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Chlorido(η^4-cycloocta-1,5-diene)(*N*,*N*′-diethylthiourea-κS)rhodium(I)

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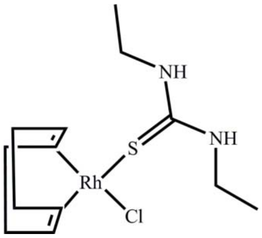
Key indicators: single-crystal X-ray study; T = 293 K; mean σ(C–C) = 0.003 Å; R factor = 0.016; wR factor = 0.042; data-to-parameter ratio = 16.1.

In the title rhodium(I) complex, [RhCl(C_8H_{12})($C_5H_{12}N_2S$)], *N*,*N*′-diethylthiourea acts as a monodenate *S*-donor ligand. The rhodium(I) coordination sphere is completed by the Cl atom and the COD [= 1,5-cyclooctadiene] ligand interacting through the π-electrons of the double bonds. If the midpoints of these two bonds are taken into account, the Rh atom exhibits a distorted square-planar coordination. The *syn* conformation of the *N*,*N*′-diethylthiourea ligand with respect to the Cl atom is stabilized by an intramolecular N—H$\cdots$Cl hydrogen bond. A weak intermolecular N—H$\cdots$Cl interaction links molecules along the *a* axis.

Related literature

For coordination modes of thiourea and thiourea-based ligands, see: Wilkinson (1987); Gibson *et al.* (1994); Robinson *et al.* (2000). For the application of thioureas as ligands for metal precursors in asymmetric catalysis, see: Breuzard *et al.* (2000). For related Rh(I) complexes containing thiourea ligands, see: Cauzzi *et al.* (1995, 1997). For structural data of the *N*,*N*′-diethylthiourea ligand, see: Ramnathan *et al.* (1995).



Experimental

Crystal data

[RhCl(C_8H_{12})($C_5H_{12}N_2S$)]
M_r = 378.76
Triclinic, $P\overline{1}$
a = 7.295 (5) Å
b = 8.705 (5) Å
c = 12.602 (5) Å
α = 101.727 (5)°
β = 102.058 (5)°
γ = 94.765 (5)°
V = 759.7 (7) Å^3
Z = 2
Mo $K\alpha$ radiation
μ = 1.42 mm^{-1}
T = 293 K
0.60 × 0.24 × 0.16 mm

Data collection

Bruker–Nonius Kappa APEXII CCD diffractometer
Absorption correction: multi-scan (*SADABS*; Bruker, 2001)
$T_{\min}$ = 0.540, $T_{\max}$ = 0.710
12830 measured reflections
2656 independent reflections
2585 reflections with $I > 2\sigma(I)$
$R_{\rm int}$ = 0.015

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.016
$wR(F^2)$ = 0.042
S = 0.97
2656 reflections
165 parameters
H-atom parameters constrained
$\Delta\rho_{\max}$ = 0.41 e Å^{-3}
$\Delta\rho_{\min}$ = −0.35 e Å^{-3}

Table 1
Hydrogen-bond geometry (Å, °).

D—H$\cdots A$	D—H	H$\cdots A$	$D\cdots A$	D—H$\cdots A$
N1—H1$\cdots$Cl1	0.86	2.39	3.152 (3)	148
N2—H2$\cdots$Cl1$^{\rm i}$	0.86	2.89	3.356 (3)	116

Symmetry code: (i) $x + 1, y, z$.

Data collection: *COLLECT* (Nonius, 1998); cell refinement: *DIRAX/LSQ* (Duisenberg, 1992); data reduction: *EVALCCD* (Duisenberg *et al.*, 2003); program(s) used to solve structure: *SIR2004* (Burla *et al.*, 2005); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *XP* in *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *WinGX* (Farrugia, 1999).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: NG5033).

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Chlorido(η^4-cycloocta-1,5-diene)(*N*,*N*′-diethylthiourea-*κS*)rhodium(I)

Giovanna Brancatelli, Dario Drommi, Giuseppe Bruno and Felice Faraone

S1. Comment

Thiourea and thiourea-based ligands form complexes with a number of transition metals (Wilkinson, 1987; Gibson *et al.*, 1994; Robinson *et al.*, 2000) and their application as ligands for metal catalyst in styrene hydroformylation has been recently shown (Breuzard *et al.*, 2000).

In order to investigate the coordination chemistry of symmetrically substituted thiourea derivatives as ligands for metal complexes applicable in asymmetric catalysis, the reaction between chloro(η^4-1,5-cyclooctadiene)rhodium(I) dimer and *N,N*′-diethylthiourea has been performed in dichloromethane. The obtained crystals were identified as the title compound by single-crystal X-ray diffraction. Figure 1 shows that in the compound (I) structure the *N,N*′-diethylthiourea acts as a monodenate S-donor ligand. Therefore the rhodium(I) coordination sphere is completed by a chlorine atom and COD [= 1,5-cyclooctadiene] ligand interacting with the metal center through the π-electrons of the double bonds. If the midpoints of these two bonds are taken into account the rhodium atom displays a distorted square planar coordination, as evidenced by the angles at Rh(1) [*M*(2)—Rh(1)—S(1) 86.4 (8)°, *M*(1)—Rh(1)—Cl(1) 88.9 (8)°, *M*(2)—Rh(1)—*M*(1) 87.8 (1)°, S(1)—Rh(1)—Cl(1) 96.97 (3)°]. In the thiourea moiety the distance S(1)—C(1) [1.732 (2) Å] is slightly longer than that found in the crystallographic structure of the *N,N*′-diethythiourea [1.707 (3) Å] (Ramnathan *et al.*, 1995). This lengthening of the S—C bond is consistent with the decreasing double bond character due to the coordination at the metal center. Further the C(1)—S(1)—Rh(1) bond angle value [115.00 (8)°] indicates that the thiourea sulfur is bound to rhodium(I) primarily *via* a lone pair in a non-bonding sp^2 sulfur orbital. C(1)—N(1) and C(1)—N(2) bond lengths [1.331 (3)Å and 1.343 (3) Å] are almost equivalent as expected for symmetrically substituted thiourea molecules. The value of Rh—S bond [2.403 (1) Å] is comparable with those found in similar complexes (Cauzzi *et al.*, 1995, 1997). The *syn* conformation of the substituent on the sulfur with respect to the chlorine atom is stabilized by the intramolecular N(1)—H(1)···Cl(1) hydrogen bonding interaction.

The crystal packing arrangement is stabilized by van der Walls forces and the very weak intermolecular N(2)—H(2)···Cl(1) A hydrogen interaction along the *a* axis (Fig. 2) between the thioamide N(2) and the Cl(1) A of the neighbor complex molecule generated by applying the crystallographic (x + 1, y, z) symmetry operation.

S2. Experimental

The compound was prepared by reacting [Rh(COD)(μ-Cl)]$_2$ (0.050 g, 0.10 mmol) with the *N,N*′-diethylthiourea ligand (0.0264 g, 0.2 mmol) in CH_2Cl_2 solution at room temperature for 30 min. After evaporation of the solvent in vacuo, the residue was dissolved in dichloromethane. Recrystallization from CH_2Cl_2/hexane gave orange crystals of the complex.

S3. Refinement

Several H atoms were located in a difference Fourier map and placed in idealized positions using the riding-model technique with C—H = 0.93Å and N—H = 0.86Å for aliphatic and thioamide H atoms, respectively.

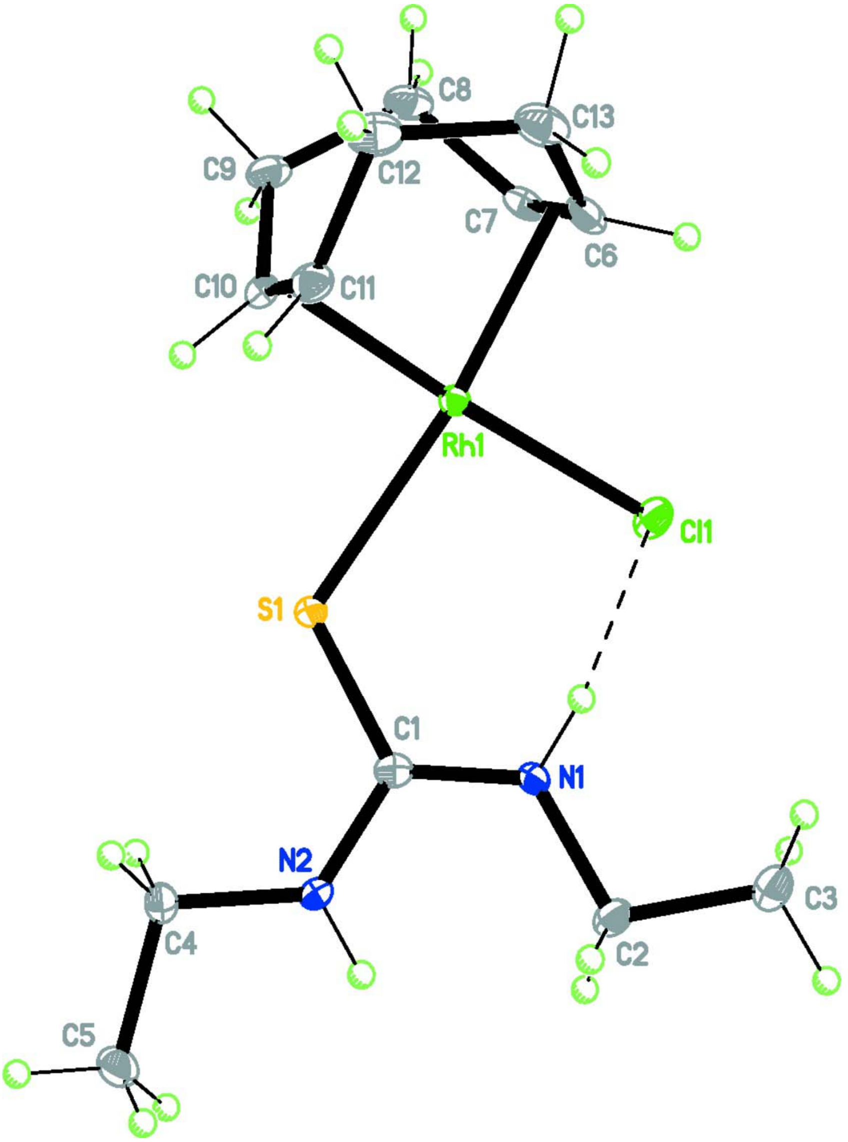

Figure 1

ORTEP view of compound (I) showing atomic labeling scheme and displacement ellipsoids at 50% probability for non-H atoms.

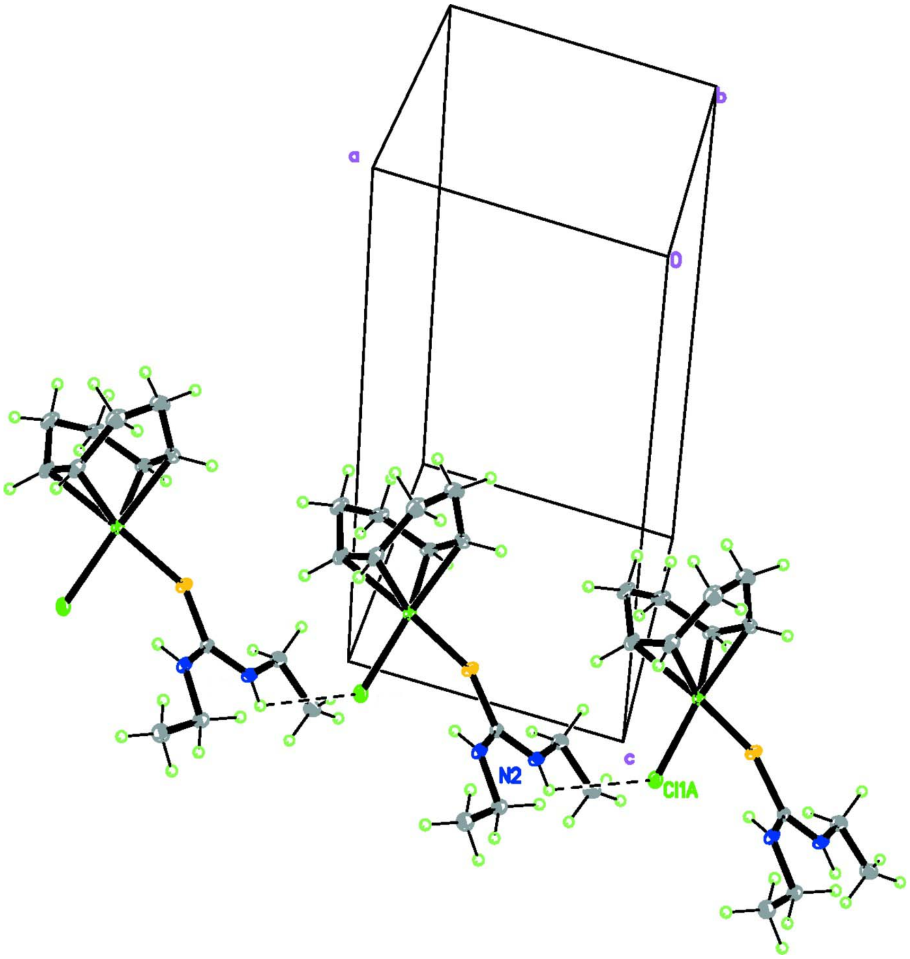

Figure 2

View of the molecular rows along the *a* axis generated by N(2)—H(2)···Cl(1) intermolecular interaction.

Chlorido(η^4-cycloocta-1,5-diene)(*N*,*N*′- diethylthiourea-κS)rhodium(I)

Crystal data

$[RhCl(C_8H_{12})(C_5H_{12}N_2S)]$
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Triclinic, $P\overline{1}$
Hall symbol: -P 1
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b = 8.705 (5) Å
c = 12.602 (5) Å
α = 101.727 (5)°
β = 102.058 (5)°
γ = 94.765 (5)°
V = 759.7 (7) Å^3

Z = 2
$F(000)$ = 388
D_x = 1.656 Mg m^{-3}
Mo $K\alpha$ radiation, λ = 0.71069 Å
Cell parameters from 93 reflections
θ = 5.3–22.3°
μ = 1.42 mm^{-1}
T = 293 K
Plate, orange
0.60 × 0.24 × 0.16 mm

Data collection

Bruker–Nonius Kappa APEXII CCD diffractometer
Graphite monochromator
ω scans

Absorption correction: multi-scan (*SADABS*; Bruker, 2001)
T_{min} = 0.540, T_{max} = 0.710
12830 measured reflections

2656 independent reflections	$h = -8 \rightarrow 8$
2585 reflections with $I > 2\sigma(I)$	$k = -10 \rightarrow 10$
$R_{\text{int}} = 0.015$	$l = -14 \rightarrow 14$
$\theta_{\text{max}} = 25°$, $\theta_{\text{min}} = 3.4°$	

Refinement

Refinement on F^2	0 restraints
Least-squares matrix: full	H-atom parameters constrained
$R[F^2 > 2\sigma(F^2)] = 0.016$	$w = 1/[\sigma^2(F_o^2) + (0.0121P)^2 + 1.6925P]$ where $P = (F_o^2 + 2F_c^2)/3$
$wR(F^2) = 0.042$	$(\Delta/\sigma)_{\text{max}} = 0.01$
$S = 0.97$	$\Delta\rho_{\text{max}} = 0.41$ e Å^{-3}
2656 reflections	$\Delta\rho_{\text{min}} = -0.35$ e Å^{-3}
165 parameters	

Special details

Geometry. All s.u.'s (except the s.u. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell s.u.'s are taken into account individually in the estimation of s.u.'s in distances, angles and torsion angles; correlations between s.u.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell s.u.'s is used for estimating s.u.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R-factor wR and goodness of fit S are based on F^2, conventional R-factors R are based on F, with F set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. R-factors based on F^2 are statistically about twice as large as those based on F, and R- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	y	z	U_{iso}*/U_{eq}
C1	0.6141 (3)	0.3714 (2)	0.14001 (16)	0.0101 (4)
C2	0.6911 (3)	0.6516 (2)	0.13529 (18)	0.0130 (4)
H2A	0.678	0.6272	0.0552	0.016*
H2B	0.8245	0.6631	0.1708	0.016*
C3	0.6120 (3)	0.8048 (2)	0.17068 (19)	0.0156 (4)
H3A	0.6701	0.8867	0.143	0.023*
H3B	0.6381	0.8348	0.2505	0.023*
H3C	0.4778	0.7899	0.1409	0.023*
C4	0.8012 (3)	0.1765 (2)	0.05213 (18)	0.0137 (4)
H4A	0.6878	0.1021	0.0172	0.016*
H4B	0.8673	0.1429	0.1169	0.016*
C5	0.9280 (3)	0.1800 (3)	−0.02977 (18)	0.0170 (5)
H5A	0.8697	0.2283	−0.088	0.026*
H5B	0.9458	0.0738	−0.0614	0.026*
H5C	1.0484	0.24	0.0086	0.026*
C6	0.1778 (3)	0.4127 (3)	0.43845 (18)	0.0150 (4)
H6	0.1543	0.5227	0.4421	0.018*
C7	0.0334 (3)	0.3011 (3)	0.36521 (18)	0.0143 (4)
H7	−0.0725	0.3469	0.3272	0.017*
C8	−0.0210 (3)	0.1390 (3)	0.38650 (19)	0.0172 (5)
H8A	−0.1559	0.108	0.3576	0.021*
H8B	0.0066	0.1461	0.4662	0.021*
C9	0.0857 (3)	0.0111 (3)	0.33158 (18)	0.0165 (4)
H9A	0.1004	−0.0689	0.3752	0.02*

H9B	0.0104	−0.0398	0.2578	0.02*
C10	0.2795 (3)	0.0767 (2)	0.32166 (18)	0.0133 (4)
H10	0.3337	0.0055	0.2693	0.016*
C11	0.4153 (3)	0.1802 (3)	0.40867 (18)	0.0146 (4)
H11	0.5463	0.1679	0.4052	0.018*
C12	0.3872 (3)	0.2309 (3)	0.52713 (18)	0.0189 (5)
H12A	0.5085	0.2448	0.5797	0.023*
H12B	0.3058	0.1479	0.542	0.023*
C13	0.2989 (3)	0.3865 (3)	0.54479 (18)	0.0193 (5)
H13A	0.222	0.3849	0.5988	0.023*
H13B	0.3993	0.4745	0.5753	0.023*
N1	0.5883 (2)	0.52268 (19)	0.16738 (14)	0.0110 (3)
H1	0.5065	0.5463	0.2062	0.013*
N2	0.7506 (2)	0.3357 (2)	0.08595 (14)	0.0121 (4)
H2	0.8134	0.4123	0.0699	0.014*
S1	0.47935 (7)	0.22159 (6)	0.17216 (4)	0.01184 (11)
Cl1	0.18218 (7)	0.52815 (6)	0.21639 (4)	0.01583 (11)
Rh1	0.27945 (2)	0.309577 (18)	0.295374 (13)	0.00866 (6)

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.0092 (10)	0.0124 (10)	0.0074 (9)	−0.0004 (8)	0.0005 (8)	0.0020 (8)
C2	0.0123 (10)	0.0113 (10)	0.0160 (11)	−0.0002 (8)	0.0036 (8)	0.0049 (8)
C3	0.0151 (11)	0.0120 (10)	0.0190 (11)	0.0001 (8)	0.0026 (9)	0.0041 (9)
C4	0.0173 (11)	0.0115 (10)	0.0147 (11)	0.0053 (8)	0.0064 (9)	0.0040 (8)
C5	0.0199 (11)	0.0212 (11)	0.0150 (11)	0.0092 (9)	0.0090 (9)	0.0078 (9)
C6	0.0155 (11)	0.0165 (11)	0.0138 (11)	0.0033 (9)	0.0081 (9)	0.0004 (9)
C7	0.0115 (10)	0.0195 (11)	0.0137 (10)	0.0037 (8)	0.0074 (8)	0.0027 (9)
C8	0.0150 (11)	0.0210 (12)	0.0159 (11)	−0.0029 (9)	0.0072 (9)	0.0035 (9)
C9	0.0195 (11)	0.0161 (11)	0.0148 (11)	−0.0030 (9)	0.0067 (9)	0.0048 (9)
C10	0.0179 (11)	0.0107 (10)	0.0148 (10)	0.0038 (8)	0.0074 (9)	0.0062 (8)
C11	0.0135 (10)	0.0186 (11)	0.0147 (11)	0.0047 (9)	0.0038 (8)	0.0086 (9)
C12	0.0196 (11)	0.0260 (12)	0.0100 (10)	0.0004 (9)	0.0007 (9)	0.0055 (9)
C13	0.0198 (12)	0.0225 (12)	0.0127 (11)	−0.0033 (9)	0.0052 (9)	−0.0016 (9)
N1	0.0119 (9)	0.0085 (8)	0.0139 (9)	0.0010 (7)	0.0073 (7)	0.0012 (7)
N2	0.0140 (9)	0.0088 (8)	0.0160 (9)	0.0013 (7)	0.0076 (7)	0.0046 (7)
S1	0.0144 (3)	0.0085 (2)	0.0149 (3)	0.00119 (19)	0.0084 (2)	0.00302 (19)
Cl1	0.0110 (2)	0.0151 (3)	0.0245 (3)	0.00368 (19)	0.0050 (2)	0.0100 (2)
Rh1	0.00837 (9)	0.00896 (9)	0.00903 (9)	0.00087 (6)	0.00293 (6)	0.00200 (6)

Geometric parameters (Å, º)

C1—N1	1.331 (3)	C7—H7	0.98
C1—N2	1.342 (3)	C8—C9	1.543 (3)
C1—S1	1.732 (2)	C8—H8A	0.97
C2—N1	1.469 (3)	C8—H8B	0.97
C2—C3	1.518 (3)	C9—C10	1.519 (3)

C2—H2A	0.97	C9—H9A	0.97
C2—H2B	0.97	C9—H9B	0.97
C3—H3A	0.96	C10—C11	1.411 (3)
C3—H3B	0.96	C10—Rh1	2.120 (2)
C3—H3C	0.96	C10—H10	0.98
C4—N2	1.466 (3)	C11—C12	1.530 (3)
C4—C5	1.526 (3)	C11—Rh1	2.130 (2)
C4—H4A	0.97	C11—H11	0.98
C4—H4B	0.97	C12—C13	1.543 (3)
C5—H5A	0.96	C12—H12A	0.97
C5—H5B	0.96	C12—H12B	0.97
C5—H5C	0.96	C13—H13A	0.97
C6—C7	1.401 (3)	C13—H13B	0.97
C6—C13	1.514 (3)	N1—H1	0.86
C6—Rh1	2.149 (2)	N2—H2	0.86
C6—H6	0.98	S1—Rh1	2.4026 (10)
C7—C8	1.525 (3)	Cl1—Rh1	2.4111 (11)
C7—Rh1	2.160 (2)		

N1—C1—N2	118.06 (18)	C8—C9—H9B	109
N1—C1—S1	122.42 (16)	H9A—C9—H9B	107.8
N2—C1—S1	119.52 (16)	C11—C10—C9	124.75 (19)
N1—C2—C3	109.49 (17)	C11—C10—Rh1	71.01 (12)
N1—C2—H2A	109.8	C9—C10—Rh1	110.88 (14)
C3—C2—H2A	109.8	C11—C10—H10	114.1
N1—C2—H2B	109.8	C9—C10—H10	114.1
C3—C2—H2B	109.8	Rh1—C10—H10	114.1
H2A—C2—H2B	108.2	C10—C11—C12	123.07 (19)
C2—C3—H3A	109.5	C10—C11—Rh1	70.20 (12)
C2—C3—H3B	109.5	C12—C11—Rh1	114.30 (15)
H3A—C3—H3B	109.5	C10—C11—H11	114
C2—C3—H3C	109.5	C12—C11—H11	114
H3A—C3—H3C	109.5	Rh1—C11—H11	114
H3B—C3—H3C	109.5	C11—C12—C13	112.13 (18)
N2—C4—C5	108.76 (17)	C11—C12—H12A	109.2
N2—C4—H4A	109.9	C13—C12—H12A	109.2
C5—C4—H4A	109.9	C11—C12—H12B	109.2
N2—C4—H4B	109.9	C13—C12—H12B	109.2
C5—C4—H4B	109.9	H12A—C12—H12B	107.9
H4A—C4—H4B	108.3	C6—C13—C12	112.96 (18)
C4—C5—H5A	109.5	C6—C13—H13A	109
C4—C5—H5B	109.5	C12—C13—H13A	109
H5A—C5—H5B	109.5	C6—C13—H13B	109
C4—C5—H5C	109.5	C12—C13—H13B	109
H5A—C5—H5C	109.5	H13A—C13—H13B	107.8
H5B—C5—H5C	109.5	C1—N1—C2	123.97 (17)
C7—C6—C13	124.7 (2)	C1—N1—H1	118
C7—C6—Rh1	71.44 (12)	C2—N1—H1	118

C13—C6—Rh1	111.45 (15)	C1—N2—C4	125.27 (17)
C7—C6—H6	113.9	C1—N2—H2	117.4
C13—C6—H6	113.9	C4—N2—H2	117.4
Rh1—C6—H6	113.9	C1—S1—Rh1	115.00 (8)
C6—C7—C8	122.7 (2)	C10—Rh1—C11	38.78 (8)
C6—C7—Rh1	70.61 (12)	C10—Rh1—C6	97.75 (8)
C8—C7—Rh1	112.66 (14)	C11—Rh1—C6	81.39 (9)
C6—C7—H7	114.4	C10—Rh1—C7	82.10 (8)
C8—C7—H7	114.4	C11—Rh1—C7	90.31 (9)
Rh1—C7—H7	114.4	C6—Rh1—C7	37.95 (8)
C7—C8—C9	112.29 (17)	C10—Rh1—S1	83.28 (6)
C7—C8—H8A	109.1	C11—Rh1—S1	89.52 (7)
C9—C8—H8A	109.1	C6—Rh1—S1	163.50 (6)
C7—C8—H8B	109.1	C7—Rh1—S1	156.79 (6)
C9—C8—H8B	109.1	C10—Rh1—Cl1	159.79 (6)
H8A—C8—H8B	107.9	C11—Rh1—Cl1	160.89 (6)
C10—C9—C8	113.14 (18)	C6—Rh1—Cl1	87.71 (7)
C10—C9—H9A	109	C7—Rh1—Cl1	90.67 (6)
C8—C9—H9A	109	S1—Rh1—Cl1	96.98 (3)
C10—C9—H9B	109		

C13—C6—C7—C8	1.3 (3)	C12—C11—Rh1—C10	−118.3 (2)
Rh1—C6—C7—C8	105.03 (19)	C10—C11—Rh1—C6	113.96 (14)
C13—C6—C7—Rh1	−103.7 (2)	C12—C11—Rh1—C6	−4.32 (16)
C6—C7—C8—C9	−92.6 (2)	C10—C11—Rh1—C7	76.94 (13)
Rh1—C7—C8—C9	−11.7 (2)	C12—C11—Rh1—C7	−41.35 (16)
C7—C8—C9—C10	29.4 (3)	C10—C11—Rh1—S1	−79.85 (12)
C8—C9—C10—C11	47.8 (3)	C12—C11—Rh1—S1	161.87 (15)
C8—C9—C10—Rh1	−33.1 (2)	C10—C11—Rh1—Cl1	169.87 (14)
C9—C10—C11—C12	4.0 (3)	C12—C11—Rh1—Cl1	51.6 (3)
Rh1—C10—C11—C12	106.7 (2)	C7—C6—Rh1—C10	−66.43 (14)
C9—C10—C11—Rh1	−102.7 (2)	C13—C6—Rh1—C10	54.45 (16)
C10—C11—C12—C13	−92.5 (3)	C7—C6—Rh1—C11	−101.71 (14)
Rh1—C11—C12—C13	−11.1 (2)	C13—C6—Rh1—C11	19.17 (15)
C7—C6—C13—C12	50.9 (3)	C13—C6—Rh1—C7	120.9 (2)
Rh1—C6—C13—C12	−30.8 (2)	C7—C6—Rh1—S1	−158.94 (17)
C11—C12—C13—C6	27.4 (3)	C13—C6—Rh1—S1	−38.1 (3)
N2—C1—N1—C2	4.4 (3)	C7—C6—Rh1—Cl1	94.03 (13)
S1—C1—N1—C2	−175.86 (15)	C13—C6—Rh1—Cl1	−145.09 (15)
C3—C2—N1—C1	174.37 (18)	C6—C7—Rh1—C10	113.53 (14)
N1—C1—N2—C4	178.38 (18)	C8—C7—Rh1—C10	−4.79 (15)
S1—C1—N2—C4	−1.3 (3)	C6—C7—Rh1—C11	75.50 (14)
C5—C4—N2—C1	166.75 (19)	C8—C7—Rh1—C11	−42.81 (16)
N1—C1—S1—Rh1	−9.74 (19)	C8—C7—Rh1—C6	−118.3 (2)
N2—C1—S1—Rh1	169.97 (13)	C6—C7—Rh1—S1	164.99 (12)
C9—C10—Rh1—C11	120.9 (2)	C8—C7—Rh1—S1	46.7 (2)
C11—C10—Rh1—C6	−65.76 (14)	C6—C7—Rh1—Cl1	−85.41 (13)
C9—C10—Rh1—C6	55.16 (16)	C8—C7—Rh1—Cl1	156.27 (15)

C11—C10—Rh1—C7	−100.44 (14)	C1—S1—Rh1—C10	−160.61 (10)
C9—C10—Rh1—C7	20.48 (15)	C1—S1—Rh1—C11	−122.23 (10)
C11—C10—Rh1—S1	97.64 (13)	C1—S1—Rh1—C6	−66.0 (2)
C9—C10—Rh1—S1	−141.43 (15)	C1—S1—Rh1—C7	148.12 (16)
C11—C10—Rh1—Cl1	−170.40 (13)	C1—S1—Rh1—Cl1	39.75 (8)
C9—C10—Rh1—Cl1	−49.5 (3)		

Hydrogen-bond geometry (Å, º)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
N1—H1···Cl1	0.86	2.39	3.152 (3)	148
N2—H2···Cl1^{i}	0.86	2.89	3.356 (3)	116

Symmetry code: (i) *x*+1, *y*, *z*.