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## Structure Reports

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# 4,4'-Dibromo-2,2'-[ethane-1,2-diylbis-[(methylimino)methylene]]diphenol

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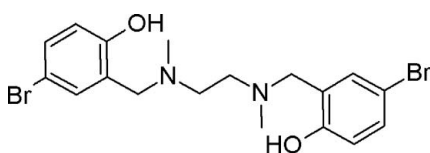
Received 6 May 2011; accepted 6 May 2011

 Key indicators: single-crystal X-ray study;  $T = 120$  K; mean  $\sigma(\text{C}-\text{C}) = 0.003$  Å;  $R$  factor = 0.021;  $wR$  factor = 0.075; data-to-parameter ratio = 14.2.

The asymmetric unit of the title compound,  $\text{C}_{18}\text{H}_{22}\text{Br}_2\text{N}_2\text{O}_2$ , contains one half-molecule that is related to the other half by a center of inversion located at the mid-point of the central C—C bond. The hydroxy (phenolic) groups are linked to the N atoms by O—H···N hydrogen bonds, which generate  $S(6)$  rings.

## Related literature

For the synthesis, see: Rivera *et al.* (2010). For the uses of tetrahydrosalens in coordination chemistry, see: Atwood (1997). For a related structure, see: Nazarenko *et al.* (2000). For reference bond lengths, see: Allen *et al.* (1987).



## Experimental

### Crystal data

 $\text{C}_{18}\text{H}_{22}\text{Br}_2\text{N}_2\text{O}_2$ 
 $M_r = 458.2$ 

 Orthorhombic, *Pbca*
 $a = 15.9282$  (3) Å

 $b = 6.1123$  (2) Å

 $c = 18.3315$  (4) Å

 $V = 1784.72$  (8) Å<sup>3</sup>
 $Z = 4$ 

 Cu  $K\alpha$  radiation

 $\mu = 5.87$  mm<sup>-1</sup>
 $T = 120$  K

 $0.36 \times 0.06 \times 0.05$  mm

### Data collection

 Oxford Diffraction Xcalibur diffractometer with an Atlas (Gemini ultra Cu) detector  
 Absorption correction: multi-scan (*CrysAlis PRO*; Oxford)

 Diffraction, 2009)  
 $T_{\min} = 0.611$ ,  $T_{\max} = 1$   
 24526 measured reflections  
 1591 independent reflections  
 1482 reflections with  $I > 3\sigma(I)$   
 $R_{\text{int}} = 0.028$ 

### Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.021$ 
 $wR(F^2) = 0.075$ 
 $S = 1.52$ 

1591 reflections

112 parameters

H atoms treated by a mixture of independent and constrained refinement

 $\Delta\rho_{\text{max}} = 0.20$  e Å<sup>-3</sup>
 $\Delta\rho_{\text{min}} = -0.32$  e Å<sup>-3</sup>
**Table 1**

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{O4}-\text{H4}o\cdots\text{N2}$	0.81 (2)	1.86 (2)	2.6051 (19)	154 (2)

Data collection: *CrysAlis PRO* (Oxford Diffraction, 2009); cell refinement: *CrysAlis PRO*; data reduction: *CrysAlis PRO*; program(s) used to solve structure: *SIR2002* (Burla *et al.*, 2003); program(s) used to refine structure: *JANA2006* (Petříček *et al.*, 2006); molecular graphics: *DIAMOND* (Brandenburg & Putz, 2005); software used to prepare material for publication: *JANA2006*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HB5874).

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## supporting information

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**4,4'-Dibromo-2,2'-{ethane-1,2-diylbis[(methylimino)methylene]}diphenol**

Augusto Rivera, Jicli José Rojas, Jaime Ríos-Motta, Michal Dušek and Karla Fejfarová

**S1. Comment**

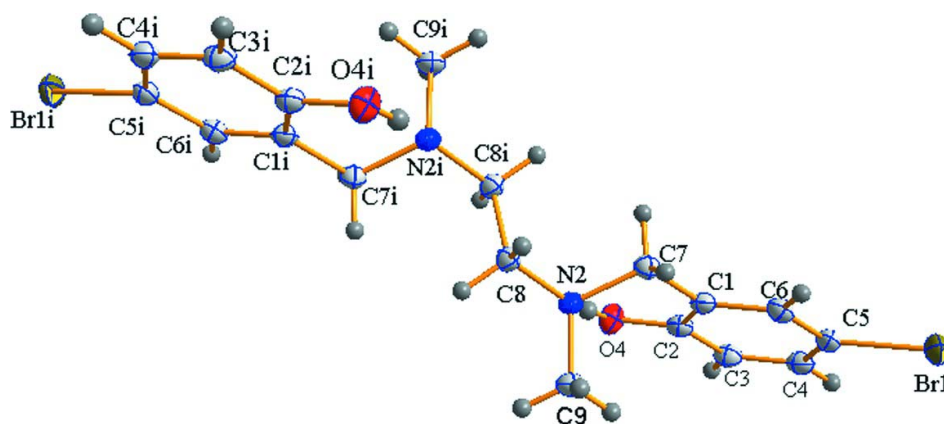
Recently, we reported the synthesis of a new class of ligands by a ring-opening reduction of bis-1,3-benzoxazines with sodium borohydride (Rivera *et al.*, 2010), and the products of these reactions are referred to as N,N'-disubstituted tetrahydro-salens (Atwood, 1997). Here we report the crystal structure of title compound (**I**). The C(sp<sup>3</sup>)—X bond distances and angles in (**I**) are within normal ranges (Allen *et al.*, 1987) and comparable with a related structure (Nazarenko, *et al.*, 2000). The C—N bonds in the N—CH<sub>2</sub>CH<sub>2</sub>—N segment are anti to each other, with a torsion angle of 180°. The observed conformation is stabilized by the short intramolecular hydrogen bonds O—H⋯N (Table 1), and these interactions generate S(6) ring motifs.

**S2. Experimental**

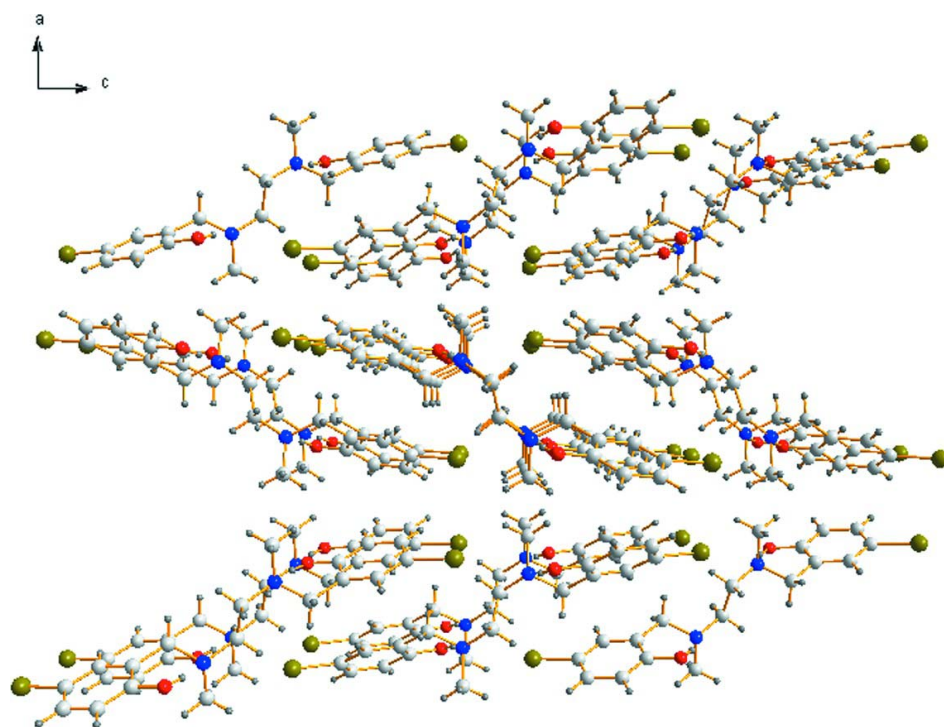
Sodium borohydride (3.0 mmol, 0.11 g) was added to a solution of 3,3'-ethylene-bis-(3,4-dihydro-6-bromo-2*H*-1,3-benzoxazine) (1 mmol) in ethanol (15 ml), and the mixture was stirred magnetically for 30 min at room temperature. After completion of the reaction, the mixture was poured into ice-cold water, neutralized with ammonium chloride (12 ml), and extracted with CHCl<sub>3</sub> (3 times 10 cm<sup>3</sup>). The combined extracts were dried over anhydrous Na<sub>2</sub>SO<sub>4</sub> and evaporated. The solid obtained was purified by recrystallization from ethanol to yield colourless needles of (**I**).

**S3. Refinement**

All hydrogen atoms were discernible in difference Fourier maps and could be refined to reasonable geometry. According to common practice H atoms bonded C atoms were kept in ideal positions with C—H distance 0.96 Å during the refinement. The methyl H atoms were allowed to rotate freely about the adjacent C—C bonds. The hydroxy hydrogen was found in difference Fourier maps and its coordinates were refined freely. The isotropic atomic displacement parameters of hydrogen atoms were evaluated as 1.2×*U*<sub>eq</sub> of the parent atom.

**Figure 1**

The molecule of the title compound. Displacement ellipsoids are drawn at 50% probability level. Atoms with suffix *i* are generated by the symmetry operation (1-*x*, -*y*, 2-*z*).

**Figure 2**

The packing for (I).

#### 4,4'-Dibromo-2,2'-(ethane-1,2-diylbis[(methylimino)methylene])diphenol

##### Crystal data

$C_{18}H_{22}Br_2N_2O_2$

$M_r = 458.2$

Orthorhombic, *Pbca*

Hall symbol: -P 2ac 2ab

$a = 15.9282(3) \text{ \AA}$

$b = 6.1123(2) \text{ \AA}$

$c = 18.3315(4) \text{ \AA}$

$V = 1784.72(8) \text{ \AA}^3$

$Z = 4$

$F(000) = 920$

$D_x = 1.705 \text{ Mg m}^{-3}$

Cu  $K\alpha$  radiation,  $\lambda = 1.5418 \text{ \AA}$

Cell parameters from 16599 reflections  
 $\theta = 2.8\text{--}66.9^\circ$   
 $\mu = 5.87\text{ mm}^{-1}$

$T = 120\text{ K}$   
 Needle, colourless  
 $0.36 \times 0.06 \times 0.05\text{ mm}$

*Data collection*

Oxford Diffraction Xcalibur  
 diffractometer with an Atlas (Gemini ultra Cu)  
 detector  
 Radiation source: Enhance Ultra (Cu) X-ray  
 Source  
 Mirror monochromator  
 Detector resolution:  $10.3784\text{ pixels mm}^{-1}$   
 Rotation method data acquisition using  $\omega$  scans  
 Absorption correction: multi-scan  
 (CrysAlis PRO; Oxford Diffraction, 2009)

$T_{\min} = 0.611$ ,  $T_{\max} = 1$   
 24526 measured reflections  
 1591 independent reflections  
 1482 reflections with  $I > 3\sigma(I)$   
 $R_{\text{int}} = 0.028$   
 $\theta_{\max} = 67.1^\circ$ ,  $\theta_{\min} = 4.8^\circ$   
 $h = -18 \rightarrow 18$   
 $k = -7 \rightarrow 7$   
 $l = -21 \rightarrow 21$

*Refinement*

Refinement on  $F^2$   
 $R[F^2 > 2\sigma(F^2)] = 0.021$   
 $wR(F^2) = 0.075$   
 $S = 1.52$   
 1591 reflections  
 112 parameters  
 0 restraints  
 41 constraints

H atoms treated by a mixture of independent  
 and constrained refinement  
 Weighting scheme based on measured s.u.'s  $w =$   
 $1/[\sigma^2(I) + 0.0016I^2]$   
 $(\Delta/\sigma)_{\max} = 0.008$   
 $\Delta\rho_{\max} = 0.20\text{ e \AA}^{-3}$   
 $\Delta\rho_{\min} = -0.32\text{ e \AA}^{-3}$

*Special details*

**Experimental.** CrysAlisPro, Oxford Diffraction (2009), Empirical absorption correction using spherical harmonics, implemented in SCALE3 ABSPACK scaling algorithm.

**Refinement.** The refinement was carried out against all reflections. The conventional  $R$ -factor is always based on  $F$ . The goodness of fit as well as the weighted  $R$ -factor are based on  $F$  and  $F^2$  for refinement carried out on  $F$  and  $F^2$ , respectively. The threshold expression is used only for calculating  $R$ -factors *etc.* and it is not relevant to the choice of reflections for refinement.

The program used for refinement, Jana2006, uses the weighting scheme based on the experimental expectations, see `_refine_ls_weighting_details`, that does not force  $S$  to be one. Therefore the values of  $S$  are usually larger than the ones from the *SHELX* program.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
Br1	0.640851 (13)	0.05571 (4)	0.580303 (11)	0.03125 (11)
O4	0.61560 (9)	0.4268 (2)	0.88393 (7)	0.0270 (4)
N2	0.58909 (9)	0.0319 (2)	0.93218 (8)	0.0195 (4)
C1	0.59115 (9)	0.1285 (3)	0.80150 (9)	0.0202 (4)
C2	0.62193 (10)	0.3384 (3)	0.81629 (9)	0.0214 (5)
C3	0.66038 (12)	0.4593 (3)	0.76131 (11)	0.0242 (5)
C4	0.66647 (10)	0.3757 (3)	0.69103 (9)	0.0251 (5)
C5	0.63443 (9)	0.1714 (3)	0.67664 (10)	0.0227 (5)
C6	0.59736 (10)	0.0460 (3)	0.73124 (10)	0.0217 (5)
C7	0.54916 (10)	-0.0057 (3)	0.86063 (9)	0.0216 (4)
C8	0.53980 (10)	-0.0646 (3)	0.99219 (9)	0.0218 (5)
C9	0.67487 (11)	-0.0584 (3)	0.93342 (10)	0.0241 (5)
H3	0.682855	0.601479	0.77203	0.0291*

H4	0.692687	0.459508	0.65304	0.0301*
H6	0.576136	-0.09726	0.720273	0.026*
H7a	0.490778	0.032417	0.863354	0.0259*
H7b	0.552676	-0.158184	0.848428	0.0259*
H8a	0.573654	-0.070797	1.035465	0.0262*
H8b	0.525375	-0.212723	0.980158	0.0262*
H9a	0.701148	-0.023134	0.97905	0.0289*
H9b	0.672436	-0.214496	0.927867	0.0289*
H9c	0.706892	0.003417	0.894138	0.0289*
H4o	0.6031 (15)	0.326 (4)	0.9104 (12)	0.0323*

*Atomic displacement parameters (Å<sup>2</sup>)*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Br1	0.03214 (19)	0.0431 (2)	0.01850 (19)	0.00457 (7)	-0.00120 (6)	-0.00091 (7)
O4	0.0343 (7)	0.0206 (6)	0.0260 (7)	-0.0033 (5)	0.0042 (5)	-0.0036 (5)
N2	0.0153 (7)	0.0236 (8)	0.0195 (7)	-0.0017 (5)	0.0009 (5)	0.0002 (5)
C1	0.0144 (7)	0.0228 (8)	0.0235 (8)	0.0010 (6)	-0.0004 (6)	0.0020 (7)
C2	0.0172 (7)	0.0228 (8)	0.0243 (8)	0.0015 (6)	-0.0011 (6)	0.0006 (7)
C3	0.0210 (8)	0.0226 (9)	0.0291 (10)	-0.0022 (6)	-0.0008 (7)	0.0037 (6)
C4	0.0180 (8)	0.0298 (9)	0.0276 (9)	-0.0002 (7)	-0.0002 (6)	0.0075 (7)
C5	0.0181 (8)	0.0314 (10)	0.0185 (8)	0.0044 (6)	-0.0013 (5)	0.0004 (7)
C6	0.0190 (8)	0.0228 (9)	0.0232 (9)	0.0018 (6)	-0.0032 (6)	0.0008 (6)
C7	0.0194 (8)	0.0244 (8)	0.0210 (8)	-0.0042 (7)	-0.0012 (6)	0.0005 (7)
C8	0.0189 (8)	0.0251 (9)	0.0214 (8)	-0.0012 (6)	0.0012 (6)	0.0036 (6)
C9	0.0163 (9)	0.0289 (10)	0.0271 (8)	0.0019 (6)	0.0000 (7)	0.0004 (6)

*Geometric parameters (Å, °)*

Br1—C5	1.9051 (18)	C4—H4	0.96
O4—H4o	0.81 (2)	C5—C6	1.392 (2)
N2—C7	1.476 (2)	C6—H6	0.96
N2—C8	1.475 (2)	C7—H7a	0.96
N2—C9	1.474 (2)	C7—H7b	0.96
C1—C2	1.400 (2)	C8—C8 <sup>i</sup>	1.521 (2)
C1—C6	1.387 (2)	C8—H8a	0.96
C1—C7	1.515 (2)	C8—H8b	0.96
C2—C3	1.392 (3)	C9—H9a	0.96
C3—C4	1.389 (3)	C9—H9b	0.96
C3—H3	0.96	C9—H9c	0.96
C4—C5	1.375 (3)		
C7—N2—C8	111.79 (13)	N2—C7—C1	111.17 (13)
C7—N2—C9	110.78 (13)	N2—C7—H7a	109.4713
C8—N2—C9	109.40 (13)	N2—C7—H7b	109.4711
C2—C1—C6	119.21 (15)	C1—C7—H7a	109.4712
C2—C1—C7	120.83 (14)	C1—C7—H7b	109.4709
C6—C1—C7	119.93 (15)	H7a—C7—H7b	107.7191

C1—C2—C3	120.04 (16)	N2—C8—C8 <sup>i</sup>	112.14 (13)
C2—C3—C4	120.46 (16)	N2—C8—H8a	109.4713
C2—C3—H3	119.771	N2—C8—H8b	109.4716
C4—C3—H3	119.7725	C8 <sup>i</sup> —C8—H8a	109.4716
C3—C4—C5	119.09 (16)	C8 <sup>i</sup> —C8—H8b	109.4707
C3—C4—H4	120.4567	H8a—C8—H8b	106.6664
C5—C4—H4	120.4562	N2—C9—H9a	109.4705
Br1—C5—C4	119.68 (13)	N2—C9—H9b	109.471
Br1—C5—C6	119.00 (14)	N2—C9—H9c	109.4718
C4—C5—C6	121.31 (16)	H9a—C9—H9b	109.4713
C1—C6—C5	119.85 (16)	H9a—C9—H9c	109.4714
C1—C6—H6	120.0729	H9b—C9—H9c	109.4713
C5—C6—H6	120.073		
N2—C8—C8 <sup>i</sup> —N2 <sup>i</sup>	180		

Symmetry code: (i)  $-x+1, -y, -z+2$ .

*Hydrogen-bond geometry (Å, °)*

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
O4—H4 $\cdots$ N2	0.81 (2)	1.86 (2)	2.6051 (19)	154 (2)