



Crystal structure of 2-(4-chloro-benzamido)benzoic acid

Rodolfo Moreno-Fuquen,^{a*} Vanessa Melo^a and Javier Ellena^b

^aDepartamento de Química - Facultad de Ciencias Naturales y Exactas, Universidad del Valle, Apartado 25360, Santiago de Cali, Colombia, and ^bInstituto de Física de São Carlos, IFSC, Universidade de São Paulo, USP, São Carlos, SP, Brazil.

*Correspondence e-mail: rodimo26@yahoo.es

Received 22 September 2015; accepted 23 September 2015

Edited by E. R. T. Tiekink, University of Malaya, Malaysia

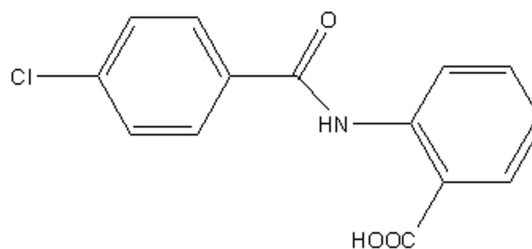
In the title molecule, C₁₄H₁₀ClNO₃, the amide C=O bond is *anti* to the *o*-carboxy substituent in the adjacent benzene ring, a conformation that facilitates the formation of an intramolecular amide-N—H···O(carbonyl) hydrogen bond that closes an *S*(6) loop. The central amide segment is twisted away from the carboxy- and chloro-substituted benzene rings by 13.93 (17) and 15.26 (15)°, respectively. The most prominent supramolecular interactions in the crystal packing are carboxylic acid-H···O(carboxyl) hydrogen bonds that lead to centrosymmetric dimeric aggregates connected by eight-membered {···HOC=O}₂ synthons.

Keywords: crystal structure; carboxylic acid; amide; hydrogen bonding.

CCDC reference: 1427117

1. Related literature

For our studies on the effects of substituents on the structures of *N*-(aryl)-amides, see: Moreno-Fuquen *et al.* (2014, 2015). For benzanilide properties, see: Nuta *et al.* (2013); Leander (1992); Ahles *et al.* (2004). For related structures, see: Saeed *et al.* (2008, 2010); Rodrigues *et al.* (2011). For hydrogen bonding, see: Desiraju & Steiner (1999), Nardelli (1995).



2. Experimental

2.1. Crystal data

C ₁₄ H ₁₀ ClNO ₃	<i>V</i> = 2513.7 (2) Å ³
<i>M_r</i> = 275.69	<i>Z</i> = 8
Monoclinic <i>C</i> 2/ <i>c</i>	Mo <i>K</i> α radiation
<i>a</i> = 26.8843 (10) Å	<i>μ</i> = 0.31 mm ⁻¹
<i>b</i> = 5.0367 (2) Å	<i>T</i> = 295 K
<i>c</i> = 20.9264 (12) Å	0.40 × 0.08 × 0.06 mm
<i>β</i> = 117.489 (2)°	

2.2. Data collection

Nonius KappaCCD diffractometer	1049 reflections with <i>I</i> > 2σ(<i>I</i>)
4248 measured reflections	<i>R</i> _{int} = 0.057
2295 independent reflections	

2.3. Refinement

<i>R</i> [<i>F</i> ² > 2σ(<i>F</i> ²)] = 0.043	H atoms treated by a mixture of independent and constrained refinement
<i>wR</i> (<i>F</i> ²) = 0.132	<i>Δρ</i> _{max} = 0.20 e Å ⁻³
<i>S</i> = 0.92	<i>Δρ</i> _{min} = -0.19 e Å ⁻³
2295 reflections	
176 parameters	

Table 1

Hydrogen-bond geometry (Å, °).

<i>D</i> —H··· <i>A</i>	<i>D</i> —H	H··· <i>A</i>	<i>D</i> ··· <i>A</i>	<i>D</i> —H··· <i>A</i>
N1—NH1···O2	0.93 (3)	1.96 (3)	2.678 (3)	133 (3)
O3—OH3···O2 ⁱ	0.82	1.83	2.645 (3)	175

Symmetry code: (i) $-x + \frac{3}{2}, -y + \frac{3}{2}, -z + 1$.

Data collection: *COLLECT* (Nonius, 2000); cell refinement: *HKL SCALEPACK* (Otwinowski & Minor, 1997); data reduction: *HKL DENZO* (Otwinowski & Minor, 1997) and *SCALEPACK*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL2014/7* (Sheldrick, 2015); molecular graphics: *ORTEP-3 for Windows* (Farrugia, 2012) and *Mercury* (Macrae *et al.*, 2006); software used to prepare material for publication: *SHELXL2014/7*.

Acknowledgements

RMF is grateful to the Universidad del Valle, Colombia, for partial financial support.

Supporting information for this paper is available from the IUCr electronic archives (Reference: TK5391).

References

- Ahles, T. A., Herndon, J. E., Small, E. J., Vogelzang, N. J., Kornblith, A. B., Ratain, M. J., Stadler, W. S., Palchak, D., Marshall, E., Wilding, G., Petrylak, D. & Holland, C. (2004). *Cancer*, **101**, 2202–2208.
- Desiraju, G. R. & Steiner, T. (1999). *The Weak Hydrogen Bond*. Oxford University Press, pp. 86–89.
- Farrugia, L. J. (2012). *J. Appl. Cryst.* **45**, 849–854.
- Leander, J. D. (1992). *Epilepsia*, **33**, 705–711.
- Macrae, C. F., Edgington, P. R., McCabe, P., Pidcock, E., Shields, G. P., Taylor, R., Towler, M. & van de Streek, J. (2006). *J. Appl. Cryst.* **39**, 453–457.
- Moreno-Fuquen, R., Azcárate, A. & Kennedy, A. R. (2014). *Acta Cryst.* **E70**, o344.
- Moreno-Fuquen, R., Azcárate, A. & Kennedy, A. R. (2015). *Acta Cryst.* **E71**, o389–o390.
- Nardelli, M. (1995). *J. Appl. Cryst.* **28**, 659.
- Nonius (2000). *COLLECT*. Nonius BV, Delft, The Netherlands.
- Nuta, D. C., Chifriuc, A. C., Draghici, C., Limban, C., Missir, A. V. & Morusciag, L. (2013). *Farmacia (Bucarest)*, **61**, 966–974.
- Otwinowski, Z. & Minor, W. (1997). *Methods in Enzymology*, Vol. 276, *Macromolecular Crystallography*, Part A, edited by C. W. Carter Jr & R. M. Sweet, pp. 307–326. New York: Academic Press.
- Rodrigues, V. Z., Kucková, L., Gowda, B. T. & Kožíšek, J. (2011). *Acta Cryst.* **E67**, o3171.
- Saeed, A., Khera, R. A., Abbas, N., Gotoh, K. & Ishida, H. (2008). *Acta Cryst.* **E64**, o2043.
- Saeed, A., Khera, R. A. & Simpson, J. (2010). *Acta Cryst.* **E66**, o214.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Sheldrick, G. M. (2015). *Acta Cryst.* **C71**, 3–8.

supporting information

Acta Cryst. (2015). E71, o856–o857 [https://doi.org/10.1107/S2056989015017879]

Crystal structure of 2-(4-chlorobenzamido)benzoic acid

Rodolfo Moreno-Fuquen, Vanessa Melo and Javier Ellena

S1. Comment

The crystal structure determination of 2-(4-chlorobenzamido)benzoic acid, (I), was investigated in a continuation of our studies on substituted *N*-phenyl benzamides which have been synthesized from picryl esters. This study was also performed to study the effect of substituents on the structures of benzanilides (Moreno-Fuquen *et al.*, 2014, 2015). Benzanilide systems are used as antimicrobial drugs (Nuta *et al.*, 2013), as anticonvulsants (Leander, 1992) or as treatment for patients with prostate carcinoma (Ahles *et al.*, 2004). Structures of similar molecules were compared with (I), *i.e.* 4-chloro-*N*-(2-methoxyphenyl)benzamide (Saeed *et al.*, 2010), 4-chloro-*N*-phenylbenzamide (Rodrigues *et al.*, 2011) and 4-chloro-*N*-(*o*-tolyl)benzamide (Saeed *et al.*, 2008). The molecular structure of (I) is shown in Fig. 1. The C=O bond is *anti* to the *o*-carboxy substituent in the benzoyl ring. The N—H and C=O bonds in the central amide group are also *anti* to each other. Comparing (I) with the three aforementioned structures reveals that significant differences in bond lengths and bond angles are not observed. The central amide segment (C1-C7(O1)-N1-C8) is twisted away from the carboxy- and chloro-substituted benzene rings by 13.93 (17) and 15.26 (15)°, respectively. Molecules of (I) are held together by intermolecular O—H⋯O hydrogen bonds of moderate strength (Desiraju & Steiner, 1999). The O3 atom is linked to O2^{*i*} atom (*i* = -*x*+3/2, -*y*+3/2, -*z*+1) with O⋯O distances of 2.645 (2), (see Table 1, Nardelli, 1995). Except for the presence of hydrogen bonding in the formation of dimer, no other significant intermolecular interactions are observed in the structure.

S2. Experimental

2,4,6-Trinitrophenyl 4-chlorobenzoate (0.060 g, 0.163 mmol) and 2-carboxyaniline (0.045 g, 0.328 mmol) were dissolved in toluene (15 ml) and stirred for 6 h under reflux. On completion of the reaction part of the solvent was evaporated and a crystalline yellow solid was obtained; m.p. 470 (1) K.

S3. Refinement

All H-atoms were located in difference Fourier maps and were positioned geometrically [C—H = 0.93 Å] and were refined using a riding-model approximation with $U_{\text{iso}}(\text{H})$ constrained to $1.2U_{\text{eq}}(\text{C})$. The O-bound H atoms was similarly fixed with O—H = 0.82 Å, and with $U_{\text{iso}}(\text{H})$ constrained to $1.5U_{\text{eq}}(\text{O})$. The N-bound H atom was found from the Fourier maps and was refined freely.

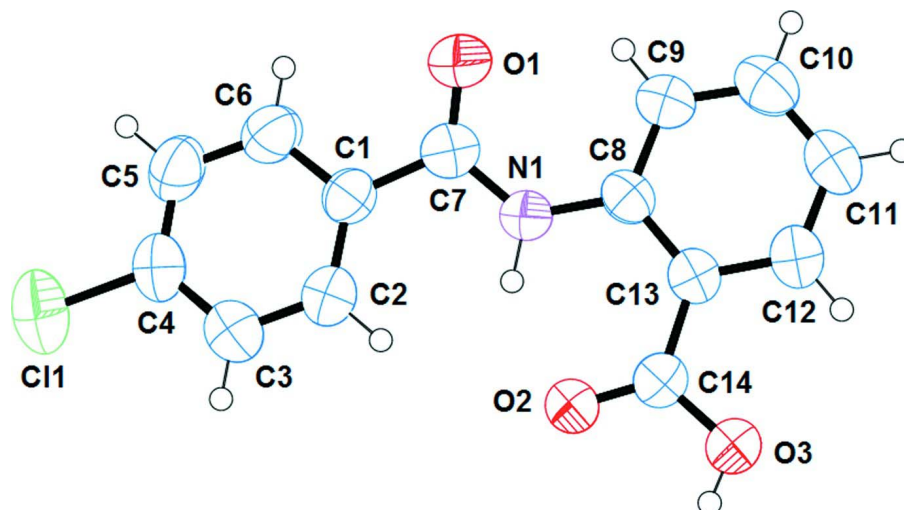


Figure 1

The molecular structure of (I) with displacement ellipsoids drawn at the 50% probability level. H atoms are shown as spheres of arbitrary radius.

2-(4-Chlorobenzamido)benzoic acid

Crystal data

$C_{14}H_{10}ClNO_3$

$M_r = 275.69$

Monoclinic, $C2/c$

$a = 26.8843$ (10) Å

$b = 5.0367$ (2) Å

$c = 20.9264$ (12) Å

$\beta = 117.489$ (2)°

$V = 2513.7$ (2) Å³

$Z = 8$

$F(000) = 1136$

$D_x = 1.457$ Mg m⁻³

Melting point: 470(1) K

Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 2553 reflections

$\theta = 3.1$ – 25.4 °

$\mu = 0.31$ mm⁻¹

$T = 295$ K

Needle, yellow

$0.40 \times 0.08 \times 0.06$ mm

Data collection

Nonius KappaCCD

diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

CCD rotation images, thick slices scans

4248 measured reflections

2295 independent reflections

1049 reflections with $I > 2\sigma(I)$

$R_{int} = 0.057$

$\theta_{max} = 25.4$ °, $\theta_{min} = 3.1$ °

$h = -31 \rightarrow 32$

$k = -6 \rightarrow 6$

$l = -25 \rightarrow 24$

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.043$

$wR(F^2) = 0.132$

$S = 0.92$

2295 reflections

176 parameters

0 restraints

Primary atom site location: structure-invariant direct methods

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring sites

H atoms treated by a mixture of independent and constrained refinement

$$w = 1/[\sigma^2(F_o^2) + (0.0632P)^2]$$

where $P = (F_o^2 + 2F_c^2)/3$
 $(\Delta/\sigma)_{\max} < 0.001$

$$\Delta\rho_{\max} = 0.20 \text{ e } \text{\AA}^{-3}$$

$$\Delta\rho_{\min} = -0.19 \text{ e } \text{\AA}^{-3}$$

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
O1	0.74640 (9)	-0.2731 (4)	0.30942 (12)	0.0824 (7)
C5	0.90834 (14)	-0.0840 (7)	0.34839 (18)	0.0786 (9)
H5	0.9265	-0.1847	0.3282	0.094*
C10	0.58970 (13)	-0.0955 (6)	0.29315 (17)	0.0739 (9)
H10	0.5647	-0.2205	0.2623	0.089*
C9	0.64347 (12)	-0.0846 (5)	0.30083 (15)	0.0623 (8)
H9	0.6543	-0.2008	0.2750	0.075*
C2	0.85453 (13)	0.2133 (6)	0.40771 (18)	0.0766 (9)
H2	0.8366	0.3155	0.4280	0.092*
C11	0.57224 (13)	0.0750 (6)	0.33029 (18)	0.0754 (9)
H11	0.5357	0.0667	0.3242	0.090*
C6	0.85319 (14)	-0.1349 (6)	0.33124 (17)	0.0730 (9)
H6	0.8344	-0.2716	0.2993	0.088*
C3	0.90974 (14)	0.2649 (6)	0.42543 (18)	0.0801 (10)
H3	0.9289	0.4004	0.4577	0.096*
C4	0.93606 (13)	0.1173 (7)	0.39566 (17)	0.0701 (9)
NH1	0.7536 (13)	0.280 (6)	0.3758 (17)	0.102 (11)*
Cl1	1.00534 (4)	0.1835 (2)	0.41722 (6)	0.1062 (4)
N1	0.73632 (10)	0.1205 (5)	0.35559 (12)	0.0592 (7)
O3	0.67865 (8)	0.6206 (3)	0.46790 (10)	0.0688 (6)
OH3	0.7015	0.7286	0.4949	0.103*
O2	0.75129 (8)	0.5079 (3)	0.45041 (10)	0.0641 (6)
C14	0.70192 (12)	0.4755 (5)	0.43671 (14)	0.0548 (7)
C8	0.68169 (11)	0.1015 (5)	0.34751 (14)	0.0529 (7)
C7	0.76554 (12)	-0.0609 (6)	0.33845 (15)	0.0599 (7)
C13	0.66432 (11)	0.2748 (5)	0.38605 (14)	0.0535 (7)
C1	0.82537 (12)	0.0120 (5)	0.36030 (15)	0.0568 (7)
C12	0.60946 (12)	0.2573 (6)	0.37640 (15)	0.0664 (8)
H12	0.5979	0.3718	0.4018	0.080*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
O1	0.0839 (16)	0.0631 (13)	0.1073 (17)	-0.0111 (11)	0.0502 (14)	-0.0225 (12)
C5	0.075 (2)	0.090 (2)	0.083 (2)	0.0100 (19)	0.047 (2)	-0.004 (2)
C10	0.063 (2)	0.071 (2)	0.072 (2)	-0.0093 (16)	0.0174 (18)	0.0023 (17)
C9	0.063 (2)	0.0595 (17)	0.0591 (19)	-0.0054 (15)	0.0242 (16)	-0.0031 (15)
C2	0.066 (2)	0.081 (2)	0.090 (2)	-0.0010 (17)	0.0417 (18)	-0.0156 (19)
C11	0.056 (2)	0.083 (2)	0.083 (2)	-0.0118 (18)	0.0279 (18)	-0.0009 (19)
C6	0.080 (2)	0.072 (2)	0.074 (2)	0.0010 (17)	0.0420 (19)	-0.0089 (16)
C3	0.066 (2)	0.085 (2)	0.089 (2)	-0.0102 (17)	0.035 (2)	-0.0121 (19)
C4	0.0581 (19)	0.083 (2)	0.073 (2)	0.0059 (17)	0.0343 (17)	0.0122 (18)
C11	0.0657 (6)	0.1390 (9)	0.1189 (8)	0.0021 (5)	0.0469 (6)	0.0093 (6)
N1	0.0576 (16)	0.0555 (15)	0.0681 (16)	-0.0068 (13)	0.0322 (13)	-0.0118 (13)
O3	0.0622 (13)	0.0726 (12)	0.0763 (13)	-0.0053 (10)	0.0360 (11)	-0.0183 (11)
O2	0.0560 (13)	0.0685 (12)	0.0694 (13)	-0.0070 (10)	0.0304 (11)	-0.0138 (10)
C14	0.0581 (19)	0.0553 (17)	0.0522 (18)	0.0026 (15)	0.0264 (16)	0.0028 (14)
C8	0.0513 (17)	0.0527 (15)	0.0506 (16)	-0.0041 (13)	0.0202 (14)	0.0055 (14)
C7	0.068 (2)	0.0570 (18)	0.0555 (18)	0.0000 (16)	0.0298 (16)	0.0024 (15)
C13	0.0515 (17)	0.0564 (16)	0.0514 (17)	-0.0007 (14)	0.0228 (14)	0.0040 (14)
C1	0.0608 (19)	0.0581 (17)	0.0554 (18)	0.0033 (15)	0.0302 (16)	0.0025 (14)
C12	0.0564 (19)	0.074 (2)	0.071 (2)	-0.0005 (16)	0.0314 (16)	0.0024 (16)

Geometric parameters (\AA , $^\circ$)

O1—C7	1.219 (3)	C6—H6	0.9300
C5—C4	1.371 (4)	C3—C4	1.360 (4)
C5—C6	1.379 (4)	C3—H3	0.9300
C5—H5	0.9300	C4—C11	1.734 (3)
C10—C11	1.378 (4)	N1—C7	1.357 (3)
C10—C9	1.380 (4)	N1—C8	1.402 (3)
C10—H10	0.9300	N1—NH1	0.93 (3)
C9—C8	1.401 (4)	O3—C14	1.315 (3)
C9—H9	0.9300	O3—OH3	0.8200
C2—C3	1.378 (4)	O2—C14	1.233 (3)
C2—C1	1.382 (4)	C14—C13	1.476 (4)
C2—H2	0.9300	C8—C13	1.406 (4)
C11—C12	1.373 (4)	C7—C1	1.501 (4)
C11—H11	0.9300	C13—C12	1.396 (4)
C6—C1	1.377 (4)	C12—H12	0.9300
C4—C5—C6	119.1 (3)	C5—C4—C11	119.4 (3)
C4—C5—H5	120.5	C7—N1—C8	128.6 (3)
C6—C5—H5	120.5	C7—N1—NH1	118 (2)
C11—C10—C9	121.4 (3)	C8—N1—NH1	113 (2)
C11—C10—H10	119.3	C14—O3—OH3	109.5
C9—C10—H10	119.3	O2—C14—O3	121.3 (3)
C10—C9—C8	120.0 (3)	O2—C14—C13	124.4 (3)

C10—C9—H9	120.0	O3—C14—C13	114.3 (3)
C8—C9—H9	120.0	C9—C8—N1	121.3 (3)
C3—C2—C1	121.0 (3)	C9—C8—C13	119.0 (3)
C3—C2—H2	119.5	N1—C8—C13	119.7 (2)
C1—C2—H2	119.5	O1—C7—N1	124.0 (3)
C12—C11—C10	119.1 (3)	O1—C7—C1	120.8 (3)
C12—C11—H11	120.5	N1—C7—C1	115.1 (3)
C10—C11—H11	120.5	C12—C13—C8	119.1 (3)
C1—C6—C5	121.5 (3)	C12—C13—C14	118.3 (3)
C1—C6—H6	119.2	C8—C13—C14	122.6 (3)
C5—C6—H6	119.2	C6—C1—C2	117.9 (3)
C4—C3—C2	119.8 (3)	C6—C1—C7	117.3 (3)
C4—C3—H3	120.1	C2—C1—C7	124.9 (3)
C2—C3—H3	120.1	C11—C12—C13	121.5 (3)
C3—C4—C5	120.7 (3)	C11—C12—H12	119.2
C3—C4—C11	119.9 (3)	C13—C12—H12	119.2
C11—C10—C9—C8	0.4 (4)	N1—C8—C13—C14	-1.4 (4)
C9—C10—C11—C12	-0.7 (5)	O2—C14—C13—C12	179.1 (3)
C4—C5—C6—C1	0.2 (5)	O3—C14—C13—C12	-0.2 (3)
C1—C2—C3—C4	0.4 (5)	O2—C14—C13—C8	-0.8 (4)
C2—C3—C4—C5	-0.4 (5)	O3—C14—C13—C8	179.9 (2)
C2—C3—C4—C11	179.4 (2)	C5—C6—C1—C2	-0.2 (4)
C6—C5—C4—C3	0.1 (5)	C5—C6—C1—C7	-179.6 (3)
C6—C5—C4—C11	-179.7 (2)	C3—C2—C1—C6	-0.1 (4)
C10—C9—C8—N1	-178.9 (2)	C3—C2—C1—C7	179.3 (3)
C10—C9—C8—C13	0.1 (4)	O1—C7—C1—C6	15.5 (4)
C7—N1—C8—C9	-18.7 (4)	N1—C7—C1—C6	-166.6 (2)
C7—N1—C8—C13	162.3 (3)	O1—C7—C1—C2	-163.9 (3)
C8—N1—C7—O1	3.2 (5)	N1—C7—C1—C2	14.0 (4)
C8—N1—C7—C1	-174.6 (2)	C10—C11—C12—C13	0.4 (4)
C9—C8—C13—C12	-0.3 (4)	C8—C13—C12—C11	0.1 (4)
N1—C8—C13—C12	178.7 (2)	C14—C13—C12—C11	-179.8 (2)
C9—C8—C13—C14	179.6 (2)		

Hydrogen-bond geometry (Å, °)

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
N1—NH1...O2	0.93 (3)	1.96 (3)	2.678 (3)	133 (3)
O3—OH3...O2 ⁱ	0.82	1.83	2.645 (3)	175

Symmetry code: (i) $-x+3/2, -y+3/2, -z+1$.