

Guanidinium 2-phenylacetate

 Graham Smith^{a*} and Urs D. Wermuth^b

^aFaculty of Science and Technology, Queensland University of Technology, GPO Box 2434, Brisbane, Queensland 4001, Australia, and ^bSchool of Biomolecular and Physical Sciences, Griffith University, Nathan, Queensland 4111, Australia
Correspondence e-mail: g.smith@qut.edu.au

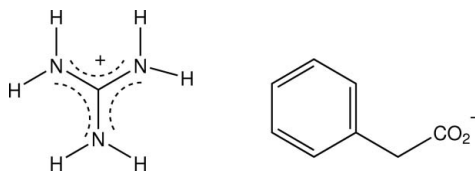
Received 25 May 2010; accepted 1 July 2010

Key indicators: single-crystal X-ray study; $T = 200$ K; mean $\sigma(\text{C}-\text{C}) = 0.003$ Å; R factor = 0.040; wR factor = 0.101; data-to-parameter ratio = 14.5.

In the structure of the title salt, $\text{CH}_6\text{N}_3^+\cdot\text{C}_8\text{H}_7\text{O}_2^-$, the guanidinium cation gives three cyclic hydrogen-bonding interactions with O-atom acceptors of three independent phenylacetate anions, one $R_2^2(8)$ and two $R_2^1(6)$, giving one-dimensional columnar structures which extend down the 4_2 axis in the tetragonal cell. Within these structures, there are solvent-accessible voids of volume 86.5 Å³.

Related literature

For the structures of simple monocyclic aromatic guanidinium carboxylates, see: Pereira Silva *et al.* (2007, 2010); Kleb *et al.* (1998); Smith & Wermuth (2010). For graph-set analysis, see: Etter *et al.* (1990).



Experimental

Crystal data

$\text{CH}_6\text{N}_3^+\cdot\text{C}_8\text{H}_7\text{O}_2^-$
 $M_r = 195.22$
Tetragonal, $P4_2/n$
 $a = 16.8418$ (10) Å
 $c = 7.8372$ (6) Å
 $V = 2223.0$ (3) Å³

$Z = 8$
Mo $K\alpha$ radiation
 $\mu = 0.09$ mm⁻¹
 $T = 200$ K
 $0.30 \times 0.25 \times 0.20$ mm

Data collection

Oxford Diffraction Gemini-S CCD-detector diffractometer
2191 independent reflections
1430 reflections with $I > 2\sigma(I)$
7477 measured reflections
 $R_{\text{int}} = 0.027$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.040$
 $wR(F^2) = 0.101$
 $S = 0.93$
2191 reflections
151 parameters
H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{\text{max}} = 0.14$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.14$ e Å⁻³

Table 1
Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{N1G}-\text{H11G}\cdots\text{O22}^{\text{i}}$	0.86 (2)	2.02 (2)	2.876 (2)	173.9 (15)
$\text{N1G}-\text{H12G}\cdots\text{O21}$	0.866 (16)	2.123 (17)	2.900 (2)	149.0 (15)
$\text{N2G}-\text{H21G}\cdots\text{O22}^{\text{ii}}$	0.834 (16)	2.219 (17)	2.9625 (19)	148.5 (15)
$\text{N2G}-\text{H22G}\cdots\text{O21}^{\text{i}}$	0.86 (2)	1.97 (2)	2.827 (2)	172.6 (15)
$\text{N3G}-\text{H31G}\cdots\text{O21}$	0.897 (16)	2.068 (16)	2.8634 (17)	147.2 (13)
$\text{N3G}-\text{H32G}\cdots\text{O22}^{\text{ii}}$	0.859 (16)	2.073 (16)	2.8520 (17)	150.5 (15)

Symmetry codes: (i) $y + \frac{1}{2}, -x + 1, z + \frac{1}{2}$; (ii) $x, y, z + 1$.

Data collection: *CrysAlis PRO* (Oxford Diffraction, 2009); cell refinement: *CrysAlis PRO*; data reduction: *CrysAlis PRO*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *PLATON* (Spek, 2009); software used to prepare material for publication: *PLATON*.

The authors acknowledge financial support from the Australian Research Council, the Faculty of Science and Technology, Queensland University of Technology and the School of Biomolecular and Physical Sciences, Griffith University.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BV2147).

References

- Etter, M. C., MacDonald, J. C. & Bernstein, J. (1990). *Acta Cryst.* **B46**, 256–262.
Kleb, D.-C., Schürmann, M., Preut, H. & Bleckmann, P. (1998). *Z. Kristallogr. New Cryst. Struct.* pp. 581–582.
Oxford Diffraction (2009). *CrysAlis PRO*. Oxford Diffraction Ltd, Yarnton, Oxfordshire, England.
Pereira Silva, P. S., Ramos Silva, M., Paixão, J. A. & Matos Beja, A. (2007). *Acta Cryst.* **E63**, o2783.
Pereira Silva, P. S., Ramos Silva, M., Paixão, J. A. & Matos Beja, A. (2010). *Acta Cryst.* **E66**, o524.
Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
Smith, G. & Wermuth, U. D. (2010). *Acta Cryst.* **E66**, o1946.
Spek, A. L. (2009). *Acta Cryst.* **D65**, 148–155.

supporting information

Acta Cryst. (2010). E66, o1947 [https://doi.org/10.1107/S1600536810025821]

Guanidinium 2-phenylacetate

Graham Smith and Urs D. Wermuth

S1. Comment

The known structures of the guanidinium salts of simple monocyclic aromatic carboxylic acids consist of those with benzoic acid (Pereira Silva *et al.*, 2007), 4-aminobenzoic acid (Pereira Silva *et al.*, 2010), 4-nitrobenzoic acid (Kleb *et al.*, 1998) and 3-nitrobenzoic acid (Smith & Wermuth, 2010). In these anhydrous structures the guanidinium cation is usually involved in cyclic hydrogen-bonding associations through $N-H\cdots O_{\text{carboxyl}}$ links [graph sets $R_2^2(8)$ or $R_2^1(6)$ (Etter *et al.*, 1990)] giving most commonly three-dimensional structures. The structure of the guanidinium salt of phenylacetic acid had not been previously reported so we carried out the 2:1 stoichiometric reaction of phenylacetic acid with guanidinium carbonate in aqueous ethanol solution, providing colourless crystals of the title compound, $\text{CH}_6\text{N}_3^+ \text{C}_8\text{H}_7\text{O}_2^-$ (I) when recrystallized from water.

In the structure of (I) (Figs. 1, 2), each guanidinium cation is involved in three cyclic hydrogen-bonding interactions with the carboxyl O-acceptors of three independent phenylacetate anions, one $R_2^2(8)$ and two $R_2^1(6)$. These result in un-associated one-dimensional columnar structures which extend down the 4_2 (c) axis in the tetragonal cell (Fig. 3). Within these columnar structures there are 86.5 \AA^3 solvent accessible voids which are large enough to accommodate water molecules but surprisingly do not, despite the sample having been obtained by recrystallization from water.

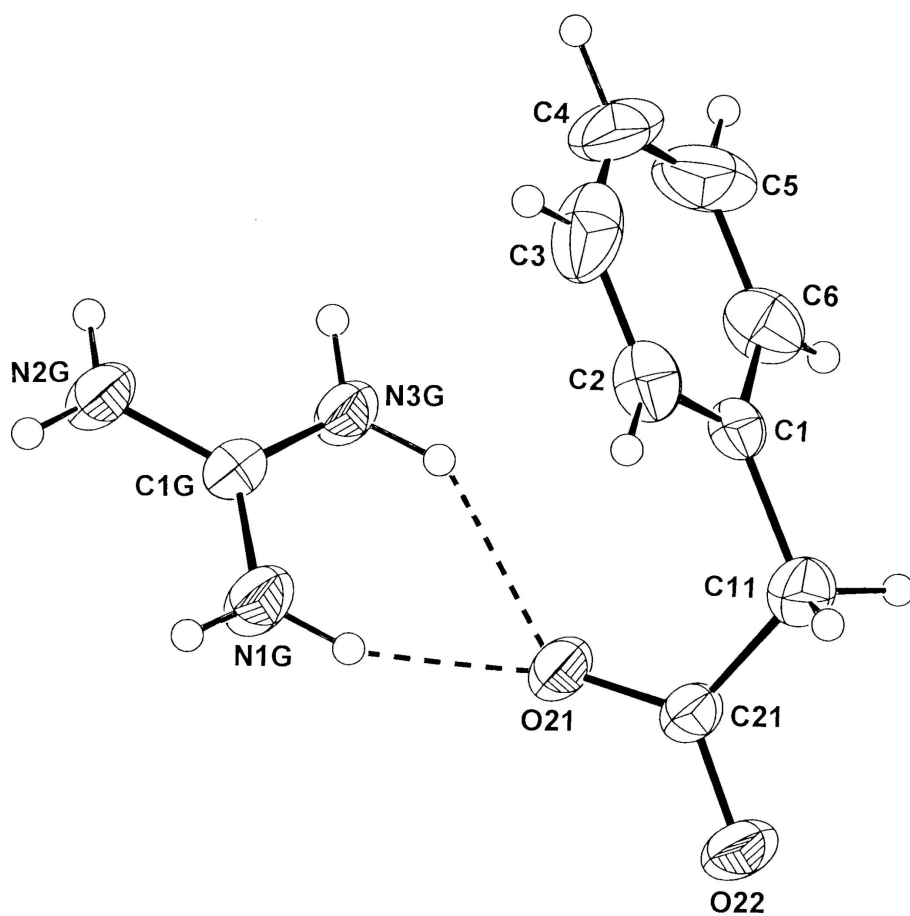
With the anion, the acetate substituent is close to normal to the plane of the benzene ring [torsion angle C2–C1–C11–C21, $86.98(18)^\circ$]. Present in the benzene ring are unexplained high unidirectional displacement parameters for three atoms [C3, C4, C5: U_{11} , 0.1009 (18), 0.185 (3), 0.1019 (18) \AA^2 respectively, *cf.* a typical value 0.0427 (9) \AA^2 for C2].

S2. Experimental

The title compound was synthesized by heating together under reflux for 10 minutes 1 mmol of phenylacetic acid and 0.5 mmol of guanidinium carbonate in 50 ml of 50% ethanol-water. After concentration to *ca* 30 ml, room temperature evaporation of the hot-filtered solution gave a colourless powder which was recrystallized from a minimum volume of water, giving on total evaporation, crystal plates of (I) (m.p. 443 K), from which a specimen suitable for X-ray analysis was cleaved.

S3. Refinement

Hydrogen atoms involved in hydrogen-bonding interactions were located by difference methods and their positional and isotropic displacement parameters were refined. The H atoms were included in the refinement in calculated positions ($C-H_{\text{aromatic}} = 0.93 \text{ \AA}$ and $C-H_{\text{aliphatic}} = 0.97 \text{ \AA}$) and treated as riding, with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$.

**Figure 1**

Molecular configuration and atom naming scheme for the guanidinium cation and the phenylacetate anion in (I). Inter-species hydrogen bonds are shown as dashed lines. Displacement ellipsoids are drawn at the 40% probability level.

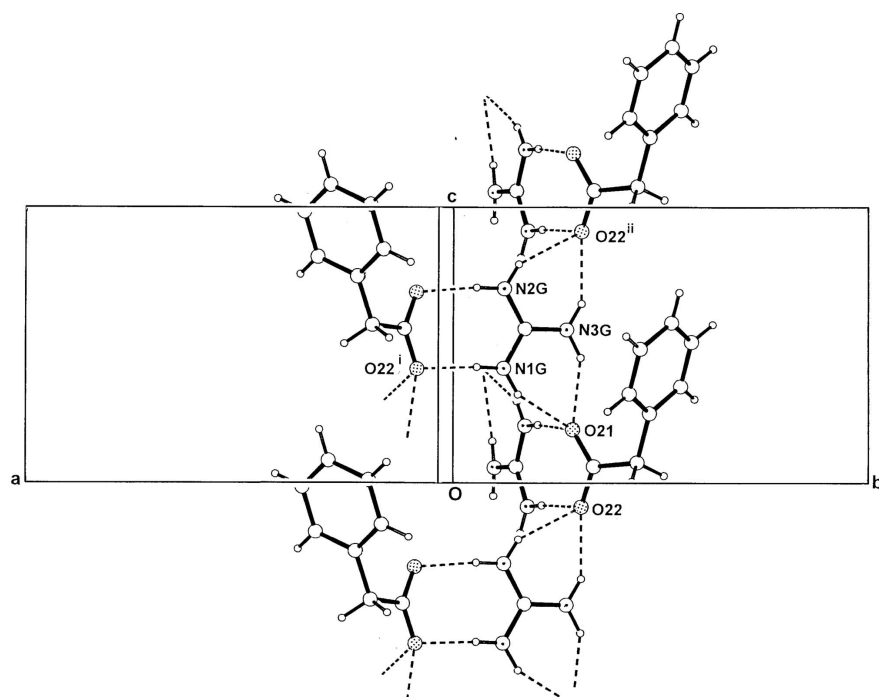


Figure 2

The hydrogen-bonding extensions of the basic asymmetric unit in the structure of (I), showing the three cyclic cation–anion hydrogen-bonding associations as dashed lines. Non-associative hydrogen atoms are deleted. For symmetry codes, see Table 1.

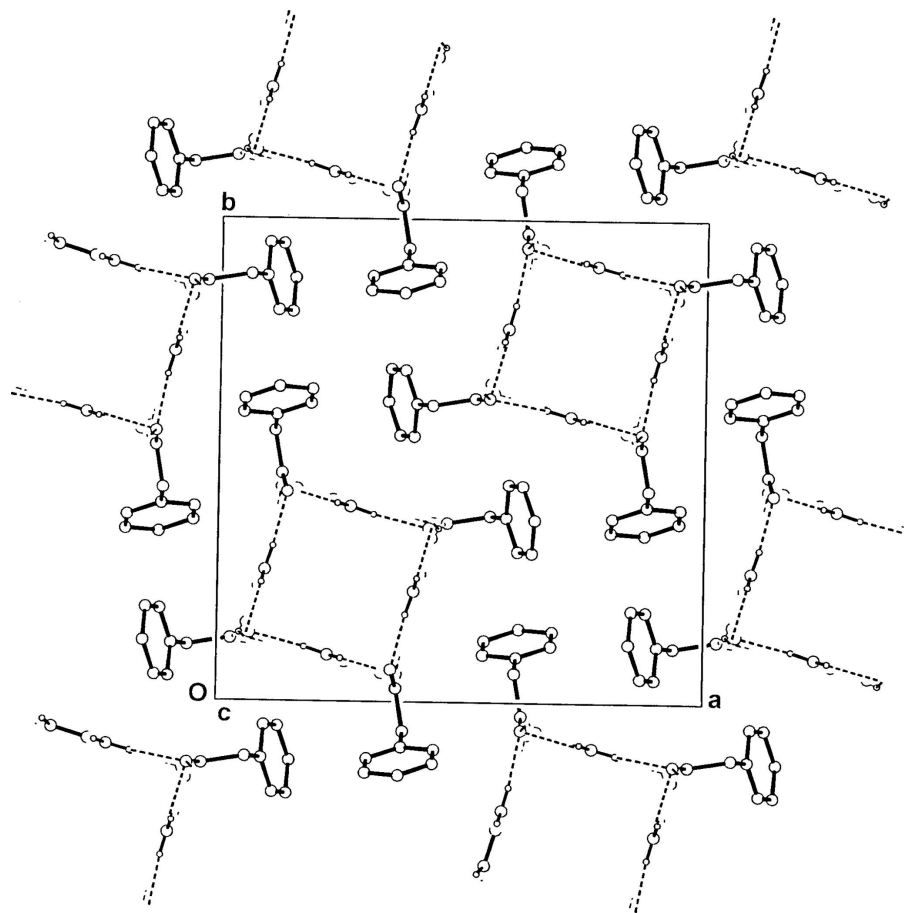


Figure 3

The hydrogen-bonded columnar structures of (I) viewed down the *c* axial direction of the tetragonal unit cell. Non-associative hydrogen atoms are deleted.

Guanidinium 2-phenylacetate

Crystal data

$\text{CH}_6\text{N}_3^+ \cdot \text{C}_8\text{H}_7\text{O}_2^-$

$M_r = 195.22$

Tetragonal, $P4_2/n$

Hall symbol: $-P\ 4bc$

$a = 16.8418\ (10)\ \text{\AA}$

$c = 7.8372\ (6)\ \text{\AA}$

$V = 2223.0\ (3)\ \text{\AA}^3$

$Z = 8$

$F(000) = 832$

$D_x = 1.167\ \text{Mg m}^{-3}$

Melting point: 443 K

Mo $K\alpha$ radiation, $\lambda = 0.71073\ \text{\AA}$

Cell parameters from 2510 reflections

$\theta = 3.1\text{--}28.6^\circ$

$\mu = 0.09\ \text{mm}^{-1}$

$T = 200\ \text{K}$

Block, colourless

$0.30 \times 0.25 \times 0.20\ \text{mm}$

Data collection

Oxford Diffraction Gemini-S CCD-detector
diffractometer

Radiation source: Enhance (Mo) X-ray source

Graphite monochromator

ω scans

7477 measured reflections

2191 independent reflections

1430 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.027$

$\theta_{\text{max}} = 26.0^\circ$, $\theta_{\text{min}} = 3.1^\circ$

$h = -20 \rightarrow 18$
 $k = -10 \rightarrow 20$

$l = -9 \rightarrow 8$

Refinement

Refinement on F^2
 Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.040$
 $wR(F^2) = 0.101$
 $S = 0.93$
 2191 reflections
 151 parameters
 0 restraints
 Primary atom site location: structure-invariant
 direct methods

Secondary atom site location: difference Fourier
 map
 Hydrogen site location: inferred from
 neighbouring sites
 H atoms treated by a mixture of independent
 and constrained refinement
 $w = 1/[\sigma^2(F_o^2) + (0.0592P)^2]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} = 0.001$
 $\Delta\rho_{\max} = 0.14 \text{ e } \text{\AA}^{-3}$
 $\Delta\rho_{\min} = -0.14 \text{ e } \text{\AA}^{-3}$

Special details

Geometry. Bond distances, angles *etc.* have been calculated using the rounded fractional coordinates. All su's are estimated from the variances of the (full) variance-covariance matrix. The cell e.s.d.'s are taken into account in the estimation of distances, angles and torsion angles

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
O21	0.64098 (7)	0.43351 (6)	0.19316 (12)	0.0534 (4)
O22	0.62981 (7)	0.44279 (6)	-0.08684 (12)	0.0518 (4)
C1	0.61757 (8)	0.59370 (8)	0.25158 (18)	0.0364 (5)
C2	0.68820 (9)	0.61788 (9)	0.3219 (2)	0.0510 (6)
C3	0.69086 (15)	0.64720 (11)	0.4853 (3)	0.0775 (9)
C4	0.6222 (2)	0.65222 (12)	0.5811 (2)	0.0918 (11)
C5	0.55179 (15)	0.62730 (13)	0.5105 (3)	0.0817 (9)
C6	0.54971 (10)	0.59874 (10)	0.3487 (2)	0.0554 (6)
C11	0.61473 (11)	0.56093 (9)	0.07399 (18)	0.0550 (6)
C21	0.62985 (8)	0.47232 (9)	0.05979 (17)	0.0378 (5)
N1G	0.77589 (11)	0.40624 (9)	0.41545 (18)	0.0537 (5)
N2G	0.77186 (10)	0.40368 (9)	0.70667 (17)	0.0517 (5)
N3G	0.66128 (8)	0.43990 (8)	0.55565 (18)	0.0445 (5)
C1G	0.73652 (9)	0.41692 (8)	0.55946 (17)	0.0381 (5)
H2	0.73470	0.61440	0.25830	0.0610*
H3	0.73900	0.66370	0.53150	0.0930*
H4	0.62360	0.67210	0.69180	0.1100*
H5	0.50520	0.63000	0.57400	0.0980*
H6	0.50160	0.58230	0.30260	0.0670*
H11	0.65390	0.58860	0.00530	0.0660*
H12	0.56290	0.57230	0.02590	0.0660*
H11G	0.8248 (12)	0.3924 (10)	0.4193 (18)	0.054 (5)*

H12G	0.7504 (10)	0.4141 (10)	0.321 (2)	0.063 (5)*
H21G	0.7451 (10)	0.4104 (9)	0.795 (2)	0.049 (5)*
H22G	0.8201 (12)	0.3872 (10)	0.710 (2)	0.061 (6)*
H31G	0.6432 (9)	0.4549 (9)	0.453 (2)	0.049 (5)*
H32G	0.6410 (10)	0.4547 (10)	0.651 (2)	0.055 (5)*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
O21	0.0885 (9)	0.0400 (6)	0.0317 (6)	0.0054 (6)	-0.0136 (5)	-0.0003 (5)
O22	0.0728 (8)	0.0548 (7)	0.0279 (6)	-0.0017 (6)	0.0036 (5)	-0.0026 (5)
C1	0.0399 (8)	0.0297 (8)	0.0397 (8)	0.0033 (7)	-0.0008 (7)	0.0028 (6)
C2	0.0427 (9)	0.0458 (10)	0.0646 (11)	-0.0026 (8)	-0.0051 (8)	0.0073 (8)
C3	0.1009 (18)	0.0509 (12)	0.0807 (15)	-0.0161 (12)	-0.0466 (13)	0.0084 (11)
C4	0.185 (3)	0.0493 (12)	0.0411 (11)	0.0059 (15)	-0.0096 (14)	-0.0145 (9)
C5	0.1019 (18)	0.0733 (14)	0.0699 (14)	0.0135 (13)	0.0366 (13)	-0.0137 (12)
C6	0.0406 (10)	0.0558 (11)	0.0698 (12)	0.0006 (8)	0.0068 (8)	-0.0051 (9)
C11	0.0797 (13)	0.0446 (9)	0.0406 (9)	0.0064 (9)	-0.0023 (8)	0.0044 (7)
C21	0.0400 (8)	0.0445 (9)	0.0288 (8)	0.0002 (7)	0.0003 (6)	0.0004 (7)
N1G	0.0462 (9)	0.0832 (11)	0.0318 (8)	0.0094 (8)	0.0017 (7)	0.0011 (7)
N2G	0.0433 (9)	0.0805 (11)	0.0312 (8)	0.0052 (8)	-0.0022 (7)	-0.0033 (7)
N3G	0.0449 (8)	0.0599 (9)	0.0288 (8)	0.0051 (6)	0.0005 (6)	0.0030 (6)
C1G	0.0415 (9)	0.0413 (8)	0.0314 (8)	-0.0044 (7)	-0.0011 (7)	-0.0003 (6)

Geometric parameters (Å, °)

O21—C21	1.2470 (17)	C1—C2	1.373 (2)
O22—C21	1.2522 (17)	C2—C3	1.373 (3)
N1G—C1G	1.321 (2)	C3—C4	1.381 (4)
N2G—C1G	1.317 (2)	C4—C5	1.374 (4)
N3G—C1G	1.325 (2)	C5—C6	1.357 (3)
N1G—H11G	0.86 (2)	C11—C21	1.518 (2)
N1G—H12G	0.866 (16)	C2—H2	0.9300
N2G—H21G	0.834 (16)	C3—H3	0.9300
N2G—H22G	0.86 (2)	C4—H4	0.9300
N3G—H31G	0.897 (16)	C5—H5	0.9300
N3G—H32G	0.859 (16)	C6—H6	0.9300
C1—C6	1.376 (2)	C11—H12	0.9700
C1—C11	1.498 (2)	C11—H11	0.9700
C1G—N1G—H12G	117.4 (11)	O21—C21—O22	124.14 (14)
H11G—N1G—H12G	123.3 (15)	C1—C2—H2	120.00
C1G—N1G—H11G	119.3 (10)	C3—C2—H2	120.00
C1G—N2G—H22G	120.6 (11)	C2—C3—H3	120.00
H21G—N2G—H22G	122.0 (15)	C4—C3—H3	120.00
C1G—N2G—H21G	117.4 (11)	C5—C4—H4	120.00
C1G—N3G—H32G	116.5 (11)	C3—C4—H4	121.00
H31G—N3G—H32G	124.3 (15)	C4—C5—H5	120.00

C1G—N3G—H31G	115.3 (10)	C6—C5—H5	120.00
C2—C1—C6	118.65 (14)	C1—C6—H6	119.00
C2—C1—C11	120.66 (13)	C5—C6—H6	119.00
C6—C1—C11	120.68 (14)	C21—C11—H11	109.00
C1—C2—C3	120.62 (16)	C21—C11—H12	108.00
C2—C3—C4	120.1 (2)	H11—C11—H12	108.00
C3—C4—C5	119.00 (18)	C1—C11—H11	108.00
C4—C5—C6	120.5 (2)	C1—C11—H12	108.00
C1—C6—C5	121.17 (17)	N2G—C1G—N3G	120.08 (14)
C1—C11—C21	115.15 (12)	N1G—C1G—N2G	119.89 (15)
O21—C21—C11	118.63 (12)	N1G—C1G—N3G	120.02 (14)
O22—C21—C11	117.24 (12)		
C6—C1—C2—C3	-0.6 (2)	C1—C2—C3—C4	0.4 (3)
C11—C1—C2—C3	-179.37 (15)	C2—C3—C4—C5	0.2 (3)
C2—C1—C6—C5	0.3 (2)	C3—C4—C5—C6	-0.5 (3)
C11—C1—C6—C5	179.08 (16)	C4—C5—C6—C1	0.2 (3)
C2—C1—C11—C21	86.98 (18)	C1—C11—C21—O21	2.0 (2)
C6—C1—C11—C21	-91.75 (18)	C1—C11—C21—O22	-178.42 (13)

Hydrogen-bond geometry (Å, °)

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
N1G—H11G...O22 ⁱ	0.86 (2)	2.02 (2)	2.876 (2)	173.9 (15)
N1G—H12G...O21	0.866 (16)	2.123 (17)	2.900 (2)	149.0 (15)
N2G—H21G...O22 ⁱⁱ	0.834 (16)	2.219 (17)	2.9625 (19)	148.5 (15)
N2G—H22G...O21 ⁱ	0.86 (2)	1.97 (2)	2.827 (2)	172.6 (15)
N3G—H31G...O21	0.897 (16)	2.068 (16)	2.8634 (17)	147.2 (13)
N3G—H32G...O22 ⁱⁱ	0.859 (16)	2.073 (16)	2.8520 (17)	150.5 (15)

Symmetry codes: (i) $y+1/2, -x+1, z+1/2$; (ii) $x, y, z+1$.