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## Structure Reports

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## 6-(4-Nitrobenzyloxy)quinoline

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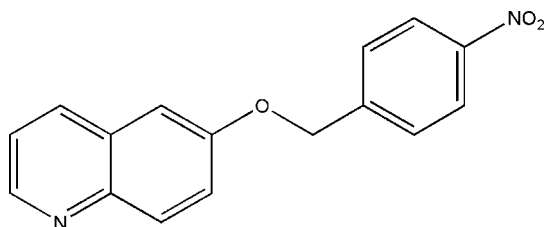
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 Key indicators: single-crystal X-ray study;  $T = 293$  K; mean  $\sigma(\text{C}-\text{C}) = 0.005$  Å;  $R$  factor = 0.065;  $wR$  factor = 0.152; data-to-parameter ratio = 13.8.

In the molecule of the title compound,  $\text{C}_{16}\text{H}_{12}\text{N}_2\text{O}_3$ , the nitrobenzene benzene ring forms a dihedral angle of  $23.8(8)^\circ$  with the plane of the quinoline ring system. The crystal structure is stabilized by an aromatic  $\pi$ - $\pi$  stacking interaction between centrosymmetrically related benzene rings [centroid-centroid distance  $3.663(2)$  Å].

## Related literature

 For related structures, see: Fu & Zhao (2007); Li & Chen (2008); Zhao (2008); Zhao *et al.* (2009).


## Experimental

## Crystal data

$\text{C}_{16}\text{H}_{12}\text{N}_2\text{O}_3$	$V = 1344.3(6)$ Å <sup>3</sup>
$M_r = 280.28$	$Z = 4$
Monoclinic, $P2_1/n$	Mo $K\alpha$ radiation
$a = 12.296(3)$ Å	$\mu = 0.10$ mm <sup>-1</sup>
$b = 8.9146(18)$ Å	$T = 293$ K
$c = 13.559(3)$ Å	$0.20 \times 0.18 \times 0.15$ mm
$\beta = 115.25(3)^\circ$	

## Data collection

Rigaku SCXmini diffractometer	11965 measured reflections
Absorption correction: multi-scan ( <i>CrystalClear</i> ; Rigaku, 2005)	2630 independent reflections
$T_{\min} = 0.976$ , $T_{\max} = 0.981$	1372 reflections with $I > 2\sigma(I)$
	$R_{\text{int}} = 0.079$

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.065$	190 parameters
$wR(F^2) = 0.152$	H-atom parameters constrained
$S = 1.02$	$\Delta\rho_{\text{max}} = 0.15$ e Å <sup>-3</sup>
2630 reflections	$\Delta\rho_{\text{min}} = -0.15$ e Å <sup>-3</sup>

Data collection: *CrystalClear* (Rigaku, 2005); cell refinement: *CrystalClear*; data reduction: *CrystalClear*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL/PC* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL/PC*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: RZ2340).

## References

- Fu, D.-W. & Zhao, H. (2007). *Acta Cryst.* **E63**, o3206.  
 Li, M. & Chen, X. (2008). *Acta Cryst.* **E64**, o2291.  
 Rigaku (2005). *CrystalClear*. Rigaku Corporation, Tokyo, Japan.  
 Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.  
 Zhao, Y.-Y. (2008). *Acta Cryst.* **E64**, o761.  
 Zhao, M. M., Li, Y. H., Wu, D. H. & Wan, Q. (2009). *Acta Cryst.* **E65**, o1261.

## supporting information

*Acta Cryst.* (2009). E65, o1795 [doi:10.1107/S1600536809025823]

**6-(4-Nitrobenzyloxy)quinoline**

**Min Min Zhao, Yong Hua Li and Yuan Zhang**

**S1. Comment**

Recently, we have reported the synthesis and crystal structure of some benzonitrile compounds (Fu & Zhao, 2007; Li & Chen, 2008; Zhao, 2008; Zhao *et al.*, 2009). As an extension of our work on the structural characterization of benzonitrile derivatives, we present herein the synthesis and crystal structure of the title compound.

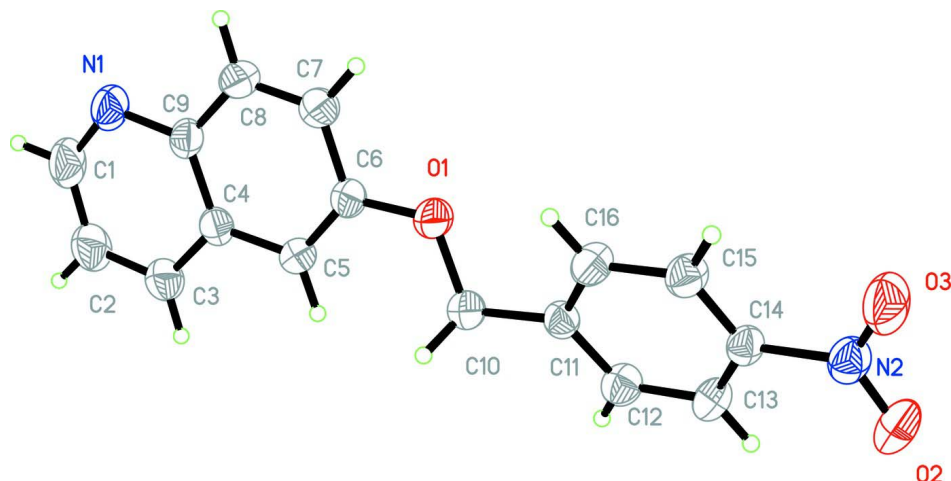
In the molecule of the title compound (Fig. 1), bond lengths and angles are within normal ranges. The C11–C16 benzene ring forms a dihedral angle of 23.8 (8)° with the plane of the quinoline ring system. No intra- or intermolecular hydrogen bonds are observed. The crystal structure is stabilized by an aromatic  $\pi$ - $\pi$  stacking interaction involving centrosymmetrically related benzene rings at (x, y, z) and (1-x, -y, 1-z), with a centroid-centroid distance of 3.663 (2) Å.

**S2. Experimental**

Quinolin-6-ol (1 g, 0.0069 mol) was added to a solution of sodium hydroxide (0.276 g, 0.0069 mol) in methanol (15 ml) and stirred for one hour. Then 4-(bromomethyl)benzonitrile (1.49 g, 0.0069 mol) was added and the mixture stirred at room temperature for 1 day. The title compound was isolated by column chromatography using petroleum ether/ethyl acetate (1:1 v/v) as eluent. Single crystals suitable for X-ray diffraction analysis were obtained by slow evaporation of an ethyl acetate/tetrahydrofuran (3:1 v/v) solution.

**S3. Refinement**

All H atoms were calculated geometrically allowed to ride on their parent atoms, with C—H = 0.93–0.97 Å, and with  $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$ .

**Figure 1**

The molecular structure of the title compound with the atom numbering scheme. Displacement ellipsoids are drawn at the 30% probability level.

### 6-(4-Nitrobenzyloxy)quinoline

#### Crystal data

$C_{16}H_{12}N_2O_3$

$M_r = 280.28$

Monoclinic,  $P2_1/n$

Hall symbol: -P 2yn

$a = 12.296$  (3) Å

$b = 8.9146$  (18) Å

$c = 13.559$  (3) Å

$\beta = 115.25$  (3)°

$V = 1344.3$  (6) Å<sup>3</sup>

$Z = 4$

$F(000) = 584$

$D_x = 1.385$  Mg m<sup>-3</sup>

Mo  $K\alpha$  radiation,  $\lambda = 0.71073$  Å

Cell parameters from 8689 reflections

$\theta = 3.0$ – $27.8$ °

$\mu = 0.10$  mm<sup>-1</sup>

$T = 293$  K

Block, colourless

$0.20 \times 0.18 \times 0.15$  mm

#### Data collection

Rigaku SCXmini  
diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

Detector resolution: 13.6612 pixels mm<sup>-1</sup>

CCD profile fitting scans

Absorption correction: multi-scan

(*CrystalClear*; Rigaku, 2005)

$T_{\min} = 0.976$ ,  $T_{\max} = 0.981$

11965 measured reflections

2630 independent reflections

1372 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.079$

$\theta_{\max} = 26.0$ °,  $\theta_{\min} = 3.3$ °

$h = -15 \rightarrow 15$

$k = -10 \rightarrow 10$

$l = -16 \rightarrow 16$

#### Refinement

Refinement on  $F^2$

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.065$

$wR(F^2) = 0.152$

$S = 1.02$

2630 reflections

190 parameters

0 restraints

Primary atom site location: structure-invariant  
direct methods

Secondary atom site location: difference Fourier  
map

Hydrogen site location: inferred from  
neighbouring sites

H-atom parameters constrained

$$w = 1/[\sigma^2(F_o^2) + (0.0563P)^2 + 0.1856P]$$

where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\max} < 0.001$

$$\Delta\rho_{\max} = 0.15 \text{ e } \text{\AA}^{-3}$$

$$\Delta\rho_{\min} = -0.15 \text{ e } \text{\AA}^{-3}$$

### Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

### Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
C1	0.3786 (4)	0.3414 (4)	-0.1994 (3)	0.0901 (11)
H1A	0.3497	0.3839	-0.2686	0.108*
C2	0.4967 (4)	0.3728 (4)	-0.1275 (3)	0.0881 (11)
H2A	0.5444	0.4336	-0.1487	0.106*
C3	0.5411 (3)	0.3132 (4)	-0.0257 (3)	0.0735 (9)
H3A	0.6202	0.3321	0.0235	0.088*
C4	0.4671 (3)	0.2230 (3)	0.0046 (2)	0.0542 (7)
C5	0.5069 (2)	0.1580 (3)	0.1090 (2)	0.0557 (7)
H5A	0.5847	0.1750	0.1616	0.067*
C6	0.4303 (3)	0.0705 (3)	0.1319 (2)	0.0555 (8)
C7	0.3126 (3)	0.0454 (3)	0.0539 (2)	0.0657 (8)
H7A	0.2613	-0.0147	0.0711	0.079*
C8	0.2730 (3)	0.1074 (3)	-0.0460 (2)	0.0650 (8)
H8A	0.1943	0.0902	-0.0967	0.078*
C9	0.3483 (3)	0.1972 (3)	-0.0747 (2)	0.0572 (8)
C10	0.5763 (2)	0.0164 (3)	0.3124 (2)	0.0597 (8)
H10A	0.5943	0.1212	0.3318	0.072*
H10B	0.6332	-0.0211	0.2861	0.072*
C11	0.5866 (2)	-0.0718 (3)	0.4102 (2)	0.0502 (7)
C12	0.6722 (3)	-0.0323 (3)	0.5119 (2)	0.0620 (8)
H12A	0.7211	0.0504	0.5189	0.074*
C13	0.6870 (3)	-0.1124 (3)	0.6032 (2)	0.0605 (8)
H13A	0.7456	-0.0855	0.6715	0.073*
C14	0.6139 (2)	-0.2321 (3)	0.5916 (2)	0.0504 (7)
C15	0.5270 (3)	-0.2749 (3)	0.4919 (2)	0.0607 (8)
H15A	0.4777	-0.3568	0.4859	0.073*
C16	0.5142 (3)	-0.1945 (3)	0.4008 (2)	0.0627 (8)
H16A	0.4565	-0.2231	0.3325	0.075*
N1	0.3048 (2)	0.2567 (3)	-0.1773 (2)	0.0782 (8)
N2	0.6285 (3)	-0.3179 (3)	0.6883 (2)	0.0655 (7)
O1	0.45682 (16)	0.0016 (2)	0.22964 (15)	0.0699 (6)
O2	0.7111 (2)	-0.2867 (3)	0.77502 (18)	0.0885 (8)

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O3	0.5578 (2)	-0.4184 (3)	0.67829 (19)	0.1026 (9)
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*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
C1	0.105 (3)	0.111 (3)	0.062 (2)	0.029 (3)	0.043 (2)	0.032 (2)
C2	0.096 (3)	0.101 (3)	0.080 (3)	0.005 (2)	0.050 (2)	0.028 (2)
C3	0.074 (2)	0.083 (2)	0.070 (2)	0.0011 (18)	0.0365 (18)	0.0121 (19)
C4	0.0594 (18)	0.0558 (17)	0.0529 (17)	0.0062 (14)	0.0293 (15)	0.0049 (15)
C5	0.0499 (17)	0.0636 (19)	0.0504 (17)	0.0014 (14)	0.0182 (14)	0.0028 (15)
C6	0.0569 (18)	0.0606 (19)	0.0488 (17)	0.0038 (15)	0.0224 (15)	0.0096 (15)
C7	0.0593 (19)	0.075 (2)	0.061 (2)	-0.0085 (16)	0.0246 (17)	0.0075 (17)
C8	0.0535 (18)	0.078 (2)	0.0564 (19)	-0.0016 (16)	0.0167 (16)	-0.0063 (17)
C9	0.0624 (19)	0.064 (2)	0.0458 (17)	0.0152 (16)	0.0241 (15)	0.0030 (15)
C10	0.0556 (18)	0.0685 (19)	0.0551 (17)	-0.0006 (15)	0.0235 (15)	0.0073 (16)
C11	0.0513 (16)	0.0530 (17)	0.0498 (17)	0.0060 (14)	0.0249 (14)	0.0037 (14)
C12	0.065 (2)	0.0585 (19)	0.0581 (19)	-0.0059 (15)	0.0220 (16)	-0.0016 (16)
C13	0.0649 (19)	0.063 (2)	0.0443 (17)	0.0002 (16)	0.0144 (15)	-0.0018 (15)
C14	0.0552 (17)	0.0486 (17)	0.0475 (17)	0.0149 (14)	0.0219 (14)	0.0077 (14)
C15	0.0610 (18)	0.0558 (18)	0.0592 (19)	-0.0042 (14)	0.0196 (16)	0.0100 (16)
C16	0.0626 (19)	0.064 (2)	0.0515 (17)	-0.0022 (16)	0.0150 (15)	0.0042 (16)
N1	0.081 (2)	0.097 (2)	0.0538 (16)	0.0222 (16)	0.0264 (15)	0.0179 (15)
N2	0.0721 (18)	0.0630 (18)	0.0608 (18)	0.0157 (15)	0.0278 (15)	0.0106 (15)
O1	0.0584 (13)	0.0896 (15)	0.0573 (13)	-0.0055 (11)	0.0204 (11)	0.0238 (11)
O2	0.0970 (18)	0.0955 (18)	0.0544 (13)	0.0180 (14)	0.0145 (13)	0.0151 (13)
O3	0.114 (2)	0.106 (2)	0.0789 (17)	-0.0218 (17)	0.0334 (15)	0.0297 (15)

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*Geometric parameters (Å, °)*

C1—N1	1.310 (4)	C10—O1	1.425 (3)
C1—C2	1.389 (4)	C10—C11	1.500 (3)
C1—H1A	0.9300	C10—H10A	0.9700
C2—C3	1.358 (4)	C10—H10B	0.9700
C2—H2A	0.9300	C11—C12	1.377 (4)
C3—C4	1.402 (4)	C11—C16	1.382 (4)
C3—H3A	0.9300	C12—C13	1.373 (4)
C4—C5	1.410 (4)	C12—H12A	0.9300
C4—C9	1.415 (4)	C13—C14	1.360 (4)
C5—C6	1.357 (4)	C13—H13A	0.9300
C5—H5A	0.9300	C14—C15	1.372 (4)
C6—O1	1.368 (3)	C14—N2	1.461 (3)
C6—C7	1.399 (4)	C15—C16	1.378 (4)
C7—C8	1.346 (4)	C15—H15A	0.9300
C7—H7A	0.9300	C16—H16A	0.9300
C8—C9	1.399 (4)	N2—O2	1.214 (3)
C8—H8A	0.9300	N2—O3	1.215 (3)
C9—N1	1.367 (3)		

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N1—C1—C2	125.1 (3)	O1—C10—H10A	110.0
N1—C1—H1A	117.5	C11—C10—H10A	110.0
C2—C1—H1A	117.5	O1—C10—H10B	110.0
C3—C2—C1	118.8 (3)	C11—C10—H10B	110.0
C3—C2—H2A	120.6	H10A—C10—H10B	108.4
C1—C2—H2A	120.6	C12—C11—C16	118.8 (3)
C2—C3—C4	119.5 (3)	C12—C11—C10	119.4 (3)
C2—C3—H3A	120.2	C16—C11—C10	121.7 (3)
C4—C3—H3A	120.2	C13—C12—C11	121.4 (3)
C3—C4—C5	122.6 (3)	C13—C12—H12A	119.3
C3—C4—C9	117.4 (3)	C11—C12—H12A	119.3
C5—C4—C9	120.0 (3)	C14—C13—C12	118.5 (3)
C6—C5—C4	119.2 (3)	C14—C13—H13A	120.8
C6—C5—H5A	120.4	C12—C13—H13A	120.8
C4—C5—H5A	120.4	C13—C14—C15	122.1 (3)
C5—C6—O1	125.3 (3)	C13—C14—N2	119.0 (3)
C5—C6—C7	120.9 (3)	C15—C14—N2	119.0 (3)
O1—C6—C7	113.8 (3)	C14—C15—C16	118.8 (3)
C8—C7—C6	120.6 (3)	C14—C15—H15A	120.6
C8—C7—H7A	119.7	C16—C15—H15A	120.6
C6—C7—H7A	119.7	C15—C16—C11	120.4 (3)
C7—C8—C9	121.1 (3)	C15—C16—H16A	119.8
C7—C8—H8A	119.5	C11—C16—H16A	119.8
C9—C8—H8A	119.5	C1—N1—C9	116.6 (3)
N1—C9—C8	119.2 (3)	O2—N2—O3	122.7 (3)
N1—C9—C4	122.6 (3)	O2—N2—C14	118.8 (3)
C8—C9—C4	118.2 (3)	O3—N2—C14	118.5 (3)
O1—C10—C11	108.5 (2)	C6—O1—C10	117.5 (2)

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