

Acta Crystallographica Section E

## Structure Reports

Online

ISSN 1600-5368

## N-(2-Formylphenyl)benzenesulfonamide

S. Thenmozhi,<sup>a</sup> S. Ranjith,<sup>a</sup> A. SubbiahPandi,<sup>a\*</sup>  
V. Dhayalan<sup>b</sup> and A. K. MohanaKrishnan<sup>b</sup><sup>a</sup>Department of Physics, Presidency College (Autonomous), Chennai 600 005, India, and <sup>b</sup>Department of Organic Chemistry, University of Madras, Guindy Campus, Chennai 600 025, India

Correspondence e-mail: a\_spandian@yahoo.com

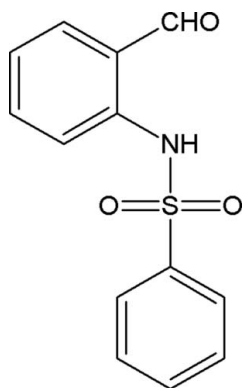
Received 28 July 2009; accepted 17 August 2009

Key indicators: single-crystal X-ray study;  $T = 293$  K; mean  $\sigma(\text{C}-\text{C}) = 0.002$  Å;  $R$  factor = 0.040;  $wR$  factor = 0.119; data-to-parameter ratio = 23.7.

In the title compound,  $\text{C}_{13}\text{H}_{11}\text{NO}_3\text{S}$ , the two aromatic rings are oriented at an angle of  $88.18(8)^\circ$ . Intramolecular  $\text{N}-\text{H}\cdots\text{O}$  and  $\text{C}-\text{H}\cdots\text{O}$  hydrogen bonds are observed, each of which generates an  $S(6)$  ring motif. In the crystal, molecules are linked into  $C(7)$  chains along  $[010]$  by intermolecular  $\text{C}-\text{H}\cdots\text{O}$  hydrogen bonds. The structure is further stabilized by intermolecular  $\text{C}-\text{H}\cdots\pi$  interactions involving the sulfonyl-bound phenyl ring.

## Related literature

For the biological activity of sulfonamides, see: Zareef *et al.* (2007); Chohan *et al.* (2007); Brown (1971); Pomarnacka & Kozlarska-Kedra (2003); Sethu Sankar *et al.* (2002). For related structures, see: Bassindale (1984); Cotton & Stokley (1970); Usha *et al.* (2005); Zhu *et al.* (2008). For hydrogen-bond motifs, see: Bernstein *et al.* (1995).



## Experimental

## Crystal data

$\text{C}_{13}\text{H}_{11}\text{NO}_3\text{S}$	$a = 7.7656(2)$ Å
$M_r = 261.29$	$b = 9.0080(2)$ Å
Triclinic, $P\bar{1}$	$c = 9.5855(2)$ Å

$\alpha = 86.293(1)^\circ$   
 $\beta = 77.912(1)^\circ$   
 $\gamma = 68.826(1)^\circ$   
 $V = 611.35(2)$  Å<sup>3</sup>  
 $Z = 2$

Mo  $K\alpha$  radiation  
 $\mu = 0.26$  mm<sup>-1</sup>  
 $T = 293$  K  
 $0.21 \times 0.19 \times 0.17$  mm

## Data collection

Bruker Kappa APEXII CCD diffractometer  
 Absorption correction: multi-scan (SADABS; Sheldrick, 1996)  
 $T_{\min} = 0.768$ ,  $T_{\max} = 0.956$

15490 measured reflections  
 3960 independent reflections  
 3228 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.022$

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.040$   
 $wR(F^2) = 0.119$   
 $S = 1.02$   
 3960 reflections  
 167 parameters

H atoms treated by a mixture of independent and constrained refinement  
 $\Delta\rho_{\text{max}} = 0.31$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.31$  e Å<sup>-3</sup>

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{N1}-\text{H1}\cdots\text{O3}$	0.80 (2)	1.99 (2)	2.6751 (19)	144 (2)
$\text{C2}-\text{H2}\cdots\text{O1}$	0.93	2.46	3.0879 (18)	125
$\text{C3}-\text{H3}\cdots\text{O2}^i$	0.93	2.56	3.2691 (19)	133
$\text{C5}-\text{H5}\cdots\text{Cg1}^{ii}$	0.93	2.80	3.700 (2)	162

Symmetry codes: (i)  $x, y + 1, z$ ; (ii)  $-x + 1, -y + 2, -z + 1$ . Cg1 is the centroid of the C8–C13 ring.

Data collection: APEX2 (Bruker, 2004); cell refinement: SAINT (Bruker, 2004); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: ORTEP-3 (Farrugia, 1997); software used to prepare material for publication: SHELXL97 and PLATON (Spek, 2009).

ST and AS thank Dr Babu Varghese, SAIF, IIT-Madras, Chennai, India, for his help with the data collection.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: CI2869).

## References

- Bassindale, A. (1984). *The Third Dimension in Organic Chemistry*, ch. 1, p. 11. New York: John Wiley and Sons.
- Bernstein, J., Davis, R. E., Shimoni, L. & Chang, N.-L. (1995). *Angew. Chem. Int. Ed. Engl.* **34**, 1555–1573.
- Brown, G. M. (1971). *Adv. Biochem.* **35**, 35–40.
- Bruker (2004). SMART and SAINT. Bruker AXS Inc., Madison, Wisconsin, USA.
- Chohan, Z. H. & Shad, H. A. (2007). *J. Enz. Inhib. Med. Chem.* **23**, 369–379.
- Cotton, F. A. & Stokley, P. F. (1970). *J. Chem. Soc.* **92**, 294–302.
- Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.
- Pomarnacka, E. & Kozlarska-Kedra, I. (2003). *Farmaco*, **58**, 423–429.
- Sethu Sankar, K., Kannadasan, S., Velmurugan, D., Srinivasan, P. C., Shanmuga Sundara Raj, S. & Fun, H.-K. (2002). *Acta Cryst.* **C58**, o277–o279.
- Sheldrick, G. M. (1996). SADABS. University of Göttingen, Germany.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Spek, A. L. (2009). *Acta Cryst.* **D65**, 148–155.
- Usha, G., Selvanayagam, S., Velmurugan, D., Ravikumar, K., Jaisankar, P. & Srinivasan, P. C. (2005). *Acta Cryst.* **E61**, o1916–o1918.
- Zareef, M., Iqbal, R., De Dominguez, N. G., Rodrigues, J., Zaidi, J. H., Arfan, M. & Supuran, C. T. (2007). *J. Enz. Inhib. Med. Chem.* **22**, 301–308.
- Zhu, H.-Y., Wu, Z. & Jiang, S. (2008). *Acta Cryst.* **E64**, o596.

## supporting information

*Acta Cryst.* (2009). E65, o2210 [doi:10.1107/S1600536809032681]

## ***N*-[2-(Formylphenyl)]benzenesulfonamide**

**S. Thenmozhi, S. Ranjith, A. SubbiahPandi, V. Dhayalan and A. K. MohanaKrishnan**

### **S1. Comment**

Sulfonamides have been recognized for their wide variety of pharmacological activities such as antibacterial, antitumor, anti-carbonic anhydrase, diuretic, hypoglycaemic, antithyroid and protease inhibitory activities. Sulfonamides particularly sulfadiazine and sulfadoxine have been used clinically as antimalarial agents (Zareef *et al.*, 2007). Due to their significant pharmacological applications and widespread use in medicine, these compounds have also gained attention in bioinorganic and metal-based drug chemistry (Chohan *et al.*, 2007). Sulfonamide derivatives are well known drugs and are used to control diseases caused by bacterial infections. The antibacterial action of this group of drugs is exerted by the complete inhibition of dihydropteroate synthase enzyme towards the *p*-amino benzonate (Brown, 1971). Benzene sulfonamide derivatives are known to exhibit anticancer and HIV activities (Pomarnacka & Kozlarska-Kedra, 2003). The sulfonamides inhibit the growth of bacterial organism and are also useful for treating urinary and gastrointestinal infections (Sethu Sankar *et al.*, 2002). In view of this medicinal importance, the crystal structure determination of the title compound (Fig.1) was carried out and the results are presented here.

Atom S1 has a distorted tetrahedral configuration. The widening of angle O1—S1—O2 [119.76 (7)°] and narrowing of angle C8—S1—N1 [106.08 (6)°] from the ideal tetrahedral value are attributed to the Thorpe-Ingold effect (Bassindale, 1984). The two aromatic rings are oriented at an angle of 88.18 (8)°. The angles around atom S1 deviate significantly from the regular tetrahedral value, with the largest deviation of 119.7 (7)° for O1—S1—O2 angle. This may be due to non-bonding interactions between S—O bonds (Cotton & Stokley, 1970). The aldehyde group is coplanar with the benzene ring, as evidenced by the torsion angle C1—C6—C7—O3 of -2.2 (3)°. The geometrical parameters agree well with those reported for related sulfonamide structures (Usha *et al.*, 2005; Zhu *et al.*, 2008). Intramolecular N1—H1···O3 and C2—H2···O1 hydrogen bonds are observed, each of which generates an S(6) ring motif (Bernstein *et al.*, 1995).

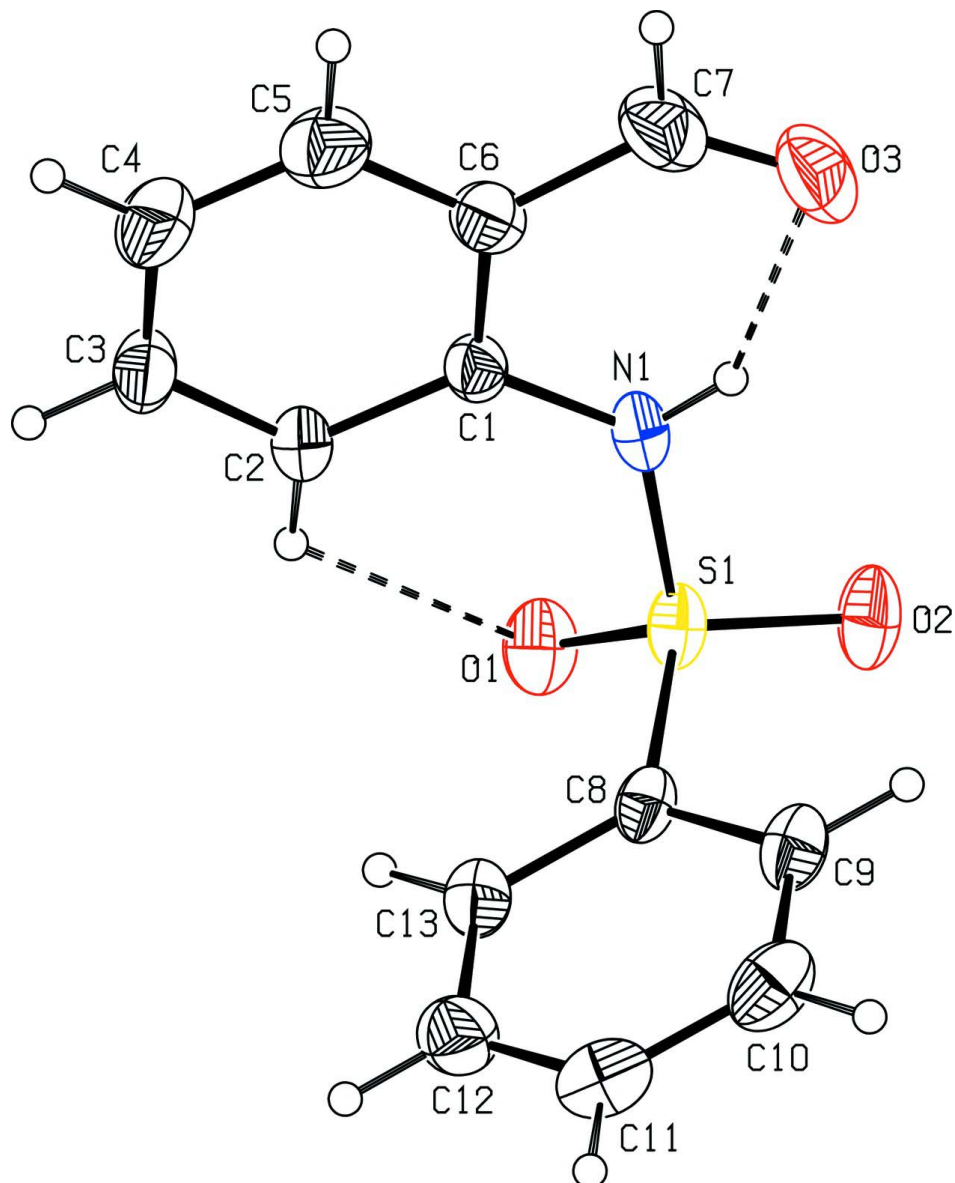
Intermolecular C—H···O hydrogen bonds involving atoms C3 and O2 link molecules into C(7) chains running along the [010] (Fig. 2). The crystal packing is further stabilized by C—H··· $\pi$  interactions involving the sulfonyl-bound phenyl ring.

### **S2. Experimental**

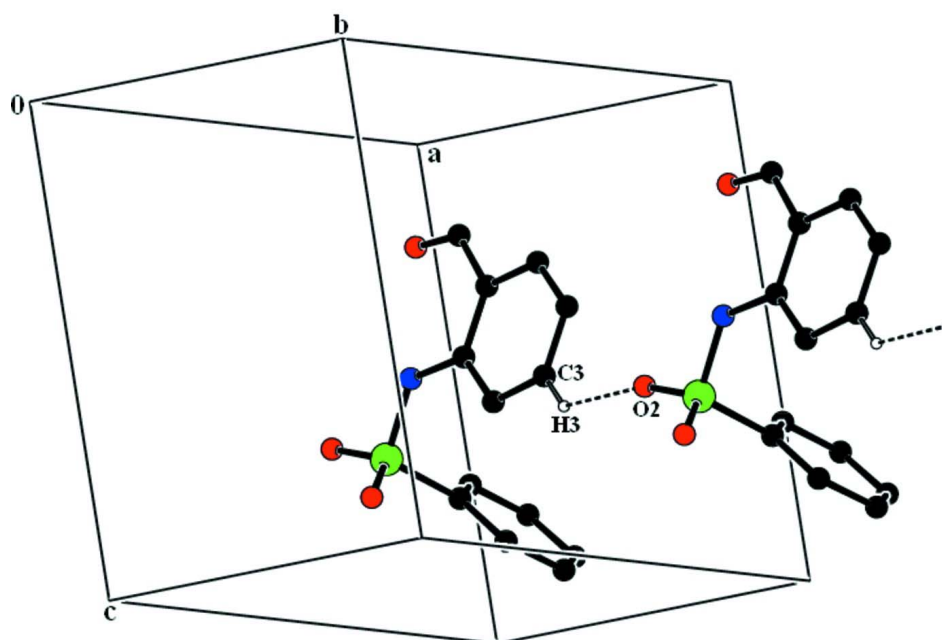
*N*-[2-(Hydroxymethyl)phenyl]benzenesulfonamide (2 g, 7.6 mmol) was added to a solution of pyridinium chlorochromate (3.25 g, 15.11 mmol) in dry dichloromethane (20 ml) at room temperature and the reaction mixture was stirred for 4 h. The solvent was removed to obtain a crude aldehyde as a white solid. Single crystals of the title compound suitable for X-ray diffraction were obtained by slow evaporation of a methanol solution.

### **S3. Refinement**

Atom H1 was located in a difference map and refined freely. All other H atoms were positioned geometrically and constrained to ride on their parent atoms, with C-H = 0.93 Å and  $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$ .

**Figure 1**

The molecular structure of the title compound, showing 30% probability displacement ellipsoids and the atomic numbering scheme. Hydrogen bonds are shown as dashed lines.

**Figure 2**

Part of a C(7) chain in the title compound.

### *N*-(2-Formylphenyl)benzenesulfonamide

#### Crystal data

$C_{13}H_{11}NO_3S$

$M_r = 261.29$

Triclinic,  $P\bar{1}$

Hall symbol:  $-P\ 1$

$a = 7.7656$  (2) Å

$b = 9.0080$  (2) Å

$c = 9.5855$  (2) Å

$\alpha = 86.293$  (1)°

$\beta = 77.912$  (1)°

$\gamma = 68.826$  (1)°

$V = 611.35$  (2) Å<sup>3</sup>

$Z = 2$

$F(000) = 272$

$D_x = 1.419$  Mg m<sup>-3</sup>

Mo  $K\alpha$  radiation,  $\lambda = 0.71073$  Å

Cell parameters from 3960 reflections

$\theta = 2.2$ – $31.3$ °

$\mu = 0.26$  mm<sup>-1</sup>

$T = 293$  K

Block, colourless

$0.21 \times 0.19 \times 0.17$  mm

#### Data collection

Bruker Kappa APEXII CCD  
diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

$\omega$  scans

Absorption correction: multi-scan  
(*SADABS*; Sheldrick, 1996)

$T_{\min} = 0.768$ ,  $T_{\max} = 0.956$

15490 measured reflections

3960 independent reflections

3228 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.022$

$\theta_{\max} = 31.3$ °,  $\theta_{\min} = 2.2$ °

$h = -11 \rightarrow 11$

$k = -13 \rightarrow 13$

$l = -13 \rightarrow 13$

Refinement

Refinement on  $F^2$   
 Least-squares matrix: full  
 $R[F^2 > 2\sigma(F^2)] = 0.040$   
 $wR(F^2) = 0.119$   
 $S = 1.02$   
 3960 reflections  
 167 parameters  
 0 restraints  
 Primary atom site location: structure-invariant  
 direct methods

Secondary atom site location: difference Fourier  
 map  
 Hydrogen site location: inferred from  
 neighbouring sites  
 H atoms treated by a mixture of independent  
 and constrained refinement  
 $w = 1/[\sigma^2(F_o^2) + (0.0582P)^2 + 0.1353P]$   
 where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\max} = 0.001$   
 $\Delta\rho_{\max} = 0.31 \text{ e } \text{\AA}^{-3}$   
 $\Delta\rho_{\min} = -0.31 \text{ e } \text{\AA}^{-3}$

Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
C1	0.19429 (17)	0.98956 (14)	0.61966 (13)	0.0409 (2)
C2	0.1595 (2)	1.10035 (15)	0.72614 (15)	0.0470 (3)
H2	0.1207	1.0779	0.8210	0.056*
C3	0.1827 (2)	1.24378 (17)	0.69056 (18)	0.0563 (3)
H3	0.1595	1.3170	0.7626	0.068*
C4	0.2392 (3)	1.2813 (2)	0.5513 (2)	0.0695 (5)
H4	0.2547	1.3783	0.5292	0.083*
C5	0.2722 (3)	1.1734 (2)	0.44619 (19)	0.0675 (4)
H5	0.3085	1.1987	0.3517	0.081*
C6	0.2529 (2)	1.02670 (18)	0.47720 (15)	0.0503 (3)
C7	0.2926 (3)	0.9189 (3)	0.35857 (18)	0.0692 (5)
H7	0.3264	0.9554	0.2677	0.083*
C8	0.36455 (18)	0.70690 (14)	0.85278 (13)	0.0413 (3)
C9	0.5123 (2)	0.57064 (16)	0.79634 (16)	0.0519 (3)
H9	0.4963	0.5055	0.7325	0.062*
C10	0.6827 (2)	0.5337 (2)	0.83658 (19)	0.0628 (4)
H10	0.7829	0.4425	0.7999	0.075*
C11	0.7062 (2)	0.6304 (2)	0.93052 (19)	0.0621 (4)
H11	0.8225	0.6047	0.9563	0.074*
C12	0.5587 (2)	0.7649 (2)	0.98675 (18)	0.0599 (4)
H12	0.5758	0.8296	1.0503	0.072*
C13	0.3853 (2)	0.80416 (16)	0.94909 (15)	0.0487 (3)
H13	0.2847	0.8941	0.9877	0.058*
N1	0.16942 (19)	0.84378 (14)	0.64708 (14)	0.0509 (3)

O1	0.00479 (14)	0.85915 (13)	0.90061 (12)	0.0576 (3)
O2	0.13222 (17)	0.60899 (13)	0.76162 (14)	0.0649 (3)
O3	0.2861 (2)	0.78653 (18)	0.36620 (14)	0.0795 (4)
S1	0.14948 (5)	0.75289 (4)	0.79852 (4)	0.04575 (11)
H1	0.199 (3)	0.790 (2)	0.5767 (19)	0.064 (5)*

*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
C1	0.0388 (6)	0.0369 (5)	0.0461 (6)	-0.0096 (4)	-0.0123 (5)	-0.0046 (5)
C2	0.0562 (7)	0.0365 (6)	0.0482 (7)	-0.0149 (5)	-0.0113 (6)	-0.0055 (5)
C3	0.0638 (9)	0.0377 (6)	0.0699 (9)	-0.0173 (6)	-0.0180 (7)	-0.0056 (6)
C4	0.0839 (12)	0.0498 (8)	0.0803 (11)	-0.0310 (8)	-0.0182 (10)	0.0112 (8)
C5	0.0769 (11)	0.0679 (10)	0.0581 (9)	-0.0294 (9)	-0.0116 (8)	0.0132 (8)
C6	0.0475 (7)	0.0533 (8)	0.0471 (7)	-0.0125 (6)	-0.0117 (6)	-0.0036 (6)
C7	0.0711 (10)	0.0802 (12)	0.0485 (8)	-0.0171 (9)	-0.0081 (7)	-0.0160 (8)
C8	0.0431 (6)	0.0316 (5)	0.0466 (6)	-0.0155 (4)	0.0004 (5)	0.0003 (4)
C9	0.0505 (7)	0.0380 (6)	0.0589 (8)	-0.0127 (5)	0.0027 (6)	-0.0044 (5)
C10	0.0474 (7)	0.0506 (8)	0.0725 (10)	-0.0056 (6)	0.0036 (7)	0.0057 (7)
C11	0.0481 (7)	0.0680 (10)	0.0680 (9)	-0.0209 (7)	-0.0118 (7)	0.0184 (8)
C12	0.0638 (9)	0.0618 (9)	0.0604 (9)	-0.0276 (7)	-0.0175 (7)	0.0038 (7)
C13	0.0518 (7)	0.0414 (6)	0.0511 (7)	-0.0160 (5)	-0.0062 (6)	-0.0039 (5)
N1	0.0662 (7)	0.0391 (5)	0.0508 (6)	-0.0199 (5)	-0.0134 (5)	-0.0108 (5)
O1	0.0440 (5)	0.0534 (6)	0.0689 (7)	-0.0164 (4)	0.0051 (5)	-0.0129 (5)
O2	0.0657 (7)	0.0436 (5)	0.0938 (8)	-0.0315 (5)	-0.0089 (6)	-0.0106 (5)
O3	0.0860 (9)	0.0802 (9)	0.0672 (8)	-0.0192 (7)	-0.0129 (7)	-0.0333 (7)
S1	0.04451 (18)	0.03489 (16)	0.0591 (2)	-0.01869 (12)	-0.00239 (14)	-0.00769 (12)

*Geometric parameters (Å, °)*

C1—C2	1.3900 (17)	C8—C9	1.3871 (18)
C1—N1	1.3957 (17)	C8—S1	1.7515 (14)
C1—C6	1.4052 (19)	C9—C10	1.375 (2)
C2—C3	1.3797 (19)	C9—H9	0.93
C2—H2	0.93	C10—C11	1.374 (3)
C3—C4	1.375 (2)	C10—H10	0.93
C3—H3	0.93	C11—C12	1.376 (2)
C4—C5	1.365 (3)	C11—H11	0.93
C4—H4	0.93	C12—C13	1.383 (2)
C5—C6	1.390 (2)	C12—H12	0.93
C5—H5	0.93	C13—H13	0.93
C6—C7	1.453 (2)	N1—S1	1.6265 (13)
C7—O3	1.208 (2)	N1—H1	0.798 (18)
C7—H7	0.93	O1—S1	1.4212 (10)
C8—C13	1.3838 (18)	O2—S1	1.4241 (10)
C2—C1—N1	123.02 (12)	C10—C9—C8	118.75 (15)
C2—C1—C6	118.86 (12)	C10—C9—H9	120.6

N1—C1—C6	118.10 (12)	C8—C9—H9	120.6
C3—C2—C1	119.78 (14)	C11—C10—C9	120.54 (15)
C3—C2—H2	120.1	C11—C10—H10	119.7
C1—C2—H2	120.1	C9—C10—H10	119.7
C4—C3—C2	121.69 (15)	C10—C11—C12	120.40 (15)
C4—C3—H3	119.2	C10—C11—H11	119.8
C2—C3—H3	119.2	C12—C11—H11	119.8
C5—C4—C3	118.78 (15)	C11—C12—C13	120.29 (16)
C5—C4—H4	120.6	C11—C12—H12	119.9
C3—C4—H4	120.6	C13—C12—H12	119.9
C4—C5—C6	121.51 (16)	C12—C13—C8	118.66 (14)
C4—C5—H5	119.2	C12—C13—H13	120.7
C6—C5—H5	119.2	C8—C13—H13	120.7
C5—C6—C1	119.37 (14)	C1—N1—S1	128.49 (9)
C5—C6—C7	117.69 (16)	C1—N1—H1	112.3 (14)
C1—C6—C7	122.94 (15)	S1—N1—H1	116.5 (13)
O3—C7—C6	126.49 (17)	O1—S1—O2	119.76 (7)
O3—C7—H7	116.8	O1—S1—N1	109.04 (7)
C6—C7—H7	116.8	O2—S1—N1	103.70 (7)
C13—C8—C9	121.35 (13)	O1—S1—C8	108.57 (7)
C13—C8—S1	120.66 (10)	O2—S1—C8	108.83 (7)
C9—C8—S1	117.98 (11)	N1—S1—C8	106.08 (6)
N1—C1—C2—C3	-178.37 (13)	C9—C10—C11—C12	0.6 (2)
C6—C1—C2—C3	0.1 (2)	C10—C11—C12—C13	-0.1 (2)
C1—C2—C3—C4	0.2 (2)	C11—C12—C13—C8	-0.9 (2)
C2—C3—C4—C5	0.3 (3)	C9—C8—C13—C12	1.3 (2)
C3—C4—C5—C6	-1.1 (3)	S1—C8—C13—C12	-178.76 (11)
C4—C5—C6—C1	1.4 (3)	C2—C1—N1—S1	-16.6 (2)
C4—C5—C6—C7	-179.47 (18)	C6—C1—N1—S1	164.92 (11)
C2—C1—C6—C5	-0.8 (2)	C1—N1—S1—O1	52.78 (14)
N1—C1—C6—C5	177.67 (14)	C1—N1—S1—O2	-178.56 (12)
C2—C1—C6—C7	-179.97 (14)	C1—N1—S1—C8	-63.97 (14)
N1—C1—C6—C7	-1.4 (2)	C13—C8—S1—O1	-19.79 (13)
C5—C6—C7—O3	178.68 (18)	C9—C8—S1—O1	160.20 (10)
C1—C6—C7—O3	-2.2 (3)	C13—C8—S1—O2	-151.70 (11)
C13—C8—C9—C10	-0.7 (2)	C9—C8—S1—O2	28.29 (13)
S1—C8—C9—C10	179.30 (11)	C13—C8—S1—N1	97.27 (11)
C8—C9—C10—C11	-0.2 (2)	C9—C8—S1—N1	-82.73 (11)

*Hydrogen-bond geometry* (Å, °)

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
N1—H1...O3	0.80 (2)	1.99 (2)	2.6751 (19)	144 (2)
C2—H2...O1	0.93	2.46	3.0879 (18)	125

---

C3—H3···O2 <sup>i</sup>	0.93	2.56	3.2691 (19)	133
C5—H5···Cg1 <sup>ii</sup>	0.93	2.80	3.700 (2)	162

---

Symmetry codes: (i)  $x, y+1, z$ ; (ii)  $-x+1, -y+2, -z+1$ .