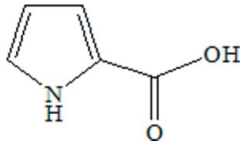
Key indicators: single-crystal X-ray study; $T$ = 173 K; mean $\sigma$(C–C) = 0.004 Å; $R$ factor = 0.063; $wR$ factor = 0.191; data-to-parameter ratio = 13.6.

In the title compound, $C_5H_5NO_2$, the pyrrole ring and its carboxyl substituent are close to coplanar, with a dihedral angle of 11.7 (3)° between the planes. In the crystal structure, adjacent molecules are linked by pairs of O—H···O hydrogen bonds to form inversion dimers. Additional N—H···O hydrogen bonds link these dimers into chains extending along the *a* axis.

## Related literature

For pyrroles sourced from marine organisms, see: Faulkner (2002). For the bioactivity of pyrrole derivatives, see: Banwell *et al.* (2006); Sosa *et al.* (2002). For related structures, see: Zeng (2006); Zeng *et al.* (2007). For graph-set motifs, see: Bernstein *et al.* (1995).




## Experimental

### Crystal data

$C_5H_5NO_2$
$M_r$ = 111.10
Monoclinic, $C2/c$
$a$ = 14.080 (3) Å
$b$ = 5.0364 (10) Å
$c$ = 14.613 (3) Å
$\beta$ = 98.969 (3)°
$V$ = 1023.6 (3) Å$^3$
$Z$ = 8
Mo $K\alpha$ radiation
$\mu$ = 0.11 mm$^{-1}$
$T$ = 173 K
0.42 × 0.40 × 0.37 mm

### Data collection

Bruker SMART 1K CCD area-detector diffractometer
Absorption correction: multi-scan (*SADABS*; Sheldrick, 1996)
$T_{min}$ = 0.954, $T_{max}$ = 0.959
2277 measured reflections
1006 independent reflections
875 reflections with $I > 2\sigma(I)$
$R_{int}$ = 0.015

### Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.063
$wR(F^2)$ = 0.191
$S$ = 1.06
1006 reflections
74 parameters
H-atom parameters constrained
$\Delta\rho_{max}$ = 0.74 e Å$^{-3}$
$\Delta\rho_{min}$ = −0.73 e Å$^{-3}$

**Table 1**
Hydrogen-bond geometry (Å, °).

| $D$—H···$A$ | $D$—H | H···$A$ | $D$···$A$ | $D$—H···$A$ |
|---|---|---|---|---|
| N1—H1*A*···O1$^{i}$ | 0.88 | 2.22 | 2.951 (3) | 141 |
| O2—H2*A*···O1$^{ii}$ | 0.84 | 2.16 | 2.986 (3) | 166 |

Symmetry codes: (i) $-x+\frac{1}{2}, -y+\frac{5}{2}, -z+1$; (ii) $-x, -y+2, -z+1$.

Data collection: *SMART* (Bruker,1999); cell refinement: *SAINT-Plus* (Bruker, 1999); data reduction: *SAINT-Plus*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

We thank the Natural Science Foundation of Guangdong Province, China (grant No. 06300581), for generously supporting this study.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: SJ2604).

## References

Banwell, M. G., Hamel, E., Hockless, D. C. R., Verdier-Pinard, P., Willis, A. C. & Wong, D. J. (2006). *Bioorg. Med. Chem.* **14**, 4627–4638.
Bernstein, J., Davis, R. E., Shimoni, L. & Chang, N.-L. (1995). *Angew. Chem. Int. Ed. Engl.* **34**, 1555–1573.
Bruker (1999). *SMART* and *SAINT-Plus*. Bruker AXS Inc., Madison, Wisconsin, USA.
Faulkner, D. J. (2002). *Nat. Prod. Rep.* **18**, 1–48.
Sheldrick, G. M. (1996). *SADABS*. University of Göttingen, Germany.
Sheldrick, G. M. (2008). *Acta Cryst.* A**64**, 112–122.
Sosa, A. C. B., Yakushijin, K. & Horne, D. A. (2002). *J. Org. Chem.* **67**, 4498–4500.
Zeng, X.-C. (2006). *Acta Cryst.* E**62**, o5505–o5507.
Zeng, X.-C., Zeng, J., Li, X. & Ling, X. (2007). *Acta Cryst.* E**63**, o3424.

# supporting information

*Acta Cryst.* (2009). E**65**, o1121 [doi:10.1107/S1600536809014044]

## 1*H*-Pyrrole-2-carboxylic acid

**Gui Hong Tang, Dong Dong Li, Gang Huang, Xing Yan Xu and Xiang Chao Zeng**

### S1. Comment

Pyrrole derivatives are well known in many marine organisms (Faulkner, 2002), some show important bioactivities, such as antitumor activity (Banwell *et al.*, 2006) and protein kinase inhibiting activity (Sosa *et al.*, 2002). This is the reason they have attracted our interest. This study is related to our previous structural investigations of methyl 2-(4,5-dibromo-1*H*-pyrrole-2-carboxamido)propionate (Zeng *et al.*, 2007) and 3-bromo-1-methyl-6,7-dihydropyrrolo[2,3-*c*]azepine- 4,8(1*H*,5*H*)-dione (Zeng, 2006). In the crystal structure, molecules of the title compound are linked through N1—H1···O1$^{i}$ hydrogen bonds to form centrosymmetric dimers (Fig. 2) of graph-set motif $R_2^2(10)$ (Bernstein *et al.*, 1995), which are linked by O2—H2···O1$^{ii}$ hydrogen bonds (another kind of centrosymmetric dimers of graph-set motif $R_2^2(8)$ are formed), generating chains extending to the *a* axis (also shown in Fig. 2).

### S2. Experimental

The commercially available 1*H*-pyrrole-2-carboxylic acid was dissolved in the mixture of EtOH (80%) and ethyl acetate (20%). Colorless monoclinic crystals suitable for X-ray analysis were obtained when the solution was exposed to the air at room temperature for about 5 d.

### S3. Refinement

All non-H atoms were refined with anisotropic displacement parameters. The H atoms were positioned geometrically [C—H = 0.95Å for CH, O—H = 0.84Å for OH, and N—H = 0.88 Å] and refined using a riding model, with $U_{iso} = 1.2U_{eq}$ ($1.5U_{eq}$ for the methyl group) of the parent atom. In the final difference Fourier map the highest peak (0.74 eÅ$^{-3}$) is 1.01Å from O2 and the deepest hole (-0.73 eÅ$^{-3}$) is 0.61Å from O2.

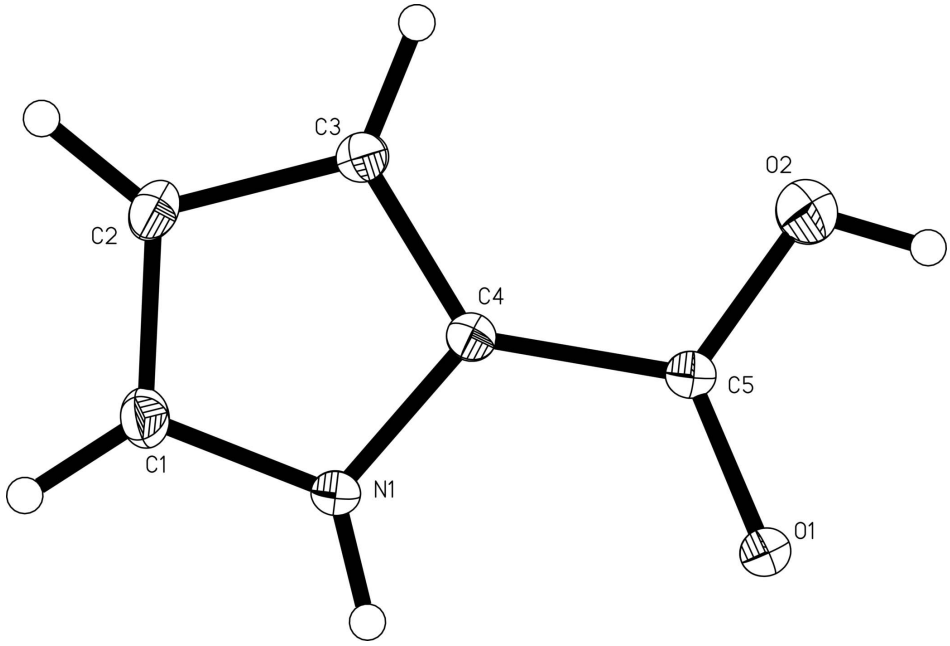



**Figure 1**

The molecular structure of the title compound, with the atom-numbering scheme. Displacement ellipsoids are drawn at the 30% probability level.

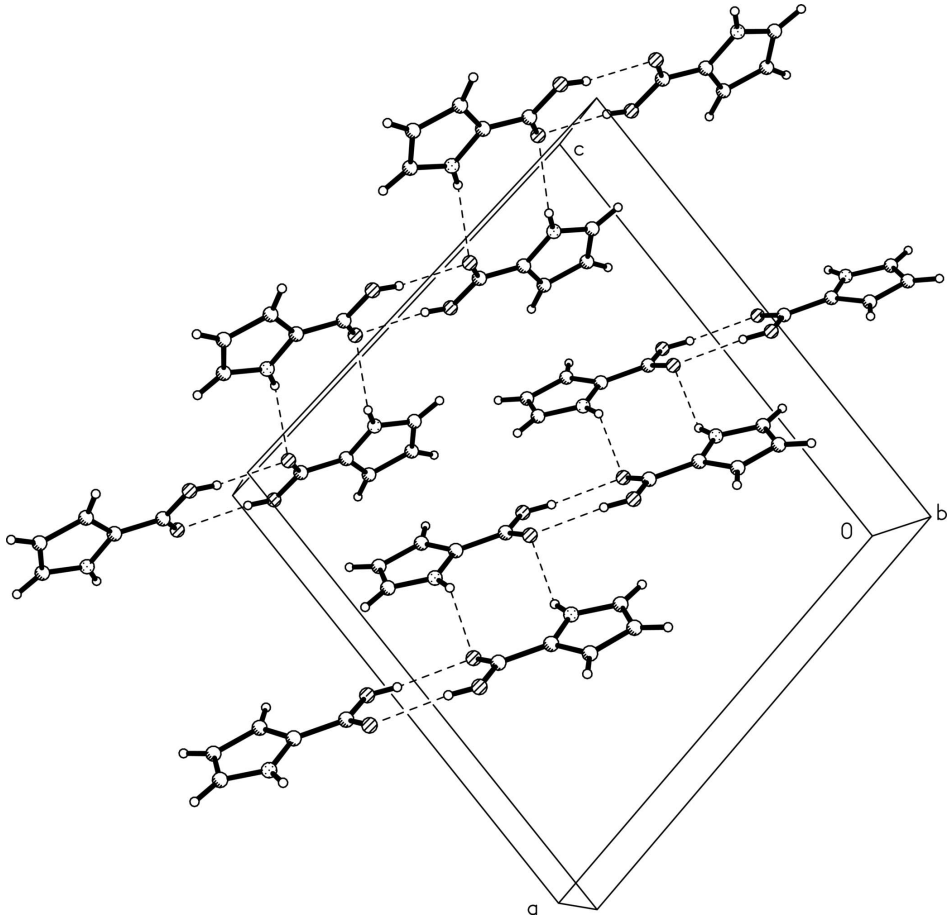



**Figure 2**

Crystal packing of (I) showing the chains formed by hydrogen bonds (dashed lines).

### 1*H*-Pyrrole-2-carboxylic acid

*Crystal data*

$C_5H_5NO_2$
$M_r = 111.10$
Monoclinic, $C2/c$
Hall symbol: -C 2yc
$a = 14.080\ (3)$ Å
$b = 5.0364\ (10)$ Å
$c = 14.613\ (3)$ Å
$\beta = 98.969\ (3)°$
$V = 1023.6\ (3)$ Å$^3$
$Z = 8$

$F(000) = 464$
$D_x = 1.442$ Mg m$^{-3}$
Melting point: 480 K
Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
Cell parameters from 1751 reflections
$\theta = 2.8–27.0°$
$\mu = 0.11$ mm$^{-1}$
$T = 173$ K
Block, colorless
0.42 × 0.40 × 0.37 mm

*Data collection*

Bruker SMART 1K CCD area-detector diffractometer
Radiation source: fine-focus sealed tube
Graphite monochromator
$\varphi$ and $\omega$ scans

Absorption correction: multi-scan (*SADABS*; Sheldrick, 1996)
$T_{min} = 0.954$, $T_{max} = 0.959$
2277 measured reflections
1006 independent reflections
875 reflections with $I > 2\sigma(I)$

$R_{int}$ = 0.015
$\theta_{max}$ = 26.0°, $\theta_{min}$ = 2.8°
$h$ = −17→13
$k$ = −6→6
$l$ = −14→18

*Refinement*

Refinement on $F^2$
Least-squares matrix: full
$R[F^2 > 2\sigma(F^2)]$ = 0.063
$wR(F^2)$ = 0.191
$S$ = 1.06
1006 reflections
74 parameters
0 restraints
Primary atom site location: structure-invariant direct methods
Secondary atom site location: difference Fourier map
Hydrogen site location: inferred from neighbouring sites
H-atom parameters constrained
$w = 1/[\sigma^2(F_o^2) + (0.1108P)^2 + 3.3345P]$ where $P = (F_o^2 + 2F_c^2)/3$
$(\Delta/\sigma)_{max}$ = 0.001
$\Delta\rho_{max}$ = 0.74 e Å$^{-3}$
$\Delta\rho_{min}$ = −0.73 e Å$^{-3}$

*Special details*

**Geometry**. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement**. Refinement of $F^2$ against ALL reflections. The weighted $R$-factor $wR$ and goodness of fit $S$ are based on $F^2$, conventional $R$-factors $R$ are based on $F$, with $F$ set to zero for negative $F^2$. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating $R$-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. $R$-factors based on $F^2$ are statistically about twice as large as those based on $F$, and $R$- factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å²)*

| | $x$ | $y$ | $z$ | $U_{iso}$*/$U_{eq}$ |
|---|---|---|---|---|
| O1 | 0.12435 (12) | 1.1503 (3) | 0.53422 (12) | 0.0223 (5) |
| C4 | 0.23786 (16) | 0.8483 (5) | 0.61313 (15) | 0.0176 (6) |
| O2 | 0.07382 (14) | 0.7350 (4) | 0.56343 (15) | 0.0373 (6) |
| H2A | 0.0220 | 0.7923 | 0.5336 | 0.056* |
| N1 | 0.31542 (14) | 1.0100 (4) | 0.61094 (15) | 0.0216 (6) |
| H1A | 0.3144 | 1.1614 | 0.5808 | 0.026* |
| C3 | 0.26837 (17) | 0.6325 (5) | 0.66849 (17) | 0.0208 (6) |
| H3 | 0.2299 | 0.4879 | 0.6828 | 0.025* |
| C5 | 0.14189 (16) | 0.9228 (5) | 0.56657 (15) | 0.0173 (6) |
| C2 | 0.36767 (18) | 0.6681 (5) | 0.69974 (17) | 0.0245 (6) |
| H2 | 0.4085 | 0.5521 | 0.7393 | 0.029* |
| C1 | 0.39405 (17) | 0.9010 (6) | 0.66242 (18) | 0.0251 (6) |
| H1 | 0.4570 | 0.9740 | 0.6712 | 0.030* |

*Atomic displacement parameters (Å²)*

| | $U^{11}$ | $U^{22}$ | $U^{33}$ | $U^{12}$ | $U^{13}$ | $U^{23}$ |
|---|---|---|---|---|---|---|
| O1 | 0.0198 (9) | 0.0190 (10) | 0.0273 (10) | −0.0004 (7) | 0.0008 (7) | 0.0048 (7) |
| C4 | 0.0184 (12) | 0.0182 (12) | 0.0167 (11) | −0.0004 (9) | 0.0039 (9) | −0.0005 (9) |
| O2 | 0.0298 (11) | 0.0331 (12) | 0.0472 (13) | −0.0029 (9) | 0.0002 (10) | 0.0044 (10) |
| N1 | 0.0191 (10) | 0.0196 (11) | 0.0253 (11) | −0.0027 (8) | 0.0013 (8) | 0.0062 (8) |
| C3 | 0.0210 (12) | 0.0198 (12) | 0.0216 (12) | 0.0003 (9) | 0.0035 (9) | 0.0020 (9) |

| | | | | | | |
|---|---|---|---|---|---|---|
| C5 | 0.0192 (12) | 0.0164 (11) | 0.0167 (11) | −0.0002 (9) | 0.0042 (9) | −0.0008 (9) |
| C2 | 0.0220 (13) | 0.0291 (14) | 0.0215 (12) | 0.0052 (10) | 0.0009 (9) | 0.0038 (10) |
| C1 | 0.0174 (12) | 0.0318 (14) | 0.0256 (13) | −0.0013 (10) | 0.0019 (9) | 0.0030 (11) |

*Geometric parameters (Å, º)*

| | | | |
|---|---|---|---|
| O1—C5 | 1.250 (3) | N1—H1A | 0.8800 |
| C4—N1 | 1.367 (3) | C3—C2 | 1.413 (3) |
| C4—C3 | 1.383 (3) | C3—H3 | 0.9500 |
| C4—C5 | 1.464 (3) | C2—C1 | 1.369 (4) |
| O2—C5 | 1.342 (3) | C2—H2 | 0.9500 |
| O2—H2A | 0.8400 | C1—H1 | 0.9500 |
| N1—C1 | 1.354 (3) | | |
| | | | |
| N1—C4—C3 | 107.8 (2) | O1—C5—O2 | 122.4 (2) |
| N1—C4—C5 | 121.3 (2) | O1—C5—C4 | 121.6 (2) |
| C3—C4—C5 | 130.8 (2) | O2—C5—C4 | 116.0 (2) |
| C5—O2—H2A | 109.5 | C1—C2—C3 | 107.2 (2) |
| C1—N1—C4 | 109.4 (2) | C1—C2—H2 | 126.4 |
| C1—N1—H1A | 125.3 | C3—C2—H2 | 126.4 |
| C4—N1—H1A | 125.3 | N1—C1—C2 | 108.6 (2) |
| C4—C3—C2 | 106.9 (2) | N1—C1—H1 | 125.7 |
| C4—C3—H3 | 126.5 | C2—C1—H1 | 125.7 |
| C2—C3—H3 | 126.5 | | |
| | | | |
| C3—C4—N1—C1 | 0.7 (3) | N1—C4—C5—O2 | 171.9 (2) |
| C5—C4—N1—C1 | 177.3 (2) | C3—C4—C5—O2 | −12.3 (4) |
| N1—C4—C3—C2 | −0.2 (3) | C4—C3—C2—C1 | −0.3 (3) |
| C5—C4—C3—C2 | −176.4 (2) | C4—N1—C1—C2 | −0.9 (3) |
| N1—C4—C5—O1 | −10.0 (3) | C3—C2—C1—N1 | 0.7 (3) |
| C3—C4—C5—O1 | 165.7 (2) | | |

*Hydrogen-bond geometry (Å, º)*

| *D*—H···*A* | *D*—H | H···*A* | *D*···*A* | *D*—H···*A* |
|---|---|---|---|---|
| N1—H1*A*···O1$^{i}$ | 0.88 | 2.22 | 2.951 (3) | 141 |
| O2—H2*A*···O1$^{ii}$ | 0.84 | 2.16 | 2.986 (3) | 166 |

Symmetry codes: (i) −*x*+1/2, −*y*+5/2, −*z*+1; (ii) −*x*, −*y*+2, −*z*+1.