

Acta Crystallographica Section E

## Structure Reports

Online

ISSN 1600-5368

## 2-Bromo-2-methyl-N-(4-nitrophenyl)-propanamide

 Rodolfo Moreno-Fuquen,<sup>a\*</sup> David E. Quintero,<sup>a</sup> Fabio Zuluaga,<sup>a</sup> Roberto L. A. Haiduke<sup>b</sup> and Alan R. Kennedy<sup>c</sup>

<sup>a</sup>Departamento de Química, Facultad de Ciencias, Universidad del Valle, Apartado 25360, Santiago de Cali, Colombia, <sup>b</sup>Instituto de Química, IQSC, Universidade de São Paulo, São Carlos, Brazil, and <sup>c</sup>WestCHEM, Department of Pure and Applied Chemistry, University of Strathclyde, 295 Cathedral Street, Glasgow G1 1XL, Scotland

Correspondence e-mail: rodimo26@yahoo.es

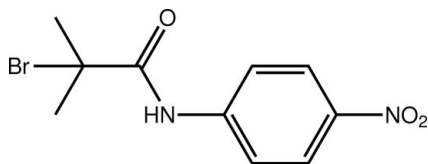
Received 8 February 2011; accepted 12 February 2011

 Key indicators: single-crystal X-ray study;  $T = 123$  K; mean  $\sigma(\text{C}-\text{C}) = 0.003$  Å;  $R$  factor = 0.032;  $wR$  factor = 0.068; data-to-parameter ratio = 17.4.

The title compound,  $\text{C}_{10}\text{H}_{11}\text{BrN}_2\text{O}_3$ , exhibits a small twist between the amide residue and benzene ring [the  $\text{C}-\text{N}-\text{C}-\text{C}$  torsion angle =  $12.7$  ( $4^\circ$ )]. The crystal structure is stabilized by weak  $\text{N}-\text{H}\cdots\text{O}$ ,  $\text{C}-\text{H}\cdots\text{Br}$  and  $\text{C}-\text{H}\cdots\text{O}$  interactions. These lead to supramolecular layers in the  $bc$  plane.

### Related literature

For initiators in ATRP (polymerization by atom transfer radical) processes, see: Matyjaszewski & Xia (2001); Pietrasik & Tsarevsky (2010). For graph-set notation of hydrogen-bond patterns, see: Etter (1990).



### Experimental

#### Crystal data

$\text{C}_{10}\text{H}_{11}\text{BrN}_2\text{O}_3$   
 $M_r = 287.11$   
 Monoclinic,  $C2/c$   
 $a = 24.1245$  (12) Å  
 $b = 5.8507$  (3) Å  
 $c = 15.4723$  (8) Å  
 $\beta = 91.837$  ( $5^\circ$ )

$V = 2182.72$  (19) Å<sup>3</sup>  
 $Z = 8$   
 Mo  $K\alpha$  radiation  
 $\mu = 3.76$  mm<sup>-1</sup>  
 $T = 123$  K  
 $0.60 \times 0.05 \times 0.05$  mm

#### Data collection

Oxford Diffraction Xcalibur E diffractometer  
 Absorption correction: analytical (*CrysAlis PRO*; Oxford Diffraction, 2009)  
 $T_{\min} = 0.444$ ,  $T_{\max} = 1.000$   
 5100 measured reflections  
 2633 independent reflections  
 2197 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.032$

#### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.032$   
 $wR(F^2) = 0.068$   
 $S = 1.06$   
 2633 reflections  
 151 parameters  
 H atoms treated by a mixture of independent and constrained refinement  
 $\Delta\rho_{\max} = 0.55$  e Å<sup>-3</sup>  
 $\Delta\rho_{\min} = -0.60$  e Å<sup>-3</sup>

**Table 1**

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{C6}-\text{H6}\cdots\text{O1}$	0.95	2.27	2.862 (3)	120
$\text{C7}-\text{H7}\cdots\text{O1}^{\text{i}}$	0.95	2.45	3.139 (3)	129
$\text{N1}-\text{H1n}\cdots\text{O3}^{\text{ii}}$	0.84 (3)	2.66 (3)	3.316 (3)	136 (2)
$\text{C10}-\text{H10}\cdots\text{Br1}^{\text{iii}}$	0.95	2.91	3.812 (2)	160

 Symmetry codes: (i)  $-x + \frac{1}{2}, -y - \frac{1}{2}, -z$ ; (ii)  $-x + \frac{1}{2}, y - \frac{1}{2}, -z + \frac{1}{2}$ ; (iii)  $x, y + 1, z$ .

Data collection: *CrysAlis CCD* (Oxford Diffraction, 2009); cell refinement: *CrysAlis CCD*; data reduction: *CrysAlis CCD*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 1997) and *Mercury* (Macrae *et al.*, 2006); software used to prepare material for publication: *SHELXL97*.

RMF is grateful to the Spanish Research Council (CSIC) for the use of a free-of-charge licence to the Cambridge Structural Database. RMF and FZ also thank the Universidad del Valle, Colombia, and the Instituto de Química de São Carlos, USP, Brasil for partial financial support. RLAH thanks CNPq for partial financial support.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: TK2720).

### References

- Etter, M. (1990). *Acc. Chem. Res.* **23**, 120–126.  
 Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.  
 Macrae, C. F., Edgington, P. R., McCabe, P., Pidcock, E., Shields, G. P., Taylor, R., Towler, M. & van de Streek, J. (2006). *J. Appl. Cryst.* **39**, 453–457.  
 Matyjaszewski, K. & Xia, J. (2001). *Chem. Rev.* **101**, 2921–2990.  
 Oxford Diffraction (2009). *CrysAlis CCD* and *CrysAlis RED*. Oxford Diffraction Ltd, Yarnton, England.  
 Pietrasik, J. & Tsarevsky, N. V. (2010). *Eur. Polym. J.* **46**, 2333–2340.  
 Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.

## supporting information

*Acta Cryst.* (2011). E67, o659 [doi:10.1107/S1600536811005320]

## 2-Bromo-2-methyl-*N*-(4-nitrophenyl)propanamide

Rodolfo Moreno-Fuquen, David E. Quintero, Fabio Zuluaga, Roberto L. A. Haiduke and Alan R. Kennedy

### S1. Comment

The title compound, (I), forms a part of a synthetic programme that the Polymer Research Group of Universidad del Valle is pursuing in order to obtain compounds that act as initiators of reactions in polymerization processes. Compound (I), Fig. 1, belongs to a series of polymeric ATRP (polymerization by atom transfer radical) initiators (Pietrasik & Tsarevsky, 2010); most initiators for ATRP processes are alkyl halides (Matyjaszewski & Xia, 2001). The C4—N1—C5—C6 torsion angle is 12.7 (4)°, indicating a small twist between the benzene ring and the amide. An intramolecular C—H···O interaction is observed (see Table 1).

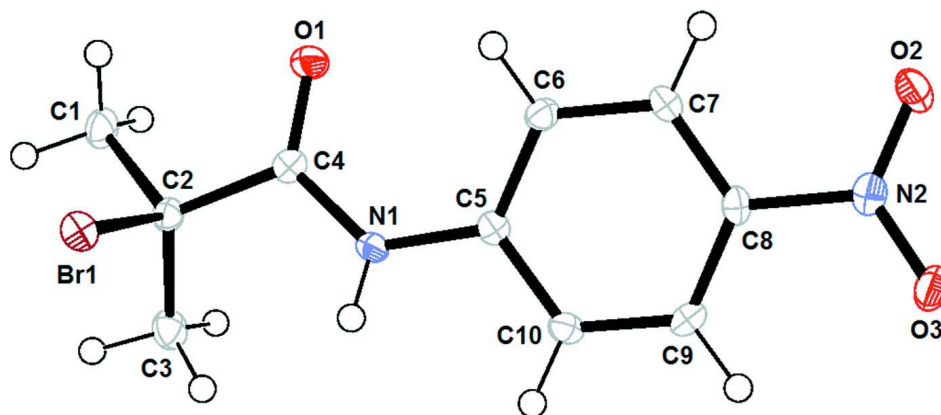
In the crystal structure, weak C—H···O and N—H···O intermolecular interactions are detected (Table 1). In a first substructure (Fig. 2), molecules are linked by C7—H···O1<sup>i</sup> contacts (*i*: -*x*+1/2, -*y*-1/2, -*z*) which leads to the formation of dimers with graphs-set notation, R<sup>2</sup><sub>2</sub>(14) (Etter, 1990). In turn, N—H···O interactions link the dimers running along the *c* axis. The N1 atom acts as hydrogen bond donor to O3<sup>ii</sup> (*ii*: -*x*+1/2, +*y*-1/2, -*z*+1/2). In a second substructure, C10—H···Br (Table 1) interactions provide further links between the dimers. Overall, the crystal structure comprises layers in the *bc* plane (Fig. 3).

### S2. Experimental

The initial reagents were purchased from Aldrich Chemical Co. and were used as received. In a 100 mL round bottom flask 4-nitroaniline (3.258 mmol, 0.450 g), triethylamine (0.653 mmol, 0.066 g) were mixed, then a solution of 2-bromo isobuturyl bromide (0.704 g) in anhydrous THF (5 mL) was added drop wise, under an argon stream. The reaction was carried out in a dry bag overnight under magnetic stirring. The solid was filtered off and dichloromethane (20 mL) added to the organic phase which was washed with brine (40 mL) followed by water (10 mL). The solution was concentrated at low pressure affording colourless crystals and recrystallized from a solution of hexane and ethyl acetate (80:20). *M.* pt. 356 (1) K.

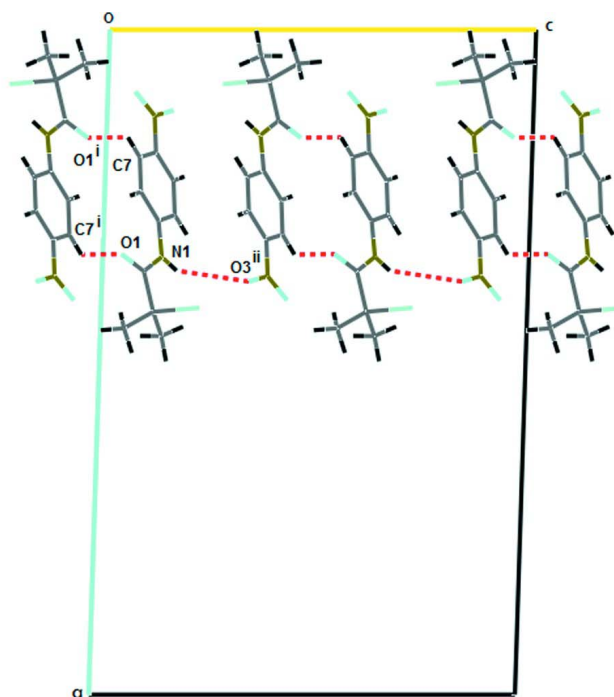
### S3. Refinement

The H-atoms were placed geometrically [C—H = 0.95 Å for aromatic & C—H = 0.98 Å for methyl], refined in the riding model approximation with  $U_{\text{iso}}(\text{H}) = 1.2-1.5U_{\text{eq}}(\text{C})$ . The N-bound H atom was refined freely.



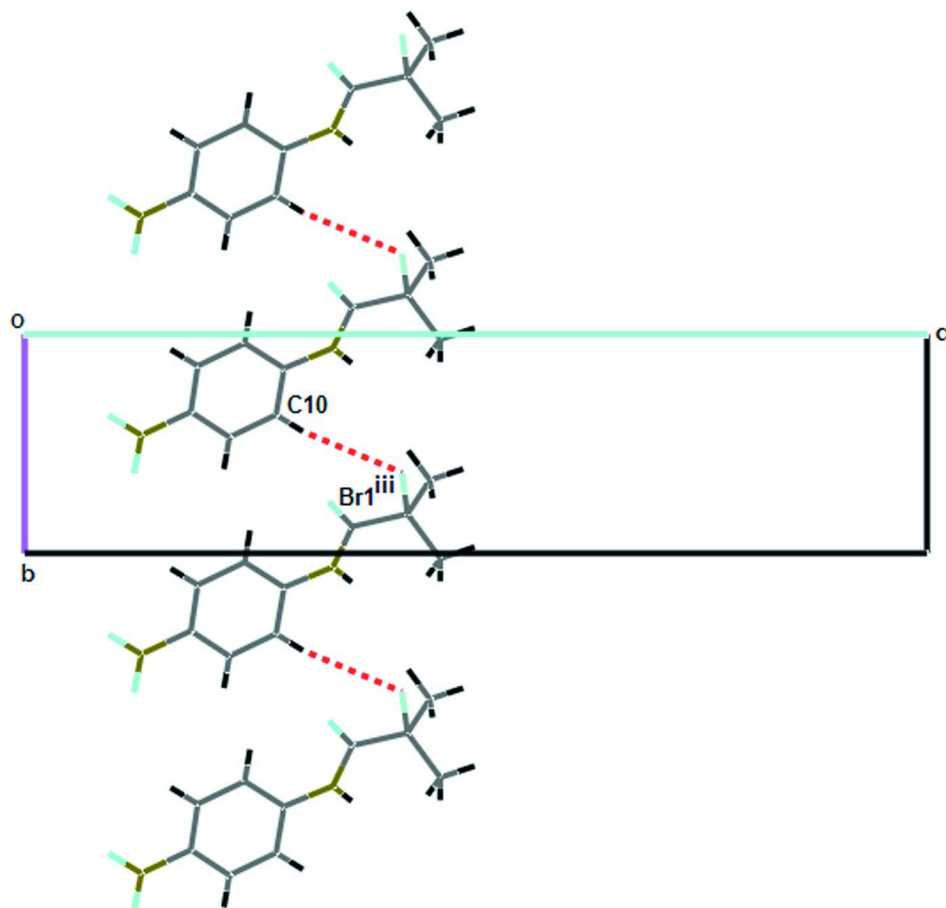
**Figure 1**

An ORTEP-3 (Farrugia, 1997) plot of (I) with the atomic labelling scheme. Displacement ellipsoids are drawn at the 50% probability level. H atoms are shown as spheres of arbitrary radius.



**Figure 2**

Part of the crystal structure of (I), showing the formation of  $R^2_2(10)$  and  $R^4_4(18)$  rings, running along  $[110]$ . Symmetry code: (i)  $-x+1, -y+1, -z$ ; (ii)  $-x+2, -y+2, -z$ .



**Figure 3**

Part of the crystal structure of (I), showing the formation of an infinite zig-zag chain of dimers along [001] direction. Symmetry code: (iii)  $x, -y+1/2+1, +z-1/2$ .

### 2-Bromo-2-methyl-N-(4-nitrophenyl)propanamide

#### Crystal data

$C_{10}H_{11}BrN_2O_3$

$M_r = 287.11$

Monoclinic,  $C2/c$

Hall symbol:  $-C 2yc$

$a = 24.1245$  (12) Å

$b = 5.8507$  (3) Å

$c = 15.4723$  (8) Å

$\beta = 91.837$  (5)°

$V = 2182.72$  (19) Å<sup>3</sup>

$Z = 8$

$F(000) = 1152$

$D_x = 1.747$  Mg m<sup>-3</sup>

Melting point: 385(1) K

Mo  $K\alpha$  radiation,  $\lambda = 0.71073$  Å

Cell parameters from 2722 reflections

$\theta = 3.2$ – $29.3$ °

$\mu = 3.76$  mm<sup>-1</sup>

$T = 123$  K

Fragment cut from needle, colourless

$0.60 \times 0.05 \times 0.05$  mm

#### Data collection

Oxford Diffraction Xcalibur E  
diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

$\omega$  scans

Absorption correction: analytical

(*CrysAlis PRO*; Oxford Diffraction, 2009)

$T_{\min} = 0.444$ ,  $T_{\max} = 1.000$

5100 measured reflections

2633 independent reflections

2197 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.032$   
 $\theta_{\text{max}} = 29.0^\circ$ ,  $\theta_{\text{min}} = 3.2^\circ$

$h = -32 \rightarrow 18$   
 $k = -6 \rightarrow 7$   
 $l = -19 \rightarrow 21$

*Refinement*

Refinement on  $F^2$   
 Least-squares matrix: full  
 $R[F^2 > 2\sigma(F^2)] = 0.032$   
 $wR(F^2) = 0.068$   
 $S = 1.06$   
 2633 reflections  
 151 parameters  
 0 restraints  
 Primary atom site location: structure-invariant  
 direct methods

Secondary atom site location: difference Fourier  
 map  
 Hydrogen site location: inferred from  
 neighbouring sites  
 H atoms treated by a mixture of independent  
 and constrained refinement  
 $w = 1/[\sigma^2(F_o^2) + (0.0213P)^2]$   
 where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\text{max}} < 0.001$   
 $\Delta\rho_{\text{max}} = 0.55 \text{ e } \text{\AA}^{-3}$   
 $\Delta\rho_{\text{min}} = -0.60 \text{ e } \text{\AA}^{-3}$

*Special details*

**Geometry.** All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted R-factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional R-factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > 2\sigma(F^2)$  is used only for calculating R-factors(gt) etc. and is not relevant to the choice of reflections for refinement. R-factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and R-factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Br1	0.418223 (9)	-0.36757 (5)	0.230618 (14)	0.01800 (9)
O1	0.33754 (6)	-0.2340 (4)	0.04409 (11)	0.0205 (4)
O2	0.09326 (6)	0.3711 (4)	0.06453 (11)	0.0219 (4)
O3	0.12126 (7)	0.6341 (3)	0.15486 (11)	0.0224 (4)
N1	0.34163 (8)	0.0629 (4)	0.13870 (13)	0.0177 (5)
N2	0.12933 (8)	0.4612 (4)	0.11152 (12)	0.0175 (5)
C1	0.44903 (10)	-0.3344 (5)	0.05532 (15)	0.0205 (6)
H1A	0.4520	-0.2507	0.0008	0.031*
H1B	0.4251	-0.4681	0.0461	0.031*
H1C	0.4860	-0.3844	0.0756	0.031*
C2	0.42421 (9)	-0.1793 (5)	0.12266 (14)	0.0156 (5)
C3	0.46064 (9)	0.0238 (5)	0.14342 (16)	0.0209 (6)
H3A	0.4986	-0.0283	0.1564	0.031*
H3B	0.4465	0.1044	0.1937	0.031*
H3C	0.4605	0.1275	0.0937	0.031*
C4	0.36323 (9)	-0.1220 (5)	0.09765 (14)	0.0143 (5)
C5	0.28757 (9)	0.1555 (5)	0.12844 (14)	0.0133 (5)
C6	0.24391 (9)	0.0415 (5)	0.08499 (14)	0.0156 (5)
H6	0.2497	-0.1040	0.0595	0.019*
C7	0.19187 (9)	0.1446 (5)	0.07971 (14)	0.0154 (5)
H7	0.1618	0.0700	0.0503	0.018*

C8	0.18422 (9)	0.3543 (5)	0.11722 (14)	0.0135 (5)
C9	0.22692 (9)	0.4702 (5)	0.16051 (14)	0.0158 (5)
H9	0.2208	0.6157	0.1857	0.019*
C10	0.27856 (9)	0.3680 (5)	0.16594 (14)	0.0150 (5)
H10	0.3083	0.4438	0.1957	0.018*
H1N	0.3624 (11)	0.132 (5)	0.1742 (17)	0.028 (8)*

*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Br1	0.01677 (13)	0.01963 (16)	0.01755 (13)	0.00024 (10)	-0.00059 (9)	0.00090 (10)
O1	0.0181 (8)	0.0222 (12)	0.0211 (9)	0.0011 (8)	-0.0017 (7)	-0.0079 (8)
O2	0.0143 (8)	0.0292 (13)	0.0218 (9)	0.0013 (8)	-0.0029 (7)	0.0000 (8)
O3	0.0204 (9)	0.0223 (12)	0.0247 (9)	0.0072 (8)	0.0027 (7)	-0.0027 (9)
N1	0.0135 (10)	0.0205 (14)	0.0187 (11)	0.0023 (9)	-0.0038 (8)	-0.0075 (10)
N2	0.0183 (10)	0.0190 (14)	0.0152 (10)	0.0039 (9)	0.0026 (8)	0.0028 (9)
C1	0.0180 (12)	0.0222 (17)	0.0214 (12)	0.0066 (11)	0.0030 (10)	-0.0025 (11)
C2	0.0151 (11)	0.0158 (15)	0.0159 (11)	0.0025 (10)	0.0004 (9)	-0.0007 (10)
C3	0.0141 (11)	0.0220 (17)	0.0265 (13)	0.0001 (11)	0.0008 (10)	0.0007 (12)
C4	0.0166 (11)	0.0144 (14)	0.0122 (11)	-0.0001 (10)	0.0032 (9)	0.0011 (10)
C5	0.0127 (11)	0.0165 (15)	0.0109 (10)	-0.0002 (10)	0.0020 (8)	0.0010 (10)
C6	0.0176 (11)	0.0145 (15)	0.0148 (11)	0.0015 (10)	0.0005 (9)	-0.0023 (10)
C7	0.0121 (11)	0.0187 (16)	0.0152 (11)	-0.0004 (10)	-0.0013 (9)	0.0010 (10)
C8	0.0119 (11)	0.0170 (15)	0.0117 (11)	0.0030 (10)	0.0024 (8)	0.0034 (10)
C9	0.0200 (12)	0.0120 (14)	0.0155 (11)	0.0010 (10)	0.0030 (9)	-0.0018 (10)
C10	0.0143 (11)	0.0177 (15)	0.0130 (11)	-0.0018 (10)	-0.0001 (8)	-0.0016 (10)

*Geometric parameters (Å, °)*

Br1—C2	2.010 (2)	C3—H3A	0.9800
O1—C4	1.211 (3)	C3—H3B	0.9800
O2—N2	1.234 (3)	C3—H3C	0.9800
O3—N2	1.232 (3)	C5—C10	1.392 (4)
N1—C4	1.366 (3)	C5—C6	1.400 (3)
N1—C5	1.416 (3)	C6—C7	1.393 (3)
N1—H1N	0.84 (3)	C6—H6	0.9500
N2—C8	1.465 (3)	C7—C8	1.373 (4)
C1—C2	1.519 (3)	C7—H7	0.9500
C1—H1A	0.9800	C8—C9	1.387 (3)
C1—H1B	0.9800	C9—C10	1.382 (3)
C1—H1C	0.9800	C9—H9	0.9500
C2—C3	1.506 (4)	C10—H10	0.9500
C2—C4	1.546 (3)		
C4—N1—C5	127.9 (2)	H3B—C3—H3C	109.5
C4—N1—H1N	117 (2)	O1—C4—N1	123.5 (2)
C5—N1—H1N	115 (2)	O1—C4—C2	121.0 (2)
O3—N2—O2	123.4 (2)	N1—C4—C2	115.4 (2)

O3—N2—C8	118.4 (2)	C10—C5—C6	120.1 (2)
O2—N2—C8	118.2 (2)	C10—C5—N1	116.8 (2)
C2—C1—H1A	109.5	C6—C5—N1	123.2 (2)
C2—C1—H1B	109.5	C7—C6—C5	119.0 (2)
H1A—C1—H1B	109.5	C7—C6—H6	120.5
C2—C1—H1C	109.5	C5—C6—H6	120.5
H1A—C1—H1C	109.5	C8—C7—C6	119.7 (2)
H1B—C1—H1C	109.5	C8—C7—H7	120.1
C3—C2—C1	112.2 (2)	C6—C7—H7	120.1
C3—C2—C4	115.2 (2)	C7—C8—C9	122.2 (2)
C1—C2—C4	110.53 (19)	C7—C8—N2	119.3 (2)
C3—C2—Br1	108.18 (15)	C9—C8—N2	118.5 (2)
C1—C2—Br1	106.45 (18)	C10—C9—C8	118.2 (2)
C4—C2—Br1	103.51 (14)	C10—C9—H9	120.9
C2—C3—H3A	109.5	C8—C9—H9	120.9
C2—C3—H3B	109.5	C9—C10—C5	120.9 (2)
H3A—C3—H3B	109.5	C9—C10—H10	119.6
C2—C3—H3C	109.5	C5—C10—H10	119.6
H3A—C3—H3C	109.5		
C5—N1—C4—O1	0.7 (4)	C5—C6—C7—C8	-0.3 (3)
C5—N1—C4—C2	179.5 (2)	C6—C7—C8—C9	0.3 (4)
C3—C2—C4—O1	145.0 (2)	C6—C7—C8—N2	-179.9 (2)
C1—C2—C4—O1	16.6 (3)	O3—N2—C8—C7	170.6 (2)
Br1—C2—C4—O1	-97.1 (2)	O2—N2—C8—C7	-8.3 (3)
C3—C2—C4—N1	-33.9 (3)	O3—N2—C8—C9	-9.6 (3)
C1—C2—C4—N1	-162.3 (2)	O2—N2—C8—C9	171.6 (2)
Br1—C2—C4—N1	84.0 (2)	C7—C8—C9—C10	-0.4 (3)
C4—N1—C5—C10	-168.9 (2)	N2—C8—C9—C10	179.8 (2)
C4—N1—C5—C6	12.7 (4)	C8—C9—C10—C5	0.5 (3)
C10—C5—C6—C7	0.4 (3)	C6—C5—C10—C9	-0.5 (3)
N1—C5—C6—C7	178.7 (2)	N1—C5—C10—C9	-179.0 (2)

*Hydrogen-bond geometry (Å, °)*

<i>D—H...A</i>	<i>D—H</i>	<i>H...A</i>	<i>D...A</i>	<i>D—H...A</i>
C6—H6...O1	0.95	2.27	2.862 (3)	120
C7—H7...O1 <sup>i</sup>	0.95	2.45	3.139 (3)	129
N1—H1n...O3 <sup>ii</sup>	0.84 (3)	2.66 (3)	3.316 (3)	136 (2)
C10—H10...Br1 <sup>iii</sup>	0.95	2.91	3.812 (2)	160

Symmetry codes: (i)  $-x+1/2, -y-1/2, -z$ ; (ii)  $-x+1/2, y-1/2, -z+1/2$ ; (iii)  $x, y+1, z$ .