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 [(Dibenzo[*b,d*]thiophen-4-yl)tellanyl]-methanethiol

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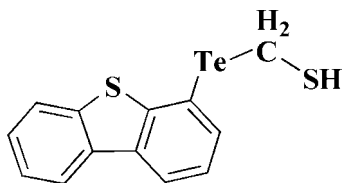
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 Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{C}-\text{C}) = 0.007$ Å; R factor = 0.034; wR factor = 0.083; data-to-parameter ratio = 17.2.

In the title compound, $\text{C}_{13}\text{H}_{10}\text{S}_2\text{Te}$, the dibenzothiophene moiety is almost planar, the maximum atomic deviation being 0.055 (5) Å. The two $\text{Te}-\text{C}$ bonds are nearly perpendicular to each other with a $\text{C}-\text{Te}-\text{C}$ bond angle of $93.0(2)^\circ$. An intermolecular $\text{C}-\text{H}\cdots\pi$ interaction is present between the methylene group and thiophene ring.

Related literature

For general background to the field-effect transistors of organotellurium derivatives, see: Inokuchi *et al.* (1987). For related structures, see: Kobayashi *et al.* (2005).



Experimental

Crystal data

 $\text{C}_{13}\text{H}_{10}\text{S}_2\text{Te}$
 $M_r = 357.93$

 Orthorhombic, $P2_12_12_1$
 $a = 5.49518(12)$ Å

 $b = 12.1422(3)$ Å

 $c = 19.0529(5)$ Å

 $V = 1271.27(6)$ Å³
 $Z = 4$

 Mo $K\alpha$ radiation
 $\mu = 2.64$ mm⁻¹
 $T = 293$ K
 $0.58 \times 0.45 \times 0.31$ mm

Data collection

 Oxford Diffraction Xcalibur Eos
 Gemini diffractometer

 Absorption correction: multi-scan
 (*CrysAlis PRO RED*; Oxford
 Diffraction, 2009)

 $T_{\min} = 0.310$, $T_{\max} = 0.495$

 5117 measured reflections
 2509 independent reflections
 2248 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.054$

Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.034$
 $wR(F^2) = 0.083$
 $S = 0.98$

2509 reflections

146 parameters

H-atom parameters constrained

 $\Delta\rho_{\max} = 0.81$ e Å⁻³
 $\Delta\rho_{\min} = -0.55$ e Å⁻³

 Absolute structure: Flack (1983),
 975 Friedel pairs

Flack parameter: 0.01 (3)

Table 1

Hydrogen-bond geometry (Å, °).

 C_g is the centroid of the thiophene ring.

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{C13}-\text{H13A}\cdots\text{Cg}^i$	0.97	2.90	3.846 (6)	166

 Symmetry code: (i) $x - \frac{1}{2}, -y + \frac{3}{2}, -z$.

Data collection: *CrysAlis PRO CCD* (Oxford Diffraction, 2009); cell refinement: *CrysAlis PRO CCD*; data reduction: *CrysAlis PRO RED* (Oxford Diffraction, 2009); program(s) used to solve structure: *SHELXTL* (Sheldrick, 2008); program(s) used to refine structure: *SHELXTL*; molecular graphics: *SHELXTL*; software used to prepare material for publication: *SHELXTL*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: XU5144).

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supporting information

Acta Cryst. (2011). E67, o519 [doi:10.1107/S160053681100273X]

[(Dibenzo[*b,d*]thiophen-4-yl)tellanyl]methanethiol**Zuo-Qin Liang and Xu-Tang Tao****S1. Comment**

The design and synthesis of high-performance organic semiconductor materials have been an active topic in the area of field-effect transistors. Recently, organotellurium compounds have received significant attention due to the high charge carrier mobility (Inokuchi *et al.*, 1987). In this paper, we report the synthesis and the crystal structure of the title compound (I).

The asymmetric unit of the title compound is shown in Fig. 1. The dibenzothiophene group possesses perfect planarity: the dihedral angle between the two phenyls is 4.11°. The two Te—C bonds are nearly perpendicular to each other with a C—Te—C bond angle of 93.0 (2)°. In the packing structure (Fig. 2), the molecules are packed into molecular columns along the *a* axis through intermolecular Te⋯ π interactions between the tellurium atom and the phenyl ring (C7—C12). The contact distance of Te⋯C9 is 3.68 (5) Å, that of Te⋯C10 is 3.47 (5) Å, and that of Te⋯C11 is 3.73 (7) Å. These contact distances are significantly shorter than the sum of the van der Waals radii of tellurium and aromatic carbon atoms (Kobayashi *et al.*, 2005). The molecular columns are connected together by intermolecular C⋯S and S⋯S interactions. The contact distances of C13⋯S1 and S1⋯S2 are 3.43 (5) Å and 3.51 (2) Å, respectively, which are obviously shorter than the sum of the corresponding van der Waals radii. There are no classic hydrogen bonds in the crystal structure.

S2. Experimental

Addition of a 1.6 *M* solution of n-BuLi in hexane (8.50 ml, 13.6 mmol) to the THF solution of dibenzothiophene (2.50 g, 13.6 mmol) at room temperature under an Ar atmosphere. The reaction mixture was stirred for 2 h, and then tellurium powder (1.70 g, 13.3 mmol) was added. After 3 h, the reaction mixture was poured into a beaker containing 200 ml cold distilled water and oxidized by passing oxygen at a moderate rate for 1 h. The organic solvent was evaporated, and the suspension was extracted with dichloromethane. The organic extracts were washed by brine, and dried by anhydrous CaCl₂. After the solvent was removed, the crude product was chromatographed on silica gel using petroleum as eluent to give the title compound (0.27 g, yield 5.56%). Single crystals suitable for X-ray diffraction were obtained by very slow evaporation of a chloroform/ethanol solution.

S3. Refinement

H atoms bonded to C atoms were fixed geometrically and allowed to ride on their attached atoms, with C—H = 0.93–0.97 Å and with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$. H atom bonded to S atom was refined using a rotating model, with S—H = 1.20 Å and with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{S})$.

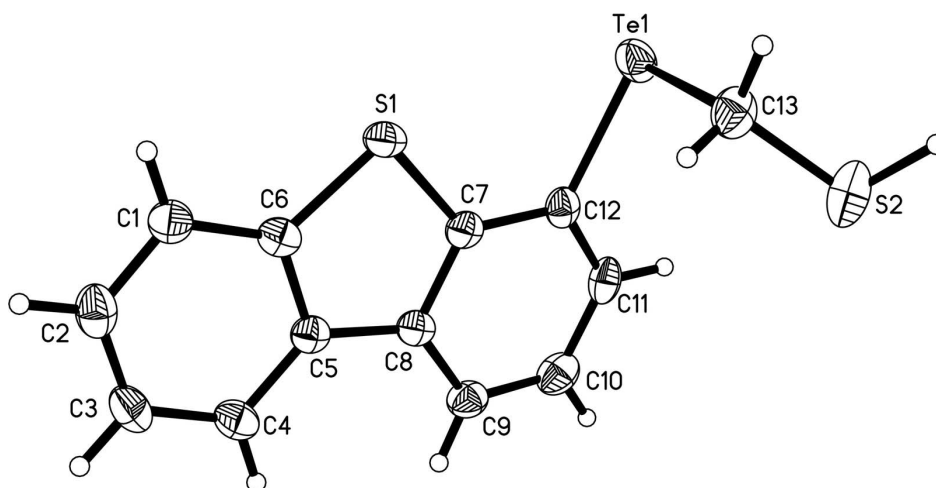


Figure 1

The molecular structure of (I), showing 30% probability displacement ellipsoids and the atom-numbering scheme.

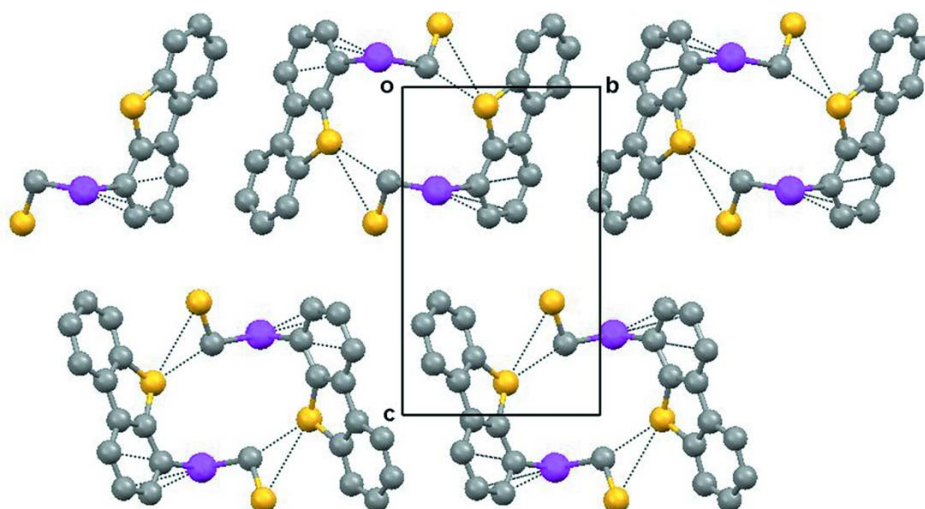


Figure 2

Stacking diagram of (I), viewed down the *a* axis, showing intermolecular Te \cdots π , C \cdots S and S \cdots S interactions (dashed lines). The hydrogen atoms have been omitted for clarity.

{8-thiatricyclo[7.4.0.0^{2,7}]trideca-1(9),2(7),3,5,10,12-hexaen-6-yltellanyl}methanethiol

Crystal data

C₁₃H₁₀S₂Te
M_r = 357.93
 Orthorhombic, *P*2₁2₁2₁
a = 5.49518 (12) Å
b = 12.1422 (3) Å
c = 19.0529 (5) Å
V = 1271.27 (6) Å³
Z = 4
F(000) = 688

D_x = 1.870 Mg m⁻³
 Mo *K* α radiation, λ = 0.7107 Å
 Cell parameters from 3403 reflections
 θ = 3.5–28.8°
 μ = 2.64 mm⁻¹
T = 293 K
 Block, colorless
 0.58 × 0.45 × 0.31 mm

Data collection

Oxford Diffraction Xcalibur Eos Gemini diffractometer
 Radiation source: Enhance (Mo) X-ray Source
 Graphite monochromator
 Detector resolution: 16.0355 pixels mm⁻¹
 ω scans
 Absorption correction: multi-scan (CrysAlis PRO RED; Oxford Diffraction, 2009)
 $T_{\min} = 0.310$, $T_{\max} = 0.495$

5117 measured reflections
 2509 independent reflections
 2248 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.054$
 $\theta_{\max} = 26.4^\circ$, $\theta_{\min} = 3.5^\circ$
 $h = -6 \rightarrow 7$
 $k = -15 \rightarrow 7$
 $l = -21 \rightarrow 24$

Refinement

Refinement on F^2
 Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.034$
 $wR(F^2) = 0.083$
 $S = 0.98$
 2509 reflections
 146 parameters
 0 restraints
 Primary atom site location: structure-invariant direct methods
 Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites
 H-atom parameters constrained
 $w = 1/[\sigma^2(F_o^2) + (0.0502P)^2]$
 where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} < 0.001$
 $\Delta\rho_{\max} = 0.81 \text{ e } \text{\AA}^{-3}$
 $\Delta\rho_{\min} = -0.55 \text{ e } \text{\AA}^{-3}$
 Absolute structure: Flack (1983), 975 Friedel pairs
 Absolute structure parameter: 0.01 (3)

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Te1	0.65007 (6)	0.83287 (3)	0.120133 (19)	0.05366 (13)
S1	0.8260 (2)	0.97156 (10)	-0.03248 (6)	0.0466 (3)
S2	1.0630 (4)	0.65872 (16)	0.17855 (9)	0.0778 (5)
H2	0.9995	0.5698	0.1991	0.093*
C12	0.9348 (8)	0.9486 (4)	0.1106 (2)	0.0401 (10)
C8	1.1849 (8)	1.0679 (3)	0.0346 (2)	0.0373 (9)
C4	1.3695 (10)	1.1778 (4)	-0.0677 (3)	0.0486 (11)
H4	1.4968	1.2065	-0.0412	0.058*
C9	1.3274 (10)	1.0953 (4)	0.0922 (3)	0.0466 (11)
H9	1.4611	1.1415	0.0868	0.056*
C2	1.1594 (11)	1.1620 (5)	-0.1780 (3)	0.0551 (12)
H2A	1.1483	1.1809	-0.2252	0.066*
C13	0.8876 (11)	0.6962 (4)	0.1038 (3)	0.0570 (15)
H13A	0.7910	0.6332	0.0895	0.068*

H13B	0.9975	0.7137	0.0656	0.068*
C3	1.3448 (11)	1.2056 (4)	-0.1376 (3)	0.0545 (12)
H3	1.4548	1.2543	-0.1579	0.065*
C6	1.0149 (8)	1.0623 (4)	-0.0788 (2)	0.0394 (10)
C7	0.9883 (8)	0.9950 (4)	0.0445 (2)	0.0381 (10)
C11	1.0743 (9)	0.9798 (4)	0.1662 (3)	0.0471 (12)
H11	1.0404	0.9519	0.2105	0.057*
C1	0.9912 (9)	1.0912 (4)	-0.1496 (3)	0.0486 (12)
H1	0.8648	1.0631	-0.1767	0.058*
C10	1.2698 (10)	1.0539 (5)	0.1574 (3)	0.0534 (13)
H10	1.3611	1.0750	0.1962	0.064*
C5	1.2002 (7)	1.1060 (4)	-0.0370 (2)	0.0380 (10)

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Te1	0.04305 (17)	0.05162 (19)	0.0663 (2)	0.00382 (16)	0.01498 (15)	0.01196 (17)
S1	0.0380 (5)	0.0542 (7)	0.0475 (6)	-0.0110 (6)	-0.0028 (5)	0.0017 (5)
S2	0.1032 (13)	0.0683 (10)	0.0621 (9)	0.0296 (10)	-0.0107 (8)	0.0080 (8)
C12	0.041 (2)	0.036 (2)	0.042 (3)	0.0116 (18)	0.006 (2)	0.004 (2)
C8	0.035 (2)	0.030 (2)	0.048 (2)	0.0031 (18)	0.001 (2)	-0.0034 (18)
C4	0.047 (3)	0.041 (2)	0.058 (3)	-0.010 (3)	0.003 (2)	-0.003 (2)
C9	0.044 (2)	0.042 (2)	0.054 (3)	-0.003 (2)	-0.008 (2)	-0.004 (2)
C2	0.066 (3)	0.055 (3)	0.045 (2)	0.011 (4)	0.010 (3)	0.006 (2)
C13	0.067 (4)	0.043 (3)	0.061 (3)	0.009 (2)	-0.006 (3)	-0.005 (2)
C3	0.058 (3)	0.045 (3)	0.060 (3)	-0.006 (3)	0.013 (3)	0.010 (2)
C6	0.036 (2)	0.036 (2)	0.046 (3)	0.0048 (19)	0.008 (2)	-0.007 (2)
C7	0.034 (2)	0.029 (2)	0.052 (3)	0.0093 (17)	0.001 (2)	-0.001 (2)
C11	0.060 (3)	0.040 (3)	0.041 (2)	0.018 (2)	0.000 (2)	0.004 (2)
C1	0.049 (3)	0.047 (3)	0.050 (3)	-0.004 (2)	-0.003 (2)	-0.001 (2)
C10	0.057 (3)	0.052 (3)	0.051 (3)	0.004 (3)	-0.015 (2)	-0.010 (2)
C5	0.036 (2)	0.033 (2)	0.045 (2)	0.0048 (17)	0.0011 (19)	-0.0028 (19)

Geometric parameters (Å, °)

Te1—C12	2.111 (5)	C9—H9	0.9300
Te1—C13	2.134 (5)	C9—C10	1.376 (8)
S1—C6	1.753 (5)	C2—H2A	0.9300
S1—C7	1.739 (5)	C2—C3	1.382 (8)
S2—H2	1.2000	C2—C1	1.374 (8)
S2—C13	1.778 (6)	C13—H13A	0.9700
C12—C7	1.412 (7)	C13—H13B	0.9700
C12—C11	1.360 (7)	C3—H3	0.9300
C8—C9	1.389 (6)	C6—C1	1.399 (7)
C8—C7	1.409 (6)	C6—C5	1.397 (6)
C8—C5	1.443 (6)	C11—H11	0.9300
C4—H4	0.9300	C11—C10	1.411 (8)
C4—C3	1.381 (8)	C1—H1	0.9300

C4—C5	1.403 (7)	C10—H10	0.9300
Te1—C13—H13A	108.6	H13A—C13—H13B	107.6
Te1—C13—H13B	108.6	C3—C4—H4	120.4
S2—C13—Te1	114.5 (3)	C3—C4—C5	119.3 (5)
S2—C13—H13A	108.6	C3—C2—H2A	119.4
S2—C13—H13B	108.6	C6—C1—H1	120.9
C12—Te1—C13	93.0 (2)	C6—C5—C8	112.0 (4)
C12—C7—S1	125.5 (3)	C6—C5—C4	118.8 (4)
C12—C11—H11	119.5	C7—S1—C6	91.0 (2)
C12—C11—C10	121.0 (5)	C7—C12—Te1	119.8 (3)
C8—C9—H9	120.1	C7—C8—C5	111.9 (4)
C8—C7—S1	112.5 (3)	C11—C12—Te1	122.4 (4)
C8—C7—C12	122.0 (4)	C11—C12—C7	117.8 (4)
C4—C3—C2	121.0 (5)	C11—C10—H10	119.5
C4—C3—H3	119.5	C1—C2—H2A	119.4
C4—C5—C8	129.2 (4)	C1—C2—C3	121.1 (5)
C9—C8—C7	118.5 (4)	C1—C6—S1	126.0 (4)
C9—C8—C5	129.6 (4)	C10—C9—C8	119.7 (5)
C9—C10—C11	121.0 (5)	C10—C9—H9	120.1
C9—C10—H10	119.5	C10—C11—H11	119.5
C2—C3—H3	119.5	C5—C4—H4	120.4
C2—C1—C6	118.3 (5)	C5—C6—S1	112.5 (4)
C2—C1—H1	120.9	C5—C6—C1	121.5 (5)
C13—S2—H2	109.5		
Te1—C12—C7—S1	4.0 (5)	C6—S1—C7—C12	178.1 (4)
Te1—C12—C7—C8	-176.4 (3)	C6—S1—C7—C8	-1.5 (3)
Te1—C12—C11—C10	177.3 (4)	C7—S1—C6—C1	-177.9 (4)
S1—C6—C1—C2	-178.6 (4)	C7—S1—C6—C5	1.9 (3)
S1—C6—C5—C8	-1.9 (5)	C7—C12—C11—C10	-1.6 (7)
S1—C6—C5—C4	178.4 (4)	C7—C8—C9—C10	-1.9 (7)
C12—Te1—C13—S2	78.3 (3)	C7—C8—C5—C4	-179.5 (5)
C12—C11—C10—C9	-1.0 (8)	C7—C8—C5—C6	0.7 (5)
C8—C9—C10—C11	2.9 (8)	C11—C12—C7—S1	-177.0 (3)
C9—C8—C7—S1	178.9 (3)	C11—C12—C7—C8	2.5 (6)
C9—C8—C7—C12	-0.7 (6)	C1—C2—C3—C4	1.1 (8)
C9—C8—C5—C4	2.6 (8)	C1—C6—C5—C8	178.0 (4)
C9—C8—C5—C6	-177.2 (4)	C1—C6—C5—C4	-1.8 (7)
C13—Te1—C12—C7	91.8 (4)	C5—C8—C9—C10	175.8 (5)
C13—Te1—C12—C11	-87.1 (4)	C5—C8—C7—S1	0.7 (5)
C3—C4—C5—C8	-178.2 (5)	C5—C8—C7—C12	-178.9 (4)
C3—C4—C5—C6	1.6 (7)	C5—C4—C3—C2	-1.3 (8)
C3—C2—C1—C6	-1.2 (8)	C5—C6—C1—C2	1.6 (7)

Hydrogen-bond geometry (Å, °)

Cg is the centroid of the thiophene ring.

<i>D</i> —H \cdots <i>A</i>	<i>D</i> —H	H \cdots <i>A</i>	<i>D</i> \cdots <i>A</i>	<i>D</i> —H \cdots <i>A</i>
C13—H13 <i>A</i> \cdots Cg ⁱ	0.97	2.90	3.846 (6)	166

Symmetry code: (i) $x-1/2, -y+3/2, -z$.