

Diaquabis(4-methylbenzoato- κ O)bis-(nicotinamide- κ N¹)manganese(II)

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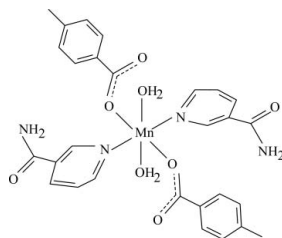
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Key indicators: single-crystal X-ray study; $T = 100$ K; mean $\sigma(\text{C}-\text{C}) = 0.002$ Å; R factor = 0.035; wR factor = 0.099; data-to-parameter ratio = 16.2.

In the mononuclear title complex, $[\text{Mn}(\text{C}_8\text{H}_7\text{O}_2)_2(\text{C}_6\text{H}_6\text{N}_2\text{O})_2(\text{H}_2\text{O})_2]$, the Mn^{II} ion is located on a crystallographic inversion center. The asymmetric unit contains one 4-methylbenzoate anion, one nicotinamide (NA) ligand and one coordinated water molecule. The four O atoms in the equatorial plane around the Mn^{II} ion form a slightly distorted square-planar arrangement, while the slightly distorted octahedral coordination is completed by the two pyridine N atoms of the NA ligands in the axial positions. The dihedral angle between the carboxylate group and the attached benzene ring is $9.01(7)^\circ$, while the pyridine and benzene rings are oriented at a dihedral angle of $42.44(5)^\circ$. In the crystal structure, intermolecular $\text{O}-\text{H}\cdots\text{O}$, $\text{N}-\text{H}\cdots\text{O}$ and $\text{C}-\text{H}\cdots\text{O}$ hydrogen bonds, and $\text{O}-\text{H}\cdots\pi$ and $\text{C}-\text{H}\cdots\pi$ interactions link the molecules into a two-dimensional network parallel to (001).

Related literature

For niacin, see: Krishnamachari (1974), and for the nicotinic acid derivative N,N -diethylnicotinamide, see: Bigoli *et al.* (1972). For related structures, see: Hökelek *et al.* (1996, 2009*a,b,c*); Hökelek & Necefoğlu (1998); Necefoğlu *et al.* (2010).



Experimental

Crystal data

$[\text{Mn}(\text{C}_8\text{H}_7\text{O}_2)_2(\text{C}_6\text{H}_6\text{N}_2\text{O})_2(\text{H}_2\text{O})_2]$
 $M_r = 605.51$
 Triclinic, $P\bar{1}$
 $a = 7.3289(2)$ Å
 $b = 10.1768(3)$ Å
 $c = 10.6292(3)$ Å
 $\alpha = 66.852(2)^\circ$
 $\beta = 78.232(4)^\circ$
 $\gamma = 70.206(3)^\circ$
 $V = 683.58(4)$ Å³
 $Z = 1$
 Mo $K\alpha$ radiation
 $\mu = 0.54$ mm⁻¹
 $T = 100$ K
 $0.38 \times 0.25 \times 0.19$ mm

Data collection

Bruker Kappa APEXII CCD area-detector diffractometer
 Absorption correction: multi-scan (SADABS; Bruker, 2005)
 $T_{\text{min}} = 0.649$, $T_{\text{max}} = 0.698$
 11619 measured reflections
 3297 independent reflections
 3022 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.028$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.035$
 $wR(F^2) = 0.099$
 $S = 1.08$
 3297 reflections
 204 parameters
 3 restraints
 H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{\text{max}} = 0.73$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.38$ e Å⁻³

Table 1

Selected bond lengths (Å).

Mn1—O2	2.1036 (11)	Mn1—N1	2.2947 (13)
Mn1—O4	2.1924 (12)		

Table 2

Hydrogen-bond geometry (Å, °).

Cg1 and Cg2 are the centroids of the C2-C7 and N1/C9-C13 rings, respectively.

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
N2—H21 \cdots O1 ⁱ	0.88 (3)	1.97 (3)	2.8456 (19)	174 (2)
N2—H22 \cdots O3 ⁱⁱ	0.86 (3)	2.09 (3)	2.952 (2)	172 (3)
O4—H42 \cdots O3 ⁱⁱⁱ	0.90 (3)	1.83 (3)	2.7071 (19)	164 (3)
C11—H11 \cdots O1 ⁱ	0.93	2.33	3.200 (2)	156
O4—H41 \cdots Cg1 ^{iv}	0.91 (2)	2.33 (2)	3.141 (2)	149 (3)
C4—H4 \cdots Cg2 ^{iv}	0.93	2.80	3.490 (2)	132

Symmetry codes: (i) $-x + 1, -y - 1, -z$; (ii) $-x + 2, -y - 1, -z$; (iii) $-x + 1, -y, -z$; (iv) $-x + 1, -y + 2, -z$.

Data collection: APEX2 (Bruker, 2007); cell refinement: SAINT (Bruker, 2007); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: ORTEP-3 for Windows (Farrugia, 1997); software used to prepare material for publication: WinGX (Farrugia, 1999) and PLATON (Spek, 2009).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: CI5069).

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supporting information

Acta Cryst. (2010). E66, m490–m491 [https://doi.org/10.1107/S1600536810011815]

Diaquabis(4-methylbenzoato- κ O)bis(nicotinamide- κ N¹)manganese(II)**Hacali Necefoğlu, Efdal Çimen, Barış Tercan, Hakan Dal and Tuncer Hökelek****S1. Comment**

As a part of our ongoing investigation on transition metal complexes of nicotinamide (NA), one form of niacin (Krishnamachari, 1974), and/or the nicotinic acid derivative *N,N*-diethylnicotinamide (DENA), an important respiratory stimulant (Bigoli *et al.*, 1972), the title compound was synthesized and its crystal structure is reported herein.

The title compound, (I), is a centrosymmetric mononuclear complex, consisting of two nicotinamide (NA) and two 4-methylbenzoate (PMB) ligands and two coordinated water molecules; the Mn^{II} ion lies on a centre of symmetry (Fig. 1). The crystal structures of some PMB and/or NA complexes of Cu^{II}, Co^{II}, Ni^{II}, Mn^{II} and Zn^{II} ions, [Cu(C₇H₅O₂)₂(C₁₀H₁₄N₂O)₂], (II) (Hökelek *et al.*, 1996), [Co(C₆H₆N₂O)₂(C₇H₄NO₄)₂(H₂O)₂], (III) (Hökelek & Necefoğlu, 1998), [Ni(C₇H₄ClO₂)₂(C₆H₆N₂O)₂(H₂O)₂], (IV) (Hökelek *et al.*, 2009a), [Ni(C₈H₇O₂)₂(C₆H₆N₂O)₂(H₂O)₂], (V) (Necefoğlu *et al.*, 2010), [Mn(C₇H₄ClO₂)₂(C₁₀H₁₄N₂O)₂(H₂O)₂], (VI) (Hökelek *et al.*, 2009b) and [Zn(C₇H₄BrO₂)₂(C₆H₆N₂O)₂(H₂O)₂], (VII) (Hökelek *et al.*, 2009c) have also been reported. In (II), the two benzoate ions are coordinated to the Cu atom as bidentate ligands, while in other structures all ligands being monodentate.

In the title compound, all ligands are monodentate. The four O atoms (O2, O4, and the symmetry-related atoms, O2', O4') in the equatorial plane around the Mn^{II} atom form a slightly distorted square-planar arrangement, while the slightly distorted octahedral coordination is completed by the two N atoms of the NA ligands (N1, N1') in the axial positions (Fig. 1). The near equality of the C1—O1 [1.2466 (19) Å] and C1—O2 [1.2690 (18) Å] bonds in the carboxylate group indicates a delocalized bonding arrangement, rather than localized single and double bonds. The average Mn—O bond length is 2.148 (1) Å (Table 1). The dihedral angle between the planar carboxylate group (O1/O2/C1) and the benzene ring A (C2—C7) is 9.01 (7)°, while that between rings A and B (N1/C9—C13) is 42.44 (5)°.

In the crystal structure, intermolecular O—H \cdots O, N—H \cdots O and C—H \cdots O hydrogen bonds (Table 2) link the molecules into a two-dimensional network parallel to the (001). Weak O—H \cdots π and C—H \cdots π interactions involving the pyridine and benzene rings are also observed (Table 2).

S2. Experimental

The title compound was prepared by the reaction of MnSO₄·H₂O (0.84 g, 5 mmol) in H₂O (10 ml) and NA (1.22 g, 10 mmol) in H₂O (10 ml) with sodium 4-methylbenzoate (1.58 g, 10 mmol) in H₂O (150 ml). The mixture was filtered and set aside to crystallize at ambient temperature for one month, giving colourless single crystals.

S3. Refinement

H atoms of the NH₂ group (H21 and H22) and water molecules (H41 and H42) were located in a difference Fourier map and refined isotropically; the water H atoms were refined with O—H and H \cdots H distance restraints. The remaining H atoms were positioned geometrically with C—H = 0.93 and 0.96 Å, for aromatic and methyl H atoms, respectively, and constrained to ride on their parent atoms, with U_{iso}(H) = xU_{eq}(C), where x = 1.5 for methyl H and x = 1.2 for aromatic H

atoms.

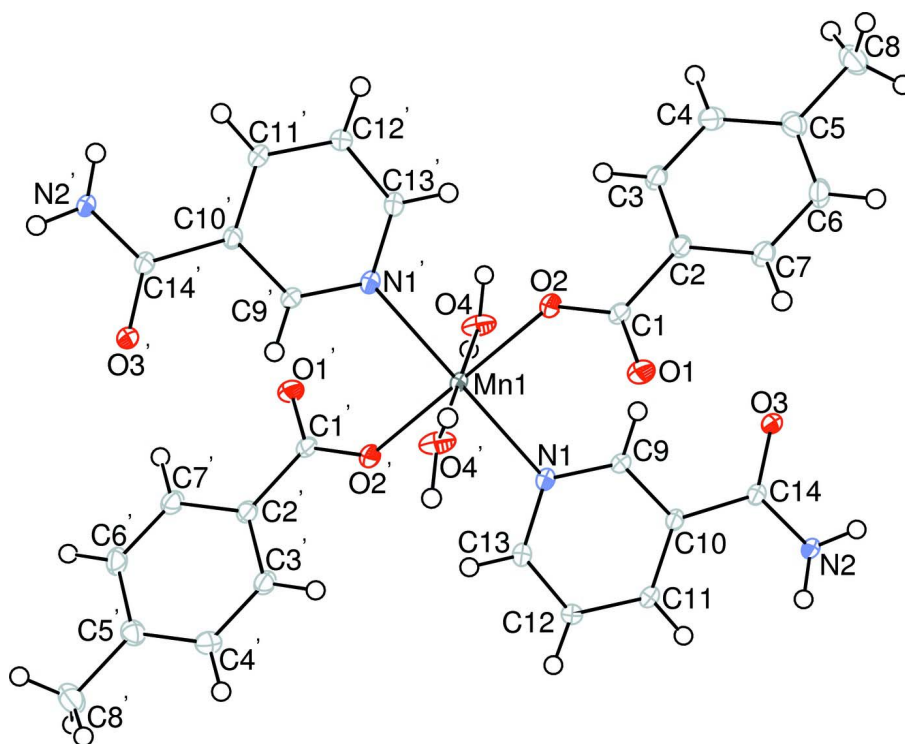


Figure 1

The molecular structure of the title compound, with the atom-numbering scheme. Displacement ellipsoids are drawn at the 50% probability level. Primed atoms are generated by the symmetry operator: (') $-x, -y, -z$.

Diaquabis(4-methylbenzoato- κ O)bis(nicotinamide- κ N¹)manganese(II)

Crystal data

$[\text{Mn}(\text{C}_8\text{H}_7\text{O}_2)_2(\text{C}_6\text{H}_6\text{N}_2\text{O})_2(\text{H}_2\text{O})_2]$

$M_r = 605.51$

Triclinic, $P\bar{1}$

Hall symbol: $-P\ 1$

$a = 7.3289$ (2) Å

$b = 10.1768$ (3) Å

$c = 10.6292$ (3) Å

$\alpha = 66.852$ (2)°

$\beta = 78.232$ (4)°

$\gamma = 70.206$ (3)°

$V = 683.58$ (4) Å³

$Z = 1$

$F(000) = 315$

$D_x = 1.471$ Mg m⁻³

Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 8087 reflections

$\theta = 2.5\text{--}28.4^\circ$

$\mu = 0.54$ mm⁻¹

$T = 100$ K

Block, colourless

$0.38 \times 0.25 \times 0.19$ mm

Data collection

Bruker Kappa APEXII CCD area-detector diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

φ and ω scans

Absorption correction: multi-scan

(*SADABS*; Bruker, 2005)

$T_{\min} = 0.649$, $T_{\max} = 0.698$

11619 measured reflections

3297 independent reflections

3022 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.028$

$\theta_{\max} = 28.4^\circ$, $\theta_{\min} = 2.1^\circ$

$h = -9 \rightarrow 9$

$k = -11 \rightarrow 13$

$l = -14 \rightarrow 14$

Refinement

Refinement on F^2
 Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.035$
 $wR(F^2) = 0.099$
 $S = 1.08$
 3297 reflections
 204 parameters
 3 restraints
 Primary atom site location: structure-invariant
 direct methods

Secondary atom site location: difference Fourier
 map
 Hydrogen site location: inferred from
 neighbouring sites
 H atoms treated by a mixture of independent
 and constrained refinement
 $w = 1/[\sigma^2(F_o^2) + (0.058P)^2 + 0.3059P]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} = 0.001$
 $\Delta\rho_{\max} = 0.73 \text{ e } \text{\AA}^{-3}$
 $\Delta\rho_{\min} = -0.38 \text{ e } \text{\AA}^{-3}$

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) etc. and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
Mn1	0.0000	0.0000	0.0000	0.01276 (11)
O1	0.24237 (17)	-0.33733 (12)	0.25988 (12)	0.0207 (2)
O2	0.16077 (17)	-0.08763 (12)	0.17238 (12)	0.0201 (2)
O3	0.78404 (16)	-0.34130 (13)	-0.01469 (13)	0.0210 (3)
O4	0.18323 (19)	0.15102 (14)	-0.09400 (13)	0.0232 (3)
H41	0.208 (4)	0.190 (3)	-0.1860 (17)	0.064 (9)*
H42	0.218 (4)	0.205 (3)	-0.058 (2)	0.042 (7)*
N1	0.21166 (18)	-0.15932 (14)	-0.10409 (13)	0.0143 (3)
N2	0.84692 (19)	-0.50868 (14)	-0.11834 (14)	0.0167 (3)
H21	0.810 (3)	-0.555 (3)	-0.159 (2)	0.028 (5)*
H22	0.958 (4)	-0.545 (3)	-0.084 (2)	0.032 (6)*
C1	0.2620 (2)	-0.21470 (17)	0.24423 (15)	0.0151 (3)
C2	0.4196 (2)	-0.21314 (16)	0.31505 (15)	0.0145 (3)
C3	0.4331 (2)	-0.07942 (16)	0.31453 (15)	0.0164 (3)
H3	0.3426	0.0099	0.2714	0.020*
C4	0.5805 (2)	-0.07856 (18)	0.37770 (16)	0.0181 (3)
H4	0.5871	0.0115	0.3767	0.022*
C5	0.7191 (2)	-0.21079 (19)	0.44279 (16)	0.0190 (3)
C6	0.7063 (2)	-0.34413 (18)	0.44183 (16)	0.0203 (3)
H6	0.7979	-0.4333	0.4839	0.024*
C7	0.5589 (2)	-0.34546 (17)	0.37902 (16)	0.0177 (3)
H7	0.5528	-0.4355	0.3795	0.021*
C8	0.8783 (3)	-0.2097 (2)	0.51214 (18)	0.0257 (4)
H8A	0.8947	-0.1115	0.4763	0.039*

H8B	0.9977	-0.2797	0.4951	0.039*
H8C	0.8437	-0.2371	0.6092	0.039*
C9	0.3975 (2)	-0.22336 (16)	-0.07444 (15)	0.0132 (3)
H9	0.4418	-0.1971	-0.0142	0.016*
C10	0.5280 (2)	-0.32707 (16)	-0.12893 (15)	0.0134 (3)
C11	0.4623 (2)	-0.36415 (17)	-0.22143 (16)	0.0170 (3)
H11	0.5453	-0.4328	-0.2602	0.020*
C12	0.2710 (2)	-0.29647 (18)	-0.25420 (17)	0.0198 (3)
H12	0.2239	-0.3182	-0.3164	0.024*
C13	0.1508 (2)	-0.19621 (17)	-0.19343 (16)	0.0172 (3)
H13	0.0221	-0.1523	-0.2153	0.021*
C14	0.7300 (2)	-0.39401 (16)	-0.08365 (15)	0.0142 (3)

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Mn1	0.01065 (17)	0.01054 (16)	0.01760 (17)	0.00071 (11)	-0.00334 (11)	-0.00758 (12)
O1	0.0252 (6)	0.0149 (5)	0.0241 (6)	-0.0030 (4)	-0.0062 (5)	-0.0092 (5)
O2	0.0216 (6)	0.0138 (5)	0.0243 (6)	0.0035 (4)	-0.0106 (5)	-0.0092 (4)
O3	0.0160 (5)	0.0184 (5)	0.0343 (6)	0.0021 (4)	-0.0096 (5)	-0.0174 (5)
O4	0.0308 (7)	0.0236 (6)	0.0233 (6)	-0.0146 (5)	0.0056 (5)	-0.0147 (5)
N1	0.0126 (6)	0.0117 (6)	0.0177 (6)	-0.0003 (4)	-0.0030 (5)	-0.0061 (5)
N2	0.0120 (6)	0.0146 (6)	0.0258 (7)	0.0020 (5)	-0.0057 (5)	-0.0125 (5)
C1	0.0149 (7)	0.0150 (7)	0.0155 (7)	0.0004 (5)	-0.0019 (5)	-0.0091 (6)
C2	0.0147 (7)	0.0137 (7)	0.0140 (6)	-0.0008 (5)	-0.0017 (5)	-0.0061 (5)
C3	0.0175 (7)	0.0129 (7)	0.0164 (7)	-0.0004 (5)	-0.0029 (6)	-0.0055 (6)
C4	0.0209 (8)	0.0181 (7)	0.0173 (7)	-0.0059 (6)	-0.0014 (6)	-0.0080 (6)
C5	0.0170 (7)	0.0245 (8)	0.0158 (7)	-0.0045 (6)	-0.0017 (6)	-0.0083 (6)
C6	0.0182 (7)	0.0177 (7)	0.0196 (7)	0.0026 (6)	-0.0066 (6)	-0.0053 (6)
C7	0.0203 (7)	0.0124 (7)	0.0186 (7)	-0.0006 (5)	-0.0041 (6)	-0.0056 (6)
C8	0.0199 (8)	0.0352 (10)	0.0240 (8)	-0.0067 (7)	-0.0061 (6)	-0.0114 (7)
C9	0.0137 (7)	0.0105 (6)	0.0160 (6)	-0.0017 (5)	-0.0027 (5)	-0.0060 (5)
C10	0.0118 (7)	0.0101 (6)	0.0174 (7)	-0.0007 (5)	-0.0031 (5)	-0.0051 (5)
C11	0.0146 (7)	0.0149 (7)	0.0224 (7)	0.0015 (5)	-0.0042 (6)	-0.0110 (6)
C12	0.0181 (8)	0.0194 (7)	0.0258 (8)	0.0016 (6)	-0.0092 (6)	-0.0142 (7)
C13	0.0137 (7)	0.0150 (7)	0.0229 (7)	0.0008 (5)	-0.0059 (6)	-0.0087 (6)
C14	0.0127 (7)	0.0113 (6)	0.0184 (7)	-0.0007 (5)	-0.0033 (5)	-0.0063 (5)

Geometric parameters (Å, °)

Mn1—O2	2.1036 (11)	C4—C3	1.387 (2)
Mn1—O2 ⁱ	2.1036 (11)	C4—C5	1.396 (2)
Mn1—O4	2.1924 (12)	C4—H4	0.93
Mn1—O4 ⁱ	2.1924 (12)	C5—C8	1.507 (2)
Mn1—N1	2.2947 (13)	C6—C5	1.396 (2)
Mn1—N1 ⁱ	2.2947 (13)	C6—C7	1.387 (2)
O1—C1	1.2466 (19)	C6—H6	0.93
O2—C1	1.2690 (18)	C7—H7	0.93

O3—C14	1.2464 (18)	C8—H8A	0.96
O4—H41	0.905 (16)	C8—H8B	0.96
O4—H42	0.906 (16)	C8—H8C	0.96
N1—C9	1.3386 (18)	C9—C10	1.391 (2)
N1—C13	1.345 (2)	C9—H9	0.93
N2—C14	1.3283 (19)	C11—C10	1.396 (2)
N2—H21	0.89 (2)	C11—C12	1.385 (2)
N2—H22	0.86 (2)	C11—H11	0.93
C1—C2	1.509 (2)	C12—H12	0.93
C2—C3	1.395 (2)	C13—C12	1.383 (2)
C2—C7	1.397 (2)	C13—H13	0.93
C3—H3	0.93	C14—C10	1.4956 (19)
O2—Mn1—O2 ⁱ	180.00 (4)	C3—C4—H4	119.5
O2—Mn1—O4	86.18 (5)	C5—C4—H4	119.5
O2 ⁱ —Mn1—O4	93.82 (5)	C4—C5—C8	120.94 (15)
O2—Mn1—O4 ⁱ	93.82 (5)	C6—C5—C4	118.20 (14)
O2 ⁱ —Mn1—O4 ⁱ	86.18 (5)	C6—C5—C8	120.86 (15)
O2—Mn1—N1 ⁱ	86.02 (4)	C5—C6—H6	119.5
O2 ⁱ —Mn1—N1 ⁱ	93.98 (4)	C7—C6—C5	120.91 (14)
O2—Mn1—N1	93.98 (4)	C7—C6—H6	119.5
O2 ⁱ —Mn1—N1	86.02 (4)	C2—C7—H7	119.6
O4—Mn1—O4 ⁱ	180.00 (5)	C6—C7—C2	120.74 (14)
O4—Mn1—N1 ⁱ	91.32 (4)	C6—C7—H7	119.6
O4 ⁱ —Mn1—N1 ⁱ	88.68 (4)	C5—C8—H8A	109.5
O4—Mn1—N1	88.68 (4)	C5—C8—H8B	109.5
O4 ⁱ —Mn1—N1	91.32 (4)	C5—C8—H8C	109.5
N1 ⁱ —Mn1—N1	180.00 (9)	H8A—C8—H8B	109.5
C1—O2—Mn1	136.86 (10)	H8A—C8—H8C	109.5
Mn1—O4—H41	122.2 (18)	H8B—C8—H8C	109.5
Mn1—O4—H42	129.2 (16)	N1—C9—C10	123.52 (13)
H42—O4—H41	107 (2)	N1—C9—H9	118.2
C9—N1—Mn1	121.52 (10)	C10—C9—H9	118.2
C9—N1—C13	117.56 (13)	C9—C10—C11	118.15 (13)
C13—N1—Mn1	120.88 (10)	C9—C10—C14	118.12 (13)
C14—N2—H21	124.4 (15)	C11—C10—C14	123.72 (13)
C14—N2—H22	113.3 (16)	C10—C11—H11	120.7
H21—N2—H22	121 (2)	C12—C11—C10	118.59 (14)
O1—C1—O2	125.89 (14)	C12—C11—H11	120.7
O1—C1—C2	118.59 (13)	C11—C12—H12	120.4
O2—C1—C2	115.51 (13)	C13—C12—C11	119.26 (14)
C3—C2—C1	120.90 (13)	C13—C12—H12	120.4
C3—C2—C7	118.51 (14)	N1—C13—C12	122.90 (14)
C7—C2—C1	120.56 (13)	N1—C13—H13	118.6
C2—C3—H3	119.7	C12—C13—H13	118.6
C4—C3—C2	120.59 (14)	O3—C14—N2	121.84 (14)
C4—C3—H3	119.7	O3—C14—C10	119.46 (13)
C3—C4—C5	121.05 (14)	N2—C14—C10	118.71 (13)

O4—Mn1—O2—C1	-124.12 (16)	O2—C1—C2—C7	-169.99 (14)
O4 ⁱ —Mn1—O2—C1	55.88 (16)	C1—C2—C3—C4	-178.83 (14)
N1 ⁱ —Mn1—O2—C1	144.28 (16)	C7—C2—C3—C4	-0.8 (2)
N1—Mn1—O2—C1	-35.72 (16)	C1—C2—C7—C6	178.70 (15)
O2—Mn1—N1—C9	-23.31 (12)	C3—C2—C7—C6	0.6 (2)
O2 ⁱ —Mn1—N1—C9	156.69 (12)	C5—C4—C3—C2	0.2 (2)
O2—Mn1—N1—C13	154.04 (12)	C3—C4—C5—C6	0.4 (2)
O2 ⁱ —Mn1—N1—C13	-25.96 (12)	C3—C4—C5—C8	-179.60 (15)
O4—Mn1—N1—C9	62.76 (11)	C5—C6—C7—C2	0.0 (2)
O4 ⁱ —Mn1—N1—C9	-117.24 (11)	C7—C6—C5—C4	-0.6 (2)
O4—Mn1—N1—C13	-119.89 (12)	C7—C6—C5—C8	179.46 (15)
O4 ⁱ —Mn1—N1—C13	60.11 (12)	N1—C9—C10—C11	1.2 (2)
Mn1—O2—C1—O1	-26.6 (3)	N1—C9—C10—C14	-177.98 (13)
Mn1—O2—C1—C2	153.01 (12)	C12—C11—C10—C9	-0.1 (2)
Mn1—N1—C9—C10	176.18 (11)	C12—C11—C10—C14	179.01 (15)
C13—N1—C9—C10	-1.3 (2)	C10—C11—C12—C13	-0.8 (2)
Mn1—N1—C13—C12	-177.22 (13)	N1—C13—C12—C11	0.8 (3)
C9—N1—C13—C12	0.2 (2)	O3—C14—C10—C9	-9.9 (2)
O1—C1—C2—C3	-172.31 (14)	O3—C14—C10—C11	170.95 (15)
O1—C1—C2—C7	9.7 (2)	N2—C14—C10—C9	170.04 (14)
O2—C1—C2—C3	8.0 (2)	N2—C14—C10—C11	-9.1 (2)

Symmetry code: (i) $-x, -y, -z$.

Hydrogen-bond geometry (\AA , $^\circ$)

Cg1 and Cg2 are the centroids of the C2-C7 and N1/C9-C13 rings, respectively.

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
N2—H21 \cdots O1 ⁱⁱ	0.88 (3)	1.97 (3)	2.8456 (19)	174 (2)
N2—H22 \cdots O3 ⁱⁱⁱ	0.86 (3)	2.09 (3)	2.952 (2)	172 (3)
O4—H42 \cdots O3 ^{iv}	0.90 (3)	1.83 (3)	2.7071 (19)	164 (3)
C11—H11 \cdots O1 ⁱⁱ	0.93	2.33	3.200 (2)	156
O4—H41 \cdots Cg1 ^v	0.91 (2)	2.33 (2)	3.141 (2)	149 (3)
C4—H4 \cdots Cg2 ^v	0.93	2.80	3.490 (2)	132

Symmetry codes: (ii) $-x+1, -y-1, -z$; (iii) $-x+2, -y-1, -z$; (iv) $-x+1, -y, -z$; (v) $-x+1, -y+2, -z$.