



Crystal structure of bis(μ -*N*-hydroxypicolinamidato)bis[bis(*N*-hydroxypicolinamide)sodium]

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Received 23 November 2016

Accepted 30 November 2016

Edited by D.-J. Xu, Zhejiang University (Yuquan Campus), China

Keywords: crystal structure; hydroxamic acids; hydrogen bonds; π - π stacking.

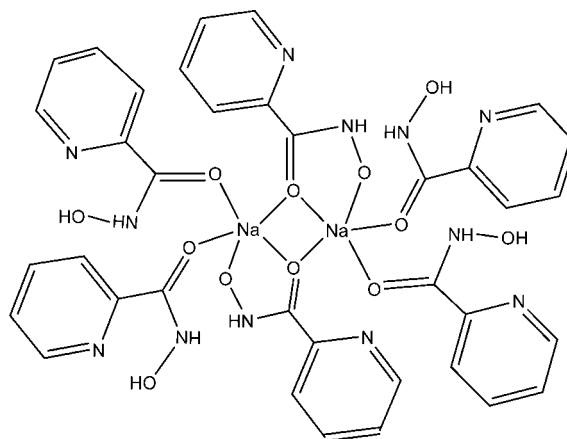
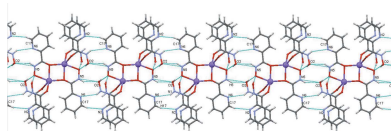
CCDC reference: 1520114

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The title compound, $[\text{Na}_2(\text{C}_6\text{H}_5\text{N}_2\text{O}_2)_2(\text{C}_6\text{H}_6\text{N}_2\text{O}_2)_4]$, is a centrosymmetric coordination dimer based on the sodium(I) salt of *N*-hydroxypicolinamide. The molecule has an $\{\text{Na}_2\text{O}_6(\mu\text{-O})_2\}$ core with two bridging carbonyl O atoms and two hydroxamate O atoms of two mono-deprotonated residues of *N*-hydroxypicolinamide, while two neutral *N*-hydroxypicolinamide molecules are coordinated in a monodentate manner to each sodium ion *via* the carbonyl O atoms [the Na—O distances range from 2.3044 (2) to 2.3716 (2) Å]. The penta-coordinated sodium ion exhibits a distorted trigonal-pyramidal coordination polyhedron. In the crystal, the coordination dimers are linked into chains along the *c* axis *via* N—H...O and N—H...N hydrogen bonds; the chains are linked into a two-dimensional framework parallel to (100) *via* weak C—H...O and π - π stacking interactions.

1. Chemical context

Hydroxamic acids as a class of organic compounds originate from Lossen's invention (Lossen, 1869). The coordination ability of hydroxamic acids has led to their extensive use in coordination and supramolecular chemistry (Świątek-Kozłowska *et al.*, 2000; Dobosz *et al.*, 1999). In particular, over the past two decades they have often been used as frameworks of metallacrowns (Golenya *et al.*, 2012a; Safyanova *et al.*, 2015; Stemmler *et al.*, 1999; Jankolovits *et al.*, 2013a,b) and as building blocks of coordination polymers (Gumienna-Kontecka *et al.*, 2007; Golenya *et al.*, 2014; Pavlishchuk *et al.*, 2010, 2011). They have also been studied intensively in biology and medicine due to their various biological activities, especially their metal-chelating ability and inhibition of a series of metalloenzymes (Codd, 2008; Griffith *et al.*, 2005; Marmion *et al.*, 2013).



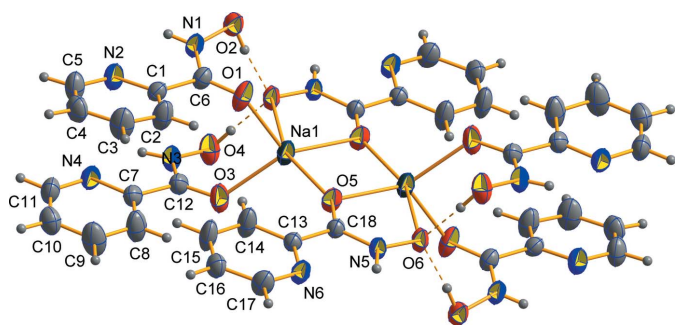


Figure 1

The centrosymmetric molecular unit of the title compound, with displacement ellipsoids drawn at the 50% probability level. H atoms are shown as spheres of undefined radius.

N-Hydroxypicolinamide (or picoline-2-hydroxamic acid, H_2PicHA) has been used extensively for the synthesis of polynuclear complexes, especially various metallacrowns (Stemmler *et al.*, 1999; Seda *et al.*, 2007; Jankolovits *et al.*, 2013a; Golenya *et al.*, 2012a; Gumienna-Kontecka *et al.*, 2013). A large number of polynuclear metal complexes based on this ligand has been investigated. The Cambridge Structural Database (Groom *et al.*, 2016) contains data on the crystal structures of over 20 coordination compounds based on *o*-PicHA. The crystal and molecular structure of *N*-hydroxypicolinamide monohydrate was the subject of two recent independent investigations (Chaiyavej *et al.*, 2015; Safyanova *et al.*, 2016).

In the course of the synthesis of hydroxamate metal complexes, especially metallacrowns, in some cases alkaline metal hydroxamates appear to be more preferable starting materials than the parent hydroxamic acids due to their better solubility in water. During our synthetic attempts using the sodium salt of *N*-hydroxypicolinamide, we noticed that the elementary analysis data differ noticeably from those expected for the monosodium salt or its hydrates, which might affect the reagent ratio in the synthesis of coordination compounds. In order to find out the reason for this deviation in the analytical data, we undertook a single crystal X-ray analysis of the sodium salt of *N*-hydroxypicolinamide. Herein we present the crystal and molecular structure of the title compound.

2. Structural commentary

The molecular structure of title compound is shown in Fig. 1. The structure determination revealed that the dinuclear hydroxamate acid salt was obtained, with the ratio of neutral and deprotonated *N*-hydroxypicolinamide being 2:1. A centrosymmetric dimeric structure is formed by non-planar subunits interconnected through the bridging carbonyl O atoms belonging to the deprotonated residues of *N*-hydroxypicolinamide [the Na— μ -O distances are Na1—O5 = 2.3044 (14) Å and Na1—O5ⁱ = 2.3558 (14) Å; symmetry code: (i) 1 - x, -y, 1 - z]. Coordination of the μ -O carbonyl and hydroxamate O atoms of the same anion lead to the formation of five-membered chelate rings [Na1—O6ⁱ = 2.3716 (14) Å

Table 1

Hydrogen-bond geometry (Å, °).

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
N1—H1...N2	0.80 (3)	2.35 (3)	2.688 (3)	106 (2)
N3—H3...N4	0.86 (3)	2.30 (3)	2.681 (3)	107 (2)
N5—H5...N6	0.84 (2)	2.25 (2)	2.670 (3)	111.1 (17)
O2—H2...O6 ⁱ	0.91 (2)	1.65 (2)	2.549 (2)	169 (2)
O4—H4...O6 ⁱ	0.91 (3)	1.66 (3)	2.5744 (19)	177 (3)
N1—H1...N6 ⁱⁱ	0.80 (2)	2.55 (3)	3.224 (2)	143 (2)
N5—H5...O2 ⁱⁱⁱ	0.84 (2)	2.35 (2)	3.058 (2)	142.6 (19)
C5—H5A...O4 ^{iv}	0.93	2.61	3.341 (3)	136
C17—H17...N2 ⁱⁱⁱ	0.93	2.60	3.330 (3)	136

Symmetry codes: (i) $-x + 1, -y, -z + 1$; (ii) $x, y, z + 1$; (iii) $x, y, z - 1$; (iv) $-x + 1, -y + 1, -z + 2$.

and O5ⁱ—Na—O6ⁱ = 70.26 (5)°]. Two neutral *N*-hydroxypicolinamide molecules coordinate in a monodentate manner to each sodium ion *via* the carbonyl O atoms [Na1—O1 = 2.3300 (16) Å and Na1—O3 = 2.3225 (15) Å]. As a result, each pentacoordinated sodium ion reveals a distorted trigonal-pyramidal coordination polyhedron ($\tau_5 = 0.50$; Addison *et al.*, 1984) with O1, O3, and O5ⁱ atoms forming the equatorial plane. The distance between the equatorial plane and the Na atom is 0.408 (1) Å and the deviation of the O—Na—O angles from ideal values are up to 23.47 (5)°. The Na—O bond lengths are in the range 2.3044 (14)—2.3716 (14) Å, which is common for pentacoordinated sodium cations (Groom *et al.*, 2016; Golenya *et al.*, 2012b; Malinkin *et al.*, 2012a,b). The central Na₂(μ -O)₂ core is virtually planar and approaches a square [the O—Na—O angles are 86.43 (5) and 93.57 (5)°].

The deprotonated hydroxamate atom O6 acts as an acceptor of two hydrogen bonds (Table 1) in which the O—H groups of the protonated hydroxamic groups of two neutral molecules of *N*-hydroxypicolinamide act as donors [O2—H2...O6(1 - x, -y, 1 - z) = 1.65 (2) Å and 169 (2)°; O4—H4...O6(1 - x, -y, 1 - z) = 1.66 (3) Å and 177 (3)°]. The nearly coplanar pyridine rings of two neutral molecules of *N*-hydroxypicolinamide coordinating to the same sodium ion reveal intramolecular stacking interactions in unusual 'head-to-head' manner [angle between planes = 10.00 (7)°, inter-centroid distance = 3.801 (1) Å, mean interplanar separation = 3.760 (1) Å, mean plane shift = 0.508 (4) Å].

The deprotonated *N*-hydroxypicolinamide residue adopts a strongly flattened conformation with a dihedral angle of only 0.6 (2)° between the hydroxamic group and the pyridine ring. At the same time, the corresponding dihedral angles in both neutral *N*-hydroxypicolinamide molecules are noticeably greater [17.5 (2) and 8.9 (2)°], indicating a deviation of the hydroxamic group from the plane of pyridine rings. The configuration about the hydroxamic C—N bond is *Z* and that about the C—C bond between the pyridine and hydroxamic groups is *E* for both the neutral and deprotonated hydroxamates. Intramolecular N—H...N attractive contacts between the hydroxamate group and the nitrogen atom of pyridine ring [2.25 (2)—2.35 (3) Å] are present in both the neutral and deprotonated *N*-hydroxypicolinamide molecules (Table 1).

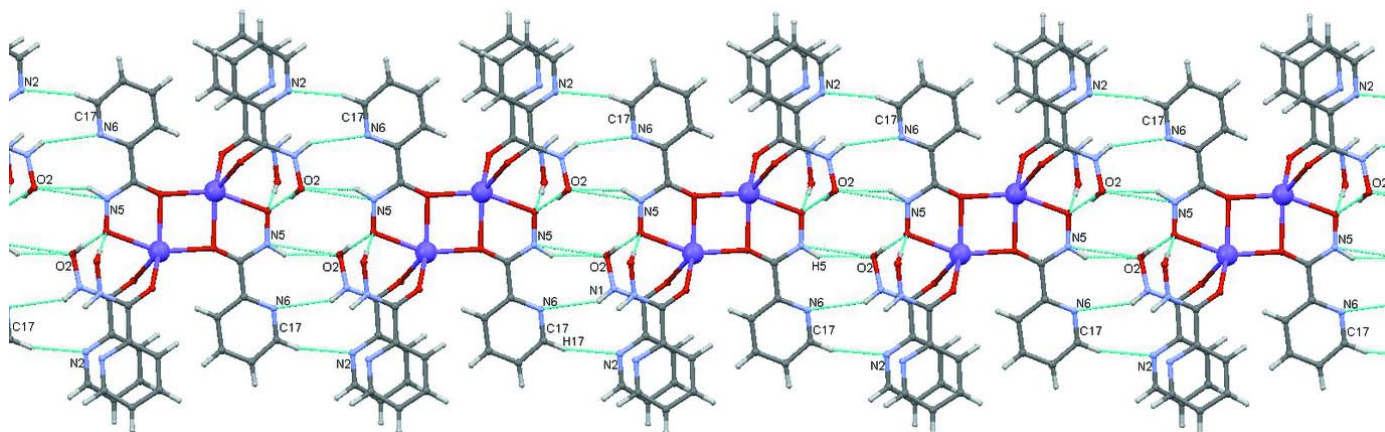


Figure 2

A packing diagram of the title compound. Hydrogen bonds are indicated by dashed lines. [The outline of the unit cell and axes should be added]

The bond lengths and angles within both the neutral and deprotonated hydroxamic groups are within normal ranges. The C–N and C–C bond lengths in the pyridine moiety are typical for 2-substituted pyridine derivatives (Moroz *et al.*, 2012; Strotmeyer *et al.*, 2003; Fritsky *et al.*, 2004).

3. Supramolecular features

In the crystal (Fig. 2), the dimeric molecules are linked into chains along the *c* axis via two pairs of classical intermolecular N5–H5···O2(*x*, *y*, *z* – 1) and N1–H1···N6(*x*, *y*, *z* – 1) hydrogen bonds supported by a pair of weak non-classical C17–H17···N2(*x*, *y*, *z* – 1) hydrogen bonds (Table 1). The chains are linked into a two-dimensional framework parallel to (100) by weaker interactions, namely a C5–H5A···O4(–*x* + 1, –*y* + 1, –*z* + 2) hydrogen bond and π – π stacking between the N4/C7–C11 pyridine ring and the deprotonated O5/C18/N5/O6 hydroxamic group [angle between planes = 4.89 (7)°, intercentroid distance = 3.766 (1) Å, mean interplanar separation = 3.385 (2) Å, mean plane shift = 1.644 (4) Å]. Intermolecular π – π stacking between the same deprotonated hydroxamic group and the N2/C1–C5 pyridine ring [angle between planes = 10.78 (8)°, intercentroid distance = 3.823 (1) Å, mean interplanar separation = 3.589 (2) Å, mean plane shift = 1.319 (4) Å] links the frameworks into a three-dimensional structure.

4. Database survey

A search of the Cambridge Structural Database (Groom *et al.*, 2016) for metal complexes based on *N*-hydroxypicolinamide revealed the crystal structures of over 20 compounds, mostly belonging to the metallacrown (MC) family. In particular, heterometallic copper(II) 15-metallacrown-5 complexes with encapsulated Gd^{III} and Eu^{III} ions (Stemmler *et al.*, 1999), Ca²⁺, Pr³⁺ and Nd³⁺ ions (Safyanova *et al.*, 2015), UO₂²⁺ (Stemmler *et al.*, 1996), and Pb²⁺ and Hg²⁺ ions (Seda *et al.*, 2007; Safyanova *et al.*, 2014) have been structurally characterized. Nickel(II) 15-metallacrown-5 complexes with Eu³⁺ (Jankolovits *et al.*, 2013b), Sm³⁺ and Pb²⁺ ions (Seda *et al.*, 2006a) in the central cavity have also been synthesized and structurally characterized.

Homo-[12-MC_{Zn(II),picHA-4}](OTf)_{1.25}(OH)_{0.75} (Jankolovits *et al.*, 2013a) and heterometallic zinc(II) 12-metallacrown-4 complexes including sandwich compounds Dy^{III}[12-MC_{Zn(II),picHA-4}]₂(OH)₃(py)₂ (Jankolovits *et al.*, 2014) and Tb^{III}[12-MC_{Zn(II),picHA-4}]₂·[24 MC_{Zn(II),picHA-8}](pyridine)₈·(triflate)₃ (Jankolovits *et al.*, 2011) have also been reported. Three structures of collapsed copper(II) metallacrowns have been reported (Golenya *et al.*, 2012a) as well as a trinuclear mixed-ligand copper(II) complex with pyridine (Seda *et al.*, 2006b) and 2,2'-dipyridine (Gumienna-Kontecka *et al.*, 2013), and mono- and binuclear complexes with platinum(II) (Griffith *et al.*, 2005). In addition, a tetranuclear Zn₄(picHA)₂(OAc)₄(DMF)₂ collapsed metallacrown complex has been structurally characterized (Jankolovits *et al.*, 2013a).

5. Synthesis and crystallization

The title compound was obtained by the reaction of *N*-hydroxypicolinamide (0.156 g, 1 mmol, dissolved in 5 ml of water) with sodium hydrogen carbonate (1 M aqueous solution, 1 ml). Colorless crystals suitable for X-ray diffraction were obtained from the resulting aqueous solution by slow evaporation at ambient temperature within 48 h (yield 78%).

6. Refinement

Crystal data, data collection and structure refinement details are summarized in Table 2. All hydrogen atoms were found in the difference Fourier maps; H atoms of pyridine rings were constrained to ride on their parent atoms with C–H = 0.93 Å and *U*_{iso} = 1.2*U*_{eq}(C), and H atoms of the N–H and O–H groups were refined isotropically.

Acknowledgements

Financial support from the European Community's Seventh Framework Program (FP7/2007–2013) under grant agreement

Table 2
Experimental details.

Crystal data	
Chemical formula	[Na ₂ (C ₆ H ₅ N ₂ O ₂) ₂ (C ₆ H ₆ N ₂ O ₂) ₄]
<i>M_r</i>	872.73
Crystal system, space group	Triclinic, <i>P</i> $\bar{1}$
Temperature (K)	298
<i>a</i> , <i>b</i> , <i>c</i> (Å)	9.7997 (7), 10.0959 (7), 11.0401 (8)
α , β , γ (°)	96.618 (6), 102.741 (6), 113.902 (7)
<i>V</i> (Å ³)	948.02 (13)
<i>Z</i>	1
Radiation type	Mo <i>K</i> α
μ (mm ⁻¹)	0.14
Crystal size (mm)	0.3 × 0.3 × 0.3
Data collection	
Diffractometer	Agilent Xcalibur Sapphire3
Absorption correction	Multi-scan (<i>CrysAlis PRO</i> ; Agilent, 2013)
<i>T</i> _{min} , <i>T</i> _{max}	0.965, 1.000
No. of measured, independent and observed [<i>I</i> > 2 σ (<i>I</i>)] reflections	12244, 5514, 3066
<i>R</i> _{int}	0.032
(<i>sin</i> θ / λ) _{max} (Å ⁻¹)	0.703
Refinement	
<i>R</i> [<i>F</i> ² > 2 σ (<i>F</i> ²)], <i>wR</i> (<i>F</i> ²), <i>S</i>	0.064, 0.144, 0.96
No. of reflections	5521
No. of parameters	300
H-atom treatment	H atoms treated by a mixture of independent and constrained refinement
$\Delta\rho_{\text{max}}$, $\Delta\rho_{\text{min}}$ (e Å ⁻³)	0.23, -0.23

Computer programs: *CrysAlis PRO* (Agilent, 2013), *SHELXT* (Sheldrick, 2015a), *SHELXL2014* (Sheldrick, 2015b), *OLEX2* (Dolomanov *et al.*, 2009).

PIRSES-GA-2013–611488 is gratefully acknowledged. KAO acknowledges for the DAAD fellowship (Leonhard-Euler-Program).

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supporting information

Acta Cryst. (2017). E73, 24-27 [https://doi.org/10.1107/S2056989016019095]

Crystal structure of bis(μ -*N*-hydroxypicolinamidato)bis[bis(*N*-hydroxypicolinamide)sodium]

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Computing details

Data collection: *CrysAlis PRO* (Agilent, 2013); cell refinement: *CrysAlis PRO* (Agilent, 2013); data reduction: *CrysAlis PRO* (Agilent, 2013); program(s) used to solve structure: SHELXT (Sheldrick, 2015a); program(s) used to refine structure: *SHELXL2014* (Sheldrick, 2015b); molecular graphics: *OLEX2* (Dolomanov *et al.*, 2009); software used to prepare material for publication: *OLEX2* (Dolomanov *et al.*, 2009).

Bis(μ -*N*-hydroxypicolinamidato)bis[bis(*N*-hydroxypicolinamide)sodium]

Crystal data

[Na₂(C₆H₅N₂O₂)₂(C₆H₆N₂O₂)₄]

$M_r = 872.73$

Triclinic, $P\bar{1}$

$a = 9.7997$ (7) Å

$b = 10.0959$ (7) Å

$c = 11.0401$ (8) Å

$\alpha = 96.618$ (6)°

$\beta = 102.741$ (6)°

$\gamma = 113.902$ (7)°

$V = 948.02$ (13) Å³

$Z = 1$

$F(000) = 452$

$D_x = 1.529$ Mg m⁻³

Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å

Cell parameters from 2896 reflections

$\theta = 3.2$ – 31.6 °

$\mu = 0.14$ mm⁻¹

$T = 298$ K

Block, clear colourless

$0.3 \times 0.3 \times 0.3$ mm

Data collection

Agilent Xcalibur Sapphire3
diffractometer

Radiation source: Enhance (Mo) X-ray Source
Graphite monochromator

Detector resolution: 16.1827 pixels mm⁻¹

ω scans

Absorption correction: multi-scan
(CrysAlis PRO; Agilent, 2013)

$T_{\min} = 0.965$, $T_{\max} = 1.000$

12244 measured reflections

5514 independent reflections

3066 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.032$

$\theta_{\max} = 30.0$ °, $\theta_{\min} = 3.0$ °

$h = -13 \rightarrow 13$

$k = -14 \rightarrow 14$

$l = -15 \rightarrow 15$

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.064$

$wR(F^2) = 0.144$

$S = 0.96$

5521 reflections

300 parameters

0 restraints

H atoms treated by a mixture of independent
and constrained refinement

$w = 1/[\sigma^2(F_o^2) + (0.0618P)^2]$

where $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} < 0.001$

$\Delta\rho_{\max} = 0.23$ e Å⁻³

$\Delta\rho_{\min} = -0.23$ e Å⁻³

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Na1	0.46101 (9)	0.10567 (8)	0.60562 (6)	0.0406 (2)
O1	0.23043 (17)	0.04309 (17)	0.66132 (12)	0.0517 (4)
O2	0.33418 (18)	0.01016 (16)	0.90258 (13)	0.0445 (4)
H2	0.426 (3)	0.053 (3)	0.883 (2)	0.066 (8)*
O3	0.56164 (19)	0.36282 (15)	0.66630 (13)	0.0555 (4)
O4	0.74368 (18)	0.39266 (18)	0.90321 (14)	0.0516 (4)
H4	0.681 (3)	0.293 (3)	0.876 (3)	0.107 (11)*
O5	0.39508 (16)	0.03347 (14)	0.38677 (12)	0.0413 (3)
O6	0.42798 (15)	−0.11024 (13)	0.17918 (12)	0.0374 (3)
N1	0.2660 (2)	0.10347 (19)	0.87255 (16)	0.0397 (4)
H1	0.250 (3)	0.145 (3)	0.930 (2)	0.075 (9)*
N2	0.1263 (2)	0.28298 (18)	0.84096 (15)	0.0445 (4)
N3	0.6408 (2)	0.45461 (19)	0.87873 (17)	0.0428 (4)
H3	0.643 (3)	0.517 (3)	0.940 (2)	0.068 (8)*
N4	0.4699 (2)	0.60614 (18)	0.85596 (16)	0.0454 (4)
N5	0.34178 (18)	−0.03212 (17)	0.17342 (15)	0.0325 (4)
H5	0.301 (2)	−0.018 (2)	0.104 (2)	0.049 (6)*
N6	0.16908 (19)	0.11506 (18)	0.13287 (14)	0.0391 (4)
C1	0.1203 (2)	0.19678 (19)	0.73700 (17)	0.0327 (4)
C2	0.0281 (3)	0.1793 (3)	0.61794 (19)	0.0515 (6)
H2A	0.0291	0.1203	0.5472	0.062*
C3	−0.0661 (3)	0.2507 (3)	0.6045 (2)	0.0596 (6)
H3A	−0.1316	0.2385	0.5248	0.072*
C4	−0.0623 (2)	0.3393 (2)	0.7094 (2)	0.0483 (5)
H4A	−0.1247	0.3889	0.7029	0.058*
C5	0.0357 (3)	0.3532 (2)	0.8243 (2)	0.0522 (6)
H5A	0.0399	0.4154	0.8954	0.063*
C6	0.2119 (2)	0.1080 (2)	0.75329 (17)	0.0337 (4)
C7	0.4699 (2)	0.52978 (19)	0.74923 (18)	0.0358 (4)
C8	0.3960 (3)	0.5331 (3)	0.6301 (2)	0.0583 (6)
H8	0.3977	0.4771	0.5579	0.070*
C9	0.3190 (3)	0.6209 (3)	0.6188 (2)	0.0679 (7)
H9	0.2696	0.6265	0.5387	0.082*
C10	0.3159 (3)	0.6994 (2)	0.7263 (2)	0.0515 (6)
H10	0.2635	0.7584	0.7214	0.062*
C11	0.3925 (3)	0.6887 (2)	0.8421 (2)	0.0502 (6)
H11	0.3905	0.7424	0.9155	0.060*
C12	0.5611 (2)	0.44067 (19)	0.76097 (18)	0.0371 (4)
C13	0.2335 (2)	0.11871 (18)	0.25376 (16)	0.0301 (4)

C14	0.2116 (3)	0.1935 (3)	0.3547 (2)	0.0540 (6)
H14	0.2580	0.1948	0.4382	0.065*
C15	0.1199 (3)	0.2665 (3)	0.3297 (2)	0.0595 (6)
H15	0.1045	0.3184	0.3964	0.071*
C16	0.0524 (2)	0.2621 (2)	0.2072 (2)	0.0431 (5)
H16	-0.0103	0.3101	0.1881	0.052*
C17	0.0792 (2)	0.1845 (2)	0.11211 (19)	0.0465 (5)
H17	0.0316	0.1803	0.0280	0.056*
C18	0.3310 (2)	0.03654 (18)	0.27692 (16)	0.0299 (4)

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Na1	0.0567 (5)	0.0543 (5)	0.0255 (4)	0.0397 (4)	0.0115 (3)	0.0068 (3)
O1	0.0587 (9)	0.0844 (10)	0.0293 (8)	0.0518 (9)	0.0114 (7)	0.0031 (7)
O2	0.0571 (10)	0.0604 (9)	0.0423 (8)	0.0450 (8)	0.0217 (7)	0.0219 (7)
O3	0.0912 (12)	0.0558 (9)	0.0339 (8)	0.0507 (9)	0.0141 (8)	0.0025 (7)
O4	0.0543 (9)	0.0553 (9)	0.0494 (9)	0.0368 (8)	0.0044 (7)	0.0007 (7)
O5	0.0575 (9)	0.0562 (8)	0.0241 (7)	0.0395 (7)	0.0098 (6)	0.0101 (6)
O6	0.0486 (8)	0.0475 (7)	0.0333 (7)	0.0374 (7)	0.0128 (6)	0.0101 (6)
N1	0.0550 (11)	0.0544 (10)	0.0294 (9)	0.0410 (9)	0.0154 (8)	0.0121 (8)
N2	0.0551 (11)	0.0550 (10)	0.0325 (9)	0.0366 (9)	0.0089 (8)	0.0040 (8)
N3	0.0558 (11)	0.0469 (10)	0.0349 (10)	0.0347 (9)	0.0103 (8)	0.0035 (8)
N4	0.0634 (12)	0.0487 (9)	0.0362 (10)	0.0343 (9)	0.0198 (8)	0.0082 (8)
N5	0.0432 (9)	0.0423 (9)	0.0234 (8)	0.0307 (8)	0.0079 (7)	0.0079 (7)
N6	0.0462 (10)	0.0531 (10)	0.0313 (9)	0.0343 (8)	0.0124 (7)	0.0101 (7)
C1	0.0354 (10)	0.0400 (10)	0.0292 (10)	0.0206 (8)	0.0130 (8)	0.0101 (8)
C2	0.0690 (15)	0.0727 (14)	0.0288 (11)	0.0515 (13)	0.0081 (10)	0.0065 (10)
C3	0.0722 (16)	0.0826 (16)	0.0386 (13)	0.0562 (14)	0.0033 (11)	0.0097 (11)
C4	0.0554 (13)	0.0542 (12)	0.0515 (14)	0.0399 (11)	0.0139 (11)	0.0153 (10)
C5	0.0673 (15)	0.0583 (13)	0.0402 (12)	0.0424 (12)	0.0105 (11)	-0.0001 (10)
C6	0.0342 (10)	0.0436 (10)	0.0273 (9)	0.0216 (9)	0.0092 (8)	0.0061 (8)
C7	0.0418 (11)	0.0306 (9)	0.0348 (10)	0.0162 (8)	0.0118 (8)	0.0049 (8)
C8	0.0881 (18)	0.0653 (14)	0.0330 (12)	0.0537 (14)	0.0065 (11)	0.0004 (10)
C9	0.0954 (19)	0.0793 (16)	0.0420 (14)	0.0639 (16)	-0.0008 (13)	0.0063 (12)
C10	0.0547 (14)	0.0495 (12)	0.0598 (15)	0.0335 (11)	0.0152 (11)	0.0112 (11)
C11	0.0666 (15)	0.0522 (12)	0.0471 (13)	0.0377 (12)	0.0257 (11)	0.0072 (10)
C12	0.0485 (12)	0.0324 (9)	0.0336 (11)	0.0202 (9)	0.0144 (9)	0.0065 (8)
C13	0.0333 (10)	0.0323 (9)	0.0279 (9)	0.0171 (8)	0.0104 (7)	0.0067 (7)
C14	0.0792 (16)	0.0755 (15)	0.0311 (11)	0.0598 (14)	0.0132 (10)	0.0079 (10)
C15	0.0878 (18)	0.0793 (16)	0.0399 (13)	0.0650 (15)	0.0215 (12)	0.0066 (11)
C16	0.0480 (12)	0.0487 (11)	0.0483 (13)	0.0333 (10)	0.0191 (10)	0.0143 (10)
C17	0.0543 (13)	0.0663 (13)	0.0350 (11)	0.0420 (11)	0.0115 (10)	0.0161 (10)
C18	0.0342 (10)	0.0327 (9)	0.0270 (9)	0.0183 (8)	0.0101 (7)	0.0063 (7)

Geometric parameters (Å, °)

Na1—O5	2.3044 (14)	C1—C2	1.367 (3)
Na1—O3	2.3225 (15)	C1—C6	1.503 (2)
Na1—O1	2.3300 (16)	C2—C3	1.377 (3)
Na1—O5 ⁱ	2.3558 (14)	C2—H2A	0.9300
Na1—O6 ⁱ	2.3716 (14)	C3—C4	1.363 (3)
Na1—Na1 ⁱ	3.3964 (13)	C3—H3A	0.9300
O1—C6	1.232 (2)	C4—C5	1.365 (3)
O2—N1	1.388 (2)	C4—H4A	0.9300
O2—H2	0.91 (2)	C5—H5A	0.9300
O3—C12	1.235 (2)	C7—C8	1.366 (3)
O4—N3	1.383 (2)	C7—C12	1.501 (3)
O4—H4	0.91 (3)	C8—C9	1.378 (3)
O5—C18	1.247 (2)	C8—H8	0.9300
O5—Na1 ⁱ	2.3558 (14)	C9—C10	1.362 (3)
O6—N5	1.3666 (18)	C9—H9	0.9300
O6—Na1 ⁱ	2.3716 (14)	C10—C11	1.372 (3)
N1—C6	1.320 (2)	C10—H10	0.9300
N1—H1	0.80 (2)	C11—H11	0.9300
N2—C1	1.333 (2)	C13—C14	1.379 (3)
N2—C5	1.339 (2)	C13—C18	1.500 (2)
N3—C12	1.319 (3)	C14—C15	1.379 (3)
N3—H3	0.86 (2)	C14—H14	0.9300
N4—C7	1.332 (2)	C15—C16	1.355 (3)
N4—C11	1.336 (2)	C15—H15	0.9300
N5—C18	1.314 (2)	C16—C17	1.372 (3)
N5—H5	0.84 (2)	C16—H16	0.9300
N6—C17	1.330 (2)	C17—H17	0.9300
N6—C13	1.334 (2)		
O5—Na1—O3	109.39 (6)	N2—C5—C4	124.16 (19)
O5—Na1—O1	107.77 (5)	N2—C5—H5A	117.9
O3—Na1—O1	98.10 (6)	C4—C5—H5A	117.9
O5—Na1—O5 ⁱ	86.43 (5)	O1—C6—N1	123.81 (17)
O3—Na1—O5 ⁱ	126.64 (6)	O1—C6—C1	121.78 (17)
O1—Na1—O5 ⁱ	125.87 (6)	N1—C6—C1	114.39 (16)
O5—Na1—O6 ⁱ	156.53 (5)	N4—C7—C8	123.26 (18)
O3—Na1—O6 ⁱ	87.61 (5)	N4—C7—C12	118.11 (17)
O1—Na1—O6 ⁱ	84.80 (5)	C8—C7—C12	118.58 (17)
O5 ⁱ —Na1—O6 ⁱ	70.26 (5)	C7—C8—C9	118.8 (2)
C6—O1—Na1	127.13 (13)	C7—C8—H8	120.6
N1—O2—H2	102.4 (15)	C9—C8—H8	120.6
C12—O3—Na1	129.55 (13)	C10—C9—C8	119.3 (2)
N3—O4—H4	103.9 (19)	C10—C9—H9	120.4
C18—O5—Na1	151.93 (12)	C8—C9—H9	120.4
C18—O5—Na1 ⁱ	114.42 (11)	C9—C10—C11	118.0 (2)
Na1—O5—Na1 ⁱ	93.57 (5)	C9—C10—H10	121.0

N5—O6—Na ⁱ	110.28 (9)	C11—C10—H10	121.0
C6—N1—O2	121.64 (15)	N4—C11—C10	124.06 (19)
C6—N1—H1	121.1 (19)	N4—C11—H11	118.0
O2—N1—H1	116.7 (19)	C10—C11—H11	118.0
C1—N2—C5	116.66 (17)	O3—C12—N3	123.70 (18)
C12—N3—O4	121.18 (16)	O3—C12—C7	121.61 (18)
C12—N3—H3	119.5 (17)	N3—C12—C7	114.69 (16)
O4—N3—H3	118.2 (17)	N6—C13—C14	122.09 (17)
C7—N4—C11	116.62 (18)	N6—C13—C18	117.42 (15)
C18—N5—O6	121.82 (15)	C14—C13—C18	120.49 (17)
C18—N5—H5	116.6 (14)	C13—C14—C15	118.87 (19)
O6—N5—H5	121.3 (14)	C13—C14—H14	120.6
C17—N6—C13	117.46 (16)	C15—C14—H14	120.6
N2—C1—C2	123.00 (17)	C16—C15—C14	119.55 (19)
N2—C1—C6	118.12 (16)	C16—C15—H15	120.2
C2—C1—C6	118.77 (16)	C14—C15—H15	120.2
C1—C2—C3	118.83 (19)	C15—C16—C17	118.02 (18)
C1—C2—H2A	120.6	C15—C16—H16	121.0
C3—C2—H2A	120.6	C17—C16—H16	121.0
C4—C3—C2	119.3 (2)	N6—C17—C16	124.00 (19)
C4—C3—H3A	120.3	N6—C17—H17	118.0
C2—C3—H3A	120.3	C16—C17—H17	118.0
C3—C4—C5	117.99 (19)	O5—C18—N5	123.20 (16)
C3—C4—H4A	121.0	O5—C18—C13	121.78 (15)
C5—C4—H4A	121.0	N5—C18—C13	115.02 (15)
Na ⁱ —O6—N5—C18	1.8 (2)	Na ⁱ —O3—C12—C7	-113.54 (18)
C5—N2—C1—C2	0.4 (3)	O4—N3—C12—O3	6.3 (3)
C5—N2—C1—C6	-175.88 (18)	O4—N3—C12—C7	-172.72 (16)
N2—C1—C2—C3	-1.8 (3)	N4—C7—C12—O3	177.27 (19)
C6—C1—C2—C3	174.5 (2)	C8—C7—C12—O3	-5.1 (3)
C1—C2—C3—C4	1.6 (4)	N4—C7—C12—N3	-3.7 (3)
C2—C3—C4—C5	-0.1 (4)	C8—C7—C12—N3	173.90 (19)
C1—N2—C5—C4	1.2 (3)	C17—N6—C13—C14	-1.3 (3)
C3—C4—C5—N2	-1.4 (4)	C17—N6—C13—C18	178.22 (17)
Na ⁱ —O1—C6—N1	-72.4 (2)	N6—C13—C14—C15	0.2 (3)
Na ⁱ —O1—C6—C1	109.31 (18)	C18—C13—C14—C15	-179.3 (2)
O2—N1—C6—O1	-7.4 (3)	C13—C14—C15—C16	0.6 (4)
O2—N1—C6—C1	170.99 (16)	C14—C15—C16—C17	-0.3 (4)
N2—C1—C6—O1	-169.34 (18)	C13—N6—C17—C16	1.7 (3)
C2—C1—C6—O1	14.2 (3)	C15—C16—C17—N6	-0.9 (3)
N2—C1—C6—N1	12.2 (3)	Na ⁱ —O5—C18—N5	175.36 (18)
C2—C1—C6—N1	-164.23 (19)	Na ⁱ —O5—C18—N5	0.2 (2)
C11—N4—C7—C8	0.0 (3)	Na ⁱ —O5—C18—C13	-5.3 (4)
C11—N4—C7—C12	177.53 (17)	Na ⁱ —O5—C18—C13	179.48 (12)
N4—C7—C8—C9	0.8 (4)	O6—N5—C18—O5	-1.4 (3)
C12—C7—C8—C9	-176.6 (2)	O6—N5—C18—C13	179.24 (14)
C7—C8—C9—C10	-1.3 (4)	N6—C13—C18—O5	-179.72 (17)

C8—C9—C10—C11	0.9 (4)	C14—C13—C18—O5	-0.2 (3)
C7—N4—C11—C10	-0.5 (3)	N6—C13—C18—N5	-0.4 (2)
C9—C10—C11—N4	0.0 (4)	C14—C13—C18—N5	179.15 (19)
Na1—O3—C12—N3	67.5 (3)		

Symmetry code: (i) $-x+1, -y, -z+1$.

Hydrogen-bond geometry (Å, °)

<i>D—H...A</i>	<i>D—H</i>	<i>H...A</i>	<i>D...A</i>	<i>D—H...A</i>
N1—H1...N2	0.80 (3)	2.35 (3)	2.688 (3)	106 (2)
N3—H3...N4	0.86 (3)	2.30 (3)	2.681 (3)	107 (2)
N5—H5...N6	0.84 (2)	2.25 (2)	2.670 (3)	111.1 (17)
O2—H2...O6 ⁱ	0.91 (2)	1.65 (2)	2.549 (2)	169 (2)
O4—H4...O6 ⁱ	0.91 (3)	1.66 (3)	2.5744 (19)	177 (3)
N1—H1...N6 ⁱⁱ	0.80 (2)	2.55 (3)	3.224 (2)	143 (2)
N5—H5...O2 ⁱⁱⁱ	0.84 (2)	2.35 (2)	3.058 (2)	142.6 (19)
C5—H5A...O4 ^{iv}	0.93	2.61	3.341 (3)	136
C17—H17...N2 ⁱⁱⁱ	0.93	2.60	3.330 (3)	136

Symmetry codes: (i) $-x+1, -y, -z+1$; (ii) $x, y, z+1$; (iii) $x, y, z-1$; (iv) $-x+1, -y+1, -z+2$.