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2,5-Dichloroanilinium 4-chlorobenzenesulfonate

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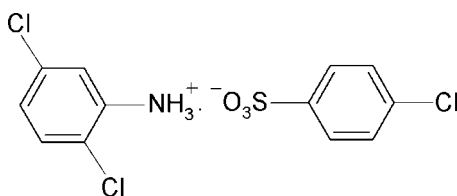
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Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{C}-\text{C}) = 0.005$ Å; R factor = 0.037; wR factor = 0.107; data-to-parameter ratio = 12.9.

In the crystal of the title compound, $\text{C}_6\text{H}_6\text{Cl}_2\text{N}^+\cdot\text{C}_6\text{H}_4\text{ClO}_3\text{S}^-$, the 2,5-dichloroanilinium cations and 4-chlorobenzenesulfonate anions are located on a crystallographic mirror plane and are connected by $\text{N}-\text{H}\cdots\text{O}$ hydrogen bonds. In the crystal, the connectivity of the hydrogen bonds leads to double chains propagating in [010].

Related literature

For the effect of substituents on the oxidative strengths of *N*-chloro, *N*-arylsulfonamides, see: Gowda *et al.* (2004a). For their effect on the structures of *N*-(aryl)-amides, see: Gowda *et al.* (2004b) and of *N*-(aryl)-methanesulfonamides, see: Gowda *et al.* (2007).



Experimental

Crystal data

$\text{C}_6\text{H}_6\text{Cl}_2\text{N}^+\cdot\text{C}_6\text{H}_4\text{ClO}_3\text{S}^-$
 $M_r = 354.62$

Monoclinic, $P2_1/m$
 $a = 9.792$ (1) Å

$b = 6.802$ (1) Å
 $c = 10.879$ (1) Å
 $\beta = 94.26$ (1)°
 $V = 722.60$ (15) Å³
 $Z = 2$

Mo $K\alpha$ radiation
 $\mu = 0.78$ mm⁻¹
 $T = 293$ K
 $0.40 \times 0.34 \times 0.24$ mm

Data collection

Oxford Diffraction Xcalibur diffractometer with a Sapphire CCD detector
Absorption correction: multi-scan (*CrysAlis RED*; Oxford)

Diffraction, 2009
 $T_{\min} = 0.745$, $T_{\max} = 0.834$
2700 measured reflections
1603 independent reflections
1439 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.012$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.037$
 $wR(F^2) = 0.107$
 $S = 1.02$
1603 reflections
124 parameters
2 restraints

H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{\max} = 0.39$ e Å⁻³
 $\Delta\rho_{\min} = -0.39$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{N1}-\text{H11N}\cdots\text{O1}^i$	0.88 (2)	1.85 (2)	2.730 (2)	176 (2)
$\text{N1}-\text{H12N}\cdots\text{O2}^{ii}$	0.89 (2)	1.88 (2)	2.753 (3)	170 (3)

Symmetry codes: (i) $-x, y + \frac{1}{2}, -z + 1$; (ii) $x, y, z + 1$.

Data collection: *CrysAlis CCD* (Oxford Diffraction, 2009); cell refinement: *CrysAlis RED* (Oxford Diffraction, 2009); data reduction: *CrysAlis RED*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *PLATON* (Spek, 2009); software used to prepare material for publication: *SHELXL97*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BT5496).

References

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supporting information

Acta Cryst. (2011). E67, o967 [doi:10.1107/S1600536811010518]

2,5-Dichloroanilinium 4-chlorobenzenesulfonate

K. Shakuntala, Sabine Foro and B. Thimme Gowda

S1. Comment

The amine and sulfonate moieties are important constituents of many important compounds. As a part of studying the substituent effects on the structures of this class of compounds (Gowda *et al.*, 2004a, 2004b, 2007), in the present work, the crystal structure of 2,5-dichloroanilinium, 4-chlorobenzenesulfonate (I) has been determined (Fig. 1). The title compound showed interesting H-bonding in its crystal structure (Fig. 2). It forms the structure through N—H···O(S) hydrogen bonding. Three H-atoms of the positively charged NH₃ group have three O atoms of the negatively charged sulfonate anion as acceptors, with each oxygen forming H-bonding with three H-atoms, one each from three positively charged NH₃ groups.

The crystal packing of (I) through N1—H11N···O1, N1—H11aN···O1a and N1—H12N···O2 hydrogen bonding (Table 1) is shown in Fig.2.

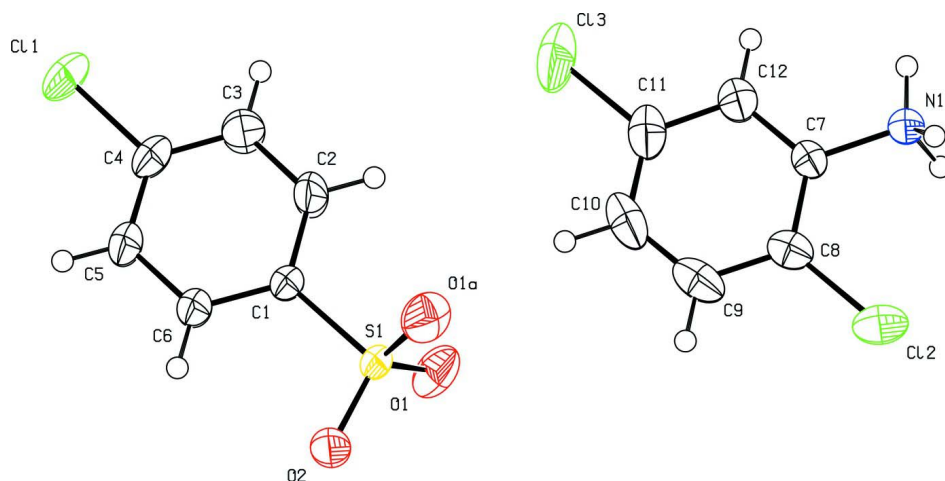
S2. Experimental

The solution of chlorobenzene (10 ml) in chloroform (40 ml) was treated dropwise with chlorosulfonic acid (25 ml) at 0 °C. After the initial evolution of hydrogen chloride subsided, the reaction mixture was brought to room temperature and poured into crushed ice in a beaker. The chloroform layer was separated, washed with cold water and allowed to evaporate slowly. The residual 4-chlorobenzenesulfonylchloride was treated with 2,5-dichloroaniline in the stoichiometric ratio and boiled for ten minutes. The reaction mixture was then cooled to room temperature and added to ice cold water (100 ml). The resultant title compound (I) was filtered under suction and washed thoroughly with cold water. It was then recrystallized to constant melting point from dilute ethanol.

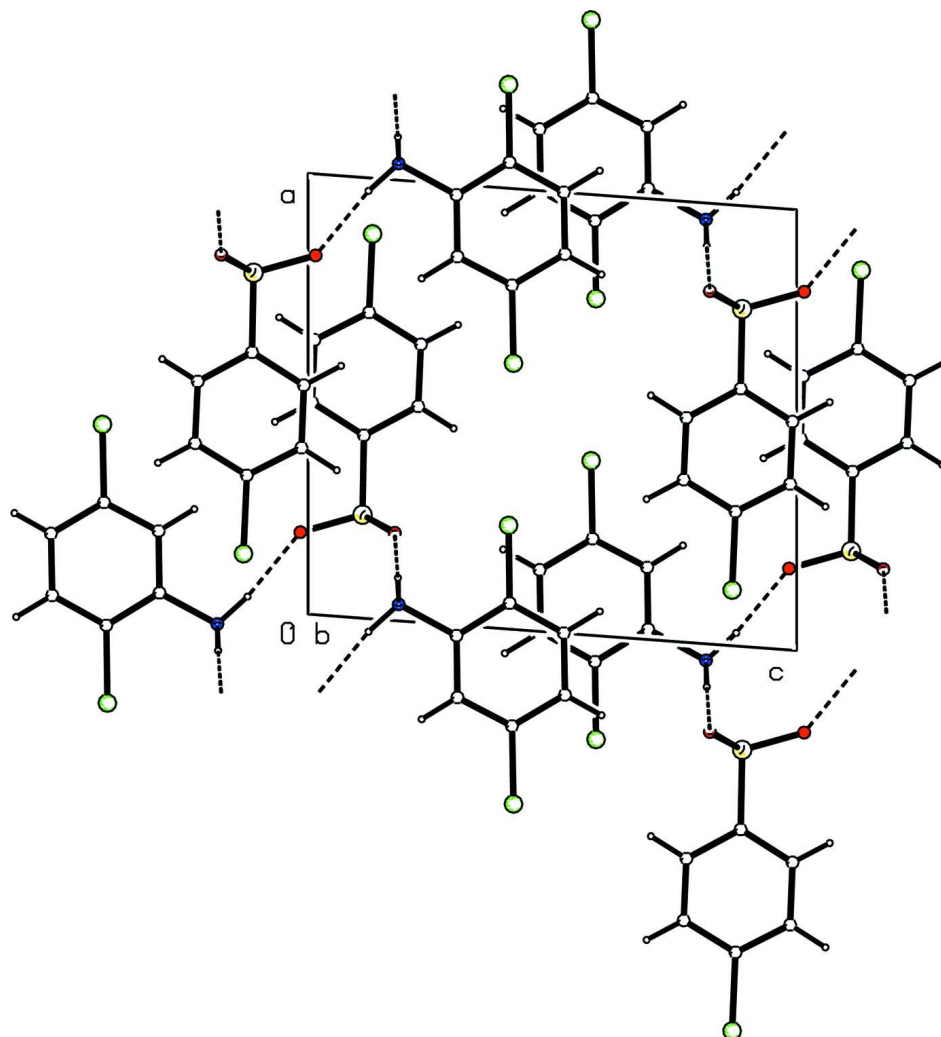
Prism like colorless single crystals used in X-ray diffraction studies were grown in ethanolic solution by slow evaporation at room temperature.

S3. Refinement

The N bounded H atoms were located in a difference map and later restrained to the distance N—H = 0.86 (2) Å. The other H atoms were positioned with idealized geometry using a riding model with C—H = 0.93 Å. All H atoms were refined with isotropic displacement parameters set to 1.2 times of the U_{eq} of the parent atom.

**Figure 1**

Molecular structure of the title compound, showing the atom labelling scheme. Displacement ellipsoids are drawn at the 50% probability level.

**Figure 2**

Molecular packing in the title compound. Hydrogen bonds are shown as dashed lines.

2,5-Dichloroanilinium 4-chlorobenzenesulfonate

Crystal data

$C_6H_6Cl_2N^+ \cdot C_6H_4ClO_3S^-$

$M_r = 354.62$

Monoclinic, $P2_1/m$

Hall symbol: -P 2yb

$a = 9.792(1) \text{ \AA}$

$b = 6.802(1) \text{ \AA}$

$c = 10.879(1) \text{ \AA}$

$\beta = 94.26(1)^\circ$

$V = 722.60(15) \text{ \AA}^3$

$Z = 2$

$F(000) = 360$

$D_x = 1.630 \text{ Mg m}^{-3}$

Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$

Cell parameters from 1811 reflections

$\theta = 2.7\text{--}27.8^\circ$

$\mu = 0.78 \text{ mm}^{-1}$

$T = 293 \text{ K}$

Prism, colourless

$0.40 \times 0.34 \times 0.24 \text{ mm}$

Data collection

Oxford Diffraction Xcalibur
diffractometer with a Sapphire CCD detector
Radiation source: fine-focus sealed tube
Graphite monochromator
Rotation method data acquisition using ω scans
Absorption correction: multi-scan
(*CrysAlis RED*; Oxford Diffraction, 2009)
 $T_{\min} = 0.745$, $T_{\max} = 0.834$

2700 measured reflections
1603 independent reflections
1439 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.012$
 $\theta_{\max} = 26.4^\circ$, $\theta_{\min} = 2.7^\circ$
 $h = -12 \rightarrow 12$
 $k = -8 \rightarrow 6$
 $l = -13 \rightarrow 10$

Refinement

Refinement on F^2
Least-squares matrix: full
 $R[F^2 > 2\sigma(F^2)] = 0.037$
 $wR(F^2) = 0.107$
 $S = 1.02$
1603 reflections
124 parameters
2 restraints
Primary atom site location: structure-invariant
direct methods
Secondary atom site location: difference Fourier
map

Hydrogen site location: inferred from
neighbouring sites
H atoms treated by a mixture of independent
and constrained refinement
 $w = 1/[\sigma^2(F_o^2) + (0.0626P)^2 + 0.4408P]$
where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} = 0.023$
 $\Delta\rho_{\max} = 0.39 \text{ e } \text{\AA}^{-3}$
 $\Delta\rho_{\min} = -0.39 \text{ e } \text{\AA}^{-3}$
Extinction correction: *SHELXL97* (Sheldrick,
2008), $F_c^* = kFc[1 + 0.001x \text{Fc}^2 \lambda^3 / \sin(2\theta)]^{-1/4}$
Extinction coefficient: 0.036 (4)

Special details

Experimental. *CrysAlis RED* (Oxford Diffraction, 2009) Empirical absorption correction using spherical harmonics, implemented in SCALE3 ABSPACK scaling algorithm.

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
C11	0.87369 (7)	0.7500	0.13209 (9)	0.0577 (3)
S1	0.23468 (6)	0.7500	0.11153 (5)	0.0274 (2)
O1	0.19906 (14)	0.5759 (2)	0.17720 (16)	0.0519 (4)
O2	0.1854 (2)	0.7500	-0.01564 (19)	0.0575 (7)
C1	0.4152 (2)	0.7500	0.1146 (2)	0.0277 (5)
C2	0.4901 (3)	0.7500	0.2257 (3)	0.0740 (15)
H2	0.4454	0.7500	0.2982	0.089*
C3	0.6311 (3)	0.7500	0.2313 (3)	0.0833 (17)
H3	0.6818	0.7500	0.3071	0.100*
C4	0.6952 (3)	0.7500	0.1251 (3)	0.0398 (7)
C5	0.6231 (3)	0.7500	0.0130 (3)	0.0406 (7)
H5	0.6684	0.7500	-0.0592	0.049*
C6	0.4811 (3)	0.7500	0.0082 (3)	0.0372 (6)

H6	0.4307	0.7500	-0.0677	0.045*
Cl2	-0.23538 (8)	0.7500	0.58953 (8)	0.0540 (3)
Cl3	0.39589 (11)	0.7500	0.57932 (12)	0.0947 (5)
N1	-0.0373 (2)	0.7500	0.8143 (2)	0.0293 (5)
H11N	-0.0933 (19)	0.852 (3)	0.8172 (19)	0.035*
H12N	0.027 (3)	0.7500	0.876 (2)	0.035*
C7	0.0232 (3)	0.7500	0.6963 (2)	0.0298 (5)
C8	-0.0594 (3)	0.7500	0.5875 (3)	0.0365 (6)
C9	0.0000 (4)	0.7500	0.4767 (3)	0.0487 (8)
H9	-0.0552	0.7500	0.4033	0.058*
C10	0.1392 (4)	0.7500	0.4734 (3)	0.0530 (9)
H10	0.1790	0.7500	0.3984	0.064*
C11	0.2195 (4)	0.7500	0.5822 (3)	0.0504 (8)
C12	0.1633 (3)	0.7500	0.6941 (3)	0.0425 (7)
H12	0.2190	0.7500	0.7672	0.051*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Cl1	0.0194 (3)	0.0852 (7)	0.0683 (6)	0.000	0.0028 (3)	0.000
S1	0.0183 (3)	0.0362 (4)	0.0279 (3)	0.000	0.0028 (2)	0.000
O1	0.0376 (8)	0.0493 (9)	0.0692 (10)	-0.0112 (7)	0.0078 (7)	0.0157 (8)
O2	0.0271 (10)	0.114 (2)	0.0306 (10)	0.000	-0.0029 (8)	0.000
C1	0.0217 (11)	0.0327 (13)	0.0290 (12)	0.000	0.0029 (9)	0.000
C2	0.0256 (15)	0.167 (5)	0.0293 (14)	0.000	0.0043 (12)	0.000
C3	0.0268 (15)	0.187 (5)	0.0349 (16)	0.000	-0.0058 (13)	0.000
C4	0.0198 (12)	0.0510 (17)	0.0487 (16)	0.000	0.0037 (11)	0.000
C5	0.0269 (13)	0.0555 (18)	0.0406 (15)	0.000	0.0113 (11)	0.000
C6	0.0257 (13)	0.0556 (17)	0.0304 (12)	0.000	0.0033 (10)	0.000
Cl2	0.0422 (4)	0.0676 (6)	0.0495 (5)	0.000	-0.0150 (3)	0.000
Cl3	0.0499 (6)	0.1572 (13)	0.0820 (7)	0.000	0.0370 (5)	0.000
N1	0.0267 (10)	0.0344 (12)	0.0268 (10)	0.000	0.0009 (8)	0.000
C7	0.0338 (13)	0.0290 (13)	0.0267 (12)	0.000	0.0040 (10)	0.000
C8	0.0450 (16)	0.0314 (14)	0.0323 (13)	0.000	-0.0029 (11)	0.000
C9	0.072 (2)	0.0452 (17)	0.0279 (14)	0.000	-0.0037 (14)	0.000
C10	0.078 (2)	0.0502 (19)	0.0336 (15)	0.000	0.0208 (15)	0.000
C11	0.0479 (18)	0.057 (2)	0.0484 (17)	0.000	0.0202 (14)	0.000
C12	0.0368 (15)	0.0563 (19)	0.0348 (14)	0.000	0.0054 (11)	0.000

Geometric parameters (Å, °)

Cl1—C4	1.744 (3)	Cl2—C8	1.725 (3)
S1—O2	1.431 (2)	Cl3—C11	1.730 (4)
S1—O1 ⁱ	1.4390 (16)	N1—C7	1.453 (3)
S1—O1	1.4390 (16)	N1—H11N	0.884 (15)
S1—C1	1.766 (2)	N1—H12N	0.886 (18)
C1—C6	1.367 (4)	C7—C12	1.373 (4)
C1—C2	1.366 (4)	C7—C8	1.383 (4)

C2—C3	1.378 (4)	C8—C9	1.377 (4)
C2—H2	0.9300	C9—C10	1.367 (5)
C3—C4	1.355 (5)	C9—H9	0.9300
C3—H3	0.9300	C10—C11	1.371 (5)
C4—C5	1.363 (4)	C10—H10	0.9300
C5—C6	1.387 (4)	C11—C12	1.374 (4)
C5—H5	0.9300	C12—H12	0.9300
C6—H6	0.9300		
O2—S1—O1 ⁱ	113.77 (8)	C5—C6—H6	119.9
O2—S1—O1	113.77 (8)	C7—N1—H11N	109.0 (14)
O1 ⁱ —S1—O1	110.72 (14)	C7—N1—H12N	111 (2)
O2—S1—C1	106.49 (12)	H11N—N1—H12N	112.4 (18)
O1 ⁱ —S1—C1	105.65 (8)	C12—C7—C8	120.5 (3)
O1—S1—C1	105.65 (8)	C12—C7—N1	119.2 (2)
C6—C1—C2	119.6 (2)	C8—C7—N1	120.3 (2)
C6—C1—S1	121.2 (2)	C9—C8—C7	119.4 (3)
C2—C1—S1	119.2 (2)	C9—C8—C12	119.9 (2)
C1—C2—C3	120.6 (3)	C7—C8—C12	120.7 (2)
C1—C2—H2	119.7	C10—C9—C8	120.6 (3)
C3—C2—H2	119.7	C10—C9—H9	119.7
C4—C3—C2	119.2 (3)	C8—C9—H9	119.7
C4—C3—H3	120.4	C11—C10—C9	119.1 (3)
C2—C3—H3	120.4	C11—C10—H10	120.4
C3—C4—C5	121.4 (3)	C9—C10—H10	120.4
C3—C4—C11	119.3 (2)	C10—C11—C12	121.6 (3)
C5—C4—C11	119.3 (2)	C10—C11—C13	119.6 (3)
C4—C5—C6	119.0 (3)	C12—C11—C13	118.9 (3)
C4—C5—H5	120.5	C11—C12—C7	118.8 (3)
C6—C5—H5	120.5	C11—C12—H12	120.6
C1—C6—C5	120.2 (3)	C7—C12—H12	120.6
C1—C6—H6	119.9		
O2—S1—C1—C6	0.0	C4—C5—C6—C1	0.0
O1 ⁱ —S1—C1—C6	-121.30 (8)	C12—C7—C8—C9	0.000 (1)
O1—S1—C1—C6	121.30 (8)	N1—C7—C8—C9	180.0
O2—S1—C1—C2	180.0	C12—C7—C8—C12	180.0
O1 ⁱ —S1—C1—C2	58.70 (8)	N1—C7—C8—C12	0.0
O1—S1—C1—C2	-58.70 (8)	C7—C8—C9—C10	0.000 (1)
C6—C1—C2—C3	0.0	C12—C8—C9—C10	180.0
S1—C1—C2—C3	180.0	C8—C9—C10—C11	0.0
C1—C2—C3—C4	0.0	C9—C10—C11—C12	0.000 (1)
C2—C3—C4—C5	0.0	C9—C10—C11—C13	180.0
C2—C3—C4—C11	180.0	C10—C11—C12—C7	0.000 (1)
C3—C4—C5—C6	0.0	C13—C11—C12—C7	180.0
C11—C4—C5—C6	180.0	C8—C7—C12—C11	0.0

C2—C1—C6—C5	0.0	N1—C7—C12—C11	180.0
S1—C1—C6—C5	180.0		

Symmetry code: (i) $x, -y+3/2, z$.

Hydrogen-bond geometry (Å, °)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
N1—H11N \cdots O1 ⁱⁱ	0.88 (2)	1.85 (2)	2.730 (2)	176 (2)
N1—H12N \cdots O2 ⁱⁱⁱ	0.89 (2)	1.88 (2)	2.753 (3)	170 (3)

Symmetry codes: (ii) $-x, y+1/2, -z+1$; (iii) $x, y, z+1$.