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KNb<sub>1.75</sub>V<sub>0.25</sub>PS<sub>10</sub>

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Key indicators: single-crystal X-ray study;  $T = 290$  K; mean  $\sigma(S-S) = 0.001$  Å; disorder in main residue;  $R$  factor = 0.019;  $wR$  factor = 0.042; data-to-parameter ratio = 22.8.

The title compound, potassium diniobium vanadium phosphorus decasulfide, KNb<sub>1.75</sub>V<sub>0.25</sub>PS<sub>10</sub>, was obtained by reaction of the elements with a eutectic mixture of KCl/LiCl. It is isostructural with the quaternary KNb<sub>2</sub>PS<sub>10</sub>, but the Nb sites are occupied by statistically disordered Nb (87.5%) and V (12.5%) atoms. The structure is composed of anionic  $\infty^1[M_2PS_{10}]^-$  chains ( $M = Nb/V$ ) separated from each other by  $K^+$  ions. The chain is composed of  $[MS_8]$  distorted bicapped trigonal prisms and  $[PS_4]$  tetrahedra. There are no interchain bonding interactions. The crystal used for the X-ray analysis was a racemic twin.

## Related literature

For related ternary compounds, see: Brec *et al.* (1983*a,b*). For related quaternary compounds, see: Goh *et al.* (2002); Do & Yun (1996); Kim & Yun (2002); Kwak *et al.* (2007); Bang *et al.* (2008); Do & Yun (2009). For related pentanary compounds, see: Kwak & Yun (2008); Dong *et al.* (2005*a,b*); Park & Yun (2010). For typical Nb<sup>4+</sup>—Nb<sup>4+</sup> bond lengths, see: Angenault *et al.* (2000)

## Experimental

## Crystal data

KNb<sub>1.75</sub>V<sub>0.25</sub>PS<sub>10</sub> $M_r = 2264.39$ Orthorhombic,  $Pca2_1$  $a = 12.9696$  (3) Å $b = 7.5229$  (2) Å $c = 13.3248$  (4) Å $V = 1300.09$  (6) Å<sup>3</sup> $Z = 1$ Mo  $K\alpha$  radiation $\mu = 3.73$  mm<sup>-1</sup> $T = 290$  K $0.36 \times 0.06 \times 0.06$  mm

## Data collection

Rigaku R-AXIS RAPID  
diffractometer  
Absorption correction: multi-scan  
(*ABSCOR*; Higashi, 1995)  
 $T_{\min} = 0.649$ ,  $T_{\max} = 1.000$ 11855 measured reflections  
2970 independent reflections  
2859 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.028$ 

## Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.019$  $wR(F^2) = 0.042$  $S = 1.07$ 

2970 reflections

130 parameters

1 restraint

 $\Delta\rho_{\max} = 0.49$  e Å<sup>-3</sup> $\Delta\rho_{\min} = -0.28$  e Å<sup>-3</sup>

Absolute structure: Flack (1983),

1417 Friedel pairs

Flack parameter: 0.47 (4)

Data collection: *RAPID-AUTO* (Rigaku, 2006); cell refinement: *RAPID-AUTO*; data reduction: *RAPID-AUTO*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: locally modified version of *ORTEP* (Johnson, 1965); software used to prepare material for publication: *STRUCTURE TIDY* (Gelato & Parthé, 1987) and *WinGX* (Farrugia, 1999).

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: FJ2392).

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## supporting information

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**KNb<sub>1.75</sub>V<sub>0.25</sub>PS<sub>10</sub>****Jaemin Yu and Hoseop Yun****S1. Comment**

Ternary group 5 metal thiophosphates have been reported to have mostly low-dimensional structures. Especially the Nb<sub>2</sub>PS<sub>10</sub> phase has a two-dimensional layered structure (Brec *et al.*, 1983a) and V<sub>2</sub>PS<sub>10</sub> adopts a one-dimensional chain structure (Brec *et al.*, 1983b). Due to empty spaces and the orbitals which can accommodate electrons, they have been of potential importance as cathode materials for secondary batteries and a number of quaternary alkali metal Nb thiophosphates, ANb<sub>2</sub>PS<sub>10</sub> (A=monovalent metals) have been investigated. Among them are NaNb<sub>2</sub>PS<sub>10</sub> (Goh *et al.*, 2002), KNb<sub>2</sub>PS<sub>10</sub> (Do & Yun, 1996), RbNb<sub>2</sub>PS<sub>10</sub> (Kim & Yun, 2002), CsNb<sub>2</sub>PS<sub>10</sub> (Kwak *et al.*, 2007), TINb<sub>2</sub>PS<sub>10</sub> (Bang *et al.*, 2008), Ag<sub>0.88</sub>Nb<sub>2</sub>PS<sub>10</sub> (Do & Yun, 2009), K<sub>0.34</sub>Cu<sub>0.5</sub>Nb<sub>2</sub>PS<sub>10</sub> (Kwak & Yun, 2008), K<sub>0.5</sub>Ag<sub>0.5</sub>Nb<sub>2</sub>PS<sub>10</sub> (Dong *et al.*, 2005a), Rb<sub>0.38</sub>Ag<sub>0.5</sub>Nb<sub>2</sub>PS<sub>10</sub> (Dong *et al.*, 2005b), and Cs<sub>0.5</sub>Ag<sub>0.5</sub>Nb<sub>2</sub>PS<sub>10</sub> (Park & Yun, 2010). It is interesting that no V analogue of these phases has been discovered yet. As a result of efforts to find new phases in this family, we report the synthesis and characterization of a new mixed-metal quinary thiophosphate, KNb<sub>2-x</sub>V<sub>x</sub>PS<sub>10</sub> (x=0.25).

The structure of KNb<sub>2-x</sub>V<sub>x</sub>PS<sub>10</sub> is isostructural with the quaternary KNb<sub>2</sub>PS<sub>10</sub> and detailed description of the structure is given previously (Do & Yun, 1996). The title compound is made up of the bicapped trigonal bipyramidal [*M*<sub>2</sub>S<sub>12</sub>] unit (*M*=Nb/V) and the tetrahedral [PS<sub>4</sub>] group. The *M* sites are occupied by the statistically disordered Nb (~87.5%) and V (~12.5%) atoms. The bicapped bipyramidal [*M*<sub>2</sub>S<sub>12</sub>] units and its neighboring tetrahedral [PS<sub>4</sub>] groups are given in Figure 1. These [*M*<sub>2</sub>S<sub>12</sub>] units are linked together to form the one-dimensional chains by sharing the S<sub>2</sub><sup>2-</sup> prism edge. The one-dimensional chain composed of *M*, P, and S extends along [100] and can be described as ∞<sup>1</sup>[*M*<sub>2</sub>PS<sub>10</sub><sup>-1</sup>].

The *M* atoms associate in pairs with *M*—*M* interactions alternating in the sequence of one short (2.8851 (3) Å) and one long (3.7590 (3) Å) distances. The short distance is typical of Nb<sup>4+</sup>—Nb<sup>4+</sup> bonding interactions (Angenault *et al.*, 2000). There are no interchain bonding interactions except the van der Waals forces and the K<sup>+</sup> ions in this van der Waals gap stabilize the structure through the electrostatic interactions (Figure 2). Finally, the classical charge balance of this phase can be represented by [K<sup>+</sup>][*M*<sup>4+</sup>]<sub>2</sub>[PS<sub>4</sub><sup>3-</sup>][S<sub>2</sub><sup>2-</sup>]<sub>3</sub> and this is consistent with the highly resistive and diamagnetic nature of the compound.

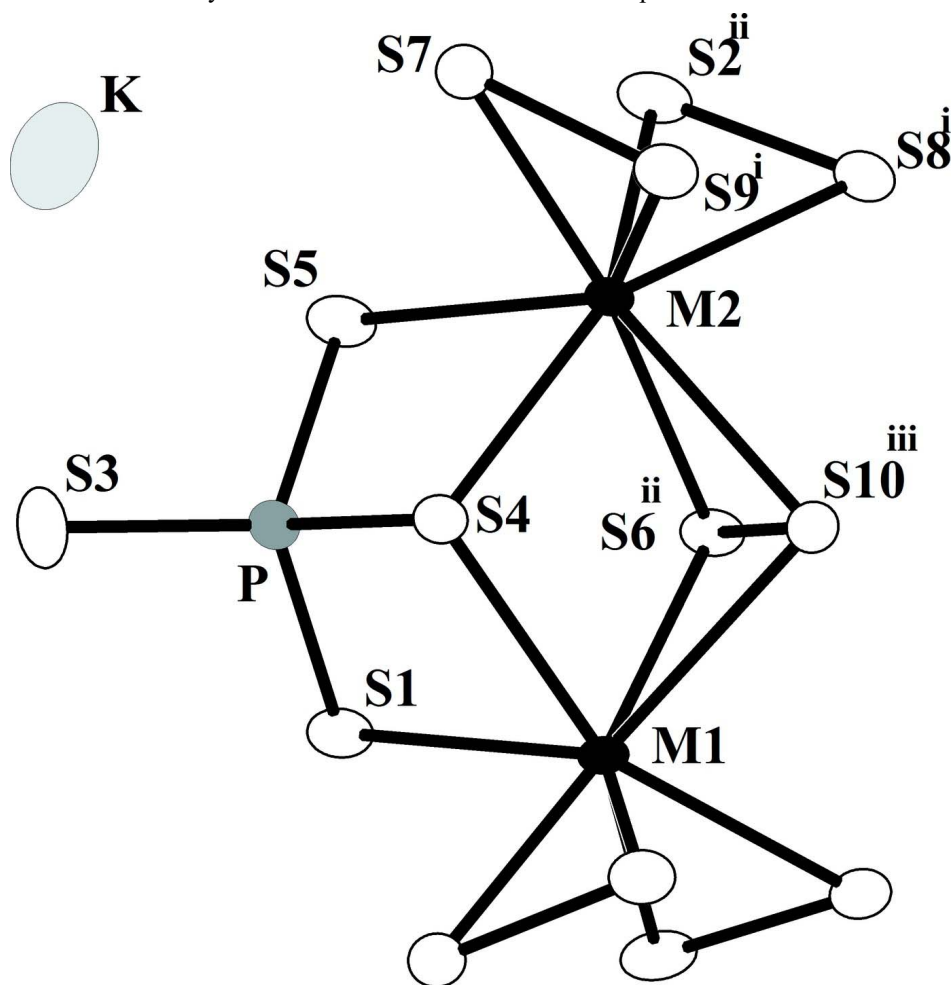
**S2. Experimental**

The compound KNb<sub>2-x</sub>V<sub>x</sub>PS<sub>10</sub> was prepared by the reaction of the elemental Nb, V, P, and S with the use of the reactive alkali metal halides. A combination of the pure elements, Nb powder (CERAC 99.9%), V powder (CERAC 99.5%), P powder (CERAC 99.95%), and S powder (Aldrich 99.999%) were mixed in a fused silica tube in a molar ratio of Nb: V: P: S = 1:1:1:5 with the eutectic mixture of KCl/LiCl. The mass ratio of the reactants and the halides flux was 2:1. The tube was evacuated to 0.133 Pa, sealed and heated gradually (50 K/h) to 650 K, where it was kept for 72 h. The tube was cooled to 423 K at 3 K/h and then was quenched to room temperature. The excess halides were removed with distilled water and black needle shaped crystals were obtained. The crystals are stable in air and water. A microprobe analysis of the crystals was made with an EDAX equipped scanning electron microscope (Jeol JSM-6700 F). Analysis of these

crystals showed only the presence of K, Nb, V, P, and S. A quantitative analysis performed with standards gave the ratio of Nb: V = 87: 13, which corresponds to  $\text{KNb}_{1.74}\text{V}_{0.26}\text{PS}_{10}$ .

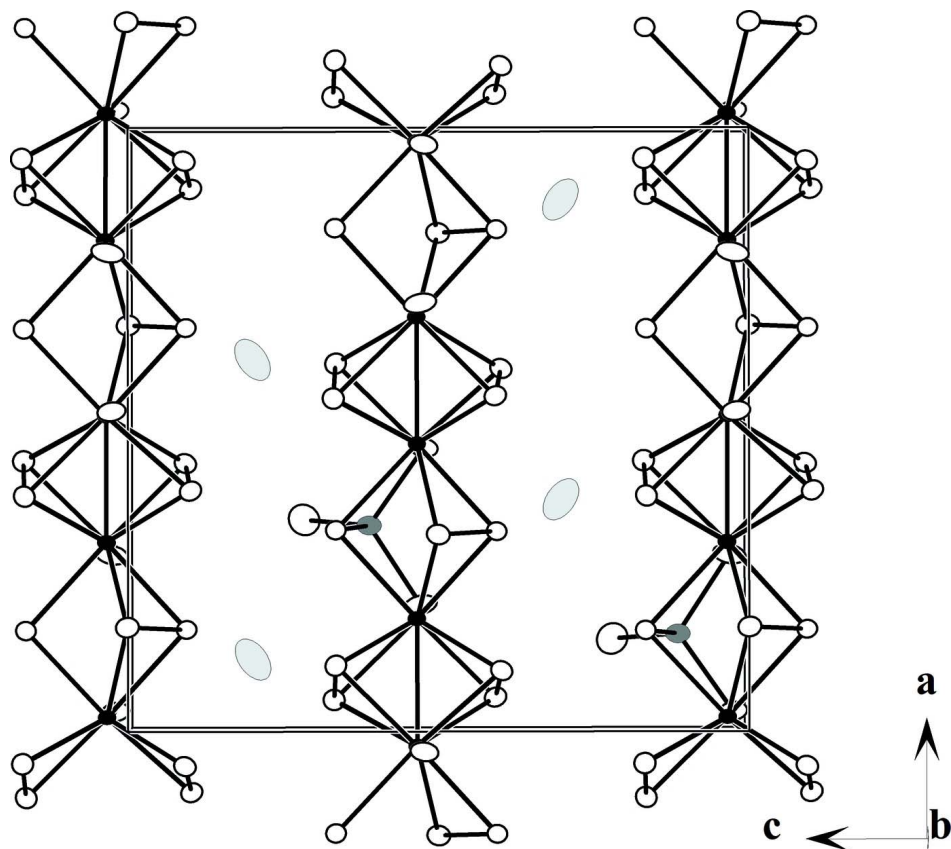
### S3. Refinement

The refinement of the model with occupational disorder on the *M* site caused significant decrease of the *R*-factor ( $wR2 = 0.042$ ) in comparison if the full occupation by either metal had been considered ( $wR2 > 0.05$ ). Also the displacement parameters in the disordered model became plausible. The disordered atoms were supposed to have the same displacement parameters. The nonstoichiometry of the K site was checked by refining the occupancy of K while those of the other atoms were fixed. With the nonstoichiometric model, the parameter remained the same. The large anisotropic displacement parameters for alkali metals are also found in the related compounds such as  $\text{KNb}_2\text{PS}_{10}$  (Do & Yun, 1996). The highest residual electron density is  $0.86 \text{ \AA}^{-3}$  from the M2 site and the deepest hole is  $0.85 \text{ \AA}^{-3}$  from the M1 site.



**Figure 1**

A view of the bicapped trigonal bipyrametric  $[M_2S_{12}]$  unit ( $M=V/Nb$ ) and its neighboring tetrahedral  $[PS_4]$  groups. Open circles are S atoms, filled circle are Nb atoms, dark and pale gray circles are P and K atoms, respectively. Displacement ellipsoids are drawn at the 50% probability level. [Symmetry code: (i)  $1 - x, -y, -1/2 + z$ ; (ii)  $0.5 - x, -y, -1/2 + z$ ; (iii)  $-1/2 + x, -y, z$ ]



**Figure 2**

View of the  $\text{KNb}_{2-x}\text{V}_x\text{PS}_{10}$  down the  $b$  axis showing the one-dimensional nature of the compound. Atoms are as marked in Fig. 1.

### potassium diniobium vanadium phosphorus decasulfide

#### Crystal data

$\text{KNb}_{1.75}\text{V}_{0.25}\text{PS}_{10}$

$M_r = 2264.39$

Orthorhombic,  $Pca2_1$

Hall symbol:  $P\ 2c\ -2ac$

$a = 12.9696\ (3)\ \text{\AA}$

$b = 7.5229\ (2)\ \text{\AA}$

$c = 13.3248\ (4)\ \text{\AA}$

$V = 1300.09\ (6)\ \text{\AA}^3$

$Z = 1$

$F(000) = 1086$

$D_x = 2.892\ \text{Mg m}^{-3}$

Mo  $K\alpha$  radiation,  $\lambda = 0.71073\ \text{\AA}$

Cell parameters from 11171 reflections

$\theta = 3.1\text{--}27.5^\circ$

$\mu = 3.73\ \text{mm}^{-1}$

$T = 290\ \text{K}$

Needle, black

$0.36 \times 0.06 \times 0.06\ \text{mm}$

#### Data collection

Rigaku R-AXIS RAPID  
diffractometer

$\omega$  scans

Absorption correction: multi-scan  
(*ABSCOR*; Higashi, 1995)

$T_{\min} = 0.649$ ,  $T_{\max} = 1.000$

11855 measured reflections

2970 independent reflections

2859 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.028$

$\theta_{\max} = 27.5^\circ$ ,  $\theta_{\min} = 3.1^\circ$

$h = -16 \rightarrow 16$

$k = -9 \rightarrow 9$

$l = -17 \rightarrow 17$

Refinement

Refinement on  $F^2$   
 Least-squares matrix: full  
 $R[F^2 > 2\sigma(F^2)] = 0.019$   
 $wR(F^2) = 0.042$   
 $S = 1.07$   
 2970 reflections  
 130 parameters  
 1 restraint

$w = 1/[\sigma^2(F_o^2) + (0.0125P)^2 + 1.1461P]$   
 where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\max} = 0.002$   
 $\Delta\rho_{\max} = 0.49 \text{ e } \text{\AA}^{-3}$   
 $\Delta\rho_{\min} = -0.28 \text{ e } \text{\AA}^{-3}$   
 Absolute structure: Flack (1983), **1417 Friedel pairs**  
 Absolute structure parameter: 0.47 (4)

Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$	Occ. (<1)
K1	0.38307 (10)	0.50472 (16)	0.30096 (10)	0.0642 (3)	
Nb1	0.023833 (17)	0.05293 (3)	0.03457 (3)	0.01512 (9)	0.861 (4)
V1	0.023833 (17)	0.05293 (3)	0.03457 (3)	0.01512 (9)	0.139 (4)
Nb2	0.313453 (17)	0.07166 (3)	0.03513 (3)	0.01521 (10)	0.889 (4)
V2	0.313453 (17)	0.07166 (3)	0.03513 (3)	0.01521 (10)	0.111 (4)
P1	0.15973 (6)	0.40030 (12)	0.11232 (7)	0.02169 (18)	
S1	0.03047 (5)	0.39473 (11)	0.02326 (8)	0.0258 (2)	
S2	0.05595 (7)	0.15141 (14)	0.40819 (7)	0.0255 (2)	
S3	0.15066 (9)	0.58306 (14)	0.21803 (9)	0.0408 (3)	
S4	0.16690 (6)	0.13978 (11)	0.16577 (6)	0.01913 (18)	
S5	0.29187 (5)	0.41184 (10)	0.02883 (11)	0.03005 (19)	
S6	0.33126 (6)	0.05595 (12)	0.40601 (7)	0.02107 (19)	
S7	0.44830 (7)	0.13335 (13)	0.16897 (6)	0.02327 (19)	
S8	0.60124 (7)	0.10543 (14)	0.39868 (7)	0.0257 (2)	
S9	0.60983 (7)	0.11862 (12)	0.66562 (6)	0.02271 (19)	
S10	0.67360 (6)	0.15938 (11)	0.00015 (6)	0.02081 (17)	

Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
K1	0.0831 (8)	0.0470 (6)	0.0625 (8)	0.0035 (6)	-0.0317 (7)	0.0031 (5)
Nb1	0.01049 (12)	0.01887 (15)	0.01601 (14)	-0.00123 (9)	-0.00013 (16)	0.00103 (15)
V1	0.01049 (12)	0.01887 (15)	0.01601 (14)	-0.00123 (9)	-0.00013 (16)	0.00103 (15)
Nb2	0.01068 (13)	0.01820 (15)	0.01675 (14)	0.00116 (9)	0.00020 (16)	-0.00022 (15)
V2	0.01068 (13)	0.01820 (15)	0.01675 (14)	0.00116 (9)	0.00020 (16)	-0.00022 (15)
P1	0.0173 (4)	0.0182 (4)	0.0296 (5)	0.0017 (3)	-0.0002 (3)	-0.0028 (4)
S1	0.0170 (3)	0.0225 (4)	0.0380 (5)	0.0025 (3)	-0.0034 (4)	0.0057 (4)
S2	0.0177 (4)	0.0362 (5)	0.0225 (4)	-0.0038 (4)	-0.0021 (4)	0.0077 (4)
S3	0.0463 (6)	0.0299 (5)	0.0462 (6)	0.0077 (5)	-0.0039 (5)	-0.0184 (5)
S4	0.0164 (4)	0.0221 (4)	0.0188 (4)	0.0005 (3)	0.0004 (3)	-0.0008 (4)

S5	0.0178 (3)	0.0214 (4)	0.0510 (5)	-0.0009 (3)	0.0070 (5)	0.0067 (5)
S6	0.0153 (4)	0.0292 (5)	0.0187 (4)	-0.0010 (3)	0.0004 (3)	0.0031 (3)
S7	0.0203 (4)	0.0279 (5)	0.0216 (4)	0.0006 (4)	-0.0003 (4)	-0.0050 (4)
S8	0.0178 (4)	0.0374 (5)	0.0219 (4)	0.0055 (4)	0.0024 (3)	0.0087 (4)
S9	0.0200 (4)	0.0274 (5)	0.0208 (4)	0.0013 (4)	-0.0012 (3)	-0.0034 (4)
S10	0.0187 (4)	0.0185 (4)	0.0252 (4)	0.0001 (3)	-0.0012 (3)	0.0014 (3)

*Geometric parameters (Å, °)*

Nb1—S8 <sup>i</sup>	2.4631 (10)	Nb2—S6 <sup>i</sup>	2.5490 (9)
Nb1—S7 <sup>ii</sup>	2.4760 (9)	Nb2—S10 <sup>ii</sup>	2.5551 (8)
Nb1—S2 <sup>iii</sup>	2.5039 (10)	Nb2—S5	2.5758 (8)
Nb1—S9 <sup>i</sup>	2.5098 (9)	Nb2—S4	2.6279 (8)
Nb1—S6 <sup>i</sup>	2.5431 (9)	Nb2—V1 <sup>v</sup>	2.8851 (3)
Nb1—S10 <sup>ii</sup>	2.5562 (8)	Nb2—Nb1 <sup>v</sup>	2.8851 (3)
Nb1—S1	2.5772 (9)	P1—S3	1.9718 (13)
Nb1—S4	2.6319 (8)	P1—S5	2.0451 (13)
Nb1—Nb2 <sup>ii</sup>	2.8851 (3)	P1—S1	2.0544 (13)
Nb1—V2 <sup>ii</sup>	2.8851 (3)	P1—S4	2.0874 (13)
Nb2—S9 <sup>iv</sup>	2.4622 (9)	S2—S8 <sup>ii</sup>	2.0235 (15)
Nb2—S2 <sup>i</sup>	2.4678 (9)	S6—S10 <sup>vi</sup>	2.0498 (12)
Nb2—S8 <sup>iv</sup>	2.5110 (10)	S7—S9 <sup>iv</sup>	2.0405 (14)
Nb2—S7	2.5406 (9)		
S8 <sup>i</sup> —Nb1—S7 <sup>ii</sup>	111.22 (3)	S9 <sup>iv</sup> —Nb2—S7	48.11 (3)
S8 <sup>i</sup> —Nb1—S2 <sup>iii</sup>	48.07 (4)	S2 <sup>i</sup> —Nb2—S7	87.95 (3)
S7 <sup>ii</sup> —Nb1—S2 <sup>iii</sup>	88.59 (3)	S8 <sup>iv</sup> —Nb2—S7	107.57 (3)
S8 <sup>i</sup> —Nb1—S9 <sup>i</sup>	91.43 (3)	S9 <sup>iv</sup> —Nb2—S6 <sup>i</sup>	138.35 (3)
S7 <sup>ii</sup> —Nb1—S9 <sup>i</sup>	48.31 (3)	S2 <sup>i</sup> —Nb2—S6 <sup>i</sup>	93.10 (3)
S2 <sup>iii</sup> —Nb1—S9 <sup>i</sup>	107.66 (3)	S8 <sup>iv</sup> —Nb2—S6 <sup>i</sup>	79.11 (3)
S8 <sup>i</sup> —Nb1—S6 <sup>i</sup>	89.43 (3)	S7—Nb2—S6 <sup>i</sup>	171.57 (3)
S7 <sup>ii</sup> —Nb1—S6 <sup>i</sup>	141.69 (3)	S9 <sup>iv</sup> —Nb2—S10 <sup>ii</sup>	91.16 (3)
S2 <sup>iii</sup> —Nb1—S6 <sup>i</sup>	81.83 (3)	S2 <sup>i</sup> —Nb2—S10 <sup>ii</sup>	121.84 (3)
S9 <sup>i</sup> —Nb1—S6 <sup>i</sup>	167.80 (3)	S8 <sup>iv</sup> —Nb2—S10 <sup>ii</sup>	79.62 (3)
S8 <sup>i</sup> —Nb1—S10 <sup>ii</sup>	117.95 (3)	S7—Nb2—S10 <sup>ii</sup>	137.82 (3)
S7 <sup>ii</sup> —Nb1—S10 <sup>ii</sup>	94.40 (3)	S6 <sup>i</sup> —Nb2—S10 <sup>ii</sup>	47.36 (3)
S2 <sup>iii</sup> —Nb1—S10 <sup>ii</sup>	79.03 (3)	S9 <sup>iv</sup> —Nb2—S5	130.13 (4)
S9 <sup>i</sup> —Nb1—S10 <sup>ii</sup>	140.60 (3)	S2 <sup>i</sup> —Nb2—S5	79.09 (3)
S6 <sup>i</sup> —Nb1—S10 <sup>ii</sup>	47.40 (3)	S8 <sup>iv</sup> —Nb2—S5	123.47 (4)
S8 <sup>i</sup> —Nb1—S1	79.57 (3)	S7—Nb2—S5	85.19 (3)
S7 <sup>ii</sup> —Nb1—S1	128.33 (3)	S6 <sup>i</sup> —Nb2—S5	86.79 (3)
S2 <sup>iii</sup> —Nb1—S1	125.94 (3)	S10 <sup>ii</sup> —Nb2—S5	126.35 (3)
S9 <sup>i</sup> —Nb1—S1	82.37 (3)	S9 <sup>iv</sup> —Nb2—S4	86.44 (3)
S6 <sup>i</sup> —Nb1—S1	85.81 (3)	S2 <sup>i</sup> —Nb2—S4	154.57 (3)
S10 <sup>ii</sup> —Nb1—S1	125.98 (3)	S8 <sup>iv</sup> —Nb2—S4	154.40 (3)
S8 <sup>i</sup> —Nb1—S4	155.91 (3)	S7—Nb2—S4	89.85 (3)
S7 <sup>ii</sup> —Nb1—S4	86.50 (3)	S6 <sup>i</sup> —Nb2—S4	85.62 (3)
S2 <sup>iii</sup> —Nb1—S4	152.94 (3)	S10 <sup>ii</sup> —Nb2—S4	74.93 (3)

S9 <sup>i</sup> —Nb1—S4	88.62 (3)	S5—Nb2—S4	75.48 (3)
S6 <sup>i</sup> —Nb1—S4	85.65 (3)	S9 <sup>iv</sup> —Nb2—V1 <sup>v</sup>	55.30 (2)
S10 <sup>ii</sup> —Nb1—S4	74.84 (3)	S2 <sup>i</sup> —Nb2—V1 <sup>v</sup>	55.11 (2)
S1—Nb1—S4	76.56 (3)	S8 <sup>iv</sup> —Nb2—V1 <sup>v</sup>	53.78 (2)
S8 <sup>i</sup> —Nb1—Nb2 <sup>ii</sup>	55.33 (2)	S7—Nb2—V1 <sup>v</sup>	53.85 (2)
S7 <sup>ii</sup> —Nb1—Nb2 <sup>ii</sup>	55.95 (2)	S6 <sup>i</sup> —Nb2—V1 <sup>v</sup>	132.81 (2)
S2 <sup>iii</sup> —Nb1—Nb2 <sup>ii</sup>	53.95 (2)	S10 <sup>ii</sup> —Nb2—V1 <sup>v</sup>	116.74 (2)
S9 <sup>i</sup> —Nb1—Nb2 <sup>ii</sup>	53.76 (2)	S5—Nb2—V1 <sup>v</sup>	115.182 (19)
S6 <sup>i</sup> —Nb1—Nb2 <sup>ii</sup>	134.72 (2)	S4—Nb2—V1 <sup>v</sup>	138.52 (2)
S10 <sup>ii</sup> —Nb1—Nb2 <sup>ii</sup>	121.07 (2)	S9 <sup>iv</sup> —Nb2—Nb1 <sup>v</sup>	55.30 (2)
S1—Nb1—Nb2 <sup>ii</sup>	110.848 (18)	S2 <sup>i</sup> —Nb2—Nb1 <sup>v</sup>	55.11 (2)
S4—Nb1—Nb2 <sup>ii</sup>	138.22 (2)	S8 <sup>iv</sup> —Nb2—Nb1 <sup>v</sup>	53.78 (2)
S8 <sup>i</sup> —Nb1—V2 <sup>ii</sup>	55.33 (2)	S7—Nb2—Nb1 <sup>v</sup>	53.85 (2)
S7 <sup>ii</sup> —Nb1—V2 <sup>ii</sup>	55.95 (2)	S6 <sup>i</sup> —Nb2—Nb1 <sup>v</sup>	132.81 (2)
S2 <sup>iii</sup> —Nb1—V2 <sup>ii</sup>	53.95 (2)	S10 <sup>ii</sup> —Nb2—Nb1 <sup>v</sup>	116.74 (2)
S9 <sup>i</sup> —Nb1—V2 <sup>ii</sup>	53.76 (2)	S5—Nb2—Nb1 <sup>v</sup>	115.182 (19)
S6 <sup>i</sup> —Nb1—V2 <sup>ii</sup>	134.72 (2)	S4—Nb2—Nb1 <sup>v</sup>	138.52 (2)
S10 <sup>ii</sup> —Nb1—V2 <sup>ii</sup>	121.07 (2)	V1 <sup>v</sup> —Nb2—Nb1 <sup>v</sup>	0.000 (17)
S1—Nb1—V2 <sup>ii</sup>	110.848 (18)	S3—P1—S5	114.14 (6)
S4—Nb1—V2 <sup>ii</sup>	138.22 (2)	S3—P1—S1	112.22 (6)
Nb2 <sup>ii</sup> —Nb1—V2 <sup>ii</sup>	0.000 (18)	S5—P1—S1	111.74 (7)
S9 <sup>iv</sup> —Nb2—S2 <sup>i</sup>	110.37 (3)	S3—P1—S4	114.43 (6)
S9 <sup>iv</sup> —Nb2—S8 <sup>iv</sup>	91.42 (3)	S5—P1—S4	100.84 (5)
S2 <sup>i</sup> —Nb2—S8 <sup>iv</sup>	47.95 (4)	S1—P1—S4	102.37 (5)

Symmetry codes: (i)  $-x+1/2, y, z-1/2$ ; (ii)  $x-1/2, -y, z$ ; (iii)  $-x, -y, z-1/2$ ; (iv)  $-x+1, -y, z-1/2$ ; (v)  $x+1/2, -y, z$ ; (vi)  $-x+1, -y, z+1/2$ .