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Chlorido[1-diphenylphosphanyl-3-(phenylsulfanyl)propane- κ^2P,S](η^5 -penta-methylcyclopentadienyl)iridium(III) chloride monohydrate

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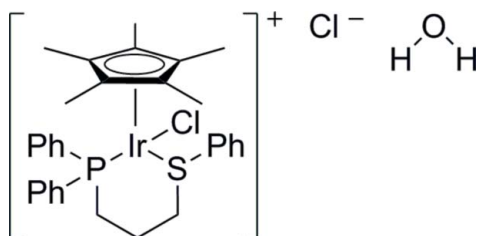
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Key indicators: single-crystal X-ray study; $T = 100$ K; mean $\sigma(C-C) = 0.010$ Å; R factor = 0.041; wR factor = 0.088; data-to-parameter ratio = 16.4.

The crystal structure of the title compound, $[\text{Ir}(\text{C}_{10}\text{H}_{15})\text{Cl}(\text{C}_{21}\text{H}_{21}\text{PS})]\text{Cl}\cdot\text{H}_2\text{O}$, consists of discrete $[\text{Ir}(\eta^5\text{-C}_5\text{Me}_5)\text{Cl}\{\text{Ph}_2\text{P}(\text{CH}_2)_3\text{SPh-}\kappa^2P,S\}]^+$ cations, chloride anions and water molecules. The Ir^{III} atom is coordinated by an $\eta^5\text{-C}_5\text{Me}_5$ ligand, a chloride and a $\text{Ph}_2\text{P}(\text{CH}_2)_3\text{SPh-}\kappa^2P,S$ ligand, leading to a three-legged piano-stool geometry. In the crystal, two water molecules and two chloride anions are linked by weak $\text{O}\cdots\text{H}\cdots\text{Cl}$ hydrogen bonding into tetramers that are located on centers of inversion. The H atoms of one of the methyl groups are disordered and were refined using a split model.

Related literature

Analogous iridium(III) complexes with $\text{Ph}_2\text{PCH}_2\text{SPh}$ ligands have been described by Valderrama *et al.* (1997). For arene ruthenium(II) complexes having ω -diphenylphosphanyl-functionalized alkyl phenyl sulfide ligands $\text{Ph}_2\text{P}(\text{CH}_2)_n\text{SPh}$ ($n = 1\text{--}3$), see: Ludwig *et al.* (2012). For an overview of the strength of hydrogen bonds, see: Steiner (2002).



Experimental

Crystal data

$[\text{Ir}(\text{C}_{10}\text{H}_{15})\text{Cl}(\text{C}_{21}\text{H}_{21}\text{PS})]\text{Cl}\cdot\text{H}_2\text{O}$
 $M_r = 752.74$
 Monoclinic, $P2_1/n$
 $a = 11.0720$ (3) Å
 $b = 8.9617$ (2) Å
 $c = 30.6266$ (7) Å
 $\beta = 100.000$ (2)°

$V = 2992.72$ (13) Å³
 $Z = 4$
 Mo $K\alpha$ radiation
 $\mu = 4.79$ mm⁻¹
 $T = 100$ K
 $0.25 \times 0.20 \times 0.20$ mm

Data collection

Oxford Diffraction Gemini S diffractometer
 Absorption correction: multi-scan (*CrysAlis PRO*; Oxford Diffraction, 2006)
 $T_{\text{min}} = 0.638$, $T_{\text{max}} = 1.000$

27405 measured reflections
 5513 independent reflections
 5098 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.047$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.041$
 $wR(F^2) = 0.088$
 $S = 1.25$
 5513 reflections
 337 parameters
 147 restraints

H atoms treated by a mixture of independent and constrained refinement
 $\Delta\rho_{\text{max}} = 1.68$ e Å⁻³
 $\Delta\rho_{\text{min}} = -1.50$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D\cdots H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{O1-H1O}\cdots\text{Cl2}$	0.95 (6)	2.27 (6)	3.216 (6)	177 (10)
$\text{O1-H2O}\cdots\text{Cl2}^i$	0.96 (8)	2.26 (11)	3.196 (6)	165 (13)

Symmetry code: (i) $-x + 1, -y + 2, -z$.

Data collection: *CrysAlis PRO* (Oxford Diffraction, 2006); cell refinement: *CrysAlis PRO*; data reduction: *CrysAlis PRO*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *DIAMOND* (Brandenburg, 2001); software used to prepare material for publication: *SHELXL97*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: NC2278).

References

- Brandenburg, K. (2001). *DIAMOND*. Crystal Impact GbR, Bonn, Germany.
 Ludwig, G., Kaluerović, G. N., Bette, M., Block, M., Paschke, R. & Steinborn, D. (2012). *J. Inorg. Biochem.* In the press.
 Oxford Diffraction (2006). *CrysAlis PRO*. Oxford Diffraction Ltd, Abingdon, England.
 Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
 Steiner, T. (2002). *Angew. Chem.* **114**, 50–80.
 Valderrama, M., Contreras, R. & Boys, D. (1997). *Polyhedron*, **16**, 1811–1817.

supporting information

Acta Cryst. (2012). E68, m858 [doi:10.1107/S1600536812021964]

Chlorido[1-diphenylphosphanyl-3-(phenylsulfanyl)propane- κ^2P,S](η^5 -penta-methylcyclopentadienyl)iridium(III) chloride monohydrate

Gerd Ludwig, Marcus Korb, Tobias Ruffer, Heinrich Lang and Dirk Steinborn

S1. Comment

Reaction of the ω -diphenylphosphanyl-functionalized alkyl phenyl sulfide $\text{Ph}_2\text{P}(\text{CH}_2)_3\text{SPh}$ with the dinuclear chlorido-bridged iridium complex $[\{\text{Ir}(\eta^5\text{-C}_5\text{Me}_5)\text{Cl}_2\}_2]$ afforded the title compound $[\text{Ir}(\eta^5\text{-C}_5\text{Me}_5)\text{Cl}\{\text{Ph}_2\text{P}(\text{CH}_2)_3\text{SPh-}\kappa P,\kappa S\}]\text{Cl}\cdot\text{H}_2\text{O}$. Similar iridium(III) and ruthenium(II) complexes with ω -diphenylphosphanyl-functionalized alkyl phenyl sulfide ligands $\text{Ph}_2\text{P}(\text{CH}_2)_n\text{SPh}$ ($n = 1\text{--}3$) have been reported by Valderrama (1997) and Ludwig (2012). Crystals of the title compound are built up from $[\text{Ir}(\eta^5\text{-C}_5\text{Me}_5)\text{Cl}\{\text{Ph}_2\text{P}(\text{CH}_2)_3\text{SPh-}\kappa P,\kappa S\}]^+$ cations, chloride anions and water molecules. The iridium atom is slightly distorted octahedrally coordinated by a $\eta^5\text{-C}_5\text{Me}_5$, a chlorido, and a chelating $\text{Ph}_2\text{P}(\text{CH}_2)_3\text{SPh-}\kappa P,\kappa S$ ligand, which leads to a three-legged piano-stool geometry of the iridium center. The Ir—Cl (2.434 (2) Å), Ir—P (2.312 (2) Å), and Ir—S bond lengths (2.350 (2) Å) are in the expected range. The six-membered IrPC_3S iridacycle adopts a chair conformation. In crystals short contacts between the oxygen atoms of the water molecules and the chloride anions ($\text{O1}\cdots\text{Cl2}$ 3.216 (5) Å; $\text{O1}\cdots\text{Cl2}(1-x, 2-y, -z)$ 3.196 (6) Å) indicate according to Steiner (2002) weak O—H \cdots Cl hydrogen bonds ($\text{H10}\cdots\text{Cl2}$ 2.27 Å; $\text{O1—H10}\cdots\text{Cl2}$ 177°; $\text{H2O}\cdots\text{Cl2}(1-x, 2-y, -z)$ 2.26 Å; $\text{O1—H2O}\cdots\text{Cl2}$ 165°).

S2. Experimental

To a methanol solution (30 ml) of $[\{\text{IrCl}_2(\eta^5\text{-C}_5\text{Me}_5)\}_2]$ (0.10 g, 0.16 mmol) $\text{Ph}_2\text{P}(\text{CH}_2)_3\text{SPh}$ (0.32 mmol) was added with stirring and then the solution was heated under reflux for three hours. The reaction mixture was cooled down to room temperature, water (2 ml) was added and after storage in a freezer at -70°C overnight the precipitate obtained was filtered off, washed with ether (3×2 ml), and dried in vacuum. Crystals of the title compound suitable for X-ray diffraction analysis were obtained from acetone solution at room temperature.

S3. Refinement

The H atoms, except those from the water molecules, were placed in calculated positions and constrained to ride on their parent atoms. Hydrogen atoms of the water molecules were found in the difference Fourier map and refined using the restraints "DFIX 0.960 0.020 H2O O1" and "DFIX 0.960 0.020 O1 H10". To modulate the anisotropic displacement parameters of the ellipsoids from the pentamethylcyclopentadienyl ligand the restraints "DELU 0.010 0.010 C1 C2 C3 C4 C5 C6 C7 C8 C9 C10", "SIMU 0.040 0.080 1.700 C1 C2 C3 C4 C5 C6 C7 C8 C9 C10" and "ISOR 0.100 C1 C2 C3 C4 C5 C6 C7 C8 C9 C10" were used. The H atoms on C6 have been refined disordered over two equally occupied positions.

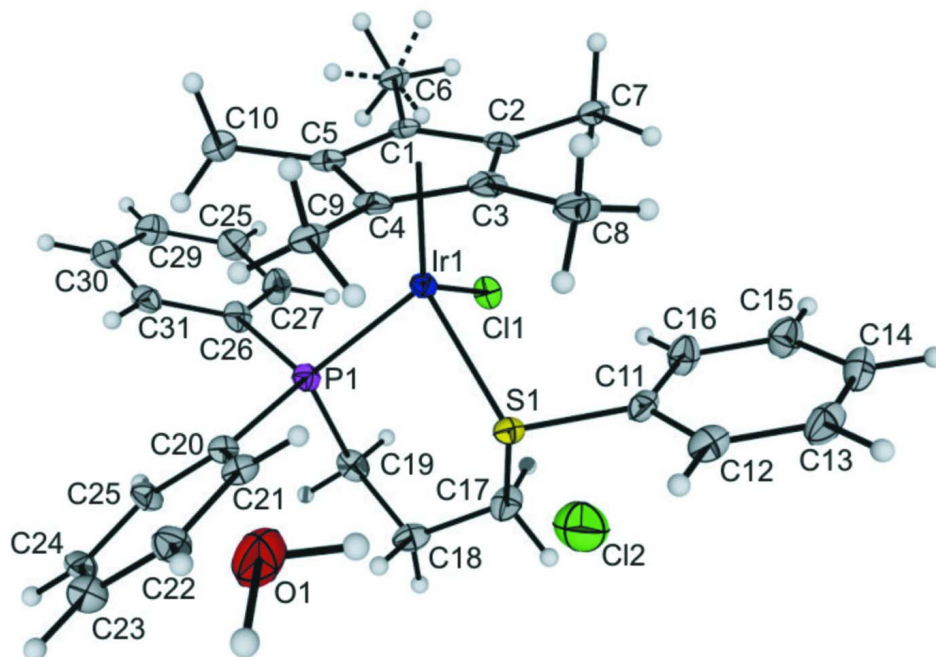


Figure 1

Structure of title compound $[\text{Ir}(\eta^5\text{-C}_5\text{Me}_5)\text{Cl}\{\text{Ph}_2\text{P}(\text{CH}_2)_3\text{SPh-}\kappa^2\text{P,}\kappa\text{S}\}]\text{Cl}\cdot\text{H}_2\text{O}$. Displacement ellipsoids are drawn at the 50% probability level and the H atoms are shown as small spheres of arbitrary radii. Disordering is shown as full and dashed bonds.

Chlorido[1-diphenylphosphanyl-3-(phenylsulfanyl)propane- $\kappa^2\text{P,S}$](η^5 -pentamethylcyclopentadienyl)iridium(III) chloride monohydrate

Crystal data

$[\text{Ir}(\text{C}_{10}\text{H}_{15})\text{Cl}(\text{C}_{21}\text{H}_{21}\text{PS})]\text{Cl}\cdot\text{H}_2\text{O}$

$M_r = 752.74$

Monoclinic, $P2_1/n$

Hall symbol: $-P\ 2_1n$

$a = 11.0720\ (3)\ \text{\AA}$

$b = 8.9617\ (2)\ \text{\AA}$

$c = 30.6266\ (7)\ \text{\AA}$

$\beta = 100.000\ (2)^\circ$

$V = 2992.72\ (13)\ \text{\AA}^3$

$Z = 4$

$F(000) = 1496$

$D_x = 1.671\ \text{Mg m}^{-3}$

Mo $K\alpha$ radiation, $\lambda = 0.71073\ \text{\AA}$

Cell parameters from 14059 reflections

$\theta = 2.9\text{--}28.1^\circ$

$\mu = 4.79\ \text{mm}^{-1}$

$T = 100\ \text{K}$

Block, yellow

$0.25 \times 0.2 \times 0.2\ \text{mm}$

Data collection

Oxford Diffraction Gemini S
diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

ω scans

Absorption correction: multi-scan

(*CrysAlis PRO*; Oxford Diffraction, 2006)

$T_{\min} = 0.638$, $T_{\max} = 1.000$

27405 measured reflections

5513 independent reflections

5098 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.047$

$\theta_{\max} = 25.5^\circ$, $\theta_{\min} = 2.9^\circ$

$h = -13 \rightarrow 13$

$k = -10 \rightarrow 10$

$l = -35 \rightarrow 37$

2 standard reflections every 50 reflections

intensity decay: none

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.041$	H atoms treated by a mixture of independent and constrained refinement
$wR(F^2) = 0.088$	$w = 1/[\sigma^2(F_o^2) + (0.P)^2 + 29.4455P]$
$S = 1.25$	where $P = (F_o^2 + 2F_c^2)/3$
5513 reflections	$(\Delta/\sigma)_{\max} = 0.001$
337 parameters	$\Delta\rho_{\max} = 1.68 \text{ e } \text{\AA}^{-3}$
147 restraints	$\Delta\rho_{\min} = -1.50 \text{ e } \text{\AA}^{-3}$
Primary atom site location: structure-invariant direct methods	

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) etc. and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$	Occ. (<1)
C1	0.4071 (6)	0.1087 (7)	0.1078 (2)	0.0180 (13)	
C2	0.5278 (6)	0.1774 (8)	0.1208 (2)	0.0196 (13)	
C3	0.5326 (6)	0.3021 (7)	0.0928 (2)	0.0183 (13)	
C4	0.4139 (6)	0.3191 (7)	0.0648 (2)	0.0175 (13)	
C5	0.3395 (6)	0.1940 (7)	0.0729 (2)	0.0192 (13)	
C6	0.3704 (6)	-0.0502 (7)	0.1281 (2)	0.0191 (8)	
H6A	0.4355	-0.0819	0.1512	0.029*	0.50
H6B	0.3577	-0.1239	0.1051	0.029*	0.50
H6C	0.2964	-0.0382	0.1401	0.029*	0.50
H6D	0.2909	-0.0808	0.1131	0.029*	0.50
H6E	0.3687	-0.0388	0.1592	0.029*	0.50
H6F	0.4300	-0.1245	0.1241	0.029*	0.50
C7	0.6279 (6)	0.1159 (7)	0.1545 (2)	0.0191 (8)	
H7A	0.6561	0.0236	0.1440	0.029*	
H7B	0.5977	0.0986	0.1816	0.029*	
H7C	0.6945	0.1859	0.1598	0.029*	
C8	0.6406 (6)	0.3992 (8)	0.0898 (3)	0.0263 (16)	
H8A	0.7006	0.3884	0.1163	0.039*	
H8B	0.6146	0.5014	0.0867	0.039*	
H8C	0.6758	0.3702	0.0646	0.039*	
C9	0.3851 (6)	0.4222 (7)	0.0260 (2)	0.0191 (8)	
H9A	0.4098	0.3770	0.0005	0.029*	
H9B	0.4284	0.5144	0.0327	0.029*	
H9C	0.2985	0.4413	0.0199	0.029*	

C10	0.2196 (6)	0.1528 (8)	0.0457 (2)	0.0249 (15)
H10A	0.2331	0.1003	0.0197	0.037*
H10B	0.1729	0.2416	0.0371	0.037*
H10C	0.1752	0.0900	0.0628	0.037*
C11	0.6084 (6)	0.5606 (7)	0.1848 (2)	0.0197 (14)
C12	0.6887 (6)	0.6611 (8)	0.1706 (2)	0.0238 (15)
H12	0.6604	0.7311	0.1488	0.029*
C13	0.8121 (6)	0.6555 (9)	0.1895 (2)	0.0273 (16)
H13	0.8668	0.7219	0.1801	0.033*
C14	0.8542 (6)	0.5532 (8)	0.2220 (2)	0.0253 (16)
H14	0.9367	0.5517	0.2348	0.030*
C15	0.7734 (7)	0.4523 (8)	0.2356 (2)	0.0275 (17)
H15	0.8021	0.3819	0.2572	0.033*
C16	0.6500 (6)	0.4559 (8)	0.2171 (2)	0.0247 (16)
H16	0.5957	0.3886	0.2263	0.030*
C17	0.3816 (6)	0.6217 (8)	0.2063 (2)	0.0219 (15)
H17A	0.3965	0.5426	0.2282	0.026*
H17B	0.4207	0.7115	0.2197	0.026*
C18	0.2441 (6)	0.6484 (8)	0.1944 (2)	0.0223 (15)
H18A	0.2282	0.7100	0.1680	0.027*
H18B	0.2177	0.7039	0.2182	0.027*
C19	0.1674 (6)	0.5070 (8)	0.1861 (2)	0.0215 (14)
H19A	0.0819	0.5325	0.1853	0.026*
H19B	0.1914	0.4392	0.2107	0.026*
C20	0.1165 (6)	0.5512 (7)	0.0942 (2)	0.0190 (14)
C21	0.1912 (6)	0.6303 (7)	0.0707 (2)	0.0200 (15)
H21	0.2737	0.6051	0.0737	0.024*
C22	0.1455 (7)	0.7462 (8)	0.0428 (2)	0.0243 (16)
H22	0.1964	0.7978	0.0268	0.029*
C23	0.0225 (7)	0.7842 (8)	0.0389 (2)	0.0269 (17)
H23	-0.0089	0.8620	0.0203	0.032*
C24	-0.0543 (6)	0.7077 (7)	0.0626 (3)	0.0256 (17)
H24	-0.1363	0.7347	0.0600	0.031*
C25	-0.0078 (6)	0.5913 (8)	0.0899 (2)	0.0209 (15)
H25	-0.0590	0.5391	0.1056	0.025*
C26	0.0716 (6)	0.2595 (8)	0.1303 (2)	0.0201 (14)
C27	0.0810 (6)	0.1559 (9)	0.1647 (2)	0.0269 (16)
H27	0.1415	0.1680	0.1896	0.032*
C28	0.0021 (7)	0.0352 (8)	0.1624 (3)	0.0300 (18)
H28	0.0100	-0.0335	0.1854	0.036*
C29	-0.0886 (6)	0.0174 (8)	0.1256 (3)	0.0279 (17)
H29	-0.1429	-0.0624	0.1243	0.034*
C30	-0.0991 (6)	0.1177 (8)	0.0907 (3)	0.0256 (16)
H30	-0.1592	0.1042	0.0658	0.031*
C31	-0.0201 (6)	0.2374 (8)	0.0931 (2)	0.0223 (15)
H31	-0.0278	0.3047	0.0697	0.027*
P1	0.18229 (15)	0.41058 (19)	0.13450 (6)	0.0167 (4)
S1	0.45073 (15)	0.57108 (18)	0.15886 (5)	0.0171 (3)

Cl1	0.38749 (14)	0.24277 (19)	0.20829 (5)	0.0202 (3)
Ir1	0.38099 (2)	0.33401 (3)	0.133078 (8)	0.01399 (8)
O1	0.3710 (5)	0.8621 (6)	-0.00804 (19)	0.0332 (13)
Cl2	0.61713 (16)	0.8410 (2)	0.06493 (6)	0.0287 (4)
H1O	0.442 (5)	0.853 (10)	0.014 (2)	0.05 (3)*
H2O	0.376 (14)	0.960 (7)	-0.020 (5)	0.15 (6)*

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.020 (3)	0.012 (3)	0.022 (4)	0.002 (2)	0.004 (3)	-0.008 (2)
C2	0.022 (3)	0.016 (3)	0.021 (3)	0.007 (3)	0.007 (2)	-0.005 (3)
C3	0.018 (3)	0.019 (3)	0.021 (3)	0.000 (2)	0.013 (2)	-0.006 (2)
C4	0.023 (3)	0.015 (3)	0.016 (3)	0.005 (3)	0.008 (2)	-0.004 (2)
C5	0.023 (3)	0.015 (3)	0.020 (3)	0.003 (3)	0.006 (3)	-0.007 (2)
C6	0.021 (2)	0.0149 (18)	0.023 (2)	-0.0029 (15)	0.0073 (16)	-0.0067 (15)
C7	0.021 (2)	0.0149 (18)	0.023 (2)	-0.0029 (15)	0.0073 (16)	-0.0067 (15)
C8	0.025 (4)	0.016 (3)	0.041 (5)	-0.004 (3)	0.016 (3)	-0.007 (3)
C9	0.021 (2)	0.0149 (18)	0.023 (2)	-0.0029 (15)	0.0073 (16)	-0.0067 (15)
C10	0.024 (3)	0.025 (4)	0.024 (4)	0.003 (3)	0.001 (3)	-0.008 (3)
C11	0.021 (4)	0.019 (3)	0.018 (4)	-0.004 (3)	0.002 (3)	-0.006 (3)
C12	0.028 (4)	0.020 (3)	0.024 (4)	-0.004 (3)	0.006 (3)	-0.004 (3)
C13	0.026 (4)	0.027 (4)	0.030 (4)	-0.012 (3)	0.010 (3)	-0.006 (3)
C14	0.019 (4)	0.035 (4)	0.021 (4)	-0.008 (3)	0.002 (3)	-0.005 (3)
C15	0.027 (4)	0.029 (4)	0.025 (4)	-0.003 (3)	-0.001 (3)	0.002 (3)
C16	0.023 (4)	0.025 (4)	0.026 (4)	-0.006 (3)	0.002 (3)	0.005 (3)
C17	0.022 (4)	0.031 (4)	0.014 (4)	-0.003 (3)	0.007 (3)	-0.008 (3)
C18	0.028 (4)	0.021 (4)	0.020 (4)	0.000 (3)	0.011 (3)	-0.005 (3)
C19	0.022 (3)	0.023 (4)	0.020 (4)	0.003 (3)	0.007 (3)	0.005 (3)
C20	0.021 (3)	0.015 (3)	0.020 (4)	0.002 (3)	0.000 (3)	-0.002 (3)
C21	0.023 (4)	0.015 (3)	0.022 (4)	-0.001 (3)	0.003 (3)	-0.004 (3)
C22	0.030 (4)	0.019 (4)	0.023 (4)	0.002 (3)	0.002 (3)	0.002 (3)
C23	0.028 (4)	0.021 (4)	0.030 (4)	0.003 (3)	0.000 (3)	0.000 (3)
C24	0.016 (3)	0.018 (4)	0.040 (5)	0.006 (3)	-0.003 (3)	-0.005 (3)
C25	0.016 (3)	0.023 (4)	0.024 (4)	0.008 (3)	0.004 (3)	-0.003 (3)
C26	0.013 (3)	0.024 (4)	0.025 (4)	-0.001 (3)	0.007 (3)	0.000 (3)
C27	0.021 (4)	0.029 (4)	0.030 (4)	-0.006 (3)	0.004 (3)	0.005 (3)
C28	0.027 (4)	0.022 (4)	0.041 (5)	-0.005 (3)	0.006 (3)	0.005 (3)
C29	0.022 (4)	0.024 (4)	0.039 (5)	-0.006 (3)	0.010 (3)	-0.002 (3)
C30	0.017 (3)	0.026 (4)	0.033 (4)	0.000 (3)	0.004 (3)	-0.008 (3)
C31	0.017 (3)	0.022 (4)	0.028 (4)	0.004 (3)	0.005 (3)	-0.003 (3)
P1	0.0149 (8)	0.0171 (8)	0.0187 (9)	0.0000 (7)	0.0049 (7)	0.0000 (7)
S1	0.0191 (8)	0.0135 (8)	0.0187 (8)	-0.0004 (6)	0.0035 (6)	-0.0015 (6)
Cl1	0.0190 (8)	0.0230 (8)	0.0187 (9)	-0.0003 (7)	0.0037 (6)	0.0050 (7)
Ir1	0.01420 (13)	0.01298 (13)	0.01507 (13)	0.00024 (10)	0.00329 (9)	0.00059 (10)
O1	0.029 (3)	0.031 (3)	0.038 (3)	-0.009 (2)	0.001 (2)	-0.002 (3)
Cl2	0.0347 (10)	0.0268 (9)	0.0233 (9)	0.0056 (8)	0.0014 (7)	-0.0008 (8)

Geometric parameters (Å, °)

C1—C5	1.419 (9)	C15—C16	1.386 (10)
C1—C2	1.462 (9)	C15—H15	0.9300
C1—C6	1.634 (9)	C16—H16	0.9300
C1—Irl	2.199 (6)	C17—C18	1.521 (9)
C2—C3	1.414 (9)	C17—S1	1.815 (7)
C2—C7	1.484 (9)	C17—H17A	0.9700
C2—Irl	2.228 (6)	C17—H17B	0.9700
C3—C4	1.448 (9)	C18—C19	1.523 (9)
C3—C8	1.495 (9)	C18—H18A	0.9700
C3—Irl	2.266 (6)	C18—H18B	0.9700
C4—C5	1.438 (9)	C19—P1	1.834 (7)
C4—C9	1.494 (9)	C19—H19A	0.9700
C4—Irl	2.189 (6)	C19—H19B	0.9700
C5—C10	1.486 (9)	C20—C21	1.383 (10)
C5—Irl	2.212 (6)	C20—C25	1.407 (9)
C6—H6A	0.9600	C20—P1	1.824 (7)
C6—H6B	0.9600	C21—C22	1.384 (9)
C6—H6C	0.9600	C21—H21	0.9300
C6—H6D	0.9600	C22—C23	1.389 (10)
C6—H6E	0.9600	C22—H22	0.9300
C6—H6F	0.9600	C23—C24	1.391 (10)
C7—H7A	0.9600	C23—H23	0.9300
C7—H7B	0.9600	C24—C25	1.380 (10)
C7—H7C	0.9600	C24—H24	0.9300
C8—H8A	0.9600	C25—H25	0.9300
C8—H8B	0.9600	C26—C27	1.393 (10)
C8—H8C	0.9600	C26—C31	1.403 (10)
C9—H9A	0.9600	C26—P1	1.816 (7)
C9—H9B	0.9600	C27—C28	1.385 (10)
C9—H9C	0.9600	C27—H27	0.9300
C10—H10A	0.9600	C28—C29	1.381 (11)
C10—H10B	0.9600	C28—H28	0.9300
C10—H10C	0.9600	C29—C30	1.385 (11)
C11—C16	1.383 (10)	C29—H29	0.9300
C11—C12	1.388 (10)	C30—C31	1.378 (10)
C11—S1	1.790 (7)	C30—H30	0.9300
C12—C13	1.390 (10)	C31—H31	0.9300
C12—H12	0.9300	P1—Irl	2.3121 (17)
C13—C14	1.372 (11)	S1—Irl	2.3498 (16)
C13—H13	0.9300	Cl1—Irl	2.4337 (16)
C14—C15	1.387 (10)	O1—H1O	0.95 (2)
C14—H14	0.9300	O1—H2O	0.96 (2)
C5—C1—C2	108.4 (6)	C18—C17—S1	113.0 (5)
C5—C1—C6	128.6 (6)	C18—C17—H17A	109.0
C2—C1—C6	122.7 (6)	S1—C17—H17A	109.0

C5—C1—Ir1	71.7 (4)	C18—C17—H17B	109.0
C2—C1—Ir1	71.8 (3)	S1—C17—H17B	109.0
C6—C1—Ir1	127.6 (4)	H17A—C17—H17B	107.8
C3—C2—C1	107.3 (6)	C17—C18—C19	114.5 (6)
C3—C2—C7	127.8 (6)	C17—C18—H18A	108.6
C1—C2—C7	124.7 (6)	C19—C18—H18A	108.6
C3—C2—Ir1	73.1 (4)	C17—C18—H18B	108.6
C1—C2—Ir1	69.6 (3)	C19—C18—H18B	108.6
C7—C2—Ir1	126.5 (5)	H18A—C18—H18B	107.6
C2—C3—C4	108.5 (6)	C18—C19—P1	113.9 (5)
C2—C3—C8	127.7 (6)	C18—C19—H19A	108.8
C4—C3—C8	123.7 (6)	P1—C19—H19A	108.8
C2—C3—Ir1	70.2 (4)	C18—C19—H19B	108.8
C4—C3—Ir1	68.2 (3)	P1—C19—H19B	108.8
C8—C3—Ir1	129.5 (5)	H19A—C19—H19B	107.7
C5—C4—C3	107.6 (6)	C21—C20—C25	118.9 (6)
C5—C4—C9	124.7 (6)	C21—C20—P1	120.1 (5)
C3—C4—C9	125.7 (6)	C25—C20—P1	120.6 (5)
C5—C4—Ir1	71.8 (4)	C20—C21—C22	121.3 (6)
C3—C4—Ir1	73.9 (4)	C20—C21—H21	119.3
C9—C4—Ir1	132.3 (4)	C22—C21—H21	119.3
C1—C5—C4	107.9 (6)	C21—C22—C23	118.9 (7)
C1—C5—C10	126.3 (6)	C21—C22—H22	120.5
C4—C5—C10	125.5 (6)	C23—C22—H22	120.5
C1—C5—Ir1	70.7 (4)	C22—C23—C24	121.0 (7)
C4—C5—Ir1	70.1 (4)	C22—C23—H23	119.5
C10—C5—Ir1	130.0 (5)	C24—C23—H23	119.5
C1—C6—H6A	109.5	C25—C24—C23	119.4 (6)
C1—C6—H6B	109.5	C25—C24—H24	120.3
H6A—C6—H6B	109.5	C23—C24—H24	120.3
C1—C6—H6C	109.5	C24—C25—C20	120.4 (7)
H6A—C6—H6C	109.5	C24—C25—H25	119.8
H6B—C6—H6C	109.5	C20—C25—H25	119.8
C1—C6—H6D	109.5	C27—C26—C31	117.9 (6)
H6A—C6—H6D	141.1	C27—C26—P1	118.9 (5)
H6B—C6—H6D	56.3	C31—C26—P1	123.2 (5)
H6C—C6—H6D	56.3	C28—C27—C26	121.2 (7)
C1—C6—H6E	109.5	C28—C27—H27	119.4
H6A—C6—H6E	56.3	C26—C27—H27	119.4
H6B—C6—H6E	141.1	C29—C28—C27	119.7 (7)
H6C—C6—H6E	56.3	C29—C28—H28	120.2
H6D—C6—H6E	109.5	C27—C28—H28	120.2
C1—C6—H6F	109.5	C28—C29—C30	120.3 (7)
H6A—C6—H6F	56.3	C28—C29—H29	119.8
H6B—C6—H6F	56.3	C30—C29—H29	119.8
H6C—C6—H6F	141.1	C31—C30—C29	119.8 (7)
H6D—C6—H6F	109.5	C31—C30—H30	120.1
H6E—C6—H6F	109.5	C29—C30—H30	120.1

C2—C7—H7A	109.5	C30—C31—C26	121.1 (7)
C2—C7—H7B	109.5	C30—C31—H31	119.5
H7A—C7—H7B	109.5	C26—C31—H31	119.5
C2—C7—H7C	109.5	C26—P1—C20	106.4 (3)
H7A—C7—H7C	109.5	C26—P1—C19	104.6 (3)
H7B—C7—H7C	109.5	C20—P1—C19	99.9 (3)
C3—C8—H8A	109.5	C26—P1—Ir1	114.2 (2)
C3—C8—H8B	109.5	C20—P1—Ir1	117.5 (2)
H8A—C8—H8B	109.5	C19—P1—Ir1	112.6 (2)
C3—C8—H8C	109.5	C11—S1—C17	99.9 (3)
H8A—C8—H8C	109.5	C11—S1—Ir1	109.9 (2)
H8B—C8—H8C	109.5	C17—S1—Ir1	109.6 (2)
C4—C9—H9A	109.5	C4—Ir1—C1	63.5 (2)
C4—C9—H9B	109.5	C4—Ir1—C5	38.1 (2)
H9A—C9—H9B	109.5	C1—Ir1—C5	37.5 (2)
C4—C9—H9C	109.5	C4—Ir1—C2	63.4 (2)
H9A—C9—H9C	109.5	C1—Ir1—C2	38.5 (2)
H9B—C9—H9C	109.5	C5—Ir1—C2	63.5 (2)
C5—C10—H10A	109.5	C4—Ir1—C3	37.9 (2)
C5—C10—H10B	109.5	C1—Ir1—C3	62.5 (2)
H10A—C10—H10B	109.5	C5—Ir1—C3	62.7 (2)
C5—C10—H10C	109.5	C2—Ir1—C3	36.7 (2)
H10A—C10—H10C	109.5	C4—Ir1—P1	110.65 (17)
H10B—C10—H10C	109.5	C1—Ir1—P1	117.57 (18)
C16—C11—C12	120.8 (6)	C5—Ir1—P1	97.09 (17)
C16—C11—S1	122.1 (5)	C2—Ir1—P1	156.11 (18)
C12—C11—S1	117.1 (5)	C3—Ir1—P1	147.63 (18)
C11—C12—C13	118.9 (7)	C4—Ir1—S1	106.07 (18)
C11—C12—H12	120.6	C1—Ir1—S1	153.20 (18)
C13—C12—H12	120.6	C5—Ir1—S1	143.28 (18)
C14—C13—C12	120.8 (7)	C2—Ir1—S1	114.81 (18)
C14—C13—H13	119.6	C3—Ir1—S1	93.77 (17)
C12—C13—H13	119.6	P1—Ir1—S1	89.06 (6)
C13—C14—C15	119.8 (7)	C4—Ir1—Cl1	154.38 (17)
C13—C14—H14	120.1	C1—Ir1—Cl1	92.31 (19)
C15—C14—H14	120.1	C5—Ir1—Cl1	124.50 (18)
C16—C15—C14	120.2 (7)	C2—Ir1—Cl1	92.63 (18)
C16—C15—H15	119.9	C3—Ir1—Cl1	124.70 (18)
C14—C15—H15	119.9	P1—Ir1—Cl1	87.35 (6)
C11—C16—C15	119.5 (7)	S1—Ir1—Cl1	91.82 (6)
C11—C16—H16	120.3	H1O—O1—H2O	105 (10)
C15—C16—H16	120.3		
C5—C1—C2—C3	-1.1 (7)	C2—C1—Ir1—C5	117.3 (6)
C6—C1—C2—C3	172.6 (5)	C6—C1—Ir1—C5	-125.1 (7)
Ir1—C1—C2—C3	-63.9 (4)	C5—C1—Ir1—C2	-117.3 (6)
C5—C1—C2—C7	-176.3 (6)	C6—C1—Ir1—C2	117.7 (7)
C6—C1—C2—C7	-2.6 (10)	C5—C1—Ir1—C3	-80.2 (4)

Ir1—C1—C2—C7	120.9 (6)	C2—C1—Ir1—C3	37.1 (4)
C5—C1—C2—Ir1	62.8 (4)	C6—C1—Ir1—C3	154.8 (6)
C6—C1—C2—Ir1	-123.5 (6)	C5—C1—Ir1—P1	63.2 (4)
C1—C2—C3—C4	4.1 (7)	C2—C1—Ir1—P1	-179.5 (3)
C7—C2—C3—C4	179.1 (6)	C6—C1—Ir1—P1	-61.9 (6)
Ir1—C2—C3—C4	-57.5 (4)	C5—C1—Ir1—S1	-109.9 (5)
C1—C2—C3—C8	-173.2 (6)	C2—C1—Ir1—S1	7.3 (7)
C7—C2—C3—C8	1.8 (11)	C6—C1—Ir1—S1	125.0 (5)
Ir1—C2—C3—C8	125.2 (7)	C5—C1—Ir1—C11	151.4 (4)
C1—C2—C3—Ir1	61.6 (4)	C2—C1—Ir1—C11	-91.3 (4)
C7—C2—C3—Ir1	-123.4 (7)	C6—C1—Ir1—C11	26.4 (5)
C2—C3—C4—C5	-5.5 (7)	C1—C5—Ir1—C4	118.2 (5)
C8—C3—C4—C5	171.9 (6)	C10—C5—Ir1—C4	-120.1 (8)
Ir1—C3—C4—C5	-64.3 (4)	C4—C5—Ir1—C1	-118.2 (5)
C2—C3—C4—C9	-170.2 (6)	C10—C5—Ir1—C1	121.7 (8)
C8—C3—C4—C9	7.2 (10)	C1—C5—Ir1—C2	38.2 (4)
Ir1—C3—C4—C9	131.0 (6)	C4—C5—Ir1—C2	-80.0 (4)
C2—C3—C4—Ir1	58.8 (4)	C10—C5—Ir1—C2	159.9 (7)
C8—C3—C4—Ir1	-123.8 (6)	C1—C5—Ir1—C3	79.5 (4)
C2—C1—C5—C4	-2.3 (7)	C4—C5—Ir1—C3	-38.7 (4)
C6—C1—C5—C4	-175.6 (6)	C10—C5—Ir1—C3	-158.8 (7)
Ir1—C1—C5—C4	60.5 (4)	C1—C5—Ir1—P1	-127.1 (4)
C2—C1—C5—C10	171.1 (6)	C4—C5—Ir1—P1	114.6 (3)
C6—C1—C5—C10	-2.2 (11)	C10—C5—Ir1—P1	-5.5 (6)
Ir1—C1—C5—C10	-126.1 (7)	C1—C5—Ir1—S1	134.9 (3)
C2—C1—C5—Ir1	-62.9 (4)	C4—C5—Ir1—S1	16.6 (5)
C6—C1—C5—Ir1	123.9 (6)	C10—C5—Ir1—S1	-103.5 (6)
C3—C4—C5—C1	4.8 (7)	C1—C5—Ir1—C11	-35.5 (4)
C9—C4—C5—C1	169.7 (6)	C4—C5—Ir1—C11	-153.7 (3)
Ir1—C4—C5—C1	-60.9 (4)	C10—C5—Ir1—C11	86.2 (6)
C3—C4—C5—C10	-168.6 (6)	C3—C2—Ir1—C4	36.3 (4)
C9—C4—C5—C10	-3.8 (10)	C1—C2—Ir1—C4	-80.1 (4)
Ir1—C4—C5—C10	125.6 (6)	C7—C2—Ir1—C4	161.2 (7)
C3—C4—C5—Ir1	65.8 (4)	C3—C2—Ir1—C1	116.4 (6)
C9—C4—C5—Ir1	-129.4 (6)	C7—C2—Ir1—C1	-118.7 (8)
C16—C11—C12—C13	-0.3 (10)	C3—C2—Ir1—C5	79.1 (4)
S1—C11—C12—C13	-179.3 (5)	C1—C2—Ir1—C5	-37.2 (4)
C11—C12—C13—C14	-0.5 (11)	C7—C2—Ir1—C5	-156.0 (7)
C12—C13—C14—C15	1.1 (11)	C1—C2—Ir1—C3	-116.4 (6)
C13—C14—C15—C16	-1.0 (11)	C7—C2—Ir1—C3	124.9 (8)
C12—C11—C16—C15	0.4 (11)	C3—C2—Ir1—P1	117.4 (5)
S1—C11—C16—C15	179.3 (6)	C1—C2—Ir1—P1	1.0 (7)
C14—C15—C16—C11	0.3 (11)	C7—C2—Ir1—P1	-117.8 (5)
S1—C17—C18—C19	76.0 (7)	C3—C2—Ir1—S1	-60.0 (4)
C17—C18—C19—P1	-71.0 (7)	C1—C2—Ir1—S1	-176.4 (3)
C25—C20—C21—C22	-0.8 (10)	C7—C2—Ir1—S1	64.9 (6)
P1—C20—C21—C22	-174.4 (5)	C3—C2—Ir1—C11	-153.2 (4)
C20—C21—C22—C23	0.8 (10)	C1—C2—Ir1—C11	90.4 (4)

C21—C22—C23—C24	-0.1 (11)	C7—C2—Ir1—C11	-28.3 (6)
C22—C23—C24—C25	-0.6 (11)	C2—C3—Ir1—C4	-120.5 (5)
C23—C24—C25—C20	0.6 (10)	C8—C3—Ir1—C4	116.4 (8)
C21—C20—C25—C24	0.1 (10)	C2—C3—Ir1—C1	-39.0 (4)
P1—C20—C25—C24	173.7 (5)	C4—C3—Ir1—C1	81.4 (4)
C31—C26—C27—C28	-0.3 (11)	C8—C3—Ir1—C1	-162.1 (7)
P1—C26—C27—C28	-177.4 (6)	C2—C3—Ir1—C5	-81.5 (4)
C26—C27—C28—C29	-0.6 (12)	C4—C3—Ir1—C5	38.9 (4)
C27—C28—C29—C30	1.4 (11)	C8—C3—Ir1—C5	155.4 (7)
C28—C29—C30—C31	-1.3 (11)	C4—C3—Ir1—C2	120.5 (5)
C29—C30—C31—C26	0.3 (10)	C8—C3—Ir1—C2	-123.1 (8)
C27—C26—C31—C30	0.4 (10)	C2—C3—Ir1—P1	-137.8 (4)
P1—C26—C31—C30	177.4 (5)	C4—C3—Ir1—P1	-17.3 (5)
C27—C26—P1—C20	-163.0 (6)	C8—C3—Ir1—P1	99.1 (7)
C31—C26—P1—C20	20.1 (7)	C2—C3—Ir1—S1	128.0 (4)
C27—C26—P1—C19	-57.8 (6)	C4—C3—Ir1—S1	-111.5 (4)
C31—C26—P1—C19	125.3 (6)	C8—C3—Ir1—S1	4.9 (6)
C27—C26—P1—Ir1	65.7 (6)	C2—C3—Ir1—C11	33.2 (4)
C31—C26—P1—Ir1	-111.2 (6)	C4—C3—Ir1—C11	153.6 (3)
C21—C20—P1—C26	-142.8 (5)	C8—C3—Ir1—C11	-89.9 (6)
C25—C20—P1—C26	43.7 (6)	C26—P1—Ir1—C4	87.6 (3)
C21—C20—P1—C19	108.7 (6)	C20—P1—Ir1—C4	-38.0 (3)
C25—C20—P1—C19	-64.8 (6)	C19—P1—Ir1—C4	-153.3 (3)
C21—C20—P1—Ir1	-13.4 (6)	C26—P1—Ir1—C1	17.6 (3)
C25—C20—P1—Ir1	173.1 (5)	C20—P1—Ir1—C1	-108.1 (3)
C18—C19—P1—C26	-174.6 (5)	C19—P1—Ir1—C1	136.6 (3)
C18—C19—P1—C20	-64.6 (5)	C26—P1—Ir1—C5	50.8 (3)
C18—C19—P1—Ir1	60.9 (5)	C20—P1—Ir1—C5	-74.9 (3)
C16—C11—S1—C17	64.4 (6)	C19—P1—Ir1—C5	169.9 (3)
C12—C11—S1—C17	-116.6 (6)	C26—P1—Ir1—C2	16.9 (5)
C16—C11—S1—Ir1	-50.8 (6)	C20—P1—Ir1—C2	-108.8 (5)
C12—C11—S1—Ir1	128.2 (5)	C19—P1—Ir1—C2	135.9 (5)
C18—C17—S1—C11	176.7 (5)	C26—P1—Ir1—C3	98.9 (4)
C18—C17—S1—Ir1	-67.9 (5)	C20—P1—Ir1—C3	-26.7 (4)
C5—C4—Ir1—C1	36.8 (4)	C19—P1—Ir1—C3	-142.0 (4)
C3—C4—Ir1—C1	-78.4 (4)	C26—P1—Ir1—S1	-165.5 (3)
C9—C4—Ir1—C1	157.6 (7)	C20—P1—Ir1—S1	68.8 (2)
C3—C4—Ir1—C5	-115.3 (5)	C19—P1—Ir1—S1	-46.5 (2)
C9—C4—Ir1—C5	120.8 (8)	C26—P1—Ir1—C11	-73.7 (3)
C5—C4—Ir1—C2	80.1 (4)	C20—P1—Ir1—C11	160.7 (2)
C3—C4—Ir1—C2	-35.1 (4)	C19—P1—Ir1—C11	45.4 (2)
C9—C4—Ir1—C2	-159.1 (7)	C11—S1—Ir1—C4	-91.0 (3)
C5—C4—Ir1—C3	115.3 (5)	C17—S1—Ir1—C4	160.2 (3)
C9—C4—Ir1—C3	-124.0 (8)	C11—S1—Ir1—C1	-28.3 (5)
C5—C4—Ir1—P1	-74.6 (4)	C17—S1—Ir1—C1	-137.2 (5)
C3—C4—Ir1—P1	170.2 (3)	C11—S1—Ir1—C5	-101.6 (4)
C9—C4—Ir1—P1	46.2 (6)	C17—S1—Ir1—C5	149.6 (4)
C5—C4—Ir1—S1	-169.7 (3)	C11—S1—Ir1—C2	-23.3 (3)

C3—C4—Ir1—S1	75.0 (4)	C17—S1—Ir1—C2	-132.2 (3)
C9—C4—Ir1—S1	-49.0 (6)	C11—S1—Ir1—C3	-54.5 (3)
C5—C4—Ir1—Cl1	57.6 (6)	C17—S1—Ir1—C3	-163.4 (3)
C3—C4—Ir1—Cl1	-57.6 (6)	C11—S1—Ir1—P1	157.8 (2)
C9—C4—Ir1—Cl1	178.4 (4)	C17—S1—Ir1—P1	48.9 (2)
C5—C1—Ir1—C4	-37.4 (4)	C11—S1—Ir1—Cl1	70.5 (2)
C2—C1—Ir1—C4	79.8 (4)	C17—S1—Ir1—Cl1	-38.4 (2)
C6—C1—Ir1—C4	-162.5 (6)		

Hydrogen-bond geometry (Å, °)

<i>D—H...A</i>	<i>D—H</i>	<i>H...A</i>	<i>D...A</i>	<i>D—H...A</i>
O1—H1O...Cl2	0.95 (6)	2.27 (6)	3.216 (6)	177 (10)
O1—H2O...Cl2 ⁱ	0.96 (8)	2.26 (11)	3.196 (6)	165 (13)

Symmetry code: (i) $-x+1, -y+2, -z$.