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## Structure Reports

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**(2-Pyridyl)[5-(2-pyridylcarbonyl)-2-pyridyl]methanone**

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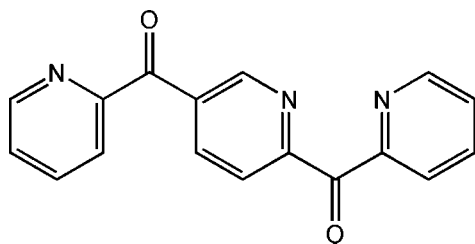
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Key indicators: single-crystal X-ray study;  $T = 293$  K; mean  $\sigma(\text{C}-\text{C}) = 0.003$  Å; disorder in main residue;  $R$  factor = 0.058;  $wR$  factor = 0.177; data-to-parameter ratio = 16.2.

In the centrosymmetric title compound,  $\text{C}_{17}\text{H}_{11}\text{N}_3\text{O}_2$ , the dihedral angle between the central and pendant pyridyl rings is  $50.29$  ( $9$ )°. In the crystal, molecules stack along the  $a$  axis by  $\pi$ - $\pi$  interactions between the pyridine rings with centroid-centroid distances of  $3.845$  ( $2$ ) Å. The N atom and one of the C atoms of the central ring are disordered by symmetry.

## Related literature

For studies on other pyridinyl-based methanone species, see: Papaefstathiou & Perlepes (2002); Dendrinou-Samara *et al.* (2003); Crowder *et al.* (2004); Chen *et al.* (2005); Wan *et al.* (2008).



## Experimental

## Crystal data

$\text{C}_{17}\text{H}_{11}\text{N}_3\text{O}_2$	$\gamma = 99.772$ ( $6$ )°
$M_r = 289.30$	$V = 333.29$ ( $19$ ) Å <sup>3</sup>
Triclinic, $P\bar{1}$	$Z = 1$
$a = 3.8453$ ( $13$ ) Å	Mo $K\alpha$ radiation
$b = 8.447$ ( $3$ ) Å	$\mu = 0.10$ mm <sup>-1</sup>
$c = 11.202$ ( $3$ ) Å	$T = 293$ K
$\alpha = 108.672$ ( $6$ )°	$0.60 \times 0.50 \times 0.29$ mm
$\beta = 97.251$ ( $6$ )°	

## Data collection

Bruker APEXII CCD area-detector diffractometer	2301 measured reflections
Absorption correction: multi-scan (SADABS; Bruker, 2007)	1623 independent reflections
$T_{\min} = 0.622$ , $T_{\max} = 1.000$	1198 reflections with $I > 2\sigma(I)$
	$R_{\text{int}} = 0.016$

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.058$	100 parameters
$wR(F^2) = 0.177$	H-atom parameters constrained
$S = 1.07$	$\Delta\rho_{\text{max}} = 0.31$ e Å <sup>-3</sup>
1623 reflections	$\Delta\rho_{\text{min}} = -0.26$ e Å <sup>-3</sup>

Data collection: APEX2 (Bruker, 2007); cell refinement: APEX2 and SAINT (Bruker, 2007); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: SHELXTL (Sheldrick, 2008); software used to prepare material for publication: SHELXTL and PLATON (Spek, 2009).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: JJ2048).

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## supporting information

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**(2-Pyridyl)[5-(2-pyridylcarbonyl)-2-pyridyl]methanone****Hong Qiang, Fan Zhang and Zi-jia Wang****S1. Comment**

Di-2-pyridylmethanone has attracted great interest in recent years as it can exist in various forms in stabilizing its metal complexes, including its neat ketone form, singly and doubly deprotonated gem-diol forms, as well as the monoanion of its hemiacetal form (Papaefstathiou *et al.*, 2002; Dendrinou-Samara *et al.*, 2003; Crowder *et al.*, 2004). Therefore, homolog compounds such as 2,6-pyridinediylbis(2-pyridyl)methanone (Chen *et al.*, 2005) and 2,6-pyridinediylbis(3-pyridyl)methanone (Wan *et al.*, 2008) were also synthesized and characterized.

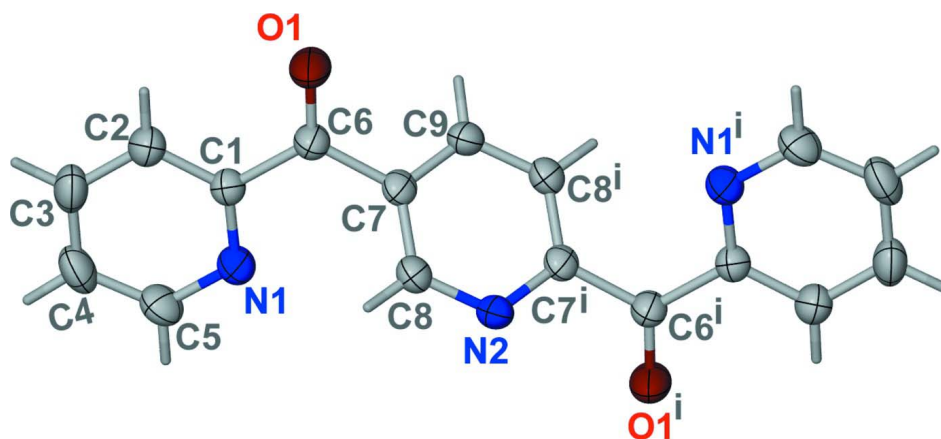
In the present study, a new member of this family, namely 2,5-pyridinediylbis(2-pyridyl)methanone (C<sub>17</sub>H<sub>11</sub>N<sub>3</sub>O<sub>2</sub>), is reported. X-ray diffraction analysis shows that the N2 and C9 atoms of the 2,5-pyridinediyl ring have an equal occupancy at the same site. Thus the molecule is centrosymmetric with two 2-pyridyl methanone groups bonding to the 2,5-pyridinediyl ring at the 2 and 5 positions, respectively. The 2-pyridyl and the center 2,5-pyridinediyl rings exhibit a dihedral angle of 50.29 (9)° (Fig. 1). Along the *a* axis, the packing between the molecules is provided by weak un-covalent interaction only: /p-electron.../p-electron ring interaction. The distance between the centroids of the proximate pyridyl rings equals 3.845 (2) Å, as shown in Fig. 2.

**S2. Experimental**

The preparation of the title compound followed the procedure previously developed for 2,6-pyridinediylbis(3-pyridyl)methanone (Wan *et al.*, 2008). The crude product was extracted with chloroform, and the combined organic extract was dried over anhydrous sodium sulfate and finally concentrated in vacuo to give a brown oil. Further purification by chromatography on silica gel (R<sub>f</sub> = 0.44, eluent: ether acetate/dichloromethane = 1:6, v/v), giving 2.96 g of light yellow powder of 2,5-pyridinediylbis(2-pyridyl)methanone in 41% yield; m.p. 108-110°C; The yellow crystals of the title compound having a average 0.40 × 0.30 × 0.20 mm dimension were obtained by slow evaporation from its solution of dichloromethane/N,N-dimethylformamide 1/1 (v/v).

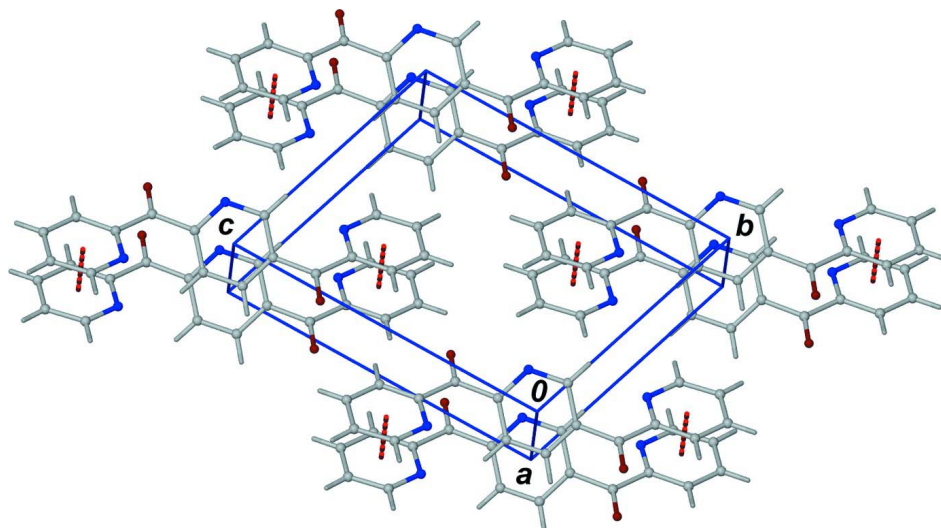
**S3. Refinement**

The hydrogen atoms were placed in idealized positions and allowed to ride on the relevant carbon atoms, with C—H = 0.93 Å and  $U_{iso}(H) = 1.2U_{eq}(C)$ .



**Figure 1**

The atom-numbering scheme of the title compound  $C_{17}H_{11}N_3O_2$ . Displacement ellipsoids are drawn at the 50% probability level and H atoms are shown as sticks of arbitrary radii. Symmetry code:  $i -x + 2, -y + 2, -z$ .



**Figure 2**

The packing illustration of the title compound,  $C_{17}H_{11}N_3O_2$ . The red-dashed lines indicate weak  $\pi \cdots \pi$  stacking interactions.

### 2,5-pyridinediylbis(2-pyridinyl)methanone

#### Crystal data

$C_{17}H_{11}N_3O_2$   
 $M_r = 289.30$   
 Triclinic,  $P\bar{1}$   
 Hall symbol:  $-P\ 1$   
 $a = 3.8453(13)\ \text{\AA}$   
 $b = 8.447(3)\ \text{\AA}$   
 $c = 11.202(3)\ \text{\AA}$   
 $\alpha = 108.672(6)^\circ$   
 $\beta = 97.251(6)^\circ$   
 $\gamma = 99.772(6)^\circ$   
 $V = 333.29(19)\ \text{\AA}^3$

$Z = 1$   
 $F(000) = 150$   
 $D_x = 1.441\ \text{Mg m}^{-3}$   
 Melting point: 401 K  
 Mo  $K\alpha$  radiation,  $\lambda = 0.71073\ \text{\AA}$   
 Cell parameters from 230 reflections  
 $\theta = 1.9\text{--}28.1^\circ$   
 $\mu = 0.10\ \text{mm}^{-1}$   
 $T = 293\ \text{K}$   
 Block, yellow  
 $0.60 \times 0.50 \times 0.29\ \text{mm}$

*Data collection*

Bruker APEXII CCD area-detector  
diffractometer  
Radiation source: fine-focus sealed tube  
Graphite monochromator  
 $\omega$ -scans  
Absorption correction: multi-scan  
*SADABS* (Bruker, 2007)  
 $T_{\min} = 0.622$ ,  $T_{\max} = 1.000$

2301 measured reflections  
1623 independent reflections  
1198 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.016$   
 $\theta_{\max} = 28.4^\circ$ ,  $\theta_{\min} = 2.0^\circ$   
 $h = -5 \rightarrow 4$   
 $k = -11 \rightarrow 11$   
 $l = -12 \rightarrow 14$

*Refinement*

Refinement on  $F^2$   
Least-squares matrix: full  
 $R[F^2 > 2\sigma(F^2)] = 0.058$   
 $wR(F^2) = 0.177$   
 $S = 1.07$   
1623 reflections  
100 parameters  
0 restraints  
Primary atom site location: structure-invariant  
direct methods

Secondary atom site location: difference Fourier  
map  
Hydrogen site location: inferred from  
neighbouring sites  
H-atom parameters constrained  
 $w = 1/[\sigma^2(F_o^2) + (0.096P)^2 + 0.0878P]$   
where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\max} < 0.001$   
 $\Delta\rho_{\max} = 0.31 \text{ e } \text{\AA}^{-3}$   
 $\Delta\rho_{\min} = -0.26 \text{ e } \text{\AA}^{-3}$

*Special details*

**Geometry.** All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > 2\sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$	Occ. (<1)
O1	1.3180 (5)	1.09147 (19)	0.33644 (14)	0.0620 (6)	
N1	0.7459 (5)	0.6838 (2)	0.16816 (16)	0.0402 (4)	
N2	0.9591 (5)	1.1260 (2)	0.10902 (16)	0.0377 (4)	0.50
C9	0.9591 (5)	1.1260 (2)	0.10902 (16)	0.0377 (4)	0.50
H9A	0.9331	1.2127	0.1798	0.045*	0.50
C1	0.9713 (5)	0.8073 (2)	0.26845 (17)	0.0339 (4)	
C2	1.0530 (6)	0.7939 (3)	0.38842 (19)	0.0435 (5)	
H2A	1.2108	0.8826	0.4556	0.052*	
C3	0.8934 (7)	0.6452 (3)	0.4057 (2)	0.0517 (6)	
H3A	0.9378	0.6331	0.4855	0.062*	
C4	0.6689 (7)	0.5157 (3)	0.3033 (2)	0.0524 (6)	
H4A	0.5633	0.4134	0.3120	0.063*	
C5	0.6028 (6)	0.5402 (3)	0.1870 (2)	0.0485 (6)	
H5A	0.4502	0.4519	0.1181	0.058*	
C6	1.1320 (5)	0.9691 (2)	0.24786 (17)	0.0372 (5)	
C7	1.0544 (5)	0.9809 (2)	0.11607 (17)	0.0333 (4)	

C8	1.0969 (5)	0.8553 (2)	0.00732 (18)	0.0369 (5)
H8A	1.1642	0.7573	0.0141	0.044*

*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
O1	0.0862 (13)	0.0422 (8)	0.0381 (8)	-0.0157 (8)	-0.0097 (8)	0.0113 (7)
N1	0.0422 (10)	0.0355 (8)	0.0394 (9)	-0.0001 (7)	0.0059 (7)	0.0136 (7)
N2	0.0467 (11)	0.0295 (8)	0.0341 (9)	0.0035 (7)	0.0096 (7)	0.0092 (7)
C9	0.0467 (11)	0.0295 (8)	0.0341 (9)	0.0035 (7)	0.0096 (7)	0.0092 (7)
C1	0.0360 (10)	0.0330 (9)	0.0334 (9)	0.0065 (7)	0.0077 (7)	0.0128 (7)
C2	0.0538 (13)	0.0404 (10)	0.0373 (10)	0.0106 (9)	0.0065 (9)	0.0157 (8)
C3	0.0712 (16)	0.0510 (12)	0.0471 (12)	0.0209 (11)	0.0180 (11)	0.0299 (10)
C4	0.0632 (15)	0.0379 (10)	0.0671 (15)	0.0110 (10)	0.0239 (12)	0.0291 (10)
C5	0.0511 (13)	0.0355 (10)	0.0534 (13)	-0.0015 (9)	0.0097 (10)	0.0140 (9)
C6	0.0421 (11)	0.0328 (9)	0.0326 (9)	0.0012 (8)	0.0041 (8)	0.0108 (7)
C7	0.0333 (10)	0.0298 (8)	0.0335 (9)	-0.0016 (7)	0.0047 (7)	0.0116 (7)
C8	0.0418 (11)	0.0302 (8)	0.0380 (10)	0.0041 (7)	0.0075 (8)	0.0132 (7)

*Geometric parameters (Å, °)*

O1—C6	1.216 (2)	C3—C4	1.372 (3)
N1—C5	1.336 (3)	C3—H3A	0.9300
N1—C1	1.341 (2)	C4—C5	1.383 (3)
N2—C8 <sup>i</sup>	1.358 (2)	C4—H4A	0.9300
N2—C7	1.360 (3)	C5—H5A	0.9300
N2—H9A	0.9207	C6—C7	1.507 (2)
C1—C2	1.386 (3)	C7—C8	1.385 (3)
C1—C6	1.501 (3)	C8—C9 <sup>i</sup>	1.358 (2)
C2—C3	1.385 (3)	C8—N2 <sup>i</sup>	1.358 (2)
C2—H2A	0.9300	C8—H8A	0.9300
C5—N1—C1	116.84 (17)	C5—C4—H4A	120.6
C8 <sup>i</sup> —N2—C7	118.59 (16)	N1—C5—C4	123.62 (19)
C8 <sup>i</sup> —N2—H9A	118.5	N1—C5—H5A	118.2
C7—N2—H9A	123.0	C4—C5—H5A	118.2
N1—C1—C2	123.47 (17)	O1—C6—C1	120.89 (17)
N1—C1—C6	117.03 (16)	O1—C6—C7	119.68 (16)
C2—C1—C6	119.48 (17)	C1—C6—C7	119.42 (15)
C3—C2—C1	118.27 (19)	N2—C7—C8	120.92 (17)
C3—C2—H2A	120.9	N2—C7—C6	116.89 (16)
C1—C2—H2A	120.9	C8—C7—C6	122.09 (16)
C4—C3—C2	119.06 (19)	C9 <sup>i</sup> —C8—C7	120.49 (17)
C4—C3—H3A	120.5	N2 <sup>i</sup> —C8—C7	120.49 (17)
C2—C3—H3A	120.5	C9 <sup>i</sup> —C8—H8A	119.8
C3—C4—C5	118.71 (18)	N2 <sup>i</sup> —C8—H8A	119.8
C3—C4—H4A	120.6	C7—C8—H8A	119.8

C5—N1—C1—C2	-1.5 (3)	C2—C1—C6—C7	177.66 (17)
C5—N1—C1—C6	-179.82 (19)	C8 <sup>i</sup> —N2—C7—C8	0.5 (3)
N1—C1—C2—C3	0.0 (3)	C8 <sup>i</sup> —N2—C7—C6	177.03 (17)
C6—C1—C2—C3	178.27 (19)	O1—C6—C7—N2	-45.2 (3)
C1—C2—C3—C4	1.6 (3)	C1—C6—C7—N2	133.4 (2)
C2—C3—C4—C5	-1.6 (4)	O1—C6—C7—C8	131.2 (2)
C1—N1—C5—C4	1.5 (3)	C1—C6—C7—C8	-50.2 (3)
C3—C4—C5—N1	0.0 (4)	N2—C7—C8—C9 <sup>i</sup>	-0.5 (3)
N1—C1—C6—O1	174.7 (2)	C6—C7—C8—C9 <sup>i</sup>	-176.85 (17)
C2—C1—C6—O1	-3.7 (3)	N2—C7—C8—N2 <sup>i</sup>	-0.5 (3)
N1—C1—C6—C7	-4.0 (3)	C6—C7—C8—N2 <sup>i</sup>	-176.85 (17)

Symmetry code: (i)  $-x+2, -y+2, -z$ .