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N-(3-Chlorobenzoyl)-2-methylbenzenesulfonamide

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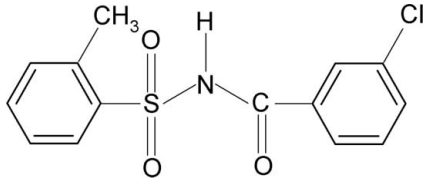
Key indicators: single-crystal X-ray study; T = 293 K; mean σ(C–C) = 0.004 Å; R factor = 0.042; wR factor = 0.113; data-to-parameter ratio = 16.3.

In the title compound, $C_{14}H_{12}ClNO_3S$, the N—H bond in the C—SO_2—NH—C(O) segment is *anti* to the C=O bond. Further, the C=O bond and the *meta*-Cl atom in the benzoyl ring are also *anti* to each other. The dihedral angle between the sulfonyl and the benzoyl benzene rings is 72.4 (1)°. In the crystal, molecules are linked by pairs of N—H···O hydrogen bonds, forming inversion dimers.

Related literature

For our studies on the effects of substituents on the structures and other aspects of *N*-(aryl)-amides, see: Bowes *et al.* (2003); Gowda *et al.* (2004), on *N*-(aryl)-methanesulfonamides, see: Jayalakshmi & Gowda (2004), on *N*-(aryl)-arylsulfonamides, see: Gowda *et al.* (2003), on *N*-(substitutedbenzoyl)-arylsulfonamides, see: Suchetan *et al.* (2010) and on *N*-chloroarylamides, see: Gowda *et al.* (1996).



Experimental

Crystal data

$C_{14}H_{12}ClNO_3S$
M_r = 309.76
Monoclinic, $C2/c$
a = 18.043 (2) Å
b = 12.046 (1) Å
c = 15.596 (2) Å
β = 118.77 (2)°
V = 2971.3 (6) Å^3
Z = 8
Mo $K\alpha$ radiation
μ = 0.40 mm^{-1}
T = 293 K
0.40 × 0.36 × 0.32 mm

Data collection

Oxford Diffraction Xcalibur diffractometer with Sapphire CCD detector
Absorption correction: multi-scan (*CrysAlis RED*; Oxford Diffraction, 2009)
T_{min} = 0.856, T_{max} = 0.882
5996 measured reflections
3024 independent reflections
2463 reflections with $I > 2\sigma(I)$
R_{int} = 0.012

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.042
$wR(F^2)$ = 0.113
S = 1.03
3024 reflections
185 parameters
1 restraint
H atoms treated by a mixture of independent and constrained refinement
$\Delta\rho_{max}$ = 0.45 e Å^{-3}
$\Delta\rho_{min}$ = −0.64 e Å^{-3}

Table 1
Hydrogen-bond geometry (Å, °).

D—H···A	D—H	H···A	D···A	D—H···A
N1—H1*N*···O2^{i}	0.82 (2)	2.06 (2)	2.876 (2)	174 (2)

Symmetry code: (i) $-x+\frac{1}{2}, -y+\frac{1}{2}, -z+2$.

Data collection: *CrysAlis CCD* (Oxford Diffraction, 2009); cell refinement: *CrysAlis CCD*); data reduction: *CrysAlis RED* (Oxford Diffraction, 2009; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *PLATON* (Spek, 2009); software used to prepare material for publication: *SHELXL97*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BT5728).

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supporting information

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N-(3-Chlorobenzoyl)-2-methylbenzenesulfonamide

P. A. Suchetan, Sabine Foro and B. Thimme Gowda

S1. Comment

Diaryl acylsulfonamides are known as potent antitumor agents against a broad spectrum of human tumor xenografts in nude mice. As part of our studies on the substituent effects on the structures and other aspects of *N*-(aryl)-amides (Bowes *et al.*, 2003; Gowda *et al.*, 2004), *N*-(aryl)-methanesulfonamides (Jayalakshmi & Gowda, 2004), *N*-(aryl)-aryl-sulfonamides (Gowda *et al.*, 2003); *N*-(substitutedbenzoyl)-arylsulfonamides (Suchetan *et al.*, 2010) and *N*-chloro-aryl-sulfonamides (Gowda *et al.*, 1996), in the present work, the crystal structure of *N*-(3-Chlorobenzoyl)- 2-methylbenzene-sulfonamide (I) has been determined (Fig.1).

The conformation of the N—H bond in the C—SO_2—NH—C(O) segment is *anti* to the C=O bond (Fig.1), similar to that observed in *N*-(benzoyl)-2-methylbenzenesulfonamide (II)(Suchetan *et al.*, 2010). Further, the conformation between the C=O bond and the *meta*-Cl in the benzoyl ring is also *anti* to each other.

The molecules are twisted at the *S* atom with the torsional angle of 66.9 (2)°, compared to the value of 68.8 (4)° in (II).

The dihedral angle between the sulfonyl benzene ring and the —SO_2—NH—C—O segment is 82.4 (1)°, compared to the value of 84.8 (1)° in (II). Furthermore, the dihedral angle between the sulfonyl and the benzoyl benzene rings is 72.4 (1)°, compared to the value of 73.9 (1)° in (II).

The packing of molecules linked by of N—H···O(S) hydrogen bonds(Table 1) is shown in Fig. 2.

S2. Experimental

The title compound was prepared by refluxing a mixture of 3-chlorobenzoic acid (0.02 mole), 2-methylbenzene-sulfonamide (0.02 mole) and excess phosphorous oxy chloride for 3 h on a water bath. The resultant mixture was cooled and poured into crushed ice. The solid, *N*-(3-chlorobenzoyl)2-methylbenzenesulfonamide, obtained was filtered, washed thoroughly with water and then dissolved in sodium bicarbonate solution. The compound was later reprecipitated by acidifying the filtered solution with dilute HCl. It was filtered, dried and recrystallized.

Prism like colourless single crystals of the title compound used in X-ray diffraction studies were obtained by slow evaporation of its toluene solution at room temperature.

S3. Refinement

The H atom of the NH group was located in a difference map and later restrained to N—H = 0.86 (2) %A. The other H atoms were positioned with idealized geometry using a riding model with C—H distances of 0.93Å (C-aromatic) and 0.96Å (C-methyl).

All H atoms were refined with isotropic displacement parameters were set at 1.2 U_{eq}(C-aromatic, N) and 1.5 U_{eq}(C-methyl).

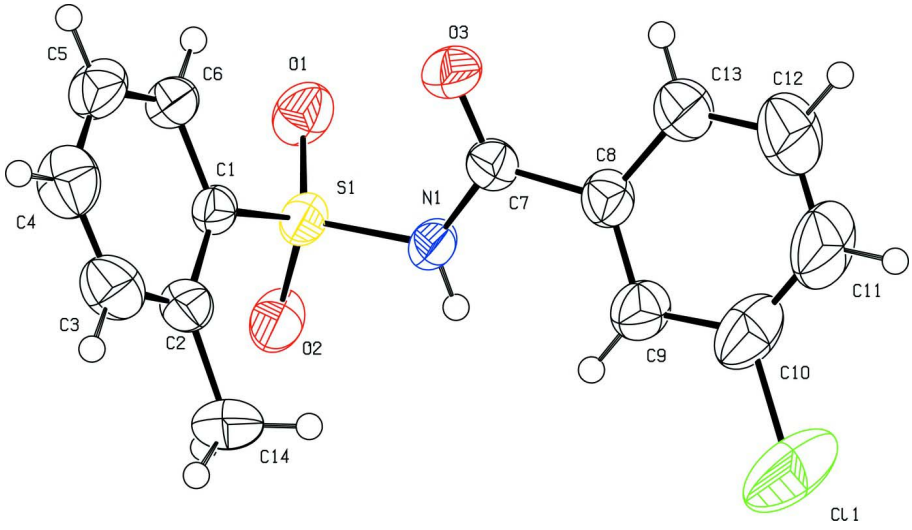

Figure 1

Molecular structure of the title compound, showing the atom-labelling scheme. Displacement ellipsoids are drawn at the 50% probability level.

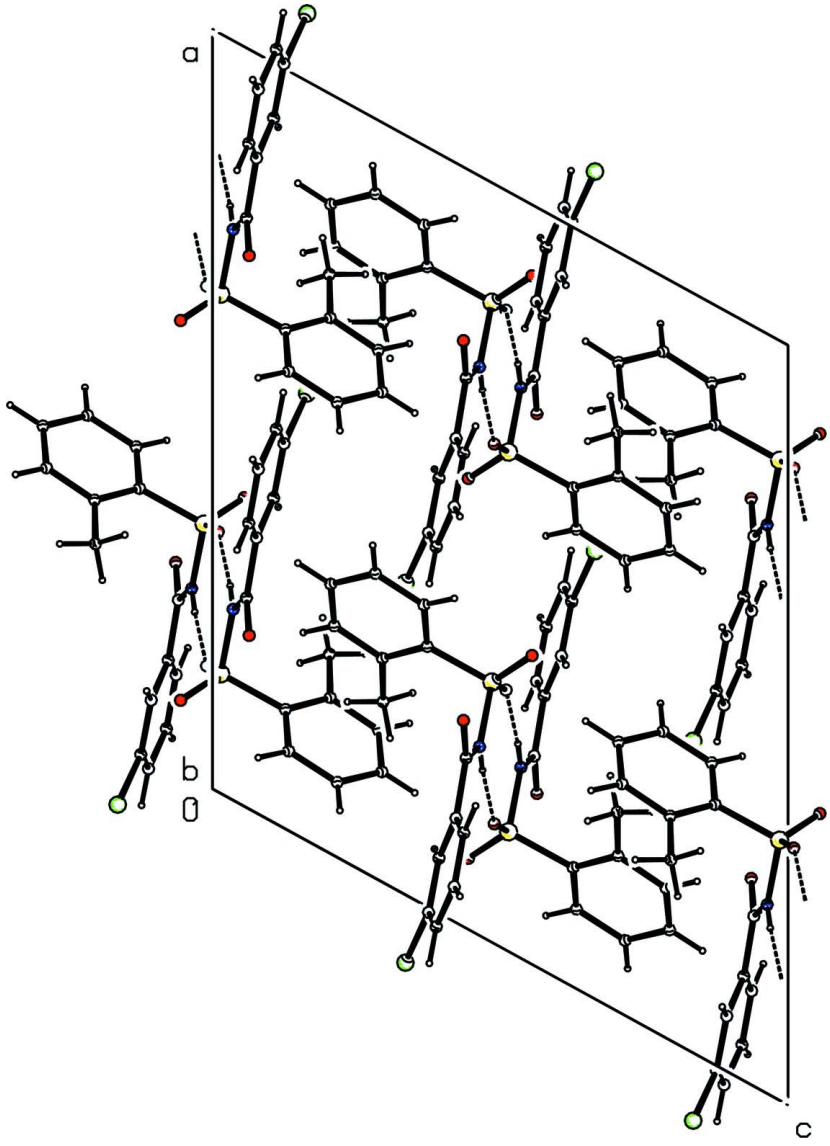

Figure 2

Molecular packing in the title compound. Hydrogen bonds are shown as dashed lines.

N-(3-Chlorobenzoyl)-2-methylbenzenesulfonamide

Crystal data

$C_{14}H_{12}ClNO_3S$
$M_r = 309.76$
Monoclinic, $C2/c$
Hall symbol: -C 2yc
$a = 18.043\ (2)$ Å
$b = 12.046\ (1)$ Å
$c = 15.596\ (2)$ Å
$\beta = 118.77\ (2)°$
$V = 2971.3\ (6)$ Å^3
$Z = 8$
$F(000) = 1280$
$D_x = 1.385$ Mg m^{-3}
Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
Cell parameters from 3176 reflections
$\theta = 2.8–27.8°$
$\mu = 0.40$ mm^{-1}
$T = 293$ K
Prism, colourless
0.40 × 0.36 × 0.32 mm

Data collection

Oxford Diffraction Xcalibur diffractometer with Sapphire CCD detector
Radiation source: fine-focus sealed tube
Graphite monochromator
Rotation method data acquisition using ω and phi scans.
Absorption correction: multi-scan (*CrysAlis RED*; Oxford Diffraction, 2009)
T_{min} = 0.856, T_{max} = 0.882
5996 measured reflections
3024 independent reflections
2463 reflections with $I > 2\sigma(I)$
R_{int} = 0.012
θ_{max} = 26.4°, θ_{min} = 2.9°
h = −22→18
k = −15→11
l = −18→19

Refinement

Refinement on F^2
Least-squares matrix: full
$R[F^2 > 2\sigma(F^2)]$ = 0.042
$wR(F^2)$ = 0.113
S = 1.03
3024 reflections
185 parameters
1 restraint
Primary atom site location: structure-invariant direct methods
Secondary atom site location: difference Fourier map
Hydrogen site location: inferred from neighbouring sites
H atoms treated by a mixture of independent and constrained refinement
$w = 1/[\sigma^2(F_o^2) + (0.0488P)^2 + 4.0385P]$ where $P = (F_o^2 + 2F_c^2)/3$
$(\Delta/\sigma)_{max} < 0.001$
$\Delta\rho_{max}$ = 0.45 e Å^{-3}
$\Delta\rho_{min}$ = −0.64 e Å^{-3}

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R-factor wR and goodness of fit S are based on F^2, conventional R-factors R are based on F, with F set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. R-factors based on F^2 are statistically about twice as large as those based on F, and R- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å²)

	x	y	z	U_{iso}*/U_{eq}
C1	0.34287 (11)	0.11403 (15)	0.87322 (14)	0.0349 (4)
C2	0.29479 (13)	0.18523 (18)	0.79536 (16)	0.0449 (5)
C3	0.30275 (17)	0.1707 (2)	0.71180 (18)	0.0604 (6)
H3	0.2716	0.2161	0.6580	0.072*
C4	0.35502 (18)	0.0918 (2)	0.70565 (19)	0.0627 (7)
H4	0.3591	0.0854	0.6486	0.075*
C5	0.40121 (16)	0.0224 (2)	0.78294 (19)	0.0546 (6)
H5	0.4363	−0.0312	0.7785	0.065*
C6	0.39508 (13)	0.03297 (18)	0.86754 (16)	0.0434 (5)
H6	0.4257	−0.0139	0.9203	0.052*
C7	0.22402 (13)	−0.03082 (17)	0.94079 (16)	0.0418 (5)
C8	0.13770 (14)	−0.06030 (18)	0.92343 (16)	0.0451 (5)
C9	0.07257 (14)	0.0159 (2)	0.89365 (18)	0.0535 (6)
H9	0.0814	0.0905	0.8859	0.064*
C10	−0.00629 (16)	−0.0220 (3)	0.8757 (2)	0.0736 (8)

C11	−0.0197 (2)	−0.1309 (3)	0.8889 (3)	0.0970 (12)
H11	−0.0729	−0.1545	0.8770	0.116*
C12	0.0461 (2)	−0.2054 (3)	0.9201 (3)	0.0981 (12)
H12	0.0372	−0.2793	0.9298	0.118*
C13	0.12381 (17)	−0.1717 (2)	0.9368 (2)	0.0633 (7)
H13	0.1678	−0.2227	0.9571	0.076*
C14	0.23804 (18)	0.2758 (2)	0.7967 (2)	0.0657 (7)
H14A	0.2719	0.3353	0.8380	0.079*
H14B	0.2025	0.2469	0.8214	0.079*
H14C	0.2036	0.3032	0.7314	0.079*
N1	0.24872 (11)	0.07910 (15)	0.96409 (14)	0.0425 (4)
H1N	0.2195 (14)	0.1279 (17)	0.9688 (18)	0.051*
O1	0.40644 (10)	0.05491 (15)	1.05526 (11)	0.0546 (4)
O2	0.34229 (10)	0.23938 (13)	1.01017 (13)	0.0547 (4)
O3	0.27146 (10)	−0.09871 (13)	0.93618 (14)	0.0578 (4)
Cl1	−0.08922 (5)	0.07097 (11)	0.83538 (9)	0.1187 (4)
S1	0.34256 (3)	0.12455 (4)	0.98556 (4)	0.03864 (16)

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.0307 (9)	0.0335 (10)	0.0424 (10)	−0.0031 (8)	0.0192 (8)	−0.0026 (8)
C2	0.0401 (11)	0.0408 (11)	0.0523 (12)	0.0002 (9)	0.0210 (10)	0.0047 (10)
C3	0.0678 (16)	0.0615 (15)	0.0487 (13)	0.0021 (13)	0.0255 (12)	0.0117 (12)
C4	0.0773 (18)	0.0704 (17)	0.0510 (14)	−0.0068 (14)	0.0393 (13)	−0.0077 (13)
C5	0.0567 (14)	0.0533 (14)	0.0659 (15)	0.0006 (11)	0.0393 (12)	−0.0115 (12)
C6	0.0397 (11)	0.0412 (11)	0.0517 (12)	0.0025 (9)	0.0239 (9)	−0.0007 (9)
C7	0.0449 (11)	0.0368 (11)	0.0473 (11)	−0.0006 (9)	0.0251 (10)	−0.0003 (9)
C8	0.0448 (12)	0.0423 (12)	0.0512 (12)	−0.0074 (9)	0.0254 (10)	−0.0072 (10)
C9	0.0428 (12)	0.0514 (13)	0.0631 (14)	−0.0037 (10)	0.0229 (11)	−0.0074 (11)
C10	0.0432 (14)	0.092 (2)	0.0811 (19)	−0.0048 (14)	0.0263 (13)	−0.0196 (17)
C11	0.0587 (18)	0.090 (2)	0.149 (3)	−0.0355 (18)	0.055 (2)	−0.034 (2)
C12	0.093 (3)	0.0636 (19)	0.155 (4)	−0.0324 (19)	0.073 (3)	−0.015 (2)
C13	0.0608 (15)	0.0442 (13)	0.092 (2)	−0.0093 (12)	0.0422 (15)	−0.0075 (13)
C14	0.0622 (16)	0.0524 (14)	0.0783 (18)	0.0205 (13)	0.0304 (14)	0.0163 (13)
N1	0.0388 (9)	0.0374 (9)	0.0605 (11)	−0.0007 (7)	0.0312 (9)	−0.0054 (8)
O1	0.0428 (8)	0.0729 (11)	0.0445 (8)	0.0079 (8)	0.0182 (7)	0.0070 (8)
O2	0.0522 (9)	0.0466 (9)	0.0766 (11)	−0.0129 (7)	0.0401 (9)	−0.0226 (8)
O3	0.0557 (10)	0.0395 (8)	0.0884 (13)	0.0041 (7)	0.0428 (9)	−0.0047 (8)
Cl1	0.0500 (4)	0.1486 (10)	0.1396 (9)	0.0240 (5)	0.0313 (5)	−0.0175 (7)
S1	0.0337 (3)	0.0404 (3)	0.0455 (3)	−0.0024 (2)	0.0219 (2)	−0.0056 (2)

Geometric parameters (Å, º)

C1—C6	1.389 (3)	C9—C10	1.388 (3)
C1—C2	1.396 (3)	C9—H9	0.9300
C1—S1	1.759 (2)	C10—C11	1.368 (5)
C2—C3	1.389 (3)	C10—Cl1	1.727 (3)

C2—C14	1.503 (3)	C11—C12	1.376 (5)
C3—C4	1.375 (4)	C11—H11	0.9300
C3—H3	0.9300	C12—C13	1.359 (4)
C4—C5	1.370 (4)	C12—H12	0.9300
C4—H4	0.9300	C13—H13	0.9300
C5—C6	1.382 (3)	C14—H14A	0.9600
C5—H5	0.9300	C14—H14B	0.9600
C6—H6	0.9300	C14—H14C	0.9600
C7—O3	1.211 (2)	N1—S1	1.6532 (17)
C7—N1	1.389 (3)	N1—H1N	0.816 (16)
C7—C8	1.490 (3)	O1—S1	1.4170 (17)
C8—C9	1.384 (3)	O2—S1	1.4361 (16)
C8—C13	1.399 (3)		

C6—C1—C2	122.21 (19)	C11—C10—C9	121.5 (3)
C6—C1—S1	116.20 (16)	C11—C10—Cl1	119.5 (2)
C2—C1—S1	121.57 (15)	C9—C10—Cl1	119.0 (3)
C3—C2—C1	115.9 (2)	C10—C11—C12	119.6 (3)
C3—C2—C14	118.8 (2)	C10—C11—H11	120.2
C1—C2—C14	125.2 (2)	C12—C11—H11	120.2
C4—C3—C2	122.4 (2)	C13—C12—C11	120.5 (3)
C4—C3—H3	118.8	C13—C12—H12	119.7
C2—C3—H3	118.8	C11—C12—H12	119.7
C5—C4—C3	120.5 (2)	C12—C13—C8	120.0 (3)
C5—C4—H4	119.8	C12—C13—H13	120.0
C3—C4—H4	119.8	C8—C13—H13	120.0
C4—C5—C6	119.3 (2)	C2—C14—H14A	109.5
C4—C5—H5	120.3	C2—C14—H14B	109.5
C6—C5—H5	120.3	H14A—C14—H14B	109.5
C5—C6—C1	119.6 (2)	C2—C14—H14C	109.5
C5—C6—H6	120.2	H14A—C14—H14C	109.5
C1—C6—H6	120.2	H14B—C14—H14C	109.5
O3—C7—N1	120.84 (19)	C7—N1—S1	122.50 (14)
O3—C7—C8	122.37 (19)	C7—N1—H1N	124.8 (18)
N1—C7—C8	116.78 (18)	S1—N1—H1N	112.7 (18)
C9—C8—C13	120.2 (2)	O1—S1—O2	118.13 (11)
C9—C8—C7	123.2 (2)	O1—S1—N1	109.57 (10)
C13—C8—C7	116.6 (2)	O2—S1—N1	103.78 (9)
C8—C9—C10	118.1 (2)	O1—S1—C1	109.27 (9)
C8—C9—H9	120.9	O2—S1—C1	109.73 (10)
C10—C9—H9	120.9	N1—S1—C1	105.56 (9)

C6—C1—C2—C3	−0.5 (3)	C8—C9—C10—Cl1	−178.3 (2)
S1—C1—C2—C3	177.92 (17)	C9—C10—C11—C12	−0.6 (6)
C6—C1—C2—C14	−179.0 (2)	Cl1—C10—C11—C12	179.4 (3)
S1—C1—C2—C14	−0.6 (3)	C10—C11—C12—C13	−0.6 (6)
C1—C2—C3—C4	−0.4 (4)	C11—C12—C13—C8	0.7 (6)
C14—C2—C3—C4	178.2 (3)	C9—C8—C13—C12	0.3 (4)

C2—C3—C4—C5	0.8 (4)	C7—C8—C13—C12	−178.5 (3)
C3—C4—C5—C6	−0.3 (4)	O3—C7—N1—S1	0.7 (3)
C4—C5—C6—C1	−0.5 (3)	C8—C7—N1—S1	179.90 (15)
C2—C1—C6—C5	0.9 (3)	C7—N1—S1—O1	−50.7 (2)
S1—C1—C6—C5	−177.54 (17)	C7—N1—S1—O2	−177.71 (18)
O3—C7—C8—C9	−156.0 (2)	C7—N1—S1—C1	66.88 (19)
N1—C7—C8—C9	24.8 (3)	C6—C1—S1—O1	7.75 (19)
O3—C7—C8—C13	22.7 (3)	C2—C1—S1—O1	−170.73 (16)
N1—C7—C8—C13	−156.5 (2)	C6—C1—S1—O2	138.74 (16)
C13—C8—C9—C10	−1.4 (4)	C2—C1—S1—O2	−39.75 (19)
C7—C8—C9—C10	177.3 (2)	C6—C1—S1—N1	−110.00 (16)
C8—C9—C10—C11	1.6 (4)	C2—C1—S1—N1	71.51 (18)

Hydrogen-bond geometry (Å, º)

D—H···*A*	*D*—H	H···*A*	*D*···*A*	*D*—H···*A*
N1—H1*N*···O2^{i}	0.82 (2)	2.06 (2)	2.876 (2)	174 (2)

Symmetry code: (i) −*x*+1/2, −*y*+1/2, −*z*+2.