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Diaquabis{2-hydroxy-5-[(pyridin-2-yl)methylideneamino]benzoato-κ^2N,N'}nickel(II) dihydrate

Meiqin Zha, Xing Li,* Yue Bing and Zhengbing Luo

Faculty of Materials Science and Chemical Engineering, Ningbo University, Ningbo, Zhejiang 315211, People's Republic of China
Correspondence e-mail: lix905@126.com

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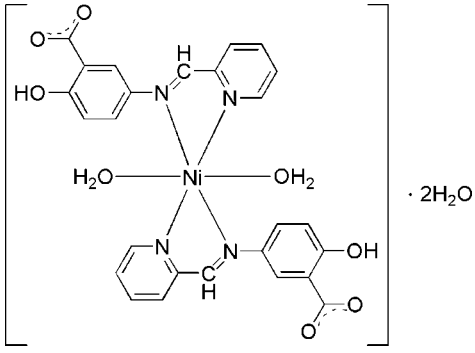
Key indicators: single-crystal X-ray study; T = 173 K; mean σ(C–C) = 0.003 Å; R factor = 0.030; wR factor = 0.074; data-to-parameter ratio = 11.4.

In the title complex, $[Ni(C_{13}H_9N_2O_3)_2(H_2O)_2]\cdot 2H_2O$, the Ni^{II} atom, located on a twofold rotation axis, is in a distorted octahedral geometry, defined by four N atoms from two 2-hydroxy-5-[(pyridin-2-yl)methylideneamino]benzoate ligands and two O atoms from two water molecules. In the crystal, intermolecular O—H···O hydrogen bonds link the complex molecules and uncoordinated water molecules into a three-dimensional network. Intramolecular O—H···O hydrogen bonds are present between the hydroxy and carboxylate groups.

Related literature

For the biological activity of Schiff base compounds, see: Ali *et al.* (2002); Cukurovali *et al.* (2002); Tarafder *et al.* (2002).



Experimental

Crystal data

$[Ni(C_{13}H_9N_2O_3)_2(H_2O)_2]\cdot 2H_2O$
M_r = 613.22
Orthorhombic, *Pbcn*
a = 15.7628 (6) Å
b = 10.5672 (3) Å
c = 15.6178 (6) Å
V = 2601.44 (16) Å^3
Z = 4
Mo $K\alpha$ radiation
μ = 0.81 mm^{-1}
T = 173 K
0.41 × 0.34 × 0.12 mm

Data collection

Bruker APEXII CCD diffractometer
Absorption correction: multi-scan (*SADABS*; Sheldrick, 1996)
T_{min} = 0.725, T_{max} = 0.907
14975 measured reflections
2294 independent reflections
1599 reflections with $I > 2\sigma(I)$
R_{int} = 0.044

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.030
$wR(F^2)$ = 0.074
S = 0.89
2294 reflections
202 parameters
4 restraints
H atoms treated by a mixture of independent and constrained refinement
$\Delta\rho_{max}$ = 0.20 e Å^{-3}
$\Delta\rho_{min}$ = −0.41 e Å^{-3}

Table 1
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
O3—H3A···O1	0.84	1.72	2.475 (2)	149
O4—H4B···O2^{i}	0.83 (2)	1.81 (2)	2.621 (2)	168 (3)
O4—H4C···O5	0.84 (2)	1.91 (2)	2.744 (3)	173 (2)
O5—H5B···O3^{ii}	0.82 (2)	2.05 (2)	2.870 (3)	178 (3)
O5—H5C···O1^{iii}	0.84 (2)	1.99 (2)	2.793 (3)	160 (3)

Symmetry codes: (i) $x, y+1, z$; (ii) $x+\frac{1}{2}, -y+\frac{1}{2}, -z$; (iii) $-x+1, -y, -z$.

Data collection: *APEX2* (Bruker, 2007); cell refinement: *SAINT* (Bruker, 2007); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HY2372).

References

Ali, M. A., Mirza, A. H., Butcher, R. J. & Tarafder, M. T. H. (2002). *Inorg. Biochem.* **92**, 141–148.
Bruker (2007). *APEX2* and *SAINT*. Bruker AXS Inc., Madison, Wisconsin, USA.
Cukurovali, A., Yilmaz, I., Ozmen, H. & Ahmedzade, M. (2002). *Transition Met. Chem.* **27**, 171–176.
Sheldrick, G. M. (1996). *SADABS*. University of Göttingen, Germany.
Sheldrick, G. M. (2008). *Acta Cryst.* A**64**, 112–122.
Tarafder, M. T. H., Jin, K. T., Crouse, K. A., Ali, A. M. & Yamin, B. M. (2002). *Polyhedron*, **21**, 2547–2554.

supporting information

Acta Cryst. (2010). E**66**, m1570 [https://doi.org/10.1107/S1600536810046118]

Diaquabis{2-hydroxy-5-[(pyridin-2-yl)methylideneamino]benzoato-κ^2N,N'}nickel(II) dihydrate

Meiqin Zha, Xing Li, Yue Bing and Zhengbing Luo

S1. Comment

Schiff base compounds have been of great interest for many years. These compounds play an important role in antitumor, antimicrobial and antiviral activities (Ali *et al.*, 2002; Cukurovali *et al.*, 2002; Tarafder *et al.*, 2002). As an extension of the work on the structural characterization of Schiff base compounds, the crystal structure of the title compound is reported here.

The title compound is a mononuclear nickel(II) complex, as shown in Fig. 1. The Ni^{II} atom, lying on a twofold rotation axis, is six-coordinated in a distorted octahedral geometry, defined by four N donors from two Schiff base ligands, two O atoms from two coordinated water molecules. The molecular formula contains two uncoordinated water molecules. The Ni—N bond lengths are 2.0754 (18) and 2.1347 (17) Å, and the Ni—O distance is 2.0380 (17) Å. Intramolecular O—H···O hydrogen bonds between the hydroxy and carboxylate groups are observed (Table 1). In the crystal, intermolecular O—H···O hydrogen bonds link the complex molecules and uncoordinated water molecules into a three-dimensional network (Fig. 2).

S2. Experimental

5-Aminosalicylic acid (1.53 g, 10 mmol), 2-pyridinecarboxaldehyde (1 ml, 10 mmol) and triethylamine (1 ml, 10 mmol) were mixed in 50 ml ethanol in a round flask. The mixture was refluxed with agitation for 4 h at 323 K to give a yellow precipitate. After filtration and washing the precipitate with ethanol, a pure Schiff base ligand, 5-[(pyridin-2-yl)methyl-eneamino]-2-hydroxybenzoic acid (yield: 2.02 g, 84%) was obtained.

A mixture of 5-[(pyridin-2-yl)methyleneamino]-2-hydroxybenzoic acid (0.024 g, 0.1 mmol), $Ni(CH_3CO_2)_2.2H_2O$ (0.025 g, 0.1 mmol) and ethanol (20 ml) was heated at 273 K for 30 min to give a red solution. After evaporating the solution at room temperature for one week, red crystals were obtained (yield: 65%).

S3. Refinement

H atoms attached to C atoms and O3 were placed in calculated positions and treated using a riding model, with C—H = 0.95 and O—H = 0.84 Å and U_{iso}(H) = 1.2U_{eq}(C) or 1.5U_{eq}(O). H atoms attached to water molecules (O4 and O5) were located in a difference Fourier map and refined isotropically.

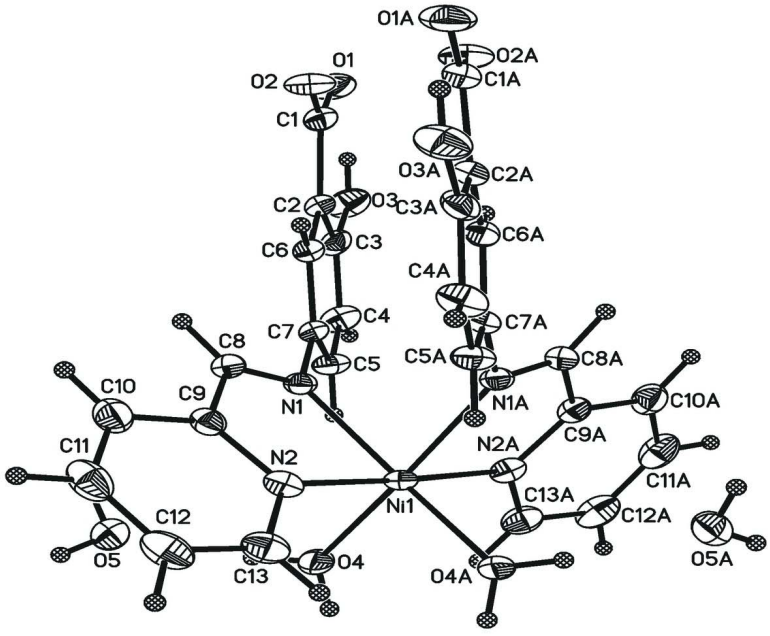

Figure 1

The molecular structure of the title compound. Displacement ellipsoids are drawn at the 30% probability level. [Symmetry code: (A) -x+1, y, -z+1/2.]

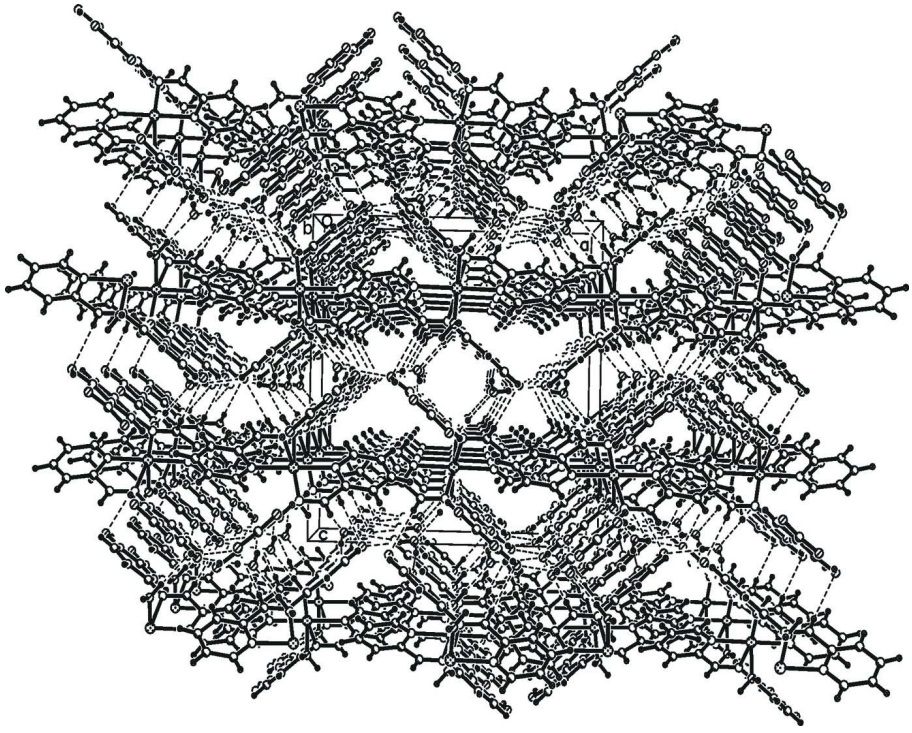

Figure 2

Three-dimensional supramolecular network in the title compound. Dashed lines denote hydrogen bonds.

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Crystal data

$[Ni(C_{13}H_9N_2O_3)_2(H_2O)_2]\cdot 2H_2O$
$M_r = 613.22$
Orthorhombic, *Pbcn*
Hall symbol: -P 2n 2ab
$a = 15.7628\ (6)$ Å
$b = 10.5672\ (3)$ Å
$c = 15.6178\ (6)$ Å
$V = 2601.44\ (16)$ Å^3
$Z = 4$
$F(000) = 1272$
$D_x = 1.566$ Mg m^{-3}
Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
Cell parameters from 14975 reflections
$\theta = 3.9–25.0°$
$\mu = 0.81$ mm^{-1}
$T = 173$ K
Platelet, red
0.41 × 0.34 × 0.12 mm

Data collection

Bruker APEXII CCD diffractometer
Radiation source: fine-focus sealed tube
Graphite monochromator
φ and ω scans
Absorption correction: multi-scan (*SADABS*; Sheldrick, 1996)
$T_{min} = 0.725$, $T_{max} = 0.907$
14975 measured reflections
2294 independent reflections
1599 reflections with $I > 2\sigma(I)$
$R_{int} = 0.044$
$\theta_{max} = 25.0°$, $\theta_{min} = 3.9°$
$h = -18\rightarrow18$
$k = -12\rightarrow12$
$l = -18\rightarrow13$

Refinement

Refinement on F^2
Least-squares matrix: full
$R[F^2 > 2\sigma(F^2)] = 0.030$
$wR(F^2) = 0.074$
$S = 0.89$
2294 reflections
202 parameters
4 restraints
Primary atom site location: structure-invariant direct methods
Secondary atom site location: difference Fourier map
Hydrogen site location: inferred from neighbouring sites
H atoms treated by a mixture of independent and constrained refinement
$w = 1/[\sigma^2(F_o^2) + (0.045P)^2]$ where $P = (F_o^2 + 2F_c^2)/3$
$(\Delta/\sigma)_{max} < 0.001$
$\Delta\rho_{max} = 0.20$ e Å^{-3}
$\Delta\rho_{min} = -0.41$ e Å^{-3}

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}
Ni1	0.5000	0.40233 (3)	0.2500	0.02879 (13)
C1	0.42670 (16)	−0.1813 (2)	0.09408 (16)	0.0418 (6)
C2	0.42072 (13)	−0.04024 (19)	0.08275 (14)	0.0314 (5)
C3	0.35591 (14)	0.0140 (2)	0.03319 (15)	0.0371 (6)
C4	0.34858 (15)	0.1436 (2)	0.02799 (15)	0.0422 (6)
H4A	0.3049	0.1798	−0.0061	0.051*
C5	0.40373 (14)	0.2213 (2)	0.07156 (14)	0.0376 (5)
H5A	0.3975	0.3106	0.0685	0.045*
C6	0.47632 (13)	0.03895 (19)	0.12562 (13)	0.0306 (5)
H6A	0.5204	0.0032	0.1593	0.037*
C7	0.46889 (13)	0.16841 (18)	0.12030 (13)	0.0281 (5)
C8	0.60516 (14)	0.2224 (2)	0.16257 (13)	0.0340 (5)
H8A	0.6237	0.1479	0.1343	0.041*

C9	0.66617 (13)	0.3041 (2)	0.20555 (14)	0.0333 (5)
C10	0.75249 (14)	0.2905 (2)	0.19473 (15)	0.0463 (6)
H10A	0.7747	0.2214	0.1629	0.056*
C11	0.80586 (17)	0.3784 (3)	0.23054 (17)	0.0543 (7)
H11A	0.8656	0.3711	0.2242	0.065*
C12	0.77113 (16)	0.4765 (3)	0.27551 (17)	0.0529 (7)
H12A	0.8067	0.5404	0.2989	0.064*
C13	0.68392 (16)	0.4830 (2)	0.28702 (16)	0.0445 (6)
H13A	0.6608	0.5500	0.3202	0.053*
N1	0.52662 (11)	0.25156 (15)	0.16309 (11)	0.0290 (4)
N2	0.63159 (11)	0.39730 (16)	0.25264 (11)	0.0347 (4)
O1	0.37216 (11)	−0.24708 (15)	0.05397 (12)	0.0591 (5)
O2	0.48289 (12)	−0.22398 (14)	0.14171 (12)	0.0561 (5)
O3	0.29873 (10)	−0.06080 (16)	−0.00716 (11)	0.0531 (5)
H3A	0.3095	−0.1370	0.0035	0.080*
O4	0.50698 (13)	0.53118 (16)	0.15316 (12)	0.0501 (5)
O5	0.62630 (14)	0.48046 (19)	0.02994 (14)	0.0614 (5)
H4B	0.4935 (15)	0.6066 (17)	0.1536 (18)	0.062 (9)*
H4C	0.5432 (14)	0.522 (2)	0.1140 (13)	0.055 (9)*
H5C	0.618 (2)	0.419 (2)	−0.0026 (18)	0.084 (11)*
H5B	0.6754 (14)	0.505 (3)	0.024 (2)	0.109 (15)*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Ni1	0.0350 (2)	0.0170 (2)	0.0344 (2)	0.000	0.00107 (18)	0.000
C1	0.0476 (16)	0.0267 (13)	0.0512 (15)	−0.0005 (11)	0.0112 (13)	−0.0075 (12)
C2	0.0363 (13)	0.0238 (12)	0.0340 (12)	0.0022 (9)	0.0041 (10)	−0.0050 (10)
C3	0.0412 (14)	0.0343 (14)	0.0359 (12)	0.0005 (10)	0.0030 (11)	−0.0131 (10)
C4	0.0479 (15)	0.0378 (14)	0.0409 (13)	0.0117 (11)	−0.0143 (11)	−0.0048 (11)
C5	0.0493 (14)	0.0233 (12)	0.0402 (13)	0.0051 (10)	−0.0040 (11)	−0.0003 (10)
C6	0.0308 (13)	0.0253 (11)	0.0357 (13)	0.0038 (9)	0.0029 (10)	0.0004 (10)
C7	0.0325 (12)	0.0212 (11)	0.0306 (12)	0.0003 (9)	0.0036 (10)	−0.0048 (9)
C8	0.0411 (15)	0.0244 (12)	0.0364 (13)	0.0011 (10)	0.0046 (11)	0.0001 (10)
C9	0.0329 (13)	0.0308 (13)	0.0362 (13)	−0.0016 (9)	−0.0001 (11)	0.0065 (11)
C10	0.0385 (15)	0.0470 (16)	0.0534 (16)	−0.0012 (11)	0.0037 (12)	0.0096 (13)
C11	0.0358 (14)	0.0617 (19)	0.0654 (19)	−0.0079 (13)	−0.0070 (13)	0.0216 (15)
C12	0.0469 (17)	0.0557 (18)	0.0563 (17)	−0.0195 (14)	−0.0180 (13)	0.0137 (14)
C13	0.0522 (16)	0.0393 (15)	0.0419 (14)	−0.0103 (12)	−0.0117 (12)	0.0048 (12)
N1	0.0346 (11)	0.0211 (9)	0.0313 (10)	0.0009 (7)	−0.0002 (8)	0.0019 (8)
N2	0.0387 (10)	0.0290 (10)	0.0365 (10)	−0.0064 (8)	−0.0045 (9)	0.0039 (9)
O1	0.0604 (12)	0.0302 (10)	0.0869 (13)	−0.0087 (8)	−0.0029 (10)	−0.0191 (9)
O2	0.0731 (13)	0.0220 (9)	0.0731 (12)	0.0054 (8)	−0.0089 (10)	0.0025 (9)
O3	0.0497 (11)	0.0443 (10)	0.0654 (12)	−0.0009 (8)	−0.0160 (9)	−0.0216 (9)
O4	0.0730 (13)	0.0260 (10)	0.0512 (11)	0.0140 (9)	0.0232 (10)	0.0116 (8)
O5	0.0665 (15)	0.0488 (13)	0.0688 (14)	−0.0066 (11)	0.0228 (12)	−0.0147 (11)

Geometric parameters (Å, º)

Ni1—O4	2.0380 (17)	C8—C9	1.456 (3)
Ni1—N2	2.0754 (18)	C8—H8A	0.9500
Ni1—N1	2.1347 (17)	C9—N2	1.345 (3)
C1—O2	1.241 (3)	C9—C10	1.378 (3)
C1—O1	1.271 (3)	C10—C11	1.372 (3)
C1—C2	1.504 (3)	C10—H10A	0.9500
C2—C6	1.384 (3)	C11—C12	1.367 (4)
C2—C3	1.404 (3)	C11—H11A	0.9500
C3—O3	1.354 (3)	C12—C13	1.388 (3)
C3—C4	1.377 (3)	C12—H12A	0.9500
C4—C5	1.376 (3)	C13—N2	1.337 (3)
C4—H4A	0.9500	C13—H13A	0.9500
C5—C7	1.395 (3)	O3—H3A	0.8400
C5—H5A	0.9500	O4—H4B	0.825 (16)
C6—C7	1.376 (3)	O4—H4C	0.843 (16)
C6—H6A	0.9500	O5—H5C	0.837 (17)
C7—N1	1.431 (3)	O5—H5B	0.823 (18)
C8—N1	1.276 (3)		

O4^{i}—Ni1—O4	96.16 (11)	C6—C7—C5	119.61 (19)
O4^{i}—Ni1—N2	93.23 (8)	C6—C7—N1	121.90 (19)
O4—Ni1—N2	88.74 (7)	C5—C7—N1	118.48 (18)
O4^{i}—Ni1—N2^{i}	88.74 (7)	N1—C8—C9	119.7 (2)
O4—Ni1—N2^{i}	93.23 (8)	N1—C8—H8A	120.2
N2—Ni1—N2^{i}	177.06 (9)	C9—C8—H8A	120.2
O4^{i}—Ni1—N1	168.85 (7)	N2—C9—C10	122.9 (2)
O4—Ni1—N1	90.93 (7)	N2—C9—C8	114.73 (19)
N2—Ni1—N1	78.29 (7)	C10—C9—C8	122.3 (2)
N2^{i}—Ni1—N1	99.48 (7)	C11—C10—C9	119.0 (2)
O4^{i}—Ni1—N1^{i}	90.93 (7)	C11—C10—H10A	120.5
O4—Ni1—N1^{i}	168.85 (7)	C9—C10—H10A	120.5
N2—Ni1—N1^{i}	99.48 (6)	C12—C11—C10	118.5 (2)
N2^{i}—Ni1—N1^{i}	78.29 (7)	C12—C11—H11A	120.7
N1—Ni1—N1^{i}	83.45 (9)	C10—C11—H11A	120.7
O2—C1—O1	125.4 (2)	C11—C12—C13	120.0 (2)
O2—C1—C2	118.4 (2)	C11—C12—H12A	120.0
O1—C1—C2	116.2 (2)	C13—C12—H12A	120.0
C6—C2—C3	118.72 (19)	N2—C13—C12	121.7 (2)
C6—C2—C1	120.2 (2)	N2—C13—H13A	119.1
C3—C2—C1	121.0 (2)	C12—C13—H13A	119.1
O3—C3—C4	119.8 (2)	C8—N1—C7	117.79 (18)
O3—C3—C2	120.2 (2)	C8—N1—Ni1	112.01 (14)
C4—C3—C2	120.0 (2)	C7—N1—Ni1	129.08 (13)
C5—C4—C3	120.8 (2)	C13—N2—C9	117.7 (2)
C5—C4—H4A	119.6	C13—N2—Ni1	127.38 (16)
C3—C4—H4A	119.6	C9—N2—Ni1	114.39 (14)

C4—C5—C7	119.7 (2)	C3—O3—H3A	109.5
C4—C5—H5A	120.2	Ni1—O4—H4B	129 (2)
C7—C5—H5A	120.2	Ni1—O4—H4C	120.0 (17)
C7—C6—C2	121.2 (2)	H4B—O4—H4C	107 (3)
C7—C6—H6A	119.4	H5C—O5—H5B	109 (3)
C2—C6—H6A	119.4		

Symmetry code: (i) $-x+1, y, -z+1/2$.

Hydrogen-bond geometry (Å, º)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
O3—H3*A*···O1	0.84	1.72	2.475 (2)	149
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O5—H5*C*···O1^{iv}	0.84 (2)	1.99 (2)	2.793 (3)	160 (3)

Symmetry codes: (ii) $x, y+1, z$; (iii) $x+1/2, -y+1/2, -z$; (iv) $-x+1, -y, -z$.