



Crystal structure of octane-1,8-diaminium 4,4'-(diazene-1,2-diyl)dibenzoate monohydrate

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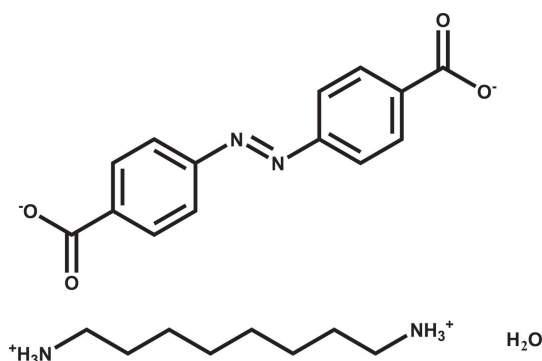
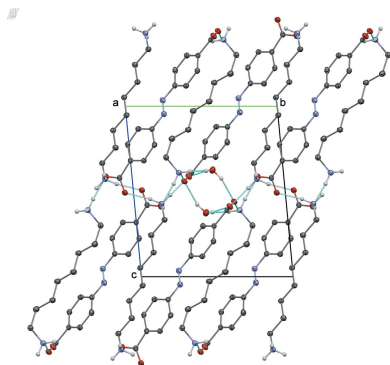
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The title salt, $C_8H_{22}N_2^{2+} \cdot C_{14}H_8N_2O_4^{2-} \cdot H_2O$, represents a pseudo-polymer ionic material, resulting from the self-organizing behavior of 4,4'-azinodibenzoate dianions and doubly protonated, 1,8-diaminium-octane cations in aqueous solution. The asymmetric unit consists of two halves of octane 1,8-diaminium cations (the complete cations are both generated by crystallographic inversion symmetry), a 4,4'-azinodibenzoate anion [dihedral angle between the aromatic rings = $10.22(4)^\circ$] and a water molecule of crystallization. One of the cations is in a fully extended linear conformation while the second one has a terminal C—C—N *gauche* conformation. In the crystal, the cations, anions and water molecules are linked into a three-dimensional network *via* a complex pattern of charge-assisted N—H...O and O—H...O hydrogen bonds.

1. Chemical context

Ionic pseudo-polymers auto-assembled from oppositely charged organic molecules are of emerging interest for many potential application fields (Webber *et al.*, 2016; Mann, 2009). It is reasonable to expect that the presence of azobenzene moieties in such materials creates the necessary precondition for achieving their reversible photosensitivity (Bushuyev *et al.*, 2016, 2018). In this context, we report the synthesis and structure of octane-1,8-diaminium 4,4'-(diazene-1,2-diyl)dibenzoate monohydrate, (I), formed by the crystallization of bianionic 4,4'-azinodibenzoic acid and bicationic 1,8-diaminooctane in aqueous solution.



2. Structural commentary

The asymmetric unit (Fig. 1) consists of two halves of octane 1,8-diaminium dications, one 4,4'-azinodibenzoic dianion and one water molecule. Bond lengths and angles of the dication and dianion are in the expected ranges. One of the octane

Table 1
Selected torsion angles ($^{\circ}$).

C2—C3—C4—C5	-177.97 (9)	C7—C8—C9—C10	-179.03 (9)
C3—C4—C5—C5 ⁱ	178.25 (10)	C8—C9—C10—C10 ⁱⁱ	178.30 (11)
N1—C2—C3—C4	177.78 (9)	N6—C7—C8—C9	-76.89 (12)

Symmetry codes: (i) $-x + 1, -y, -z + 2$; (ii) $-x, -y + 1, -z$.

Table 2
Hydrogen-bond geometry ($\text{\AA}, ^{\circ}$).

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
N1—H1A...O30 ⁱⁱⁱ	0.895 (17)	1.897 (17)	2.7796 (13)	168.3 (14)
N1—H1B...O13	0.916 (16)	1.881 (17)	2.7942 (12)	175.0 (14)
N1—H1C...O29 ^{iv}	0.918 (16)	1.970 (16)	2.8579 (12)	162.5 (13)
N6—H6A...O31 ^v	0.921 (16)	1.912 (16)	2.8296 (13)	174.0 (13)
N6—H6B...O11	0.921 (16)	1.879 (16)	2.7735 (12)	163.2 (13)
N6—H6C...O29 ^{vi}	0.947 (16)	1.849 (17)	2.7847 (12)	169.0 (14)
N6—H6C...O30 ^{vi}	0.947 (16)	2.590 (16)	3.0880 (12)	113.2 (11)
O31—H31A...O13 ^{vii}	0.89 (2)	1.88 (2)	2.7418 (11)	165.2 (17)
O31—H31B...O11	0.88 (2)	1.97 (2)	2.8184 (11)	162.2 (16)

Symmetry codes: (iii) $-x + 1, -y + 1, -z$; (iv) $x + 1, y - 1, z + 1$; (v) $-x, -y + 1, -z + 1$; (vi) $x, y - 1, z + 1$; (vii) $-x + 1, -y + 1, -z + 1$.

1,8-diaminium dication shows a fully extended *all-trans* conformation with torsion angles close to 180° (Table 1). The second cation has its two terminal torsion angles N6—C7—C8—C9 synclinal with a value of $-76.89 (12)^{\circ}$. The fully extended conformation corresponds to the most stable one, compared to the arrangement with synclinal torsion angles, as shown from DFT calculations and a database survey performed on α,ω -alkyldiaminium cations (Brozdowska & Chojnacki, 2017). The less energetically favorable *gauche* conformation is presumably stabilized by the charge-assisted hydrogen-bonded network.

The geometry of the 4,4'-azinodibenzoic dianion is characterized by the presence of two benzoic acid residues linked *via* a *trans*-configured azo group is consistent with known data (Fernández *et al.*, 2015; Sahoo *et al.*, 2012). The angle between the phenyl rings of $10.22 (4)^{\circ}$ is consistent with a small deviation from planarity of the azobenzene moiety. The carboxylate groups are inclined to the molecular mean plane at angles of $38.40 (3)^{\circ}$ (O11/C12/O13) and $16.53 (5)^{\circ}$ (O29/C28/O30).

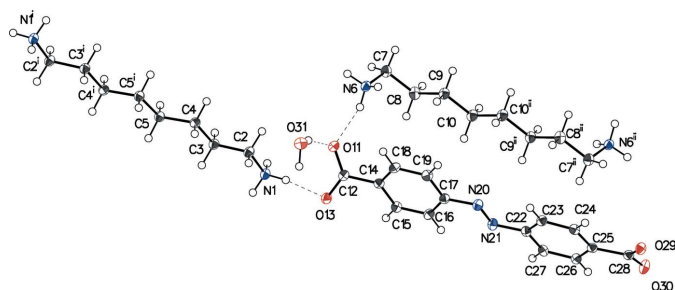


Figure 1
Molecular structure and atom-labelling scheme of (I). Displacement ellipsoids are drawn at the 50% probability level and hydrogen bonds are shown as dotted lines. [Symmetry codes: (i) $1 - x, -y, 2 - z$; (ii) $-x, 1 - y, -z$.]

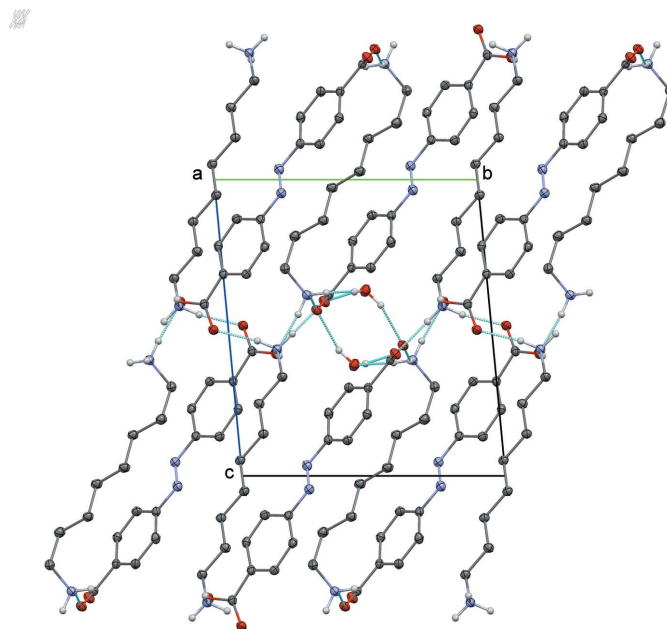


Figure 2
Projection along the *b*-axis direction showing the packing in layers consolidated by the hydrogen-bond network (dotted lines). Hydrogen atoms not involved in hydrogen bonds and hanging hydrogen bonds are omitted for clarity.

3. Supramolecular features

In addition to the electrostatic interactions, the anions and cations are connected by charge-assisted N—H...O hydrogen bonds (Table 2). The complex pattern of hydrogen bonds also includes the water molecules. Therefore, the 4,4'-azinodibenzoic dianion is linked through hydrogen bonds with three cations on one side and with two cations and two water molecules on the other side. Anions and cation stack in two-dimensional arrays in the *ab* plane separated by a zone with

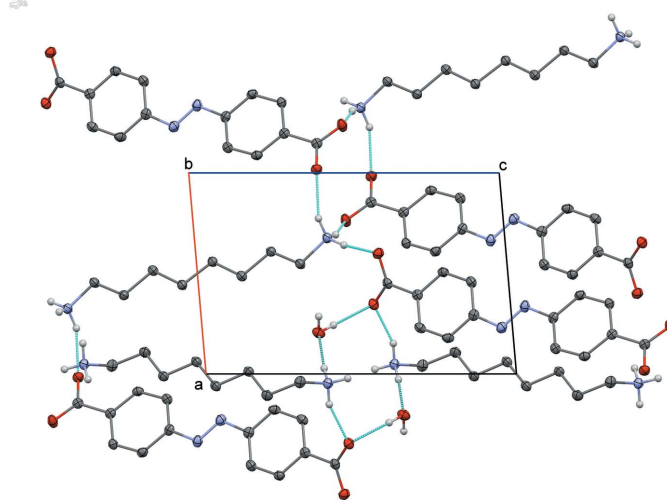


Figure 3
Partial packing view along the *b*-axis direction showing the $R_4^1(12)$ graph-set motifs. Hydrogen atoms not involved in hydrogen bonds and hanging hydrogen bonds have been omitted for clarity.

Table 3
Experimental details.

Crystal data	
Chemical formula	$C_8H_{22}N_2^{2+} \cdot C_{14}H_8N_2O_4^{2-} \cdot H_2O$
M_r	432.51
Crystal system, space group	Triclinic, $P\bar{1}$
Temperature (K)	100
a, b, c (Å)	8.3604 (1), 11.4100 (2), 12.4661 (2)
α, β, γ (°)	83.001 (1), 83.364 (1), 73.973 (1)
V (Å ³)	1130.24 (3)
Z	2
Radiation type	Cu $K\alpha$
μ (mm ⁻¹)	0.75
Crystal size (mm)	0.4 × 0.4 × 0.1
Data collection	
Diffractionmeter	Bruker APEXII CCD
Absorption correction	Multi-scan (<i>SADABS</i> ; Krause <i>et al.</i> , 2015)
T_{\min} , T_{\max}	0.525, 0.753
No. of measured, independent and observed [$I > 2\sigma(I)$] reflections	28413, 4048, 3920
R_{int}	0.046
$(\sin \theta/\lambda)_{\text{max}}$ (Å ⁻¹)	0.602
Refinement	
$R[F^2 > 2\sigma(F^2)]$, $wR(F^2)$, S	0.037, 0.096, 1.05
No. of reflections	4048
No. of parameters	313
H-atom treatment	H atoms treated by a mixture of independent and constrained refinement
$\Delta\rho_{\text{max}}$, $\Delta\rho_{\text{min}}$ (e Å ⁻³)	0.20, -0.30

Computer programs: *APEX2* and *SAINT* (Bruker, 2013), *SHELXT* (Sheldrick, 2015a), *SHELXL2018* (Sheldrick, 2015b), *OLEX2* (Dolomanov *et al.*, 2009), *Mercury* (Macrae *et al.*, 2008) and *publCIF* (Westrip, 2010).

the hydrogen-bonded network involving the ionized amino and carboxylic groups and the water molecules (Fig. 2). This network contains two 12-membered rings comprising either two cations and two anions or two cations, two anions and two water molecules (Fig. 3), according to the graph set descriptor $R_4^4(12)$ (Etter *et al.*, 1990).

4. Database survey

A search in the Cambridge Structural Database (Version 5.39 with one update; Groom *et al.*, 2016) returned 48 entries for octane-1,8-diaminium compounds. These include simple halide salts (Brisson & Brisse, 1984; van Blerk & Kruger, 2007; van Megen & Reiss, 2013); metal halide salts (Kessentini *et al.*, 2011) comprising lead halide complexes (Lemmerer & Billing, 2012; Smith *et al.*, 2017), and more complex systems where the diaminium cations are encapsulated in a macrocycle (Kim *et al.*, 2009; Yu *et al.*, 2014). A similar search for 4,4'-azinodibenzoic acid and its salts returned 43 entries, including the structure of the simple acid (Yu & Liu, 2009). The dianion has been also used as linker to prepare MOF or coordination frameworks (see, for example, Hou *et al.*, 2013, Zhang *et al.*, 2016, Guo *et al.*, 2013 and Deng *et al.*, 2015), and co-crystallized to give gelator salts (Sahoo & Dastidar, 2012; Sahoo *et al.*, 2012) or supramolecular assemblies (Beatty *et al.*, 2002; Yu *et al.*, 2011).

5. Synthesis and crystallization

Crystals of the title compound were obtained by the dropwise addition with intensive stirring of 5 ml of 0.10 *M* aqueous 1,8-octamethylenediamine into 25 ml of 0.02 *M* aqueous 4,4'-dicarboxyazobenzene disodium salt at room temperature. The final solution (pH 12.5) was allowed to partly evaporate at room temperature and atmospheric pressure. The resulting orange oblong crystals in the form of thin narrow leaves up to 1 cm long were gently removed from the liquid phase and air-dried on filter paper.

6. Refinement

Crystal data, data collection and structure refinement details are summarized in Table 3. Hydrogen atoms bound to nitrogen or oxygen atoms were located from difference syntheses and refined without any restraints. Hydrogen atoms linked to carbon atoms were included using an appropriate riding model (AFIX 43 and AFIX 23 for aromatic and methylene hydrogen atoms respectively) with C–H = 0.95–0.99 Å and $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$.

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supporting information

Acta Cryst. (2018). E74, 724-727 [https://doi.org/10.1107/S2056989018006187]

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Computing details

Data collection: *APEX2* (Bruker, 2013); cell refinement: *SAINTE* (Bruker, 2013); data reduction: *SAINTE* (Bruker, 2013); program(s) used to solve structure: *SHELXT* (Sheldrick, 2015a); program(s) used to refine structure: *SHELXL2018* (Sheldrick, 2015b); molecular graphics: *OLEX2* (Dolomanov *et al.*, 2009) and *Mercury* (Macrae *et al.*, 2008); software used to prepare material for publication: *OLEX2* (Dolomanov *et al.*, 2009) and *publCIF* (Westrip, 2010).

Octane-1,8-diaminium 4,4'-(diazene-1,2-diyl)dibenzoate monohydrate

Crystal data

$C_8H_{22}N_2^{2+} \cdot C_{14}H_8N_2O_4^{2-} \cdot H_2O$

$M_r = 432.51$

Triclinic, $P\bar{1}$

$a = 8.3604$ (1) Å

$b = 11.4100$ (2) Å

$c = 12.4661$ (2) Å

$\alpha = 83.001$ (1)°

$\beta = 83.364$ (1)°

$\gamma = 73.973$ (1)°

$V = 1130.24$ (3) Å³

$Z = 2$

$F(000) = 464$

$D_x = 1.271$ Mg m⁻³

Cu $K\alpha$ radiation, $\lambda = 1.54184$ Å

Cell parameters from 3920 reflections

$\theta = 4.1$ – 68.1 °

$\mu = 0.75$ mm⁻¹

$T = 100$ K

Block, orange

$0.4 \times 0.4 \times 0.1$ mm

Data collection

Bruker APEXII CCD

diffractometer

Radiation source: fine-focus sealed tube

Graphite monochromator

φ and ω scans

Absorption correction: multi-scan

(SADABS; Krause *et al.*, 2015)

$T_{\min} = 0.525$, $T_{\max} = 0.753$

28413 measured reflections

4048 independent reflections

3920 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.046$

$\theta_{\max} = 68.1$ °, $\theta_{\min} = 4.1$ °

$h = -10 \rightarrow 10$

$k = -13 \rightarrow 13$

$l = -14 \rightarrow 14$

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.037$

$wR(F^2) = 0.096$

$S = 1.05$

4048 reflections

313 parameters

0 restraints

Primary atom site location: structure-invariant
direct methods

Hydrogen site location: mixed

H atoms treated by a mixture of independent
and constrained refinement

$$w = 1/[\sigma^2(F_o^2) + (0.0541P)^2 + 0.3252P]$$

$$\text{where } P = (F_o^2 + 2F_c^2)/3$$

$$(\Delta/\sigma)_{\max} = 0.001$$

$$\Delta\rho_{\max} = 0.20 \text{ e } \text{\AA}^{-3}$$

$$\Delta\rho_{\min} = -0.30 \text{ e } \text{\AA}^{-3}$$

Extinction correction: SHELXL2018

(Sheldrick, 2015b),

$$F_c^* = kFc[1 + 0.001x\lambda^3/\sin(2\theta)]^{-1/4}$$

Extinction coefficient: 0.0427 (17)

Special details

Geometry. All esds (except the esd in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell esds are taken into account individually in the estimation of esds in distances, angles and torsion angles; correlations between esds in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell esds is used for estimating esds involving l.s. planes.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
C12	0.44977 (13)	0.39551 (9)	0.39040 (8)	0.0164 (2)
C14	0.39242 (13)	0.48992 (10)	0.29711 (9)	0.0163 (2)
C15	0.50609 (13)	0.50353 (10)	0.20773 (9)	0.0180 (2)
H15	0.618450	0.454880	0.207597	0.022*
C16	0.45684 (14)	0.58671 (10)	0.11990 (9)	0.0187 (2)
H16	0.534114	0.593956	0.058933	0.022*
C17	0.29240 (14)	0.66026 (10)	0.12125 (9)	0.0175 (2)
C18	0.17852 (14)	0.64902 (10)	0.21041 (9)	0.0193 (2)
H18	0.067342	0.699946	0.211663	0.023*
C19	0.22842 (14)	0.56296 (10)	0.29746 (9)	0.0185 (2)
H19	0.150301	0.553905	0.357480	0.022*
C22	0.27923 (14)	0.83834 (10)	-0.13411 (9)	0.0169 (2)
C23	0.12601 (14)	0.92810 (10)	-0.13255 (9)	0.0182 (2)
H23	0.049801	0.933790	-0.069462	0.022*
C24	0.08731 (13)	1.00833 (10)	-0.22421 (9)	0.0174 (2)
H24	-0.015163	1.070596	-0.222999	0.021*
C25	0.19653 (13)	0.99937 (9)	-0.31875 (9)	0.0162 (2)
C26	0.35000 (14)	0.91128 (10)	-0.31874 (9)	0.0173 (2)
H26	0.426383	0.905667	-0.381738	0.021*
C27	0.39101 (13)	0.83176 (10)	-0.22644 (9)	0.0178 (2)
H27	0.496160	0.772422	-0.226335	0.021*
C28	0.14659 (13)	1.08221 (10)	-0.42094 (9)	0.0168 (2)
N20	0.23031 (12)	0.75035 (8)	0.03481 (8)	0.0199 (2)
N21	0.33560 (12)	0.75025 (8)	-0.04586 (7)	0.0194 (2)
O11	0.34625 (9)	0.34229 (7)	0.44062 (6)	0.0200 (2)
O13	0.59935 (9)	0.37427 (7)	0.41176 (6)	0.0209 (2)
O29	0.01715 (9)	1.17097 (7)	-0.41274 (6)	0.0204 (2)
O30	0.23647 (10)	1.05688 (7)	-0.50725 (6)	0.0234 (2)
C2	0.54646 (13)	0.18776 (10)	0.66601 (9)	0.0181 (2)
H2A	0.442430	0.177349	0.642576	0.022*
H2B	0.521051	0.269756	0.692620	0.022*
C3	0.60840 (13)	0.08987 (10)	0.75674 (9)	0.0183 (2)
H3A	0.715478	0.098563	0.776743	0.022*
H3B	0.630505	0.008358	0.729672	0.022*

C4	0.48635 (13)	0.09539 (10)	0.85797 (8)	0.0176 (2)
H4A	0.380562	0.083043	0.839346	0.021*
H4B	0.461117	0.177451	0.884428	0.021*
C5	0.55783 (13)	-0.00224 (10)	0.94788 (8)	0.0178 (2)
H5A	0.580205	-0.084114	0.921493	0.021*
H5B	0.665684	0.008692	0.964131	0.021*
N1	0.67578 (12)	0.17874 (9)	0.57222 (8)	0.0177 (2)
H1A	0.6930 (19)	0.1068 (16)	0.5450 (12)	0.033 (4)*
H1B	0.6457 (19)	0.2412 (15)	0.5188 (12)	0.031 (4)*
H1C	0.777 (2)	0.1830 (13)	0.5905 (12)	0.028 (4)*
C7	0.04089 (14)	0.23686 (10)	0.30051 (9)	0.0206 (3)
H7A	-0.076237	0.235667	0.296577	0.025*
H7B	0.107206	0.151394	0.317135	0.025*
C8	0.10729 (14)	0.28667 (10)	0.19086 (9)	0.0222 (3)
H8A	0.134670	0.221260	0.141092	0.027*
H8B	0.212435	0.306963	0.199818	0.027*
C9	-0.01161 (13)	0.39982 (10)	0.13799 (9)	0.0194 (2)
H9A	-0.117543	0.380517	0.129346	0.023*
H9B	-0.037363	0.466406	0.186456	0.023*
C10	0.06004 (14)	0.44447 (10)	0.02760 (9)	0.0210 (3)
H10A	0.089089	0.376931	-0.020020	0.025*
H10B	0.164268	0.465977	0.036703	0.025*
N6	0.04744 (12)	0.31029 (9)	0.39017 (7)	0.0178 (2)
H6A	-0.0289 (18)	0.3860 (14)	0.3860 (11)	0.026 (3)*
H6B	0.153 (2)	0.3200 (13)	0.3926 (12)	0.029 (4)*
H6C	0.0223 (19)	0.2670 (14)	0.4569 (13)	0.033 (4)*
O31	0.20434 (10)	0.46567 (8)	0.62560 (7)	0.0235 (2)
H31A	0.281 (2)	0.5069 (17)	0.6193 (14)	0.049 (5)*
H31B	0.241 (2)	0.4140 (17)	0.5755 (15)	0.048 (5)*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C12	0.0197 (6)	0.0127 (5)	0.0160 (5)	-0.0030 (4)	0.0010 (4)	-0.0038 (4)
C14	0.0187 (5)	0.0134 (5)	0.0176 (5)	-0.0050 (4)	-0.0011 (4)	-0.0029 (4)
C15	0.0165 (5)	0.0163 (5)	0.0192 (5)	-0.0014 (4)	-0.0007 (4)	-0.0022 (4)
C16	0.0201 (6)	0.0182 (5)	0.0170 (5)	-0.0046 (4)	0.0020 (4)	-0.0024 (4)
C17	0.0209 (6)	0.0135 (5)	0.0184 (5)	-0.0048 (4)	-0.0030 (4)	-0.0012 (4)
C18	0.0160 (5)	0.0161 (5)	0.0234 (6)	-0.0009 (4)	-0.0009 (4)	-0.0015 (4)
C19	0.0188 (5)	0.0161 (5)	0.0197 (5)	-0.0048 (4)	0.0029 (4)	-0.0017 (4)
C22	0.0203 (5)	0.0137 (5)	0.0174 (5)	-0.0041 (4)	-0.0041 (4)	-0.0023 (4)
C23	0.0193 (5)	0.0183 (5)	0.0168 (5)	-0.0053 (4)	0.0003 (4)	-0.0027 (4)
C24	0.0156 (5)	0.0147 (5)	0.0210 (5)	-0.0017 (4)	-0.0020 (4)	-0.0026 (4)
C25	0.0177 (5)	0.0136 (5)	0.0183 (5)	-0.0051 (4)	-0.0028 (4)	-0.0021 (4)
C26	0.0186 (5)	0.0171 (5)	0.0164 (5)	-0.0044 (4)	-0.0005 (4)	-0.0034 (4)
C27	0.0170 (5)	0.0155 (5)	0.0198 (5)	-0.0006 (4)	-0.0036 (4)	-0.0036 (4)
C28	0.0180 (5)	0.0147 (5)	0.0186 (5)	-0.0057 (4)	-0.0017 (4)	-0.0014 (4)
N20	0.0213 (5)	0.0176 (5)	0.0197 (5)	-0.0042 (4)	-0.0027 (4)	0.0006 (4)

N21	0.0234 (5)	0.0167 (5)	0.0170 (5)	-0.0031 (4)	-0.0027 (4)	-0.0015 (4)
O11	0.0222 (4)	0.0189 (4)	0.0190 (4)	-0.0074 (3)	-0.0003 (3)	0.0013 (3)
O13	0.0196 (4)	0.0197 (4)	0.0220 (4)	-0.0042 (3)	-0.0038 (3)	0.0033 (3)
O29	0.0190 (4)	0.0174 (4)	0.0211 (4)	-0.0002 (3)	-0.0014 (3)	0.0014 (3)
O30	0.0284 (4)	0.0192 (4)	0.0169 (4)	0.0010 (3)	0.0015 (3)	0.0007 (3)
C2	0.0176 (5)	0.0178 (5)	0.0170 (5)	-0.0032 (4)	0.0020 (4)	-0.0009 (4)
C3	0.0162 (5)	0.0195 (5)	0.0176 (5)	-0.0035 (4)	0.0001 (4)	-0.0003 (4)
C4	0.0174 (5)	0.0183 (5)	0.0168 (5)	-0.0049 (4)	0.0007 (4)	-0.0018 (4)
C5	0.0178 (5)	0.0186 (6)	0.0165 (6)	-0.0044 (4)	0.0001 (5)	-0.0020 (4)
N1	0.0176 (5)	0.0169 (5)	0.0163 (5)	-0.0023 (4)	0.0001 (4)	0.0008 (4)
C7	0.0246 (6)	0.0164 (5)	0.0200 (6)	-0.0033 (4)	-0.0032 (4)	-0.0017 (4)
C8	0.0199 (6)	0.0231 (6)	0.0200 (6)	0.0002 (4)	-0.0004 (4)	-0.0035 (4)
C9	0.0178 (5)	0.0213 (6)	0.0170 (5)	-0.0026 (4)	0.0007 (4)	-0.0023 (4)
C10	0.0193 (5)	0.0241 (6)	0.0177 (6)	-0.0037 (5)	0.0023 (4)	-0.0025 (5)
N6	0.0184 (5)	0.0160 (5)	0.0179 (5)	-0.0040 (4)	-0.0016 (4)	0.0008 (4)
O31	0.0213 (4)	0.0217 (4)	0.0262 (4)	-0.0041 (3)	0.0024 (3)	-0.0056 (3)

Geometric parameters (Å, °)

C12—C14	1.5103 (14)	C2—N1	1.4907 (13)
C12—O11	1.2578 (13)	C3—H3A	0.9900
C12—O13	1.2599 (14)	C3—H3B	0.9900
C14—C15	1.4011 (15)	C3—C4	1.5244 (14)
C14—C19	1.3939 (15)	C4—H4A	0.9900
C15—H15	0.9500	C4—H4B	0.9900
C15—C16	1.3800 (15)	C4—C5	1.5289 (15)
C16—H16	0.9500	C5—C5 ⁱ	1.525 (2)
C16—C17	1.3984 (16)	C5—H5A	0.9900
C17—C18	1.3951 (15)	C5—H5B	0.9900
C17—N20	1.4325 (14)	N1—H1A	0.895 (17)
C18—H18	0.9500	N1—H1B	0.916 (16)
C18—C19	1.3909 (16)	N1—H1C	0.918 (16)
C19—H19	0.9500	C7—H7A	0.9900
C22—C23	1.4018 (16)	C7—H7B	0.9900
C22—C27	1.3920 (15)	C7—C8	1.5196 (16)
C22—N21	1.4253 (14)	C7—N6	1.4905 (14)
C23—H23	0.9500	C8—H8A	0.9900
C23—C24	1.3839 (16)	C8—H8B	0.9900
C24—H24	0.9500	C8—C9	1.5247 (15)
C24—C25	1.4008 (15)	C9—H9A	0.9900
C25—C26	1.3944 (16)	C9—H9B	0.9900
C25—C28	1.5172 (15)	C9—C10	1.5228 (15)
C26—H26	0.9500	C10—C10 ⁱⁱ	1.527 (2)
C26—C27	1.3891 (15)	C10—H10A	0.9900
C27—H27	0.9500	C10—H10B	0.9900
C28—O29	1.2656 (13)	N6—H6A	0.921 (16)
C28—O30	1.2549 (13)	N6—H6B	0.921 (16)
N20—N21	1.2575 (14)	N6—H6C	0.947 (16)

C2—H2A	0.9900	O31—H31A	0.89 (2)
C2—H2B	0.9900	O31—H31B	0.88 (2)
C2—C3	1.5176 (15)		
O11—C12—C14	118.09 (9)	H3A—C3—H3B	107.7
O11—C12—O13	124.36 (10)	C4—C3—H3A	108.8
O13—C12—C14	117.54 (9)	C4—C3—H3B	108.8
C15—C14—C12	119.29 (9)	C3—C4—H4A	109.3
C19—C14—C12	121.45 (9)	C3—C4—H4B	109.3
C19—C14—C15	119.26 (10)	C3—C4—C5	111.41 (9)
C14—C15—H15	119.6	H4A—C4—H4B	108.0
C16—C15—C14	120.82 (10)	C5—C4—H4A	109.3
C16—C15—H15	119.6	C5—C4—H4B	109.3
C15—C16—H16	120.2	C4—C5—H5A	108.9
C15—C16—C17	119.54 (10)	C4—C5—H5B	108.9
C17—C16—H16	120.2	C5 ⁱ —C5—C4	113.31 (11)
C16—C17—N20	123.64 (10)	C5 ⁱ —C5—H5A	108.9
C18—C17—C16	120.26 (10)	C5 ⁱ —C5—H5B	108.9
C18—C17—N20	116.10 (9)	H5A—C5—H5B	107.7
C17—C18—H18	120.1	C2—N1—H1A	109.6 (10)
C19—C18—C17	119.74 (10)	C2—N1—H1B	112.0 (9)
C19—C18—H18	120.1	C2—N1—H1C	113.2 (9)
C14—C19—H19	119.8	H1A—N1—H1B	109.5 (13)
C18—C19—C14	120.36 (10)	H1A—N1—H1C	107.0 (13)
C18—C19—H19	119.8	H1B—N1—H1C	105.4 (13)
C23—C22—N21	125.14 (10)	H7A—C7—H7B	107.8
C27—C22—C23	120.04 (10)	C8—C7—H7A	109.0
C27—C22—N21	114.81 (9)	C8—C7—H7B	109.0
C22—C23—H23	120.5	N6—C7—H7A	109.0
C24—C23—C22	119.07 (10)	N6—C7—H7B	109.0
C24—C23—H23	120.5	N6—C7—C8	112.95 (9)
C23—C24—H24	119.4	C7—C8—H8A	108.5
C23—C24—C25	121.19 (10)	C7—C8—H8B	108.5
C25—C24—H24	119.4	C7—C8—C9	114.99 (9)
C24—C25—C28	120.42 (9)	H8A—C8—H8B	107.5
C26—C25—C24	119.24 (10)	C9—C8—H8A	108.5
C26—C25—C28	120.32 (10)	C9—C8—H8B	108.5
C25—C26—H26	120.1	C8—C9—H9A	109.0
C27—C26—C25	119.89 (10)	C8—C9—H9B	109.0
C27—C26—H26	120.1	H9A—C9—H9B	107.8
C22—C27—H27	119.7	C10—C9—C8	112.75 (9)
C26—C27—C22	120.51 (10)	C10—C9—H9A	109.0
C26—C27—H27	119.7	C10—C9—H9B	109.0
O29—C28—C25	117.57 (9)	C9—C10—C10 ⁱⁱ	113.21 (11)
O30—C28—C25	117.37 (9)	C9—C10—H10A	108.9
O30—C28—O29	125.06 (10)	C9—C10—H10B	108.9
N21—N20—C17	112.88 (9)	C10 ⁱⁱ —C10—H10A	108.9
N20—N21—C22	114.90 (9)	C10 ⁱⁱ —C10—H10B	108.9

H2A—C2—H2B	108.1	H10A—C10—H10B	107.7
C3—C2—H2A	109.6	C7—N6—H6A	112.5 (9)
C3—C2—H2B	109.6	C7—N6—H6B	111.9 (9)
N1—C2—H2A	109.6	C7—N6—H6C	108.3 (9)
N1—C2—H2B	109.6	H6A—N6—H6B	109.4 (13)
N1—C2—C3	110.15 (9)	H6A—N6—H6C	107.9 (12)
C2—C3—H3A	108.8	H6B—N6—H6C	106.5 (13)
C2—C3—H3B	108.8	H31A—O31—H31B	102.9 (16)
C2—C3—C4	113.74 (9)		
C12—C14—C15—C16	178.35 (9)	C25—C26—C27—C22	0.71 (16)
C12—C14—C19—C18	-179.80 (9)	C26—C25—C28—O29	-170.63 (9)
C14—C15—C16—C17	1.46 (16)	C26—C25—C28—O30	9.78 (15)
C15—C14—C19—C18	-0.49 (16)	C27—C22—C23—C24	0.85 (16)
C15—C16—C17—C18	-0.48 (16)	C27—C22—N21—N20	-175.22 (9)
C15—C16—C17—N20	179.03 (9)	C28—C25—C26—C27	-176.74 (9)
C16—C17—C18—C19	-0.97 (16)	N20—C17—C18—C19	179.48 (9)
C16—C17—N20—N21	4.90 (15)	N21—C22—C23—C24	179.78 (9)
C17—C18—C19—C14	1.45 (16)	N21—C22—C27—C26	178.99 (9)
C17—N20—N21—C22	179.88 (8)	O11—C12—C14—C15	-146.46 (10)
C18—C17—N20—N21	-175.57 (9)	O11—C12—C14—C19	32.86 (14)
C19—C14—C15—C16	-0.99 (16)	O13—C12—C14—C15	32.65 (14)
C22—C23—C24—C25	1.52 (16)	O13—C12—C14—C19	-148.03 (10)
C23—C22—C27—C26	-1.97 (16)	C2—C3—C4—C5	-177.97 (9)
C23—C22—N21—N20	5.80 (15)	C3—C4—C5—C5 ⁱ	178.25 (10)
C23—C24—C25—C26	-2.76 (16)	N1—C2—C3—C4	177.78 (9)
C23—C24—C25—C28	175.60 (9)	C7—C8—C9—C10	-179.03 (9)
C24—C25—C26—C27	1.62 (15)	C8—C9—C10—C10 ⁱⁱ	178.30 (11)
C24—C25—C28—O29	11.03 (14)	N6—C7—C8—C9	-76.89 (12)
C24—C25—C28—O30	-168.56 (10)		

Symmetry codes: (i) $-x+1, -y, -z+2$; (ii) $-x, -y+1, -z$.

Hydrogen-bond geometry (\AA , $^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
N1—H1A \cdots O30 ⁱⁱⁱ	0.895 (17)	1.897 (17)	2.7796 (13)	168.3 (14)
N1—H1B \cdots O13	0.916 (16)	1.881 (17)	2.7942 (12)	175.0 (14)
N1—H1C \cdots O29 ^{iv}	0.918 (16)	1.970 (16)	2.8579 (12)	162.5 (13)
N6—H6A \cdots O31 ^v	0.921 (16)	1.912 (16)	2.8296 (13)	174.0 (13)
N6—H6B \cdots O11	0.921 (16)	1.879 (16)	2.7735 (12)	163.2 (13)
N6—H6C \cdots O29 ^{vi}	0.947 (16)	1.849 (17)	2.7847 (12)	169.0 (14)
N6—H6C \cdots O30 ^{vi}	0.947 (16)	2.590 (16)	3.0880 (12)	113.2 (11)
O31—H31A \cdots O13 ^{vii}	0.89 (2)	1.88 (2)	2.7418 (11)	165.2 (17)
O31—H31B \cdots O11	0.88 (2)	1.97 (2)	2.8184 (11)	162.2 (16)

Symmetry codes: (iii) $-x+1, -y+1, -z$; (iv) $x+1, y-1, z+1$; (v) $-x, -y+1, -z+1$; (vi) $x, y-1, z+1$; (vii) $-x+1, -y+1, -z+1$.